
EXERCÍCIOS – EQUAÇÃO DE NERST

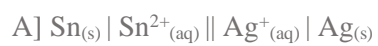
1] Calcule a ddp da pilha a partir da concentração fornecida:

A] $\text{Cu}^{2+}_{(\text{aq})}$ [0,1M] | $\text{Zn}^{2+}_{(\text{aq})}$ [1,0M] | $E^0_{\text{Red}} \text{Cu}^{2+} = +0,64\text{v}$ | $E^0_{\text{Red}} \text{Zn}^{2+} = -0,76\text{v}$



B] $\text{Cu}^{2+}_{(\text{aq})}$ [0,5M] | $\text{Zn}^{2+}_{(\text{aq})}$ [0,01M] | $E^0_{\text{Red}} \text{Cu}^{2+} = +0,64\text{v}$ | $E^0_{\text{Red}} \text{Zn}^{2+} = -0,76\text{v}$

2] Calcule a tensão produzida a 25°C pela célula:



$$[\text{Sn}^{2+}_{(aq)}] = 0,15 \text{ mol/L}$$

$$[\text{Ag}^{+}_{(aq)}] = 1,7 \text{ mol/L}$$

$$E^0_{\text{Red}} \text{Sn}^{2+}_{(aq)} = -0,14 \text{ V}$$

$$E^0_{\text{Red}} \text{Ag}^{+}_{(aq)} = +0,80 \text{ V}$$



$$E^0_{\text{Red}} \text{Cr}^{3+}_{(aq)} = -0,74 \text{ V}$$

$$E^0_{\text{Red}} \text{Pb}^{2+}_{(aq)} = -0,13 \text{ V}$$
