

- Q.12 Write down the electronic configuration of carbon and sodium. (Atomic No-11) (Atomic No-6)
- Q.13 Enlist the merits of Gaseous fuels over solid and liquid fuels.
- Q.14 Describe the types of hardness of water . Write down the causes of hardness of water.
- Q.15 Write down two properties of S-block elements and two properties of p-block elements.
- Q.16 Write down any five characteristics of a good fuel.
- Q.17 Differentiate between thermosetting. Plastic and Thermoplastic with examples.
- Q.18 Write down the composition of producer gas and three uses of it.

SECTION-C

Note: Long answer type questions. Attempt any one questions out of two questions. $1 \times 10 = 10$

- Q.19 i) Define Corrosion. Explain any one method of prevention of corrosion.
- ii) Define an Ionic bond. Explain the formation of sodium chloride (NaCl) Atomic number of Sodium is 11 and Chloride is 17.
- Q.20 Write short notes on the following:
- i) Isotopes and Isobars
- ii) Main features of Rutherford Model of Atom.

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**Level 3 / 1st. Sem. / DVOC (Ref. & Air Cond., Medical Imaging Tech., Auto. Servicing, ITM, PT, SD, AMT, FP, EMS)
Subject : Applied Chemistry**

Time : 2 Hrs.

M.M. : 50

SECTION-A

Note: Very short answer type questions. All questions are compulsory $(10 \times 1 = 10)$

- Q.1 Cathode Rays carry _____
- Q.2 Azimuthal Quantum number is devoted by _____
- Q.3 Electrons are present _____ (inside/outside) nucleus.
- Q.4 Monomer of polyethene is _____
- Q.5 Full form of LPG _____
- Q.6 Rusting of Iron is more in saline water than pure water. (True/False)
- Q.7 pH of basic solution is _____
- Q.8 Coating of iron with tin by hot dipping is called _____
- Q.9 Define Neutron.
- Q.10 An example of solid fuel _____

SECTION-B

Note: Short answer type questions. Attempt any six questions out of eight questions. $6 \times 5 = 30$

- Q.11 Define Quantum numbers. Name and give the significance of each Quantum number.