

- Q.6 Which is not a good conductor
- Aqueous solution of NaCl
 - fused NaCl
 - Solid NaCl
 - Silver
- Q.7 Charge on a proton is:-
- Positive
 - Negative
 - No charge
 - Any of above
- Q.8 The symbol of iron is _____.
- N
 - Cu
 - Fe
 - Be
- Q.9 _____ is the strong acid:-
- NaOH
 - CH_3COOH
 - HCl
 - KOH
- Q.10 _____ is used as adsorption agent
- Cellulose
 - Charcoal
 - Starch
 - Nylon

SECTION-B

Note: Objective type questions. All questions are compulsory. (10x1=10)

- Q.11 PPM stands for _____.
- Q.12 Atomic number of Carbon is _____.
- Q.13 Give an example of Strong base.
- Q.14 Oxidation involves _____ of electrons.

- Q.15 The energy of 4s orbital is _____ than that of 3d orbital.
- Q.16 SI unit of volume is _____
- Q.17 _____ water does not form lather with Soap.
- Q.18 Number of orbitals in d-subshell are _____.
- Q.19 When Acid and base react it forms _____
- Q.20 Electrolysis involves _____ at anode.

SECTION-C

Note: Short answer type questions. Attempt any twelve questions out of fifteen questions. (12x5=60)

- Q.21 Define Molarity and Molality. Give the mathematic expression for each.
- Q.22 State Heisen berg's Uncertainty Principles. Write down its mathematical expression. Give its significance also.
- Q.23 Calculate the number of moles in 392 g of Sulphuric acid (H_2SO_4)
- Q.24 Briefly explain common Ion effect with a suitable example.
- Q.25 Write down the applications of a buffer solution.
- Q.26 What are emulsion? How are they classified.
- Q.27 Write down the electronic Configuration of Carbon and Oxygen.
- Q.28 Discuss in brief the size of Atom in periods and groups.