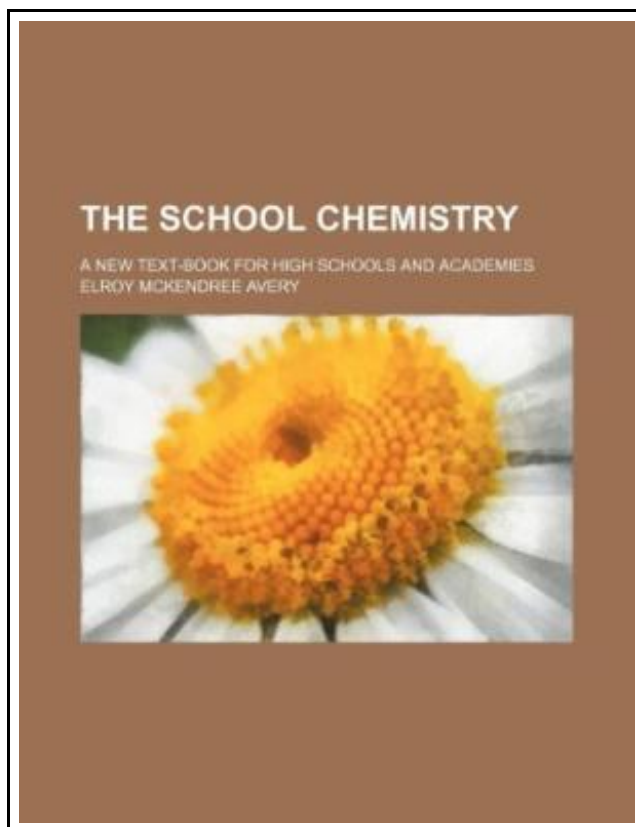


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(Ike Fadel)*

THE SCHOOL CHEMISTRY; A NEW TEXT-BOOK FOR HIGH SCHOOLS AND ACADEMIES



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Rarebookclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1904 Excerpt: .for transferring the energy of the sun to the living animal. The energy of sunlight expended in breaking up the carbon oxide that is absorbed by the leaf is stored up as potential energy in the vegetable tissue, and again set free as heat in the animal body. The sun s energy that is transformed into the chemical energy of carbon compounds that become parts of vegetable tissue is the physical source and foundation of all life. The plant breathes in carbon dioxide; the animal breathes out carbon dioxide. Each carbon particle involved may go through the same cycle of changes again and again, carrying with it on each round the energy necessary for the maintenance of vegetable and animal life. EXERCISES. 1. Assuming that a square meter of leaf will decompose a liter of carbon dioxide per hour, what weight of carbon will be assimilated in an hour by 1,000 trees, each of which has 100,000 leaves, each leaf measuring 25 sq. cm.? 2. Is $\text{Hg}(\text{CN})_2$ a mercurous or a mercuric compound? What is its name? 3. Describe the usual method of preparing chlorine, and write the reaction. Find the percentage of chlorine in the substance that furnishes it. 4. Zinc nitrate and potassium carbonat react as follows: $\text{Zn}(\text{NO}_3)_2 + \text{K}_2\text{CO}_3 = \text{ZnCO}_3 + 2 \text{KNO}_3$. How much $\text{Zn}(\text{NO}_3)_2$ is required to give 103.17 grams of ZnCO_3 ? 5. How much ZnCO_3 may be obtained from 156 grams of 6. How much K_2CO_3 is needed...

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