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Pearson Chemistry Chapter 8. In covalent bonding, electrons sharing usually occurs so that atoms attain the electron configurations of nobele gases. Combinations of atoms of the nonmetals and metalloids in groups 4A,5A, 6Aand 7A of the periodic tale are likely to forms covalent bonds. The combined atoms usually acquire a total of eight electrons or an octect.

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Single Covalent Bonds. When only one pair of electrons is shared, such as in a hydrogen molecule, it is a single covalent bond. The shared electron pair is often referred to as the bonding pair. For a hydrogen molecule, shown in Figure 8.4, each covalently bonded atom equally attracts the pair of shared electrons.

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•covalent bond: a bond formed by the sharing of electrons between atoms •molecule: a neutral group of atoms joined by covalent bonds •diatomic molecule: a molecule consisting of two atoms Glossary Terms

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Covalent Bonds (pages 165–167) 1. Define a covalent bond. 2. A molecule is a group of atoms that are joined together by one or more covalent bonds. Circle the correct answer. negative neutral positive 3. Is the following sentence true or false? In a covalent bond, the atoms are held together by the attractions between the shared electrons and the

Chapter 6 Chemical Bonds Section 6.2 Covalent Bonding

overlapping when a covalent bond forms. a.electron dotb.structural formula c.space-fillingd.electron cloud 3. Describe a molecule. 4. Is the following sentence true or false? In a covalent bond, the atoms are held together by the attractions between the shared electrons and the protons in each nucleus. 5. Circle the correct answer.

Section 6.2 Covalent Bonding - Quia

Chapter 8 Covalent Bonding and Molecular Structure 8-3 There are two types of repulsive forces between the two atoms. First, the nuclei repel because they are both positively charged. Second, the electrons repel because they are both negatively charged. The attractive forces between the two atoms result from the

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