bookdown+: Authoring Articles, Mails, Guitar Chords, Chemical Molecular Formulae and Equations with R bookdown

Peng Zhao 2017-05-09

Contents

1	1 Introduction			
2	Demos			
	2.1	Chemical formulae	7	
	2.2	Chemical equations	7	
	2.3	Structural formulae	8	

4 CONTENTS

Chapter 1

Introduction

Chapter 2

Demos

2.1 Chemical formulae

```
\begin{array}{l} {\rm H_2O,\,Sb_2O_3} \\ {\rm H^+,\,CrO_4^{2^-},\,[AgCl_2]^-,\,Y^99^+,\,Y^{99+}} \\ {\rm Fe^{II}Fe^{III}2O_4,\,2\,H_2O,\,2\,H_2O,\,0.5\,H_2O,\,\frac{1}{2}\,H_2O,\,(_{1/2})H_2O\,\,,\,n\,H_2O} \\ {}^{227}{\rm Th^+,\,^227\,_90\,Th^+,\,_1^0n^-,\,_1^01\,n^-} \\ {\rm H^3HO,\,H^3HO} \\ {\rm (NH_4)_2S,\,[\{(X_2)_3\}_2]^{3+}} \\ {\rm H_2(aq),\,CO_3^2-_{(aq)},\,NaOH(aq,\infty)} \\ {\rm OCO^-,\,NO^{(2\cdot)-}} \\ {\rm NO_x,\,\&ffamily\&bfseries\&ce\{NO_x\},\,Fe^{n+},\,\&ffamily\&bfseries\&ce\{Fe^n+\}\&ce\{\&mu-Cl\},\,\,\&ce\{[Pt(\&ta^2-C2H4)Cl3]-\} \\ {\rm KCr(SO_4)_2\cdot12\,H_2O,\,KCr(SO_4)_{2.12}H_2O,\,KCr(SO_4)_2*12\,H_2O} \\ {\rm C_6H_5-CHO,\,A-B=C\equiv D,\,\&ffamily\&bfseries\&ce\{A-B=C\#D\}} \\ {\rm A-B=C\equiv D,\,A-B-C,\,A=B=C\equiv D,\,A\cdots B\cdots C,\,A\rightarrow B\leftarrow C} \end{array}
```

2.2 Chemical equations

 $\mathbf{A} \longrightarrow \mathbf{B}$

 $\mathbf{A} \longleftarrow \mathbf{B}$

 $\mathbf{A} \longleftrightarrow \mathbf{B}$

$$\mathbf{A} \longleftarrow -> \mathbf{B}$$

$$A \rightleftharpoons B$$

$$A \stackrel{\rightharpoonup}{\longleftarrow} B$$

$$\mathbf{A} \xrightarrow{\mathbf{H}_2\mathbf{O}} \mathbf{B}$$

$$A \xrightarrow{textabove} B$$

$$A \xrightarrow{x} B$$

$$A \xrightarrow{x} B$$

$$A + B$$

$$A-B$$

$$A = B$$

$$A \pm B$$

$$SO_4^{2-} + Ba^{2+} \longrightarrow BaSO_4 \downarrow$$

$$A \downarrow B \downarrow \longrightarrow B \uparrow B \uparrow$$

$$CH_4 + 2 \left(O_2 + \frac{79}{2} N_2\right)$$

$$xNa(NH_4)HPO_4 \xrightarrow{\Delta} (NaPO_3)_x + xNH_3 \uparrow + xH_2O$$

$$\mathrm{CO}_2 + \mathrm{C} \longrightarrow 2\,\mathrm{CO}$$

$$\mathrm{Hg}^{2+} \xrightarrow{I-} \mathrm{HgI}_2 \xrightarrow{I-} [\mathrm{Hg}^{\mathrm{II}}\mathrm{I}_4]^{2-}$$

$$\operatorname{Zn}^{2+} \xrightarrow{+2\operatorname{OH}^{-}} \operatorname{Zn}(\operatorname{OH})_{2} \downarrow \xrightarrow{+2\operatorname{OH}^{-}} \left[\operatorname{Zn}(\operatorname{OH})_{4}\right]^{2-}$$

$$K = \frac{[{\rm Hg}^{2+}][{\rm Hg}]}{[{\rm Hg}_2^{2+}]}$$

$$K = \frac{[Hg^2 +][Hg]}{[Hg^2 +]}$$

2.3 Structural formulae

$$CH_3$$
 Cl
 O

$$\begin{array}{c} \text{CH}_3 \\ \text{H}_3\text{C} \\ \text{H} \end{array}$$

$$O$$
 CH_3
 CH_3
 CH_3

Bibliography