

### Set 3. Significant Figures and Calculations

**Skill 3.01: Be able to round the result of a subtraction/addition calculation**

**Skill 3.02: Be able to round the result of a multiplication/division calculation**

**Skill 3.04: Be able to round the result of a calculation involving exact numbers**

**Skill 3.05: Be able to round the result of a chain calculation**

**Skill 3.01: Be able to round the result of a subtraction/addition calculation**

#### Skill 3.01 Concepts

In addition and subtraction, the number of significant figures to the right of the decimal point in the final sum or difference is determined by the smallest number of significant figures to the right of the decimal point in any of the original numbers. Consider the following example:

$89.332 \leftarrow$  three significant figures after the decimal point  
 $+ 1.1 \leftarrow$  one significant figure after the decimal point  
 $90.432 \leftarrow$  round off to 90.4

Because there is only one significant figure to the right of the decimal in the second number you should round to this digit in the final result.

#### Skill 3.01 Example 1

Carry out the following arithmetic operations and round the result to the correct significant figures:

(a)  $11, 254.1 \text{ g} + 0.1983 \text{ g}$

(b)  $66.59 \text{ L} - 3.113 \text{ L}$

(c)  $123 \text{ mL} + 1.006 \text{ mL} - 2.6 \text{ mL}$

#### Skill 3.01 Exercise 1

**Skill 3.02: Be able to round the result of a multiplication/division calculation**

#### Skill 3.02 Concepts

In multiplication and division, the number of significant figures in the final product or quotient is determined by the original number that has the smallest number of significant figures. The following example is illustrative:

$2.8 \times 4.5039 = 12.61092 \leftarrow$  round off to 13 because 2.8 only has 2 significant figures

#### Skill 3.02 Example 1

Do the following calculations and round to the correct significant figures:

(a)  $1222\text{g}/2\text{g}$

(b)  $60.50 \text{ ft} \times 2.0000 \text{ ft}$

(c)  $1000. \text{ in}/5.0 \text{ in}$

#### Skill 3.02 Exercise 1

### Skill 3.03: Be able to round the result of a calculation involving exact numbers

#### Skill 3.03 Concepts

An exact number is obtained from a definition or by counting numbers of objects and can be considered to have an infinite number of significant figures.

#### Skill 3.03 Example 1

Which of the following are examples of exact numbers?

- |                           |  |                                  |
|---------------------------|--|----------------------------------|
| (a) There are 24 skittles | (b) mount rainier is 14, 442 feet high | (c) there are 2.54 cm in an inch |
|---------------------------|--|----------------------------------|

Because exact numbers have an infinite number of significant figures you never round to them. For example if an object has a mass of 0.2786 g, then the mass of eight such objects is:

$$0.2786 \text{ g} \times 8 = 2.339 \text{ g}$$

Notice, we DID NOT round to the product to one significant figure, because the number 8 is 8.00000..., by definition.

#### Skill 3.03 Example 2

One skittle is found to weigh 1.0 g. How much does 500 skittles weigh to the correct number of significant figures?

#### Skill 3.03 Exercise 1

### Skill 3.04: Be able to round the result of a chain calculation

#### Skill 3.04 Concepts

A chain calculation is a calculation involving more than one step. The rule we will apply in this class is to always round to the measurement with the fewest significant figures. Consider the following example:

Suppose a student collected the following data and was asked to find the density of the object.

Volume of water in graduated cylinder: 20.1 mL

Volume of water in graduated cylinder with submerged object: 25.9 mL

Mass of object: 5.75 g

Step 1: calculate the volume of the object:  $25.9 \text{ mL} - 20.1 \text{ mL} = 5.8 \text{ mL}$

Step 2: calculate the density (mass/volume):  $5.75 \text{ g} / 5.8 \text{ mL} = 0.9918 \text{ g/mL}$

Step 3: round the result to the appropriate significant figures: 0.992 g/mL

**\*\*Go back to the data, the worst measurement contains 3 significant figures, therefore, that is all that is allowed in the final reported value.**

**Skill 3.04 Example 1**

A student is to determine the density of an object using the following data:

Length of cube: 55. cm

Mass of cube: 0.201 g

Use the equations below to calculate the density of the object, then round to the appropriate significant figures

Volume of a cube = (Length)<sup>3</sup>

$$\text{Density} = \frac{\text{mass}}{\text{volume}}$$

**Skill 3.04 Example 2**

A student is to determine the density of an object using the following data:

Length of the cube: 11 cm

Mass of the cube: 950. g

**Skill 3.04 Exercise 1**