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188414

Roll No.

1st Sem / DVOC
Subject : Applied Chemistry

Time : 2 Hrs.

M.M. : 50

SECTION-A

Note: Very short questions. Attempt all ten questions.
(10x1=10)

- Q.1 A s orbital can contain maximum _____ electrons.
- Q.2 Principal Quantum no is denoted by _____.
- Q.3 The S shape of 2P orbital is _____.
- Q.4 An anion is _____ charged particle.
- Q.5 Nucleus of an atom contains _____.
- Q.6 Full form of PVC is _____.
- Q.7 The water which easily forms lather with soap is known as _____.
- Q.8 Monomer of Bakelite is _____.

Q.9 pH of Acidic Solution is_____.

Q.10 An example of liquid fuel_____.

SECTION-B

Note: Short answer type questions. Attempt any six questions out of eight questions. (6x5=30)

Q.11 Differentiate between an orbit and orbital.

Q.12 Write down the electronic configuration of Nitrogen and Sulphur (Atomic No-16) (Atomic No-7)

Q.13 State and explain Modern Periodic Law.

Q.14 Define Ionization Energy and Electron affinity.

Q.15 Enlist the qualities of drinking water.

Q.16 Write down the composition of water gas and write three uses of it.

Q.17 Differentiate between covalent bonding and Ionic bonding.

Q.18 Differentiate between Sludge and Scale formation.

SECTION-C

Note: Long answer questions. Attempt any one questions out of two questions. (1x10=10)

Q.19 i) Explain sacrificial Anodic protection method of Corrosion.

ii) Main Postulates of Bohr's Model of an atom.

Q.20 Write short notes on any two:-

i) Thermosetting and Thermoplastics

ii) Characteristics of Good fuel

iii) Electron Sea Model of metallize bonding.