

- Q.31 What is covalent bonding? Give example. (CO-2)
 Q.32 What is the difference between strong and weak electrolytes? (CO-4)
 Q.33 What are the characteristics of drinking water? (CO-3)
 Q.34 Describe in brief the cleansing action of soap. (CO-6)
 Q.35 What is electroplating? Write its any two applications. (CO-5)

SECTION-D

- Note:** Long answer type questions. Attempt any two questions out of three questions. (2x10=20)
 Q.36 (a) Write a short note on acid base titration. (CO-5)
 (b) How colloids are classified? Give examples. (CO-6)
 Q.37 Define PH of a solution. What is a PH scale? Write the applications of pH. (CO-4)
 Q.38 Explain the formation of ionic compounds with the help of an example. Write the characteristics of ionic compounds. (CO-2)

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1st Sem. Branch: DMLT
Sub : Basic Chemistry

Time : 3Hrs.

M.M. : 100

SECTION-A

Note: Multiple choice questions. All questions are compulsory (10x1=10)

- .Q.1 The symbol of Copper is (CO-1)
 a) N b) Fe
 c) Cu d) Be
 Q.2 The SI unit of Pressure is- (CO-1)
 a) Mole b) Pascal
 c) Litre d) Kelvin
 Q.3 The charge on an electron is (CO-2)
 a) Positive b) Negative
 c) Neutral d) None of the above
 Q.4 What is the molecular formula of sodium chloride? (CO-4)
 a) KOH b) NaOH
 c) HCl d) NaCl
 Q.5 The example of conductors are- (CO-5)
 a) Silver b) Copper
 c) Iron d) All of the above
 Q.6 When an atom gain an electron it forms- (CO-1)
 a) Cation b) Anion
 c) Mole d) All of the above

- Q.7 The pure water should be- (CO-3)
 a) Tasteless b) Free from ions
 c) Colourless d) All of the above
- Q.8 Which of the following is a s-block element (CO-2)
 a) Hydrogen b) Nitrogen
 c) Carbon d) Oxygen
- Q.9 Which of the following is a colloid? (CO-6)
 a) Gel b) Smog
 c) Butter d) All of the above
- Q.10 What is the maximum number of electrons in an orbital? (CO-2)
 a) 2 b) 4
 c) 6 d) 8

SECTION-B

- Note:** Objective type questions. All questions are compulsory. (10x1=10)
- Q.11 The SI unit of time is _____ (CO-1)
- Q.12 A _____ is the mixture of solute and solvent. (CO-1)
- Q.13 Covalent bond is formed by _____ (sharing/transfer) of electrons. (CO-2)
- Q.14 (Soft/Hard) _____ water forms leather with soap solution. (CO-3)
- Q.15 What is unit of Normality? (CO-4)
- Q.16 Adsorption is a _____ (bulk/surface) phenomenon. (CO-6)
- Q.17 Give an example of a strong electrolyte. (CO-5)

- Q.18 The temporary hardness can be removed by boiling. (True/False) (CO-3)
- Q.19 Give an example of a natural buffer. (CO-4)
- Q.20 The mixture of oil in water forms _____ (emulsion/solution) (CO-6)

SECTION-C

- Note:** Short answer type questions. Attempt any twelve questions out of fifteen questions. (12x5=60)
- Q.21 Write any four disadvantages of hard water. (CO-3)
- Q.22 What are the molecular formula of Silver chloride and sodium hydroxide. (CO-1)
- Q.23 What are quantum numbers? (CO-2)
- Q.24 What is the difference between temporary and permanent hardness of water? (CO-3)
- Q.25 What is ionization? How does NaCl ionizes in water to make aqueous, solution? (CO-4)
- Q.26 Calculate the molecular mass of NaOH. The atomic weight of (H=1, Na=23, and O=16) (CO-1)
- Q.27 Write any two differences between Lyophilic and lyophobic colloids. (CO-6)
- Q.28 Define Molarity of a solution. Write its unit and formula. How does molarity varies with temperature. (CO-1)
- Q.29 What are the applications of electrolysis? (CO-5)
- Q.30 What is the difference between compound and mixture? (CO-1)