Boyle's Law

Boyle's Law states that the volume of a gas varies inversely with its pressure if temperature is held constant. If one goes up, the other goes down. We use the formula:

$$P_1 \times V_1 = P_2 \times V_2$$

Solve the following problems (assuming constant temperature.)

1. A sample of oxygen gas occupies a volume of 250mL at 740 torr pressure. What volume will it occupy at 800 torr pressure?

2. A sample of carbon dioxide occupies a volume of 3.50 Liters