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Solutions to exercises 1A.1(b) The perfect gas law [1A.5] is $pV = nRT$, implying that the pressure would be $nRT/p = V$. All quantities on the right are given to us except n , which can be computed from the given mass of Ar. $25 \text{ g} / 39.95 \text{ g mol}^{-1} = 0.626 \text{ mol}$. $(0.626 \text{ mol}) \times (8.31 \times 10^{-2} \text{ dm}^3 \text{ bar K}^{-1} \text{ mol}^{-1}) \times (30 + 273) \text{ K} = 10.5 \text{ bar dm}^3$. So no, the sample would not exert a pressure of 2.0 bar.

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