

## Density

We can compare the weights of things made of different materials.

For a fair test, we weigh objects made of different materials which have the \_\_\_\_\_.

1 cm<sup>3</sup> of iron is heavier than 1 cm<sup>3</sup> of plastic.

This is because the \_\_\_\_\_ has more \_\_\_\_\_ even though the \_\_\_\_\_ is the same.

Iron is more \_\_\_\_\_ than plastic.

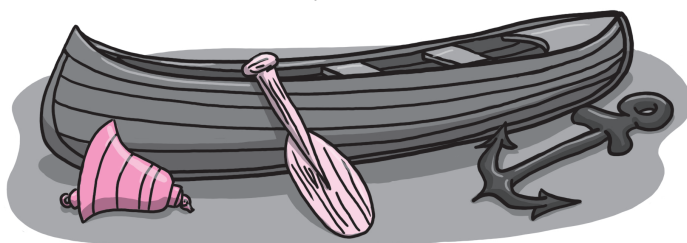
Glass marbles \_\_\_\_\_ in water. This happens because glass is \_\_\_\_\_ dense than water.

A cork \_\_\_\_\_ in water. This happens because cork is \_\_\_\_\_ than water.

- 1 "Iron is heavier than wood." This is not always true. Give an example of something made of wood, and something made of iron where the wooden object is heavier.

(a) Object made of wood:

(b) Object made of iron:



- 2 Rewrite "Iron is heavier than wood" to make it scientifically correct.

(a) Rewrite it using the word **dense**.

(b) Rewrite it without using the words **dense** or **density**. Explain what you meant in (a).

- 3 Here is a list of materials. Number them in order of density. Put 1 against the least dense material, and 5 against the most dense.

air

wood

water

helium

steel

**Density** tells us the mass of \_\_\_\_\_ of material. It enables materials to be compared.

- 4 A 100 cm<sup>3</sup> block of iron has a mass of 790 g.

(a) Mass of 1 cm<sup>3</sup> of iron =  ÷  =  grams

(b) Complete the sentence: The density of iron (in g/cm<sup>3</sup>) is .

(c) 1000 g of salt has a volume of 500 cm<sup>3</sup>. Work out its density using an equation.

$$\begin{array}{rcccl} \text{mass (g)} & = & \text{density (g/cm}^3\text{)} & \times & \text{volume (cm}^3\text{)} \\ \hline \text{_____} & = & \text{_____} & \times & \text{500} \end{array}$$

(d) Work out the density of aluminium if 540 g has a volume of  $200 \text{ cm}^3$ .

(e) 8 kg of rice has a volume of  $10\,000 \text{ cm}^3$ . Work out the density in  $\text{g/cm}^3$ . ( $1 \text{ kg} = 1000 \text{ g}$ )

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5 Calculate the density of these materials in  $\text{g/cm}^3$ .

(a) Uranium:  $20 \text{ cm}^3$  has a mass of 380 g.

(b) Lead:  $60 \text{ cm}^3$  has a mass of 660 g.

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6 The volumes of liquids are measured in  $\text{m}\ell$  (millilitres).  $1 \text{ m}\ell = \underline{\hspace{1cm}} \text{ cm}^3$ .  
Calculate the density in  $\text{g/cm}^3$  of

(a) Olive oil, if  $750 \text{ m}\ell$  has a mass of 675 g,

(b) Water, if  $350 \text{ m}\ell$  has a mass of 350 g.

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7 Silver has a density of  $10 \text{ g/cm}^3$ .

(a) Complete the sentence: The mass of  $1 \text{ cm}^3$  of silver is  grams.

(b) Work out the mass of  $15 \text{ cm}^3$  of silver using an equation.

$$\begin{array}{ccccc} \text{mass (g)} & = & \text{density (g/cm}^3\text{)} & \times & \text{volume (cm}^3\text{)} \\ \hline \text{ } & = & 10 & \times & \text{ } \end{array}$$

(c) Work out the mass of  $45 \text{ cm}^3$  of silver using an equation.

$$\begin{array}{ccccc} \text{mass (g)} & = & \text{density (g/cm}^3\text{)} & \times & \text{volume (cm}^3\text{)} \\ \hline \text{ } & = & 10 & \times & \text{ } \end{array}$$

(d) Work out the mass of  $100 \text{ cm}^3$  of silver.

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8 Calculate the mass in grams of

(a)  $200 \text{ cm}^3$  of flour with a density of  $0.8 \text{ g/cm}^3$ ,

(b)  $60 \text{ cm}^3$  of cheese with a density of  $1.1 \text{ g/cm}^3$ .

9 Jelly has a density of  $1.5 \text{ g/cm}^3$ .

(a) What is the mass of  $1 \text{ cm}^3$  of jelly?

(b) If you divide 180 g of jelly into 1.5 g pieces, how many pieces would you have?

(c) What is the volume of 180 g of jelly? Count the  $1 \text{ cm}^3$  (1.5 g) pieces.

(d) Work out the volume of 800 g of salt using an equation.

$$\begin{array}{ccccc} \text{mass (g)} & = & \text{density (g/cm}^3\text{)} & \times & \text{volume (cm}^3\text{)} \\ \boxed{\quad\quad\quad} & = & \boxed{2} & \times & \boxed{\quad\quad\quad} \end{array}$$

(e) Work out the volume of 1600 g of rice using an equation.

$$\begin{array}{ccccc} \text{mass (g)} & = & \text{density (g/cm}^3\text{)} & \times & \text{volume (cm}^3\text{)} \\ \boxed{\quad\quad\quad} & = & \boxed{0.8} & \times & \boxed{\quad\quad\quad} \end{array}$$

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10 Calculate the volume in  $\text{cm}^3$  of

(a) 39 g of chocolate with a density of  $1.3 \text{ g/cm}^3$ ,

(b) 112 g of treacle with a density of  $1.4 \text{ g/cm}^3$ .

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11 Complete the word equations using **density**, **mass** and **volume**.

(a) density =

(b) mass =

(c) volume =

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12 Rewrite your word equations using symbols.

$\rho$  (rho) is the density,  $m$  is the mass and  $V$  is the volume.

(a)  $\rho =$

(b)  $m =$

(c)  $V =$

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13 A gold bar is a  $15 \text{ cm} \times 8 \text{ cm} \times 6 \text{ cm}$  rectangular block. The density of gold is  $19 \text{ g/cm}^3$ .

(a) Calculate the volume in  $\text{cm}^3$ .

(b) Calculate the mass in grams.

(c) If each gram of gold is worth £50, calculate the cost of the bar.

