

JEE-Main Examination – JANUARY 2026

( 22<sup>ND</sup> JANUARY 2026)

TIME: 9:00 A.M. TO 12:00 NOON

## MATHEMATICS TEST PAPER WITH SOLUTION

1. If the sum of the first four terms of an A.P. is 6 and the sum of its first six terms is 4 , then the sum of its first twelve terms is
1. -22
  2. -20
  3. -24
  4. -26

**Ans. (1)****Sol.** Sum of first 4 term  $S_4 = 6$ 

$$\frac{4}{2}(2a + 3d) = 6 \Rightarrow 2a + 3d = 3$$

$$\text{Sum of first 6 terms } S_6 = 4$$

$$\frac{6}{2}(2a + 5d) = 4 \Rightarrow 2a + 5d = \frac{4}{3} \dots \dots$$

$$\text{eq. (2)} - \text{eq. (1)}$$

$$(2a + 5d) - (2a + 3d) = \frac{4}{3} - 3$$

$$\Rightarrow d = -\frac{5}{6}$$

$$\therefore 2a + 3\left(-\frac{5}{6}\right) = 3 \Rightarrow a = \frac{11}{4}$$

$$S_{12} = \frac{12}{2} \left\{ 2 \times \frac{11}{4} + (12 - 1) \left(-\frac{5}{6}\right) \right\}$$

$$S_{12} = 6 \left(-\frac{22}{6}\right) = -22$$

2. Let  $f(x) = x^{2025} - x^{2000}$ ,  $x \in [0,1]$  and the minimum value of the function  $f(x)$  in the interval  $[0,1]$  be  $(80)^{80}(n)^{-81}$ . Then  $n$  is equal to

1. - 41
2. -81
3. -80
4. -40

**Ans. (2)****Sol.**

$$f(x) = x^{2025} - x^{2000}$$

$$f'(x) = 0 \Rightarrow x = \left(\frac{2000}{2025}\right)^{1/25} = \alpha \text{ ( say )}$$

$$\therefore f(0) = 0, f(1) = 0, f(\alpha) = \left(\frac{80}{81}\right)^{80} \cdot \frac{-1}{81} = 80^{80} \cdot (-81)^{-81}$$

3. If  $A = \begin{bmatrix} 2 & 3 \\ 3 & 5 \end{bmatrix}$ , then the determinant of the matrix  $(A^{2025} - 3 A^{2024} + A^{2023})$  is

1. 24
2. 16

3. 28

4. 12

**Ans. (2)****Sol.**

$$A = \begin{bmatrix} 2 & 3 \\ 3 & 5 \end{bmatrix} \Rightarrow A^2 = \begin{bmatrix} 13 & 21 \\ 21 & 34 \end{bmatrix}$$

$$|A^{2025} - 3A^{2024} + A^{2023}|$$

$$= |A^{2023}(A^2 - 3A + I)|$$

$$= |A|^{2023}|A^2 - 3A + I|$$

$$= 1 \cdot \begin{vmatrix} 8 & 12 \\ 12 & 20 \end{vmatrix} = 160 - 144 = 16$$

4. If a random variable  $x$  has the probability distribution

$x$	0	1	2	3	4	5	6	7
$P(x)$	0	$2k$	$k$	$3k$	$2k^2$	$2k$	$k^2 + k$	$7k^2$

then  $P(3 < x \leq 6)$  is equal to

1. 0.34
2. 0.33
3. 0.64
4. 0.22

**ANS (2)****Sol**

$$\sum P(x_i) = 1$$

$$\Rightarrow 9k + 10k^2 = 1$$

$$\Rightarrow 10k^2 + 9k - 1 = 0 \Rightarrow k = \frac{1}{10}$$

$$P(3 < x \leq 6) = 3k + 3k^2$$

$$= \frac{3}{10} + \frac{3}{100} = 0.33$$

$$= 0.33$$

5. If the chord joining the points  $P_1(x_1, y_1)$  and  $P_2(x_2, y_2)$  on the parabola  $y^2 = 12x$  subtends a right angle at the vertex of the parabola, then  $x_1x_2 - y_1y_2$  is equal to

1. 284
2. 292
3. 280
4. 288

**ANS (4)****Sol**

$$(x_1y_1) = (3t_1^2, 6t_1) \& (x_2y_2) = (3t_2^2, 6t_2)$$

$$t_1t_2 = -4$$

$$x_1x_2 = 9(t_1t_2)^2, y_1y_2 = 36t_1t_2$$

$$x_1x_2 - y_1y_2 = 9(16) - 36(-4)$$

$$= 144 + 144$$

$$= 288$$

6. Let the relation R on the set  $M = \{1, 2, 3, \dots, 16\}$  be given by  $R = \{(x, y) : 4y = 5x - 3, x, y \in M\}$ .

Then the minimum number of elements required to be added in R, in order to make the relation symmetric, is equal to

1. 4
2. 1
3. 2
4. 3

**ANS (3)**

**Sol**  $R = \{(3, 3), (7, 8), (11, 13)\}$  to make it symmetric  $(8, 7), (13, 11)$  must be added.

7. Two distinct numbers  $a$  and  $b$  are selected at random from  $1, 2, 3, \dots, 50$ . The probability, that their product  $ab$  is divisible by 3, is

1.  $\frac{561}{1225}$
2.  $\frac{272}{1225}$
3.  $\frac{664}{1225}$
4.  $\frac{8}{25}$

**ANS (3)**

**Sol** Req. probability =  $1 - (\text{product not divisible by 3})$

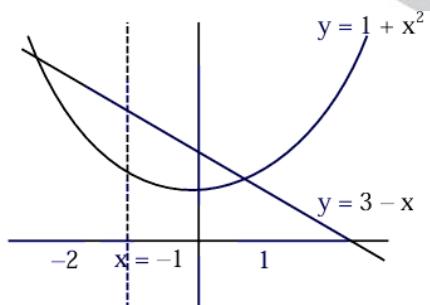
Multiple of 3 = 16

Not multiple of 3 = 34

$$= 1 - \frac{^{34}C_2}{^{50}C_2} = \frac{664}{1225}$$

8. Let the line  $x = -1$  divide the area of the region  $\{(x, y) : 1 + x^2 \leq y \leq 3 - x\}$  in the ratio  $m:n, \gcd(m, n) = 1$ . Then  $m+n$  is equal to

1. 27
2. 28
3. 25
4. 26

**ANS (1)****Sol**

$$\begin{aligned} \frac{m}{n} &= \frac{\int_{-1}^1 [(3-x) - (1+x^2)] dx}{\int_{-2}^1 [(3-x) - (1+x^2)] dx} = \frac{20}{7} \\ \therefore m+n &= 20+7 \\ &= 27 \end{aligned}$$

9. The number of solutions of  $\tan^{-1} 4x + \tan^{-1} 6x = \frac{\pi}{6}$ , where  $-\frac{1}{2\sqrt{6}} < x < \frac{1}{2\sqrt{6}}$ , is equal to

1. 3
2. 2

3. 1

4. 0

**ANS (3)****Sol**

$$\begin{aligned}\tan^{-1} 4x + \tan^{-1} 6x &= \frac{\pi}{6} \\ \Rightarrow \tan^{-1} \left( \frac{4x+6x}{1+24x^2} \right) &= \frac{\pi}{6} \\ \Rightarrow \frac{10x}{1-24x^2} &= \frac{1}{\sqrt{3}} \\ \Rightarrow 24x^2 + 10\sqrt{3}x - 1 &= 0 \\ x &= \frac{-10\sqrt{3} \pm \sqrt{300+96}}{48} \\ x &= \frac{\sqrt{396}-10\sqrt{3}}{48}\end{aligned}$$

Only 1 solution in  $\left(-\frac{1}{2\sqrt{6}}, \frac{1}{2\sqrt{6}}\right)$ 

- 10.** Let  $f: [1, \infty) \rightarrow \mathbb{R}$  be a differentiable function. If  $6 \int_1^x f(t)dt = 3xf(x) + x^3 - 4$  for all  $x \geq 1$ , then the value of  $f(2) - f(3)$  is

1. - 4

2. 4

3. -3

4. 3

**ANS (4)****Sol**  $6 \int_1^x f(t)dt = 3xf(x) + x^3 - 4$ 

Diff. both side

$$\begin{aligned}6f(x) &= 3xf'(x) + 3f(x) + 3x^2 \\ 3f(x) &= 3xf'(x) + 3x^2 \\ x \frac{dy}{dx} - y &= -x^2 \\ \frac{x \frac{dy}{dx} - 9}{x^2} &= -1 \\ \Rightarrow \frac{d}{dx} \left( \frac{y}{x} \right) &= -1 \\ \frac{y}{x} &= -x + C \\ \Rightarrow f(x) &= -x^2 + Cx \\ \text{at } x = 1, y = 1 &\Rightarrow C = 2 \\ f(x) &= -x^2 + 2x \\ f(2) - f(3) &= 3\end{aligned}$$

- 11.** If the image of the point  $P(1, 2, a)$  in the line  $\frac{x-6}{3} = \frac{y-7}{2} = \frac{z-1}{2}$  is  $Q(5, b, c)$ , then  $a^2 + b^2 + c^2$  is equal to

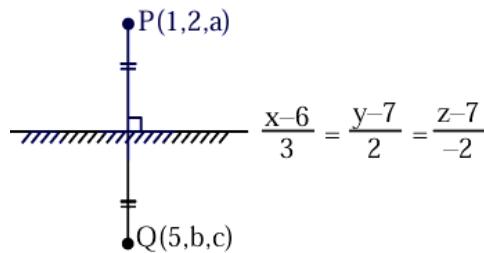
1. 298

2. 293

3. 264

4. 283

**ANS (1)****Sol**



Point  $M \equiv \left( 3, \frac{b}{2} + 1, \frac{c+a}{2} \right)$  satisfies the line

$$\begin{aligned}\frac{3-6}{3} &= \frac{\frac{b}{2}+1-7}{2} = \frac{\frac{c+a}{2}-7}{-2} \\ -1 &= \frac{b-12}{4} = \frac{c+a-14}{-4} \\ \Rightarrow b &= 8 \dots (1) \text{ & } c+a = 18 \dots (2)\end{aligned}$$

Now  $PQ \perp L$

$$\begin{aligned}\Rightarrow (4i + (b-2)j + (c-a)k) \cdot (3i + 2j - 2k) &= 0 \\ \Rightarrow 12 + 2(b-2) - 2(c-a) &= 0 \\ \Rightarrow 6 + (b-2) - (c-a) &= 0 \\ \Rightarrow b - c + a + 4 &= 0 \\ \Rightarrow 8 - c + a + 4 &= 0 \\ \Rightarrow c + a &= 12 \dots (3)\end{aligned}$$

From (2) & (3)

$$c = 15 \text{ & } a = 3$$

$$\text{So } a^2 + b^2 + c^2 = 9 + 64 + 225 = 298$$

- 12.** If the line  $\alpha x + 2y = 1$ , where  $\alpha \in \mathbb{R}$ , does not meet the hyperbola  $x^2 - 9y^2 = 9$ , then a possible value of  $\alpha$  is:

1. 0.7
2. 0.6
3. 0.8
4. 0.5

**ANS (3)**

**Sol**  $y = \frac{1-\alpha x}{2}$

Put this in equation of hyperbola

$$\begin{aligned}\therefore x^2 - 9\left(\frac{1-\alpha x}{2}\right)^2 &= 9 \\ (4-9\alpha^2)x^2 + 18\alpha x - 45 &= 0\end{aligned}$$

$\because$  line does not intersect hyperbola

$$\begin{aligned}\therefore D &< 0 \\ \Rightarrow \alpha^2 - \frac{5}{9} &> 0 \\ \Rightarrow \alpha \in \left(-\infty, -\frac{\sqrt{5}}{3}\right) \cup \left(\frac{\sqrt{5}}{3}, \infty\right)\end{aligned}$$

Here  $\frac{\sqrt{5}}{3} \approx 0.74$

13. The coefficient of  $x^{48}$  in  $(1+x) + 2(1+x)^2 + 3(1+x)^3 + \dots + 100(1+x)^{100}$  is equal to

1.  ${}_{100}C_{49} - {}_{100}C_{50}$
2.  $100 \cdot {}_{100}C_{49} - {}_{100}C_{48}$
3.  ${}_{100}C_{50} + {}_{101}C_{49}$
4.  $100 \cdot {}_{101}C_{49} - {}_{101}C_{50}$

**ANS (4)**

**Sol** Let  $1+x = r$

$$\therefore S = 1.r + 2.r^2 + 3.r^3 + \dots + 100r^{100} \dots \dots (1)$$

$$rS = 1.r^2 + 2.r^3 + \dots + 99r^{100} + 100r^{101} \dots \dots (2)$$

(1) - (2) gives

$$S = -\frac{(1+x)^{101}}{x^2} + \frac{1}{x^2} + \frac{100(1+x)^{101}}{x}$$

$\therefore$  coefficient  $x^{48}$  in  $S$

$$= -\text{coefficient of } x^{48} \text{ in } \frac{(1+x)^{101}}{x^2} + 100.$$

$$\text{Coefficient of } x^{48} \text{ in } \frac{(1+x)^{101}}{x}$$

$$= 100 {}_{101}C_{49} - {}_{101}C_{50}$$

14. Let the solution curve of the differential equation  $xdy - ydx = \sqrt{x^2 + y^2} dx, x > 0, y(1) = 0$ , be  $y = y(x)$ . Then  $y(3)$  is equal to

1. 1
2. 4
3. 2
4. 6

**ANS (2)**

**Sol**

$$\begin{aligned} \frac{xdy - ydx}{x^2} &= \frac{\sqrt{x^2 + y^2}}{x^2} dx \\ d\left(\frac{y}{x}\right) &= \sqrt{1 + \frac{y^2}{x^2}} \cdot \frac{1}{x} dx \\ \int \frac{d\left(\frac{y}{x}\right)}{\sqrt{1 + \left(\frac{y}{x}\right)^2}} &= \int \frac{1}{x} dx \\ \Rightarrow \ln \left( \frac{y}{x} + \sqrt{1 + \frac{y^2}{x^2}} \right) &= \ln x + \ln k = \ln kx \\ \Rightarrow y + \sqrt{x^2 + y^2} &= kx^2 \\ 0 + 1 &= k \\ \Rightarrow y + \sqrt{x^2 + y^2} &= x^2 \\ y + \sqrt{9 + y^2} &= 9 \\ y &= 4 \end{aligned}$$

15. The number of distinct real solutions of the equation  $x|x+4| + 3|x+2| + 10 = 0$  is

1. 2

2. 0

3. 1

4. 3

**ANS (3)****Sol** Case I  $x < -4$ 

$$\begin{aligned}x(-(x+4)) + 3(-(x+2)) + 10 &= 0 \\x^2 + 7x - 4 &= 0 \\ \Rightarrow x = -\frac{7 + \sqrt{65}}{2} \text{ or } -\frac{7 - \sqrt{65}}{2} &\end{aligned}$$

Reject                      Accept

Case II  $-4 \leq x < -2$ 

$$\begin{aligned}x(x+4) + 3(-(x+2)) + 10 &= 0 \\x^2 + x + 4 &= 0\end{aligned}$$

D &lt; 0                      No solution

Case III  $x \geq -1$ 

$$\begin{aligned}x(x+4) + 3(x+2) + 10 &= 0 \\x^2 + 7x + 16 &= 0\end{aligned}$$

D &lt; 0                      No solution

 $\Rightarrow$  No. of solution = 1.

- 16.** The value of  $\int_{-\frac{\pi}{2}}^{\frac{\pi}{2}} \left( \frac{1}{[x]+4} \right) dx$ , where  $[ \cdot ]$  denotes the greatest integer function, is

1.  $\frac{7}{60}(3\pi - 1)$ 2.  $\frac{7}{60}(\pi - 3)$ 3.  $\frac{1}{60}(21\pi - 1)$ 4.  $\frac{1}{60}(\pi - 7)$ **ANS (1)****Sol**

$$\begin{aligned}I &= \int_{-\pi/2}^{\pi/2} \frac{1}{[x]+4} dx \\I &= \int_{-\pi/2}^{-1} \frac{dx}{2} + \int_{-1}^0 \frac{dx}{3} + \int_0^1 \frac{dx}{4} + \int_1^{\pi/2} \frac{dx}{5} \\&= \frac{1}{2} \left( -1 + \frac{\pi}{2} \right) + \frac{1}{3}(1) + \frac{1}{4}(1) + \left( \frac{\pi}{2} - 1 \right) \frac{1}{5} \\&= \frac{7\pi}{20} - \frac{7}{60} = \frac{7}{60}(3\pi - 1)\end{aligned}$$

- 17.** If the domain of the function  $f(x) = \sin^{-1} \left( \frac{5-x}{3+2x} \right) + \frac{1}{\log_e(10-x)}$  is  $(-\infty, \alpha] \cup [\beta, \gamma) - \{\delta\}$ , then

6( $\alpha + \beta + \gamma + \delta$ ) is equal to

1. 70

2. 66

3. 67

4. 68

**ANS (1)**

**Sol**

$$\begin{aligned} -1 &\leq \frac{5-x}{2x+3} \leq 10-x > 0, 10-x \neq 1 \\ \left| \frac{5-x}{2x+3} \right| &\geq 1 \& x < 10 \& x \neq 9 \\ (5-x)^2 - (2x+3)^2 &\leq 0 \& x < 10 \& 4x \neq 9 \\ (x+8)(3x-2) &\geq 0 \& x < 10 \& x \neq 9 \\ \Rightarrow (-\infty, -8] \cup \left[ \frac{2}{3}, 10 \right) - \{9\} & \\ \Rightarrow (\alpha + \beta + \gamma + \delta) &= 6 \left( -8 + \frac{2}{3} + 10 + 9 \right) \\ &= 70 \end{aligned}$$

- 18.** Let the set of all values of  $r$ , for which the circles  $(x+1)^2 + (y+4)^2 = r^2$  and  $x^2 + y^2 - 4x - 2y - 4 = 0$  intersect at two distinct points be the interval  $(\alpha, \beta)$ . Then  $\alpha\beta$  is equal to
1. 24
  2. 21
  3. 25
  4. 20

**ANS (3)**

**Sol**

$$\begin{aligned} (x-2)^2 + (y-1)^2 &= 3^2 \& (x+1)^2 + (y+4)^2 = r^2 \\ |r_1 - r_2| &< c_1 c_2 < r_1 + r_2 \\ |r - 3| &< \sqrt{(2+1)^2 + (1+4)^2} < r + 3 \\ |r - 3| &< \sqrt{34} \& r + 3 > \sqrt{34} \\ -\sqrt{34} < r - 3 &< \sqrt{34} \& r > \sqrt{34} - 3 \\ \text{i.e. } r &= (3 - \sqrt{34}, 3 + \sqrt{34}) \cap (\sqrt{34} - 3, \infty) \\ \text{i.e. } r &\in (\sqrt{34} - 3, \sqrt{34} + 3) \\ \therefore \alpha\beta &= (\sqrt{34} - 3)(\sqrt{34} + 3) \\ &= 34 - 9 \\ &= 25 \end{aligned}$$

- 19.** Let  $P(\alpha, \beta, \gamma)$  be the point on the line  $\frac{x-1}{2} = \frac{y+1}{-3} = z$  at a distance  $4\sqrt{14}$  from the point  $(1, -1, 0)$  and nearer to the origin. Then the shortest distance, between the lines  $\frac{x-\alpha}{1} = \frac{y-\beta}{2} = \frac{z-\gamma}{3}$  and  $\frac{x+5}{2} = \frac{y-10}{1} = \frac{z-3}{1}$ , is equal to

1.  $2\sqrt{\frac{7}{4}}$

2.  $4\sqrt{\frac{5}{7}}$

3.  $7\sqrt{\frac{5}{4}}$

4.  $4\sqrt{\frac{7}{5}}$

**ANS (4)**

**Sol** Let  $P(2\lambda + 1, -3\lambda - 1, \lambda)$

Then  $4\lambda^2 + 9\lambda^2 + \lambda^2 = 16 \cdot 14 \Rightarrow \lambda = \pm 4 \Rightarrow -4$  (nearer to origin)

$$\therefore P(-7, 11, -4)$$

$$\therefore \text{Shortest distance} = \frac{\left| \begin{matrix} 2 & -1 & 7 \\ 1 & 2 & 3 \\ 2 & 1 & 1 \end{matrix} \right|}{\sqrt{i^2 + j^2 + k^2}} = \frac{\frac{28}{\sqrt{1+25+9}}}{\sqrt{5}} = \frac{4\sqrt{7}}{\sqrt{5}}$$

- 20.** Let  $\overrightarrow{AB} = 2\hat{i} + 4\hat{j} - 5\hat{k}$  and  $\overrightarrow{AD} = \hat{i} + 2\hat{j} + \lambda\hat{k}, \lambda \in \mathbb{R}$ . Let the projection of the vector  $\vec{v} = \hat{i} + \hat{j} + \hat{k}$  on the diagonal  $\overrightarrow{AC}$  of the parallelogram ABCD be of length one unit. If  $\alpha, \beta$ , where  $\alpha > \beta$ , be the roots of the equation  $\lambda^2 x^2 - 6\lambda x + 5 = 0$ , then  $2\alpha - \beta$  is equal to

1. 4
2. 6
3. 1
4. 3

**ANS (4)**

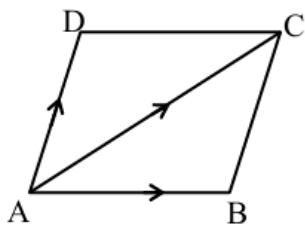
**Sol**

$$\overrightarrow{AC} = 3\hat{i} + 6\hat{j} + (\lambda - 5)\hat{k}$$

$$\vec{v} \cdot \overrightarrow{AC} = 1 \Rightarrow 3 + 6 + \lambda - 5 = \sqrt{9 + 36 + (\lambda - 5)^2}$$

$$\Rightarrow \lambda^2 + 8\lambda + 16 = \lambda^2 - 10\lambda + 70$$

$$\Rightarrow \lambda = \frac{54}{18} = 3$$



$$\therefore \text{Quadratic : } 9x^2 - 18x + 5 = 0 \Rightarrow x = \frac{1}{3}, \frac{5}{3}$$

$$\therefore 2\alpha - \beta = \frac{10-1}{3} = 3$$

## SECTION-B

### NUMERICAL VALUE

- 21.** If  $\int (\sin x)^{-\frac{11}{2}} (\cos x)^{-\frac{5}{2}} dx = -\frac{p_1}{q_1}(\cot x)^{\frac{9}{2}} - \frac{p_2}{q_2}(\cot x)^{\frac{5}{2}} - \frac{p_3}{q_3}(\cot x)^{\frac{1}{2}} + \frac{p_4}{q_4}(\cot x)^{-\frac{3}{2}} + C$ , where  $p_i$  and  $q_i$  are positive integers with  $\gcd(p_i, q_i) = 1$  for  $i = 1, 2, 3, 4$  and C is the constant of integration, then  $\frac{15p_1p_2p_3p_4}{q_1q_2q_3q_4}$  is equal to \_\_\_\_\_.

**ANS (16)**

$$\begin{aligned} \int (\tan x)^{-11/2} \cdot \sec^8 x dx \\ = \int (\tan x)^{-11/2} (1 + \tan^2 x) 3 \sec^2 x dx \end{aligned}$$

Put  $\tan x = t$

$$\begin{aligned}
 & \Rightarrow \int t^{-11/2}(1+t^2)^3 dt = \int t^{-11/2}(1+t^6+3t^2+3(4))dt \\
 & = \int (t^{-11/2} + t^{1/2} + 3 \cdot t^{-7/2} + 3 \cdot t^{-3/2})dt \\
 & = -\frac{2}{9}(\cot x)^{9/2} - \frac{6}{5}(\cot)^{5/2} - 6(\cot x)^{1/2} + \frac{2}{3}(\cot x)^{-3/2} + C \\
 & \Rightarrow p_1 = 2, p_2 = 6, p_3 = 6, p_4 = 2 \\
 & \& q_1 = 9, q_2 = 5, q_3 = 1, q_4 = 3 \\
 & \frac{15p_1p_2p_3p_4}{q_1q_2q_3q_4} = \frac{15 \cdot 2 \cdot 6 \cdot 6 \cdot 2}{9 \cdot 5 \cdot 1 \cdot 3} = 16
 \end{aligned}$$

22. Let  $\alpha = \frac{-1+i\sqrt{3}}{2}$  and  $\beta = \frac{-1-i\sqrt{3}}{2}, i = \sqrt{-1}$ .  
If  $(7 - 7\alpha + 9\beta)^{20} + (9 + 7\alpha - 7\beta)^{20} + (-7 + 9\alpha + 7\beta)^{20} + (14 + 7\alpha + 7\beta)^{20} = m^{10}$ , then  $m$  is \_\_\_\_

**ANS** (49)

**Sol**

$$\begin{aligned}
 & (9 + 7\omega - 7\omega^2) + \omega^{20}(9 + 7\omega - 7\omega^2)^{20} + \\
 & \omega^{40}(9 + 7\omega - 7\omega^2)^{20} + (14 + 7(\omega + \omega^2))^{20} \\
 & (9 + 7\omega - 7\omega^2)^{20}(1 + \omega + \omega^2) + (14 - 7)^{20} = 7^{20} \\
 & = (49)^{10}
 \end{aligned}$$

Hence,  $M = 49$

23. If  $\frac{\cos^2 48^\circ - \sin^2 12^\circ}{\sin^2 24^\circ - \sin^2 6^\circ} = \frac{\alpha + \beta\sqrt{5}}{2}$ , where  $\alpha, \beta \in \mathbb{N}$ , then  $\alpha + \beta$  is equal to \_\_\_\_.

**ANS** (4)

**Sol**

$$\begin{aligned}
 \text{Use } & \sin(A+B)\sin(A-B) = \\
 & \sin^2 A - \sin^2 B \\
 & \cos(A+B)\cos(A-B) = \cos^2 A - \sin^2 B \\
 & \frac{\cos 60^\circ \cos 36^\circ}{\sin 30^\circ \sin 18^\circ} = \frac{\sqrt{5}+1}{\sqrt{5}-1} \times \frac{\sqrt{5}+1}{\sqrt{5}+1} = \frac{(\sqrt{5}+1)^2}{4} \\
 & = \frac{3+\sqrt{5}}{2} \\
 & \alpha = 3; \beta = 1
 \end{aligned}$$

24. Let ABC be a triangle. Consider four points  $p_1, p_2, p_3, p_4$  on the side AB, five points  $p_5, p_6, p_7, p_8, p_9$  on the side BC, and four points  $p_{10}, p_{11}, p_{12}, p_{13}$  on the side AC. None of these points is a vertex of the triangle ABC. Then the total number of pentagons, that can be formed by taking all the vertices from the points  $p_1, p_2, \dots, p_{13}$ , is \_\_\_\_

**ANS** (660)

**Sol** Case 1:

2 from AB, 2 from BC, 1 from AC

$$\binom{4}{2} \cdot \binom{5}{2} \cdot \binom{4}{1} = 6 \cdot 10 \cdot 4 = 240$$

Case 2:

2 from AB, 1 from BC, 2 from AC

$$\binom{4}{2} \cdot \binom{5}{1} \cdot \binom{4}{2} = 6 \cdot 5 \cdot 6 = 180$$

Case 3 :

1 from AB, 2 from BC, 2 from AC

$$\binom{4}{1} \cdot \binom{5}{2} \cdot \binom{4}{2} = 4 \cdot 10 \cdot 6 = 240$$

- 25.** Let  $A$  be a  $3 \times 3$  matrix such that  $A + A^T = 0$ . If  $A \begin{bmatrix} 1 \\ -1 \\ 0 \end{bmatrix} = \begin{bmatrix} 3 \\ 3 \\ 2 \end{bmatrix}$ ,  $A^2 \begin{bmatrix} 1 \\ -1 \\ 0 \end{bmatrix} = \begin{bmatrix} -3 \\ 19 \\ -24 \end{bmatrix}$  and  $\det(\text{adj}(2\text{adj}(A + I))) = (2)^\alpha \cdot (3)^\beta \cdot (11)^\gamma$ ,  $\alpha, \beta, \gamma$  are non-negative integers, then  $\alpha + \beta + \gamma$  is equal to \_\_\_\_\_

**ANS** (18)

**Sol**  $A \begin{bmatrix} 1 \\ -1 \\ 0 \end{bmatrix} = \begin{bmatrix} 3 \\ 3 \\ 2 \end{bmatrix}$  and  $A \begin{bmatrix} 3 \\ 3 \\ 2 \end{bmatrix} = \begin{bmatrix} -3 \\ 19 \\ -24 \end{bmatrix}$

$$\det(\text{adj}(2\text{adj}(A + I))) = |2\text{adj}(A + I)|^2 = 64|\text{adj}(A + I)|^2$$

$$= 64|A + I|^4$$

$$\text{Let } A = \begin{bmatrix} 0 & a & b \\ -a & 0 & c \\ -b & -c & 0 \end{bmatrix} \Rightarrow \begin{cases} -a = 3 \\ -b + c = 2 \\ 3a + 2b = -3 \end{cases}$$

$$\Rightarrow \begin{cases} a = -3 \\ b = 3 \\ c = 5 \end{cases}$$

$$\therefore |A + I| = \begin{vmatrix} 1 & -3 & 3 \\ 3 & 1 & 5 \\ -3 & -5 & 1 \end{vmatrix} = 44$$

### PHYSICS TEST PAPER WITH SOLUTION

- 26.** The escape velocity from a spherical planet  $A$  is 10 km/s. The escape velocity from another planet  $B$  whose density and radius are 10% of those of planet  $A$ , is \_\_\_ m/s.

1.  $200\sqrt{5}$
2. 1000
3.  $100\sqrt{10}$
4.  $1000\sqrt{2}$

**Ans.** (3)

**Sol.**

$$V_e = \sqrt{\frac{2GM}{R}} = \sqrt{\frac{2G \times \rho \times \frac{4\pi R^3}{3}}{R}} \Rightarrow V_e \propto \sqrt{\rho} \times R$$

$$\frac{(V_e)_B}{(V_e)_A} = \sqrt{\frac{\rho_B}{\rho_A} \times \frac{R_B}{R_A}} = \sqrt{\frac{0.1\rho_A}{\rho_A} \times \left(\frac{0.1R_A}{R_A}\right)}$$

$$\frac{(V_e)_B}{(V_e)_A} = \frac{1}{10} \times \frac{1}{\sqrt{10}}$$

$$(V_e)_B = \frac{10 \times 1000}{10\sqrt{10}} = 100\sqrt{10} \text{ m/sec}$$

- 27.** A solid sphere of mass 5 kg and radius 10 cm is kept in contact with another solid sphere of mass 10 kg and radius 20 cm. The moment of inertia of this pair of spheres about the tangent passing through the point of contact is \_\_\_ kg.m<sup>2</sup>.

1. 0.18

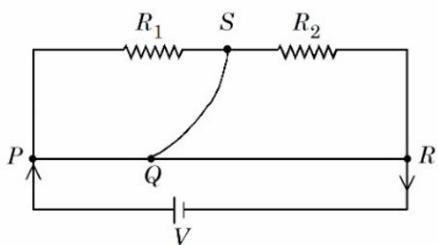
2. 0.63  
3. 0.72  
4. 0.36

**Ans. (2)**

**Sol.**

$$\begin{aligned} I &= \frac{7}{5} [m_1 R_1^2 + m_2 R_2^2] \\ &= \frac{7}{5} [5(10)^2 + 10 \times (20)^2] \times 10^{-4} \\ I &= 63 \times 10^{-2} \text{ kg m}^2 \\ I &= 0.63 \text{ kg m}^2 \end{aligned}$$

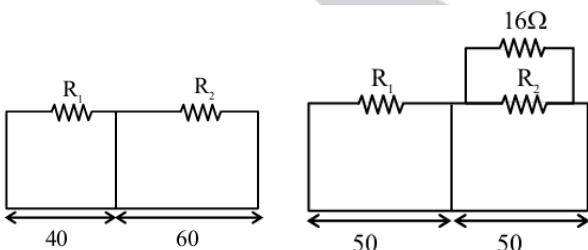
**28.** A meter bridge with two resistances  $R_1$  and  $R_2$  as shown in figure was balanced (null point) at 40 cm from the point  $P$ . The null point changed to 50 cm from the point  $P$ , when  $16\Omega$  resistance is connected in parallel to  $R_2$ . The values of resistances  $R_1$  and  $R_2$  are \_\_\_\_.



1.  $R_2 = 16\Omega, R_1 = \frac{16}{3}\Omega$   
2.  $R_2 = 8\Omega, R_1 = \frac{16}{3}\Omega$   
3.  $R_2 = 4\Omega, R_1 = \frac{4}{3}\Omega$   
4.  $R_2 = 12\Omega, R_1 = \frac{12}{3}\Omega$

**Ans. (2)**

**Sol.**



$$\begin{aligned} \frac{R_1}{R_2} &= \frac{40}{60} = \frac{2}{3} \\ \frac{R_1}{\left(\frac{R_2 \times 16}{R_2 + 16}\right)} &= \frac{50}{50} \Rightarrow R_1 = \frac{16R_2}{16+R_2} \dots \\ \frac{2}{3} R_2 &= \frac{16R_2}{16+R_2} \\ \frac{32}{3} + \frac{2R_2}{3} &= 16 \\ \frac{2R_2}{3} &= 16 - \frac{32}{3} = \frac{16}{3} \\ R_2 &= 8\Omega \end{aligned}$$

By equation (1)

$$R_1 = \frac{2}{3} R_2 = \frac{16}{3} \Omega$$

- 29.** Electric field in a region is given by  $\vec{E} = Ax\hat{i} + By\hat{j}$ , where  $A = 10 \text{ V/m}^2$  and  $B = 5 \text{ V/m}^2$ . If the electric potential at a point (10,20) is 500 V , then the electric potential at origin is \_\_\_ V.

1. 2000
2. 500
3. 0
4. 1000

**Ans.** (1)

**Sol.**

$$\vec{E} = 10x\hat{i} + 5y\hat{j}$$

$$V_{\text{at}(10,20)} = 500 \text{ V}$$

$$\Delta V = \int \vec{E} \cdot d\vec{r}$$

$$500 - V_0 = - \int_{(0,0)}^{(10,20)} (10x\hat{i} + 5y\hat{j}) \cdot (dx\hat{i} + dy\hat{j})$$

$$500 - V_0 = - \left[ 5x^2 + \frac{5y^2}{2} \right]_{(0,0)}^{(10,20)}$$

$$500 - V_0 = \left( 500 + 5 \times \frac{400}{2} \right) - (0 - 0)$$

$$500 - V_0 = 500 + 1000$$

$$V_0 = 2000 \text{ V}$$

- 30.** Match the LIST-I with LIST-II

List-I		List-II	
A.	Spring constant	I.	$ML^2 T^{-2} K^{-1}$
B.	Thermal conductivity	II.	$ML^0 T^{-2}$
C.	Boltzmann constant	III.	$ML^2 T^{-3} A^{-2}$
D.	Inductive reactance	IV.	$MLT^{-3} K^{-1}$

Choose the correct answer from the options given below:

1. A-I, B-IV, C-II, D-III
2. A-III, B-II, C-IV, D-I
3. A-II, B-I, C-IV, D-III
4. A-II, B-IV, C-I, D-III

**Ans.** (4)

**Sol.** (A)

$$F = Kx$$

$$[MLT^{-2}] = [K][L]$$

$$[K] = ML^0 T^{-2}$$

(B) Thermal conductivity

$$\frac{dQ}{dt} = \frac{kA}{\ell} \Delta T$$

$$ML^2 T^{-3} = \frac{[k]L^2 K}{L}$$

$$[K] = MLT^{-3} K^{-1}$$

(C) Boltzman constant

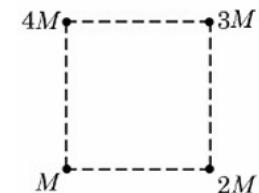
$$[K] = ML^2 T^{-2} K^{-1}$$

(D) Inductive reactance

$$\frac{[V]}{[I]} = \frac{ML^2 T^{-3} A^{-1}}{A}$$

$$= ML^2 T^{-3} A^{-2}$$

- 31.** Net gravitational force at the center of a square is found to be  $F_1$  when four particles having mass  $M, 2M, 3M$  and  $4M$  are placed at the four corners of the square as shown in figure and it is  $F_2$  when the positions of  $3M$  and  $4M$  are interchanged. The ratio  $\frac{F_1}{F_2}$  is  $\frac{\alpha}{\sqrt{5}}$ . The value of  $\alpha$  is \_\_\_\_.

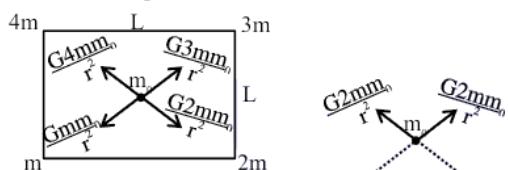


1. 3
2.  $2\sqrt{5}$
3. 1
4. 2

**Ans. (4)**

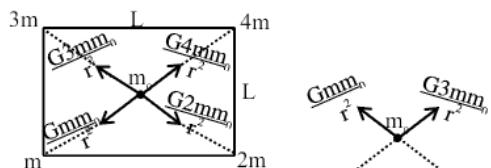
**Sol.**

Initial configuration



$$F = 2\sqrt{2} \frac{Gmm_0}{r^2}$$

New configuration



$$F' = \sqrt{10} \frac{Gmm_0}{r^2} \Rightarrow \frac{F}{F'} = 2\sqrt{2} \cdot \frac{1}{\sqrt{10}} = \frac{2}{\sqrt{5}}$$

$$\therefore \alpha = 2$$

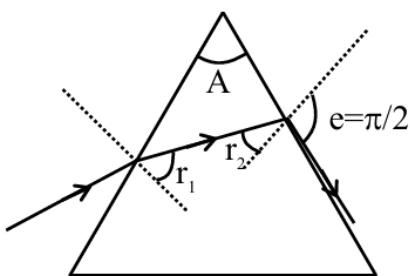
**32.** Consider an equilateral prism (refractive index  $\sqrt{2}$ ). A ray of light is incident on its one surface at a certain angle  $i$ . If the emergent ray is found to graze along the other surface then the angle of refraction at the incident surface is close to \_\_\_\_.

1.  $20^\circ$
2.  $15^\circ$
3.  $30^\circ$
4.  $40^\circ$

**Ans. (2)**

**Sol.** Equilateral prism.

$$A = 60^\circ$$



$$\mu \sin r_2 = 1 \cdot \sin e = 1$$

$$\sin r_2 = \frac{1}{\mu} = \frac{1}{\sqrt{2}}$$

$$r_2 = 45^\circ$$

$$\therefore r_1 = A - r_2 = 15^\circ$$

**33.** Given below are two statements:

Statement I: Pressure of a fluid is exerted only on a solid surface in contact as the fluid-pressure does not exist everywhere in a still fluid.

Statement II: Excess potential energy of the molecules on the surface of a liquid, when compared to interior, results in surface tension.

In the light of the above statements, choose the correct answer from the options given below

1. Statement I is true but Statement II is false
2. Statement I is false but Statement II is true
3. Both Statement I and Statement II are false
4. Both Statement I and Statement II are true

**Ans. (2)**

**Sol.** According to pascal's law pressure at any point in liquid at rest is same in all direction.

It exist at every point in the liquid not just at boundaries. So statement (1) is false.

For interior molecule net cohesive forces are zero statement (2) is correct.

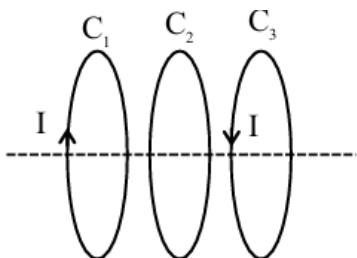
**34.** Three identical coils  $C_1, C_2$  and  $C_3$  are closely placed such that they share a common axis.  $C_2$  is exactly midway.  $C_1$  carries current  $I$  in anti-clockwise direction while  $C_3$  carries current  $I$  in clockwise direction. An induced current flows through  $C_2$  will be in clockwise direction when

1.  $C_1$  and  $C_3$  move with equal speeds towards  $C_2$
2.  $C_1$  moves away from  $C_2$  and  $C_3$  moves towards  $C_2$

3.  $C_1$  and  $C_3$  move with equal speeds away from  $C_2$
4.  $C_1$  moves towards  $C_2$  and  $C_3$  moves away from  $C_2$

**Ans.** (4)

**Sol.**



Magnetic field through the coil is

$$\vec{B} = (B_{C_2} - B_{C_1})\hat{i}$$

$$\phi = (B_{C_2} - B_{C_1})A$$

$$\varepsilon = \frac{-d\phi}{dt}$$

Find the direction according to Lenz's law

If coil move away then magnetic field decreases & vice versa

- 35.** A projectile is thrown upward at an angle  $60^\circ$  with the horizontal. The speed of the projectile is 20 m/s when its direction of motion is  $45^\circ$  with the horizontal. The initial speed of the projectile is \_\_\_\_ m/s.
1. 40
  2.  $40\sqrt{2}$
  3.  $20\sqrt{2}$
  4.  $20\sqrt{3}$

**Ans.** (3)

**Sol.**

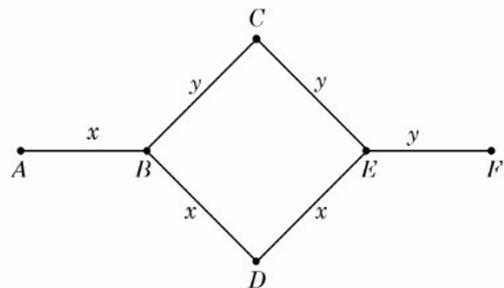
$$u \cos 60^\circ = \frac{20}{\sqrt{2}}$$

$$\frac{u}{2} = \frac{20}{\sqrt{2}}$$

$$u = \frac{40}{\sqrt{2}}$$

$$u = 20\sqrt{2} \text{ m/s}$$

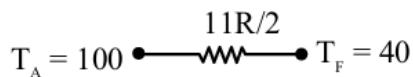
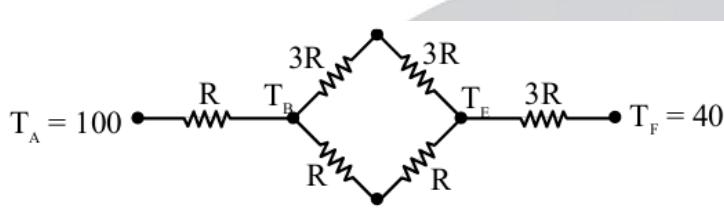
- 36.** Rods  $x$  and  $y$  of equal dimensions but of different materials are joined as shown in figure. Temperatures of end points  $A$  and  $F$  are maintained at  $100^\circ\text{C}$  and  $40^\circ\text{C}$  respectively. Given the thermal conductivity of rod  $x$  is three times of that of rod  $y$ , the temperature at junction points  $B$  and  $E$  are (close to):



1.  $80^{\circ}\text{C}$  and  $70^{\circ}\text{C}$  respectively
2.  $89^{\circ}\text{C}$  and  $73^{\circ}\text{C}$  respectively
3.  $80^{\circ}\text{C}$  and  $60^{\circ}\text{C}$  respectively
4.  $60^{\circ}\text{C}$  and  $45^{\circ}\text{C}$  respectively

**Ans.** (2)

**Sol.** Let  $[R = \frac{\ell}{3KA}]$



$$\left[ H = \frac{100 - 40}{\frac{11R}{2}} \right]$$

... (1)

$$H = \frac{100 - T_B}{R}$$

... (2)

$$H = \frac{T_E - 40}{3R}$$

... (3)

using (1) and (2)

$$120 = 1100 - 11T_A$$

$$T_B = 89^{\circ}\text{C}$$

using (1) and (3)

$$T_E = 73^{\circ}\text{C}$$

- 37.**  $7.9\text{MeV}\alpha$ -particle scatters from a target material of atomic number 79. From the given data the estimated diameter of nuclei of the target material is (approximately) \_\_\_\_ m.

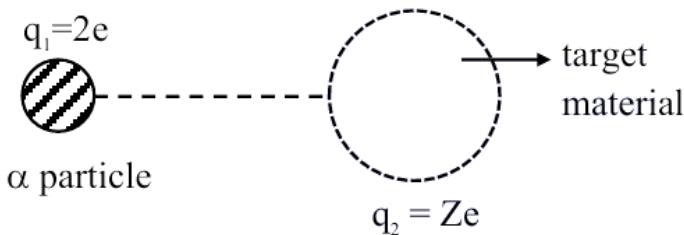
$$\left[ \frac{1}{4\pi\epsilon_0} = 9 \times 10^9 \text{Nm}^2/\text{C}^2 \text{ and electron charge } = 1.6 \times 10^{-19} \text{C} \right]$$

1.  $2.88 \times 10^{-14}$

2.  $1.69 \times 10^{-12}$

3.  $1.44 \times 10^{-13}$

4.  $5.76 \times 10^{-14}$

**Ans.** (4)**Sol.** By mechanical energy conservation

$(Me)_i = (Me)_f$

$PE_i + KE_i = PE_f + KE_f$

$0 + 7.9 \times 10^6 \times 1.6 \times 10^{-19} = \frac{k(2e)(Ze)}{r} + 0$

$r = \frac{9 \times 10^9 \times 2 \times (1.6 \times 10^{-19})^2 \times 79}{7.9 \times 10^6 \times 1.6 \times 10^{-19}} = 2.88 \times 10^{-14} \text{ m}$

For diameter  $\Rightarrow D = 2r = 5.76 \times 10^{-14} \text{ m}$

- 38.** The minimum frequency of photon required to break a particle of mass 15.348 amu into  $4\alpha$  particles is

\_\_\_\_ kHz. [mass of He nucleus = 4.002 amu, 1 amu =  $1.66 \times 10^{-27}$  kg,  $h = 6.6 \times 10^{-34}$  J.s and  $c = 3 \times 10^8 \text{ m/s}$ ]

1.  $9 \times 10^{20}$
2.  $9 \times 10^{19}$
3.  $14.94 \times 10^{20}$
4.  $14.94 \times 10^{19}$

**Ans.** (4)**Sol.**  $hv = (4 \times 4.002 - 15.348) \times 1.66 \times 10^{-27} \times (3 \times 10^8)^2$   
 $v = 14.94 \times 10^{19} \text{ kHz}$ 

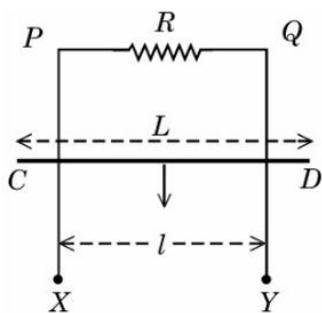
- 39.** A thin convex lens of focal length 5 cm and a thin concave lens of focal length 4 cm are combined together (without any gap) and this combination has magnification  $m_1$  when an object is placed 10 cm before the convex lens. Keeping the positions of convex lens and object undisturbed a gap of 1 cm is introduced between the lenses by moving the concave lens away, which lead to a change in magnification of total lens system to  $m_2$ . The value of  $\left|\frac{m_1}{m_2}\right|$  is \_\_\_\_.

1.  $\frac{3}{2}$
2.  $\frac{5}{9}$
3.  $\frac{25}{27}$
4.  $\frac{5}{27}$

**Ans.** (Dropped)

- 40.**  $XPQY$  is a vertical smooth long loop having a total resistance  $R$  where  $PX$  is parallel to  $QY$  and separation between them is  $l$ . A constant magnetic field  $B$  perpendicular to the plane of the loop exists in the entire space. A rod  $CD$  of length  $L$  ( $L > l$ ) and mass  $m$  is made to slide down from rest

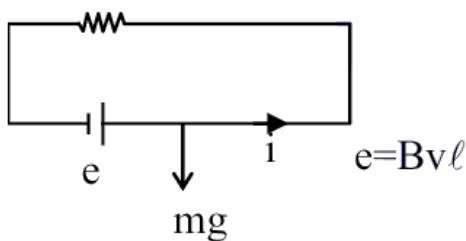
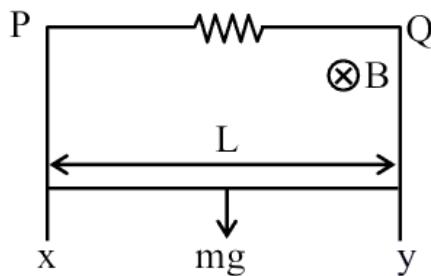
under the gravity as shown in figure. The terminal speed acquired by the rod is \_\_\_ m/s. ( $g =$  acceleration due to gravity)



1.  $\frac{8mgR}{B^2l^2}$
2.  $\frac{2mgR}{B^2L^2}$
3.  $\frac{mgR}{B^2l^2}$
4.  $\frac{2mgR}{B^2l^2}$

**Ans.** (3)

**Sol.**



at equilibrium (Or for terminal velocity)

$$mg = iB\ell \Rightarrow mg = \left(\frac{Bv\ell}{R}\right) B\ell$$

$$V = \frac{mgR}{B^2\ell^2}$$

41. The volume of an ideal gas increases 8 times and temperature becomes  $(1/4)^{\text{th}}$  of initial temperature during a reversible change. If there is no exchange of heat in this process ( $\Delta Q = 0$ ) then identify the gas from the following options (Assuming the gases given in the options are ideal gases):
1.  $\text{CO}_2$
  2.  $\text{NH}_3$

3. He

4. O<sub>2</sub>

**Ans. (3)**

**Sol.**

$$PV^\gamma = \text{constant}$$

$$TV^{\gamma-1} = \text{constant}$$

$$TV^{\gamma-1} = \left(\frac{T}{4}\right) (8V)^{(\gamma-1)}$$

$$4 = 8^{(\gamma-1)}$$

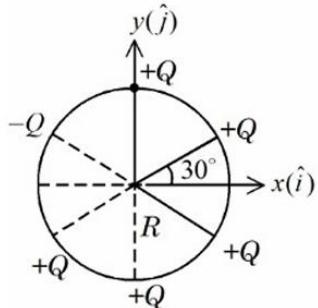
$$2^2 = 2^{3\gamma-3}$$

$$2 = 3(\gamma - 1)$$

$$\gamma = \frac{5}{3}$$

Gas is a monoatomic gas Answer is He .

- 42.** Six point charges are kept 60° apart from each other on the circumference of a circle of radius  $R$  as shown in figure. The net electric field at the center of the circle is \_\_\_\_\_. ( $\epsilon_0$  is permittivity of free space)



$$1. \frac{Q}{4\pi\epsilon_0 R^2} (\sqrt{3}\hat{i} - \hat{j})$$

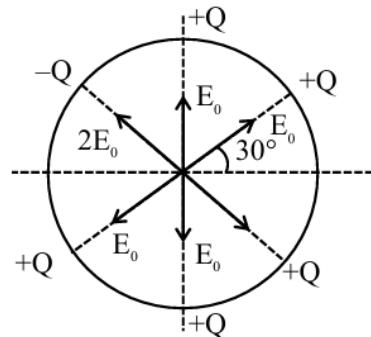
$$2. -\frac{Q}{4\pi\epsilon_0 R^2} (\sqrt{3}\hat{i} - \hat{j})$$

$$3. -\frac{5Q}{8\pi\epsilon_0 R^2} (\hat{i} + \sqrt{3}\hat{j})$$

$$4. -\left(\frac{5Q}{8\pi\epsilon_0 R^2}\right) (\hat{i} - 3\hat{j})$$

**Ans. (2)**

**Sol.** Let  $\frac{kQ}{r^2} = E_0$



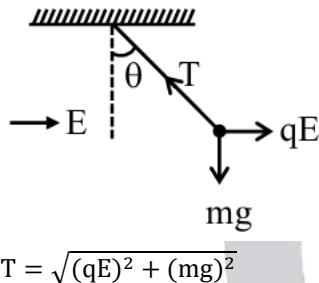
$$\begin{aligned}\vec{E}_{\text{net}} &= 2E_0 \cos 30^\circ (-\hat{i}) + 2E_0 \sin 30^\circ (\hat{j}) \\ &= \frac{2kQ}{r^2} \left[ \frac{\sqrt{3}}{2} (-\hat{i}) + \frac{1}{2} \hat{j} \right] \\ &= \frac{-1Q}{4\pi\epsilon_0 r^2} (\sqrt{3}\hat{i} - \hat{j})\end{aligned}$$

- 43.** A simple pendulum has a bob with mass  $m$  and charge  $q$ . The pendulum string has negligible mass. When a uniform and horizontal electric field  $\vec{E}$  is applied, the tension in the string changes. The final tension in the string, when pendulum attains an equilibrium position is \_\_\_\_ . (g: acceleration due to gravity)

1.  $m g + qE$
2.  $\sqrt{m^2 g^2 + q^2 E^2}$
3.  $\sqrt{m^2 g^2 - q^2 E^2}$
4.  $mg - qE$

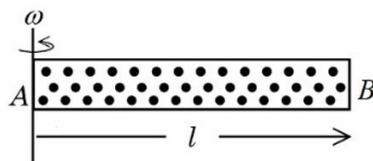
**Ans. (2)**

**Sol.**



$$T = \sqrt{(qE)^2 + (mg)^2}$$

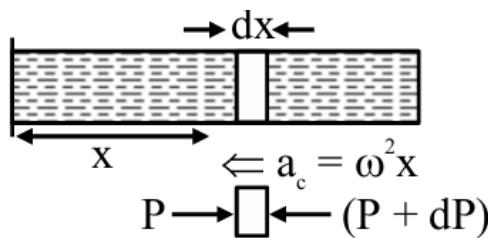
- 44.** A cylindrical tube  $AB$  of length  $l$ , closed at both ends contains an ideal gas of 1 mol having molecular weight  $M$ . The tube is rotated in a horizontal plane with constant angular velocity  $\omega$  about an axis perpendicular to  $AB$  and passing through the edge at end  $A$ , as shown in the figure. If  $P_A$  and  $P_B$  are the pressures at  $A$  and  $B$  respectively, then  
(Consider the temperature is same at all points in the tube)



1.  $P_B = P_A \exp(M\omega^2 l^2 / 3RT)$
2.  $P_B = P_A \exp(M\omega^2 l^2 / 2RT)$
3.  $P_B = P_A \exp(M\omega^2 l^2 / RT)$
4.  $P_B = P_A$

**Ans. (2)**

**Sol.**



$$A[(P + dP) - P] = (dm)(\omega^2 x)$$

$$dP = \frac{(dm)}{A} \omega^2 x$$

$$dP = \frac{(\rho)(A)(dx)\omega^2 x}{A}$$

also  $[PM = \rho RT]$

$$\rho = \frac{PM}{RT}$$

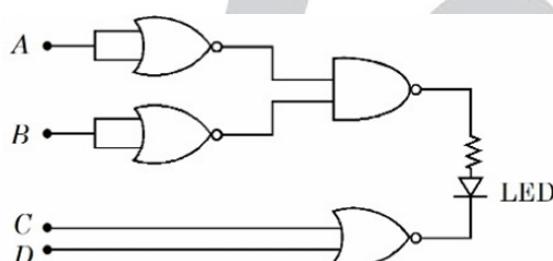
$$dP = \left( \frac{PM}{RT} \right) \omega^2 x dx$$

$$\int_{P_A}^{P_B} \frac{dP}{P} = \frac{\omega^2 M}{RT} \int_0^\ell x dx$$

$$\ell n \left( \frac{P_B}{P_A} \right) = \frac{\omega^2 \ell^2 M}{2RT}$$

$$P_B = P_A e^{\frac{M \omega^2 \ell^2}{2RT}}$$

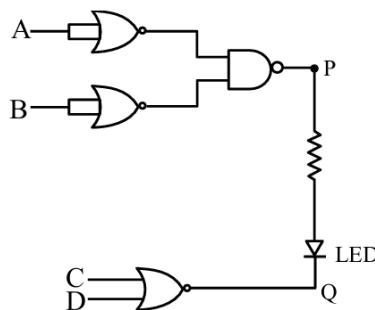
- 45.** Find the correct combination of A, B, C and D inputs which can cause the LED to glow.



1. 1101
2. 1000
3. 0100
4. 0011

**Ans.** (1)

**Sol.**



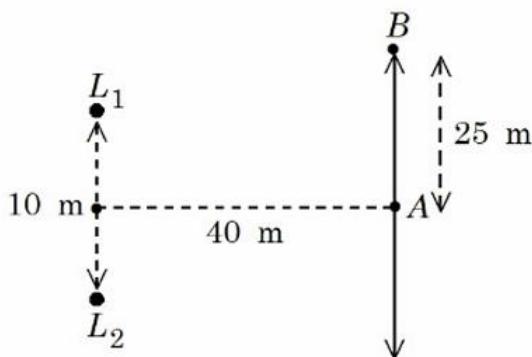
LED will glow in forward biasing :

P higher potential – 1

Q lower potential – 0

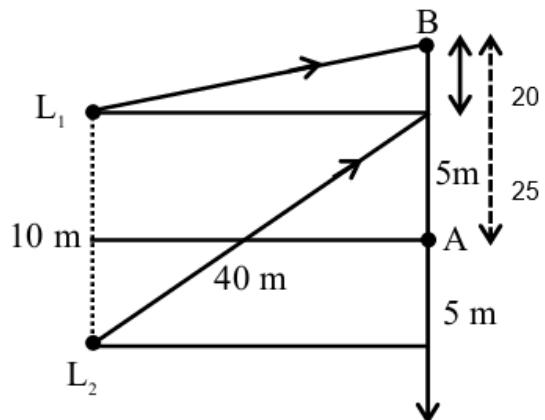
**SECTION-B****NUMERICAL VALUE**

- 46.** Two loudspeakers ( $L_1$  and  $L_2$ ) are placed with a separation of 10 m, as shown in figure. Both speakers are fed with an audio input signal of same frequency with constant volume. A voice recorder, initially at point  $A$ , at equidistance to both loud speakers, is moved by 25 m along the line  $AB$  while monitoring the audio signal. The measured signal was found to undergo 10 cycles of minima and maxima during the movement. The frequency of the input signal is \_\_\_ Hz  
(Speed of sound in air is 324 m/s and  $\sqrt{5} = 2.23$ )



**Ans. (600)**

**Sol.** Point B will 10<sup>th</sup> maxima



$$\Delta x = L_2 B - L_1 B$$

$$L_1 B = \sqrt{20^2 + 40^2} = 20\sqrt{5} \text{ m} = 44.6 \text{ m}$$

$$L_2 B = \sqrt{40^2 + 30^2} = 50 \text{ m}$$

$$\Delta x = 50 - 44.6 = 5.4 \text{ m}$$

$$\Delta x = n\lambda$$

$$5.4 = 10 \times \lambda$$

$$\lambda = 0.54 \text{ m}$$

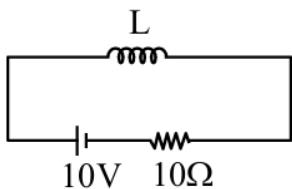
$$V = f\lambda$$

$$f = \frac{324}{0.54} = 600 \text{ Hz}$$

- 47.** Inductance of a coil with  $10^4$  turns is 10 mH and it is connected to a dc source of 10 V with internal resistance of  $10\Omega$ . The energy density in the inductor when the current reaches  $\left(\frac{1}{e}\right)$  of its maximum value is  $\alpha\pi \times \frac{1}{e^2} \text{ J/m}^3$ . The value of  $\alpha$  is \_\_\_ . ( $\mu_0 = 4\pi \times 10^{-7} \text{ Tm/A}$ ).

**Ans. (20)**

**Sol.**  $L = 10 \times 10^{-3} \text{H}$   
 $N = 10^4$



$$I_0 = \frac{10}{10} = 1 \text{ A (max current)}$$

$$I = \frac{1}{e}$$

$$E_d = \frac{B^2}{2\mu_0}$$

$$B = \mu_0 n I$$

$$L = \mu_0 n^2 \pi R^2 \ell$$

$$E_d = \frac{\mu_0 n^2 I^2}{2}$$

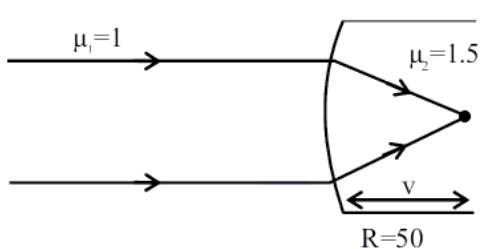
$$= \frac{4\pi \times 10^{-7} \times 10^8 \times \frac{1}{e^2}}{2}$$

$$= \frac{20\pi}{e^2}$$

- 48.** A parallel beam of light travelling in air (refractive index 1.0) is incident on a convex spherical glass surface of radius of curvature 50 cm. Refractive index of glass is 1.5. The rays converge to a point at a distance  $x$  cm from the centre of the curvature of the spherical surface. The value of  $x$  is \_\_\_\_ cm.

**Ans. (100)**

**Sol.**



$$\frac{\mu_2 - \mu_1}{v} = \frac{\mu_2 - \mu_1}{R} \Rightarrow \frac{1.5 - 1}{v} = \frac{1.5 - 1}{50}$$

$$V = 150 \text{ cm}$$

$x \rightarrow$  measure from center

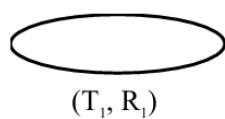
$$x = V - R$$

$$= 150 - 50 = 100 \text{ cm}$$

- 49.** A circular disc has radius  $R_1$  and thickness  $T_1$ . Another circular disc made of the same material has radius  $R_2$  and thickness  $T_2$ . If the moment of inertia of both discs are same and  $\frac{R_1}{R_2} = 2$  then  $\frac{T_1}{T_2} = \frac{1}{\alpha}$ . The value of  $\alpha$  is \_\_\_\_.

**Ans. (16)**

**Sol.**



$$m_1 = \pi R_1^2 T_1 \rho \quad m_2 = \pi R_2^2 T_2 \rho$$

$$I_1 = \frac{m_1 R_1^2}{2} \quad I_2 = \frac{m_2 R_2^2}{2}$$

$$I_1 = I_2$$

$$\frac{\pi R_1^2 T_1 \rho R_1^2}{2} = \frac{\pi R_2^2 T_2 \rho R_2^2}{2} \Rightarrow \frac{T_1}{T_2} = \frac{1}{16}$$

- 50.** The electric field of a plane electromagnetic wave, travelling in an unknown nonmagnetic medium is given by,

$$E_y = 20 \sin(3 \times 10^6 x - 4.5 \times 10^{14} t) \text{ V/m}$$

(where  $x, t$  and other values have S.I. units). The dielectric constant of the medium is \_\_\_\_ (speed of light in free space is  $3 \times 10^8 \text{ m/s}$ )

**Ans.** (4)

**Sol.**

$$\begin{aligned} n &= \frac{c}{v} \\ v &= \frac{\omega}{k} = \frac{4.5 \times 10^{14}}{3 \times 10^6} = \frac{3}{2} \times 10^8 \\ n &= 2 \\ n &= \sqrt{\mu_r \epsilon_r} (\mu_r = 1) \\ 2 &= \sqrt{\epsilon_r} \\ \epsilon_r &= 4 \end{aligned}$$

### CHEMISTRY TEST PAPER WITH SOLUTION

- 51.** Given below are two statements:

**Statement I:** The Henry's law constant  $K_H$  is constant with respect to variations in solution's concentration over the range for which the solution is ideally dilute.

**Statement II:**  $K_H$  does not differ for the same solute in different solvents.

In the light of the above statements, choose the correct answer from the options given below

5. Both Statement I and Statement II are true
6. Both Statement I and Statement II are false
7. Statement I is true but Statement II is false
8. Statement I is false but Statement II is true

**Ans.** (1)

**Sol.**  $K_H$  depends on the nature of gas and solvent.

- 52.** As compared with chlorocyclohexane, which of the following statements correctly apply to chlorobenzene?

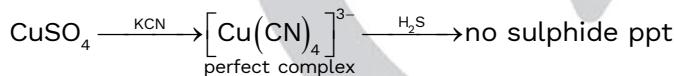
- A. The magnitude of negative charge is more on chlorine atom.  
 B. The C-Cl bond has partial double bond character.  
 C. C-C—I bond is less polar.  
 D. C-Cl bond is longer due to repulsion between delocalised electrons of the aromatic ring and lone pairs of electrons of chlorine.  
 E. The C-Cl bond is formed using  $sp^2$  hybridised orbital of carbon.
1. A, C and E Only
  2. A, D and E Only
  3. B, C and E Only
  4. B, C and D Only

**Ans. (4)****Sol.** Chlorocyclohexane is more polar due to  $-I$  effect of  $-Cl$ ,

Whereas chlorobenzene has  $-I > +M$ , so it is less polar & also has partial double bond character.

- 53.** A first row transition metal (M) does not liberate  $H_2$  gas from dilute HCl. 1 mol of aqueous solution of  $MSO_4$  is treated with excess of aqueous KCN and then  $H_2S(g)$  is passed through the solution. The amount of MS (metal sulphide) formed from the above reaction is \_\_\_\_\_ mol.

1. 3
2. 2
3. 1
4. 0

**Ans. (3)****Sol.**  $Cu \xrightarrow{dil.HCl} \text{no reaction}$ 

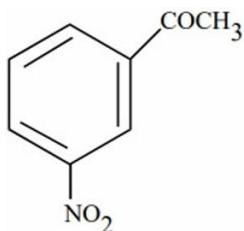
- 54.** A'p'-block element (E) and hydrogen form a binary cation  $(EH_x)^+$ , while  $EH_3$  on treatment with  $K_2HgI_4$  in alkaline medium gives a precipitate of basic mercury(II)amido-iodine. Given below are first ionisation enthalpy values ( $\text{kJ mol}^{-1}$ ) for first element each from group 13, 14, 15 and 16. Identify the correct first ionisation enthalpy value for element E.

1. 1086
2. 801
3. 1312
4. 1402

**Ans. (4)****Sol.** Element E is N, the species is  $NH_4^+$ , among B, C, N and O, N has highest first ionization energy.

- 55.** Given below are two statements:

**Statement I:** Benzene is nitrated to give nitrobenzene, which on further treatment with  $CH_3COCl/AlCl_3$  will give



**Statement II:** -NO<sub>2</sub> group is a m-directing, and deactivating group.

In the light of the above statements, choose the most appropriate answer from the options given below

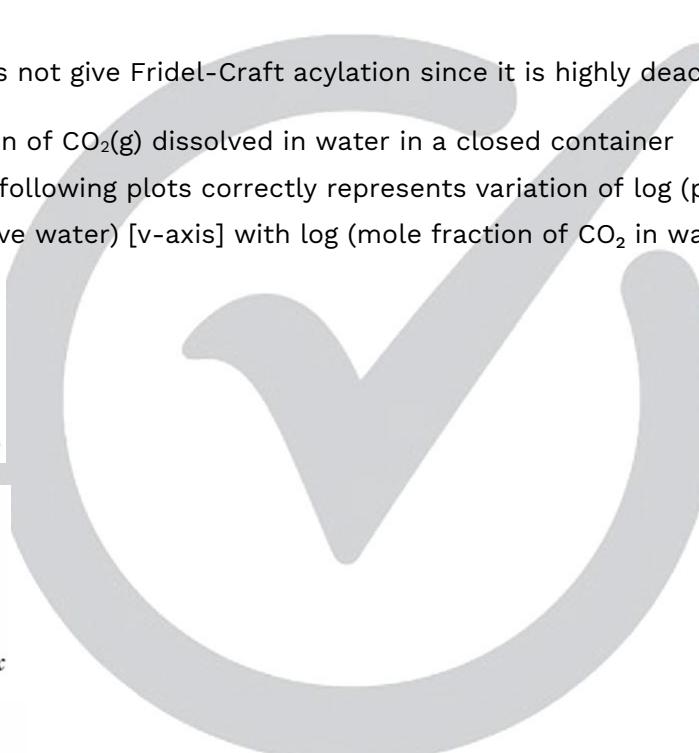
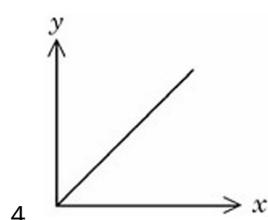
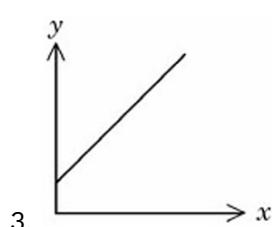
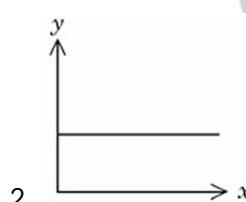
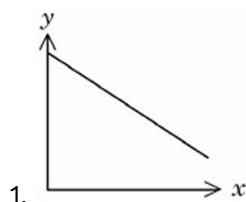
1. Both Statement I and Statement II are correct
2. Statement I is correct but Statement II is incorrect
3. Both Statement I and Statement II are incorrect
4. Statement I is incorrect but Statement II is correct

**Ans. (D)**

**Sol.** Nitrobenzene does not give Friedel-Crafts acylation since it is highly deactivated ring.

**56.** Consider a solution of CO<sub>2</sub>(g) dissolved in water in a closed container

Which one of the following plots correctly represents variation of log (partial pressure of CO<sub>2</sub> in vapour phase above water) [y-axis] with log (mole fraction of CO<sub>2</sub> in water) [x-axis] at 25°C?



**Ans. (1)**

**Sol.** From Henry's law:

$$P(g) = K_H \cdot X(g)$$

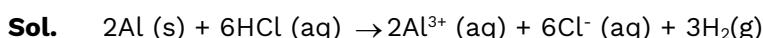
$$\log P(g) = \log K_H + \log X(g)$$

- 57.** In the reaction.



1. 12 L HCl(aq) is consumed for every 6L H<sub>2</sub>(g) produced.
2. 67.2 L H<sub>2</sub>(g) at STP is produced for every mole of Al that reacts
3. 11.2 L H<sub>2</sub>(g) at STP is produced for every mole of HCl consumed.
4. 33.6 L H<sub>2</sub>(g) is produced regardless of temperature and pressure for every mole of Al that reacts.

**Ans.** (3)



Mole of H, produced

$$= 2 \times \text{mole of HCl used}$$

$$= \frac{2}{3} \times \text{mol of Al used}$$

- 58.** Match the LIST-I with LIST-II

<b>List-I Reagents</b>	<b>List-II Name of Reaction involving carbonyl compounds</b>
A. NH <sub>2</sub> - NH <sub>2</sub> , KOH	I. Tollen's Test
B. Ag(NH <sub>3</sub> ) <sub>2</sub> OH	II. Clemmensen Reduction
C. Aq. CuSO <sub>4</sub> , Sodium Potassium tartarate, KOH	III. Wolff-Kishner Reduction
D. Zn-Hg, HC1	IV. Fehling's Test

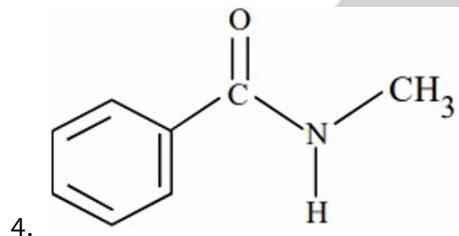
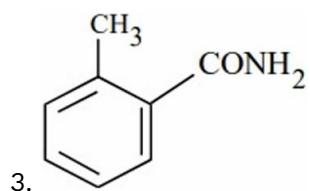
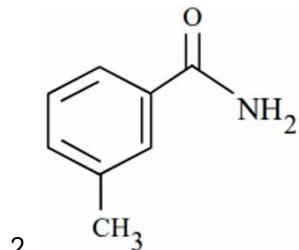
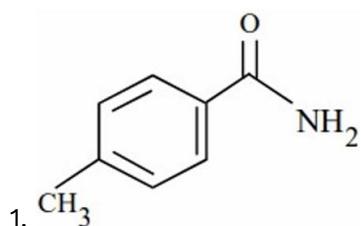
Choose the correct answer from the options given below:

1. A-III, B-IV, C-I, D-II
2. A-III, B-I, C-IV, D-II
3. A-II, B-I, C-IV, D-III
4. A-IV, B-III, C-II, D-I

**Ans.** (2)

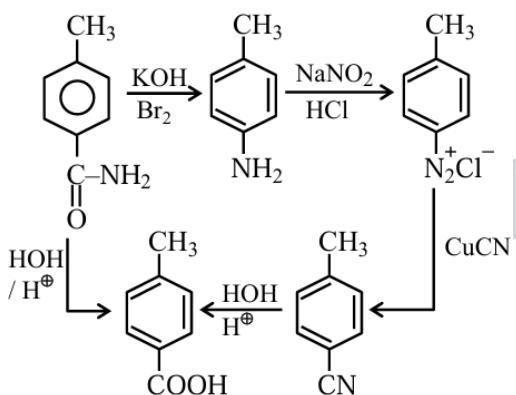
**Sol.** Theoretical (NCERT Based)

- 59.** 'A' is a neutral organic compound (M. F: C<sub>8</sub>H<sub>9</sub>ON). On treatment with aqueous Br<sub>2</sub>/HO<sup>(-)</sup>, 'A' forms a compound 'B' which is soluble in dilute acid. 'B' on treatment with aqueous NaNO<sub>2</sub>/HC1 (0-5 °C) produces a compound 'C' which on treatment with CuCN/NaCN produces 'D'. Hydrolysis of 'D' produces 'E' which is also obtainable from the hydrolysis of 'A'. 'E' on treatment with acidified KMnO<sub>4</sub> produces 'F'. 'F' contains two different types of hydrogen atoms. The structure of 'A' is



**Ans. (2)**

**Sol.**



**60. Given below are two statements:**

**Statement I:** Sucrose is dextrorotatory. However, sucrose upon hydrolysis gives a solution having mixture of products. This solution shows laevorotation.

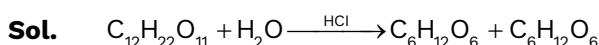
**Statement II:** Hydrolysis of sucrose gives glucose and fructose. Since the laevorotation of glucose is more than the dextrorotation of fructose, the resulting solution becomes laevorotatory.

In the light of the above statements, choose the correct answer from the options given below

1. Both Statement I and Statement II are true
2. Statement I is false but Statement II is true
3. Statement I is true but Statement II is false

4. Both Statement I and Statement II are false

**Ans.** (3)



D - Glucose + D fructose

$$\begin{aligned} [\alpha]_{D\text{-sucrose}} &= +66.5^\circ, [\alpha]_{D\text{-Glucose}} = +52.5^\circ \\ [\alpha]_{D\text{-Fructose}} &= -92.4^\circ \end{aligned}$$

$\Rightarrow$  Sucrose is dextrorotatory and hydrolysed product is laevorotatory.

- 61.** Given below are two statements:

Statement I: The halogen that makes longest bond with hydrogen in  $HX$ , has the smallest covalent radius in its group.

Statement II: A group 15 element's hydride  $EH_3$  has the lowest boiling point among corresponding hydrides of other group 15 elements. The maximum covalency of that element E is 4 .

In the light of the above statements, choose the correct answer from the options given below

1. Statement I is true but Statement II is false
2. Both Statement I and Statement II are false
3. Both Statement I and Statement II are true
4. Statement I is false but Statement II is true

**Ans.** (2)

**Sol.**  $HF < HCl < HBr < HI$  (bond length order)

$F < Cl < Br < I$  (radius order)

$PH_3 < AsH_3 < NH_3 < SbH_3 < BiH_3$  (Boiling point order)

Maximum possible covalency of phosphorous is 6

- 62.** The energy required by electrons, present in the first Bohr orbit of hydrogen atom to be excited to second Bohr orbit is \_\_\_\_  $Jmol^{-1}$ .

Given:  $R_H = 2.18 \times 10^{-11}$  ergs.

1.  $1.635 \times 10^{-18}$
2.  $9.835 \times 10^5$
3.  $1.635 \times 10^{-11}$
4.  $9.835 \times 10^{12}$

**Ans.** (2)

**Sol.**

$$E_n = -R_H \times \frac{Z^2}{n^2}$$

$$\begin{aligned} \Delta E &= 2.18 \times 10^{-11} \times 10^{-7} \times 1^2 \left[ \frac{1}{1^2} - \frac{1}{2^2} \right] \\ &= 1.635 \times 10^{-18} \text{ Joule/atom} \\ &= 1.635 \times 10^{-18} \times 6.02 \times 10^{23} \text{ Joule/mole} \\ &= 9.835 \times 10^5 \text{ Joule/mole} \end{aligned}$$

- 63.** Given below are two statements:

Statement I: Phenol on treatment with  $CHCl_3$ / aq. KOH under refluxing condition, followed by acidification produces *p*-hydroxy benzaldehyde as the major product and *o*-hydroxy benzaldehyde as the minor product.

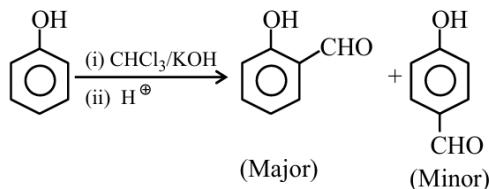
Statement II: The mixture of *p*-hydroxybenzaldehyde and *o* hydroxybenzaldehyde can be easily separated through steam distillation.

In the light of the above statements, choose the correct answer from the options given below

1. Statement I is true but Statement II is false
  2. Both Statement I and Statement II are false
  3. Both Statement I and Statement II are true
  4. Statement I is false but Statement II is true

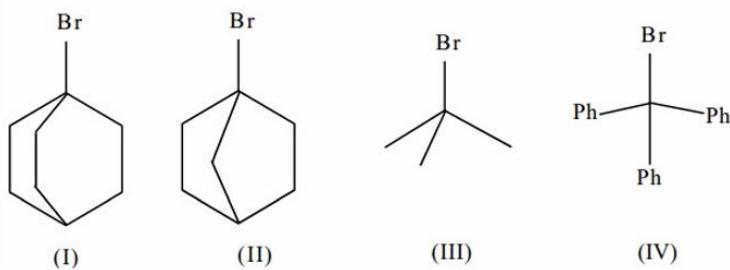
**Ans. (4)**

**Sol.**



(can be separated by steam distillation)

- 64.** The correct order of the rate of reaction of the following reactants with nucleophile by  $S_N1$  mechanism is :  
(Given : Structures I and II are rigid)



1. I < II < III < IV
  2. IV < III < II < I
  3. II < I < III < IV
  4. III < I < II < IV

**Ans. (3)**

**Sol.** Rate of  $S_N1$   $\propto$  Stability of  $C^+$  formed

(I) and (II) are unstable due to Bredt's rule, (I) has more + I effect.

(II) < (I) < (III) < (IV)

- 65.** A  $\rightarrow$  product (First order reaction).

Three sets of experiment were performed for a reaction under similar experimental conditions:

Run 1 ⇒ 100 mL of 10 M solution of reactant A

Run 2 ⇒ 200 mL of 10 M solution of reactant A

Run 3 ⇒ 100 mL of 10 M solution of reactant A + 100 mL of H<sub>2</sub>O added.

The correct variation of rate of reaction is

1. Run 1 < Run 2 < Run 3
  2. Run 3 < Run 1 = Run 2
  3. Run 1 = Run 2 = Run 3
  4. Run 3 < Run1 < Run2

**Ans. (2)**

**Sol.** For I<sup>st</sup> order reaction

$$\text{Rate} = k[A]$$

with decrease in concentration of A , rate of reaction decreases.

**66.** Consider the transition metal ions Mn<sup>3+</sup>, Cr<sup>3+</sup>, Fe<sup>3+</sup> and Co<sup>3+</sup> and all form low spin octahedral complexes. The correct decreasing order of unpaired electrons in their respective d-orbitals of the complexes is

1. Mn<sup>3+</sup> > Fe<sup>3+</sup> > Co<sup>3+</sup> > Cr<sup>3+</sup>
2. Cr<sup>3+</sup> > Fe<sup>3+</sup> > Co<sup>3+</sup> > Mn<sup>3+</sup>
3. Fe<sup>3+</sup> > Co<sup>3+</sup> > Mn<sup>3+</sup> > Cr<sup>3+</sup>
4. Cr<sup>3+</sup> > Mn<sup>3+</sup> > Fe<sup>3+</sup> > Co<sup>3+</sup>

**Ans.** (4)

**Sol.** Co<sup>3+</sup> → 3 d<sup>6</sup> ⇒ t<sub>2g</sub><sup>2,2,2</sup>e<sub>g</sub><sup>0,0</sup>, unpaired electron = 0

Fe<sup>3+</sup> → 3 d<sup>5</sup> ⇒ t<sub>2g</sub><sup>2,2,1</sup>e<sub>g</sub><sup>0,0</sup> unpaired electron = 1

Cr<sup>3+</sup> → 3 d<sup>3</sup> ⇒ t<sub>2g</sub><sup>1,1,1</sup>e<sub>g</sub><sup>0,0</sup> unpaired electron = 3

Mn<sup>3+</sup> → 3 d<sup>4</sup> ⇒ t<sub>2g</sub><sup>2,1,1</sup>e<sub>g</sub><sup>0,0</sup> unpaired electron = 2

**67.** Two p-block elements X and Y form fluorides of the type EF<sub>3</sub>. The fluoride compound XF<sub>3</sub> is a Lewis acid and YF<sub>3</sub> is a Lewis base. The hybridizations of the central atoms of XF<sub>3</sub> and YF<sub>3</sub> respectively are

1. sp<sup>2</sup> and sp<sup>3</sup>
2. Both sp<sup>3</sup>
3. Both sp<sup>2</sup>
4. sp<sup>3</sup> and sp<sup>2</sup>

**Ans.** (1)

**Sol.** XF<sub>3</sub> = BF<sub>3</sub>; sp<sup>2</sup>  
YF<sub>3</sub> = NF<sub>3</sub>; sp<sup>3</sup>

**68.** Match the LIST-I with LIST-II

List-I Thermodynamic Process		List-II Magnitude in kJ	
A.	Work done in reversible, isothermal expansion of 2 mol of ideal gas from 2dm <sup>3</sup> to 20dm <sup>3</sup> at 300 K .	I.	4
B.	Work done in irreversible isothermal expansion of 1 mol ideal gas from 1 m <sup>3</sup> to 3 m <sup>3</sup> at 300 K against a constant pressure of 3 kPa .	II.	11.5
C.	Change in internal energy for adiabatic expansion of a 1 mol ideal gas with change of temperature = 320 K and $\bar{C}_V = \frac{3}{2}R$ .	III.	6
D.	Change in enthalpy at constant pressure of 1 mol ideal gas with change of temperature = 337 K and $\bar{C}_P = \frac{5}{2}R$ .	IV.	7

Choose the correct answer from the options given below:

1. A-I, B-II, C-III, D-IV
2. A-III, B-II, C-IV, D-I

3. A-II, B-I, C-III, D-IV  
4. A-II, B-III, C-I, D-IV

**Ans. (4)**

**Sol.** Option (A)

$$W = -nRT \ln \frac{V_2}{V_1}$$

$$= \frac{-2 \times 8.314 \times 300}{1000} \times \ln \left( \frac{20}{2} \right) \text{ kJ}$$

$$= -11.5 \text{ kJ}$$

Option (B)

$$\begin{aligned}W &= -P_{\text{ext}} [V_2 - V_1] \\&= -3[3 - 1] \\&= -6 \text{ kJ}\end{aligned}$$

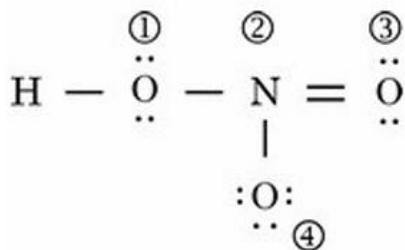
Option (C)

$$\begin{aligned}\Delta U &= nC_v\Delta T \\ &= 1 \times \frac{3}{2} \times \frac{8.314 \times 320}{1000} 100 \text{ kJ} \\ &= 3.99\end{aligned}$$

Option (D)

$$\begin{aligned}\Delta H &= nC_p\Delta T \\ &= 1 \times \frac{5}{2} \times \frac{8.314 \times 337}{1000} \text{ kJ} \\ &= 7 \text{ kJ}\end{aligned}$$

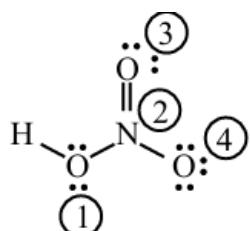
**69.** The formal charges on the atoms marked as (1) to (4) in the Lewis representation of  $\text{HNO}_3$  molecule respectively are



1. 0, 0, -1, +1
  2. +1,0,0, -1
  3. 0, -1,0, +1
  4. 0,+1,0, -1

**Ans. (4)**

**Sol.** Consider the structure of  $\text{HNO}_3$

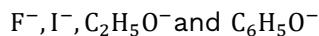


- 1: -0  
2: -(+1)  
3: -0  
4: -(-1)

Formal charge = valence e's - non bonding e's

$$-\left(\frac{\text{bonding electrons}}{2}\right)$$

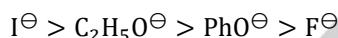
- 70.** The correct order of reactivity of  $\text{CH}_3\text{Br}$  in methanol with the following nucleophiles is



1.  $\text{I}^- > \text{C}_2\text{H}_5\text{O}^- > \text{F}^- > \text{C}_6\text{H}_5\text{O}^-$
2.  $\text{I}^- > \text{F}^- > \text{C}_6\text{H}_5\text{O}^- > \text{C}_2\text{H}_5\text{O}^-$
3.  $\text{I}^- > \text{C}_2\text{H}_5\text{O}^- > \text{C}_6\text{H}_5\text{O}^- > \text{F}^-$
4.  $\text{I}^- > \text{C}_6\text{H}_5\text{O}^- > \text{F}^- > \text{C}_2\text{H}_5\text{O}^-$

**Ans. (3)**

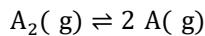
**Sol.** Order of nucleophilicity:



## SECTION-B

### NUMERICAL VALUE

- 71.** Dissociation of a gas  $\text{A}_2$  takes place according to the following chemical reaction. At equilibrium, the total pressure is 1 bar at 300 K .



The standard Gibbs energy of formation of the involved substances has been provided below:

Substance	$\Delta G_f^\circ/\text{kJmol}^{-1}$
$\text{A}_2$	-100.00
A	-50.832

The degree of dissociation of  $\text{A}_2(\text{g})$  is given by  $(x \times 10^{-2})^{1/2}$  where  $x = \underline{\hspace{2cm}}$ . (Nearest integer).

[Given:  $R = 8 \text{ J mol}^{-1} \text{ K}^{-1}$ ,  $\log 2 = 0.3010$ ,  $\log 3 = 0.48$  ]

Assume degree of dissociation is not negligible.

**Ans. (33)**

**Sol.**

$$-1.664 \times 10^3 = -8.3 \times 300 \ln K_p$$

$$\ln K_p = 0.693$$

$$K_p = 2$$

$$2 = \frac{4\alpha^2 P_0}{1-\alpha^2}$$

$$\alpha = \frac{1}{\sqrt{3}}$$

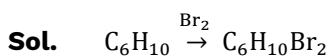
$$\begin{aligned} \alpha &= \left( \frac{100}{3} \times 10^{-2} \right)^{1/2} \\ &= (33.33 \times 10^{-2})^{1/2} \end{aligned}$$

- 72.** The cycloalkene (X) on bromination consumes one mole of bromine per mole of (X) and gives the product (Y) in which C: Br ratio is 3: 1. The percentage of bromine in the product (Y) is  $\underline{\hspace{2cm}}$  %. (Nearest

integer)

(Given : molar mass in gmol<sup>-1</sup>: H: 1, C: 12, O: 16, Br: 80 )

**Ans. (66)**



Molecular mass of  $C_6H_{10}Br_2$  is :

$$12 \times 6 + 10 + 160$$

$$72 + 10 + 160 = 242$$

$$\% \text{ of Br} = \frac{160}{242} \times 100$$

$$\% \text{ of Br} = 66.11\% \approx 66\%$$

- 73.** Sodium fusion extract of an organic compound (Y) with  $CHCl_3$  and chlorine water gives violet color to the  $CHCl_3$  layer. 0.15 g of (Y) gave 0.12 g of the silver halide precipitate in Carius method. Percentage of halogen in the compound (Y) is

\_\_\_\_\_. (Nearest integer)

(Given : molar mass gmol<sup>-1</sup>: C: 12, H: 1, Cl: 35.5, Br: 80, I: 127 )

**Ans. (43)**

**Sol.** Iodine gives violet colour

$$\begin{aligned}\% \text{ of I} &= \frac{\text{Atomic weight of I}}{\text{Molecular weight of AgI}} \times \frac{m}{W} \times 100 \\ &= \frac{127}{235} \times \frac{0.12}{0.15} \times 100 \\ &\% \text{ of I} = 43.23\% \approx 43\%\end{aligned}$$

- 74.** Consider the following electrochemical cell at 298 K



If the reaction quotient at a given time is  $10^6$ , then the cell EMF ( $E_{cell}$ ) is \_\_\_\_  $\times 10^{-1}$  V (Nearest integer).

Given the standard half-cell reduction potential as

$$E_{Bi_2O_3/Bi, OH^-}^\circ = -0.44 \text{ V} \text{ and } E_{Sn(OH)_6^{2-}/HSnO_2^-, OH^-}^\circ = -0.90 \text{ V}$$

**Ans. (4)**

**Sol.**  $E_{cell}^\circ = -0.44 - (-0.90)$   
 $= +0.46 \text{ V}$

Applying Nernst equation :-

$$E_{cell} = E_{cell}^\circ - \frac{0.06}{n} \log Q$$

$$E_{cell} = 0.46 - \frac{0.06}{6} \log 10^6$$

$$E_{cell} = 4 \times 10^{-1}$$

$$x = 4$$

- 75.** The temperature at which the rate constants of the given below two gaseous reactions become equal is \_\_\_\_ K. (Nearest integer)

$$X \rightarrow Y \quad k_1 = 10^6 e^{\frac{-30000}{T}}$$

$$P \rightarrow Q \quad k_2 = 10^4 e^{\frac{-24000}{T}}$$

Given:  $\ln 10 = 2.303$

**Ans. (1303)**

**Sol.**

$$10^4 e^{-\frac{24000}{T}} = 10^6 e^{-\frac{30000}{T}}$$

$$e^{\frac{6000}{T}} = 100$$

$$\frac{6000}{T} = 2 \ln 10$$

$$T = \frac{6000}{2 \times 2.303}$$

$$T = 1302.64 \text{ K}$$

$$T \approx 1303 \text{ K}$$

