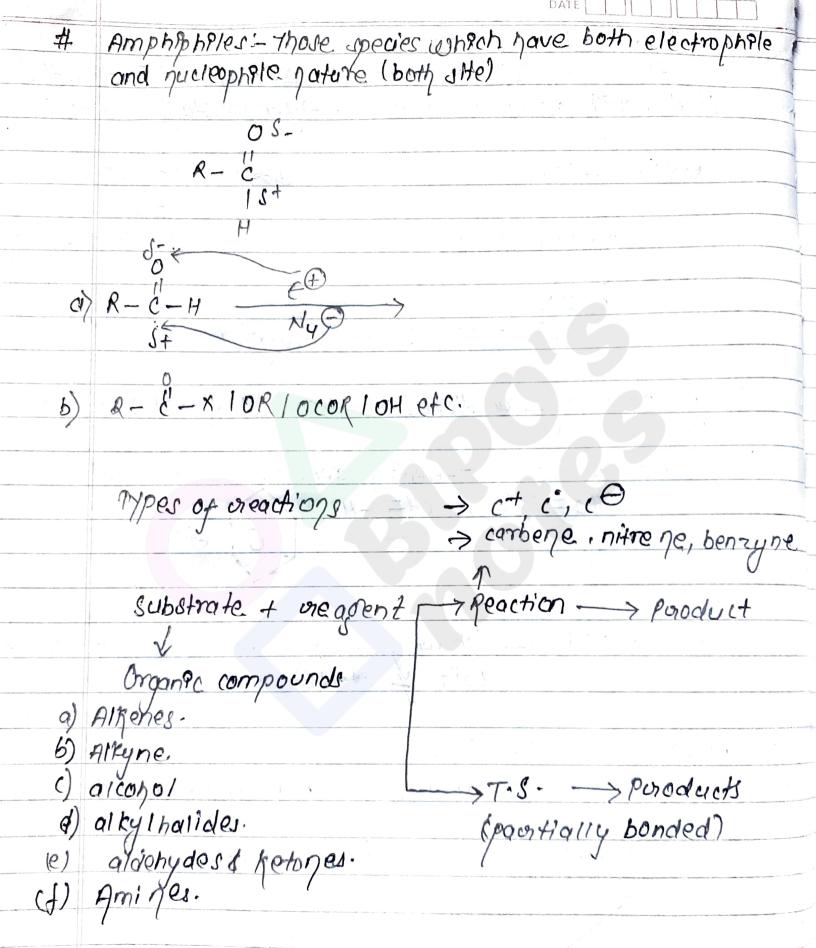
	DATE
	Haloalkanes and Haloanenes
	General 9deas
	$S + R \longrightarrow Pajoduct(P)$
	A Types of R
	Substrate Reagent
	Types of reagent
	\parallel Plectrophilic \sim propert(electrophile) \rightarrow
	electrophiles are electron loving species.
2	electrophiles are electron dipicient species and atto
	electrophiles are electron loving species. electrophiles are electron dipicient species and atto ch at the electron rich centre.
<i>a</i>)	All positively charged species are electrophile
	eg:- CH3, NO2, -CI, BY, H
	(A) (B)
1	Except:- NIHy, H30
(b)	compounds having Promplete octet
ler.	CHg, Alciz, Bfz etc.
	CHg, AICIZ, BF3 etc.
7	07 > 00 > 01 > 00
	BI37BBr3>BC1, 7BF3 Back bonding
}	Die il about 100 la la character
	electrophPhelic character increase
	0.1 0.1
(2)	
X	$SiF_{4} + 2F \longrightarrow SiF_{4}, -2$ $PCI_{-} + CI \longrightarrow PCI_{-}$
	PAGE
	classnate

<i>y</i>		DATE
- -	/	Made: All polansing functional group one electrophile or nu
		Note: All polarising functional group one electrophile or nu leophile.
	α	$\begin{array}{cccccccccccccccccccccccccccccccccccc$
		$\begin{array}{cccccccccccccccccccccccccccccccccccc$
		At 1°
	6	$CH_3 - C \equiv N$ $S = \sqrt{4}$
		$\begin{array}{cccccccccccccccccccccccccccccccccccc$
		$\int \int $
		and the second of the second o
	14)	Note: All lewes aceds are electrophile but all electrophis
		e are not lewis acid
		foreg:-
		toreg:- CHz 98 electrophPle but not lewis and.
	-	- Stranschoug the printed Langerth (Light Algorithm)
	<i>Q</i> .	Nucleo pholes ->
	4	e ofth species and attack at the electron defi
		clent centre
	<i>a</i> >	All -vely charged species are Nucleaphiles.
/		early in the second which is a second of
		H, CHg, CI, NO2, Br, CN-etc
	b)	compounds having lone pair of electrons are ju
		cleophiles: eg:-
		H20, NH3, PH3, PC13, RNH2
		R-OH etc
	<u>c)</u>	All organometallic compounds are nucleophile
		Classmate

	Organometalle compounds:-	4 / / · . · ·	in the same	
	U The compound	1 4n which.	motal %	directly
	bonded with carbon are known	as organom	netall9c c	ompound.
	S- S+			
	e.9:- R-Mgx	Bas.	· · · · · · · · · · · · · · · · · · ·	7+
	e.g:- R-Mgx		R2 CL	1 68
	Granord creagent		99111	mann's weg
	Granord creagent			gent
	RZN		1 1	4 %
	R2ZAI R2Zn -> frand lan	nd uneagent		
		0	7	
4)	All polanising kingtingal am	up one plan	trophPle	and nucle
	All polanising functional gro	7		-70
2	Compound & hondra & dontra	density and	nucleon	h9/e.
()	Compounds having n-electron (H2=CH2, CH=C-	H	1	
			etc.	
			palan, unuga esalleren - viilinenenule utanilisea	
	and the state of t		1 y 1	t ett planskallen flemetijden, met han in 2000 hannet in normaler in dieser til til
	Nucleanhalanti. Pondana.	1 Dunlanh	plo to de	inate. Ina
	Nucleoph?/Pc?ty:- Pendency of ex	Jadamo	,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,	101)6
	par of e	I CO OLY		
7.	Ambedant and the			
٥,	Ambadent jucleophile ->	011		
	NucleophPles w	more the	an one di	eviar site:
		\rightarrow .		
	C = N O - N	l=0 nitni	40	
	y node		. 9	



(1)	Addition opeachions:>
->	T-bond broken -> T-bond formed.
	\mathcal{H}
•	weaken bond Combon bond Stronger bond
—)	10=0 +112
	Coopera bood Changer panal
	weaken bond
->	Those reaction in which we add the reagent by breaking the T-bonds and forming new T-bonds.
	the T-bonds and tomong new T-bonds.
	Addition oreactions
	Constitution of the second of
	EAR (electro philic addition reaction)
	eg:-
	Alkene e Addition
	Electrical Control of the Control of
7	NAR (-Nucleophilic addition reaction)
	OH OH
	$CH_3 - C - H = \frac{CH_3 - C - H}{V \cdot S \cdot S} \rightarrow CH_3 - C - H \rightarrow CH_3 - C - H$
	CH3-C-H - 7013-CH3 - 1
	(S+) (N)
	complete octet
	Radical addition reaction (RAR) -
	AH = CH H-Br CH2 - CH2
	H202 1 1 11 0 m
-	Radical addition reaction (RAR) \rightarrow $(H_2 = CH_2, H_2O_2, H_$
	CH3-CH2-BY
	classmate PAGE []

	substitution seaction - when we replace an atom or group
	substitution seaction - when we replace an atom or group with other by breaking and toming of a 5 bond.
	NSR (SNI) RXN; MucleophPle substitution openchion.
/	eg:-
	C CO AN
	$R \rightarrow X + NY \xrightarrow{SN} R - NY + R$
	OXI)
=-,	
	-> ESR or se mn: - electrophplac substitution reaction.
	eg:-
	E
	$\begin{array}{c} F \\ \hline \\$
	$\begin{array}{c} \downarrow \\ \uparrow \\ \uparrow \\$
	-> RSR or ESER TXN -> RadPool Substitution Ynx con-
	eg:-Hydrogenation of alkene
	J. J
(iii)	elimprotion geoction: 5-bond > TI-bond
	break formed
-	OH OH
	g = H
	CH2 - CH2 910-KOH CH2 = CH2 + H20 + BT
-	
	BY
-	
	when we elimphote some group or atoms form molecule the reaction 9s called elimphon reaction.
	THE reaction 43 called letter that beaches,
-	
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	placemet a

(IV) Oxidation and reduction reaction
$$\Rightarrow$$
 Addition of H or removal of O .

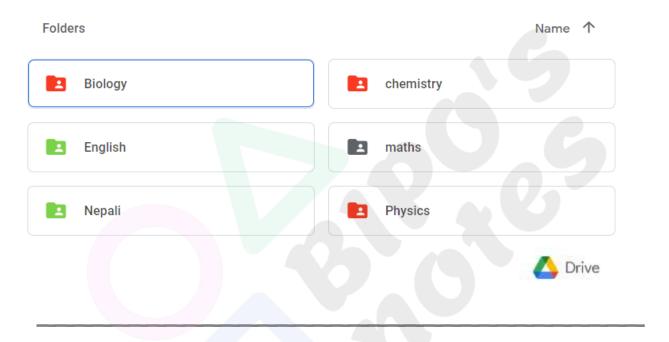
$$R - CH_2 - OH \xrightarrow{EOJ} R - C - H \xrightarrow{EOJ} R - C - OH$$

$$H + H + H$$

Bipin Khatri

(Bipo)

Class 12 complete notes and paper collection.



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