

Units and Values

Avogadro's number: $N_A = 6.023 \times 10^{23}$
Boltzmann's cosntant: $k = 1.381 \times 10^{-23} J/K$,
 $R = 8.31J/Kmol$, $nR = Nk$.
Calorie: 1cal = 4.2J (amount of heat needed to raise the temperature of a gram of water by 1°C.)
Absolute zero: 0K = −273.15°C.

Definitions

Isotherm: constant temperature
Isobar: constant pressure
Isothermal compression: so slow that temperature does not rise
Adiabatic compression: so fast heat does not escape
fundamental assumption of statistical mechanics: In an isolated system in thermal equilibrium, all accessible microstates are equally probable
thermodynamic limit: Let number of particles go to infinity

Important formulas

$$\Omega_{total}(N, n) = \Omega_A(N, n)\Omega_B(N, n)$$
$$C_V \rightarrow 0 \quad \text{as} \quad T \rightarrow 0$$
$$C_V \rightarrow \text{constant} \quad \text{as} \quad T \rightarrow \infty$$
$$\Delta S = 0 \quad \rightarrow \quad \text{reversible}$$
$$\Delta S \geq 0 \quad \rightarrow \quad \text{irreversible}$$

Important facts

- Macrostates with maximum multiplicity are the most likely and they correspond to the average values
- ΔS is path independent, choose the simplest path between the states to compute

Ideal Gas

Low density gas $PV = NkT$, ignore particle interactions.
 $\langle v^2 \rangle = 3kT/m$. Compression $W = NkT \ln V_i/V_f$. For indistinquishable particles (we get $1/N!$)

$$\Omega(U, V, N) = f(N)V^N U^{3N/2} \approx \frac{V^N}{N!(3N/2)!} (2\pi mU/h^2)^{3N/2}$$
$$S = Nk \left[\ln \left(\frac{V}{N} \left(\frac{4\pi mU}{3Nh^2} \right)^{3/2} \right) + \frac{5}{2} \right]$$
$$\Delta S = Nk \ln V_f/V_i \quad (U, N \text{ fixed})$$

Equipartition theorem

For each quadratic degree of freedom (kinetic v^2 , spring k^2 , rotational ω^2 , etc.) contributes to $1/2kT$ to the average energy at equilibrium at temperature T . For N particles with f degrees of freedom $U = NfkT/2$. Safest to apply in changes of energy. Coutning f : one for each dimension, 2 from rotation (as one axis is symmetric), vibratinge counts twice (frozen out at room temperature) in a solid 3 vibrational directions $f+=6$. Equipartition theorem only works for kinetic energy in liquids. Use $f = 3$ for a monatmoic gas and $f = 5$ for a diatomic gas. Derived by Boltzmannfactors and converting to integral

$$Z = \frac{1}{\Delta q} \sum_q e^{-\beta cq^2} \Delta q$$

and calculating energy from partition function

Heat and Work

$$\Delta U = Q + W \tag{1}$$

Assuming quasistatic compression $W = -P\delta V$. Use pressure from POV of gas, negative work means gas is doing work, positive work means the outside is doing work on the gas.

$$W = - \int_{V_i}^{V_f} P(V) dV \text{quasistatic}$$

Heat Capacity

Heat capacity: amount of heat needed to raise temperature per degree temperature increase $C = Q/\Delta T$.
Specific heat capacity: Heat capacity per unit mass $c = C/m$.

$$C = (\Delta U - W)/\Delta T, C_V = \left(\frac{\partial U}{\partial T}\right)_V, C_P = \left(\frac{\partial U}{\partial T}\right)_P + P \left(\frac{\partial V}{\partial T}\right)$$

At the phase transition you can put heat into system without increasing temperature, $C = \infty$. Latent heat $L = Q/m$ is the heat needed to accomplish phase transition per mass, assume constant pressure latm, and no other work.

Adiabatic compression

So fast no heat flows out, still quasistatic, $\Delta U = W$.

$$V^\gamma P = \text{constant}, \gamma = (f + 2)/f$$

Two state system

Multiplicity of the macrostate with N particles where N_\uparrow point up

$$\Omega(N, n) = \frac{N!}{N_\uparrow!(N - N_\uparrow)!} = \frac{N!}{N_\uparrow!N_\downarrow!} \tag{2}$$

Einstein solid

$N/3$ atoms as harmonic oscillators which can oscillate in 3 dimensions with $q = (U_n - N\hbar\omega/2)/\hbar\omega$ units of energy

$$\Omega(N, q) = \binom{q + N - 1}{q} = \frac{(q + N - 1)!}{q!(N - 1)!} \approx \frac{(q + N)!}{q!N!} N_\downarrow! \tag{3}$$

If a composite system: $q = q_A + q_B$ and $N = N_a + N_b$ and $\Omega = \Omega_A \Omega_B$.

$$\text{low } T \quad \Omega \approx \left(\frac{Ne}{q}\right)^q, \quad q \ll N$$

$$\text{high } T \quad \Omega \approx \left(\frac{qe}{N}\right)^N, \quad N \ll q$$

$$S = kN \left[\ln \left(\frac{q}{N} + 1 \right) \right], \quad q = U/\hbar\omega$$

thermodynamic potentials

- Internal Energy: U** Total internal thermal energy of the system. The change in the stored energy equal the sum of «energies in transit» $\Delta U = Q + W$ (first law of thermodynamic). The infinitesimal change in internal energy: $dU = \delta Q + \delta W$. For infinitesimal reversible process: $dU = TdS - PdV$.

- Enthalpy: $H = U + PV, \Delta F \leq W$** (at constant T). Energy needed to create the system out of nothing, and make room for it

- Helmholtz free energy: $F = U - TS$** , (constant pressure). Total energy needed to create the system, minus the heat you get from enviornment. Thermodynamic identity: $dF = -SdT - PdV + \mu dN$. Minimized in thermal bath, with only energy exchange.
 $\mu = \left(\frac{\partial F}{\partial N}\right)_{T,V}, S = -\left(\frac{\partial F}{\partial T}\right)_{V,N}, P = -\left(\frac{\partial F}{\partial V}\right)_{T,N}$.

- Gibbs free energy: $G = U - TS + PV, \Delta G \leq W_{other}$** . Create system in constant temperature and pressure. Minimized when change in volume with constant pressure. $\mu = \frac{\partial G}{\partial N}, V = \frac{\partial G}{\partial P}, -S = \frac{\partial G}{\partial T}$

- Chemical potential $\mu = \left(\frac{\partial G}{\partial N}\right)_{T,P}$** . The energy increase by adding a particle in to the system with constant temperature and pressure.

At constant U and T , S increases.

At constant V and T , F decreases.

At constant T and P , G decreases.

Reversible process

At every step of the path the system is at equilibrium. For infinitesimal reversible process: $dU = TdS - PdV$.
Clausius equality: $dS = \delta Q_{rev}/T$. Irreversible heat transfer is smaller than the reversible heat exchange at a given T

$$\frac{\delta Q_{irrev}}{T} < \frac{\delta Q_{rev}}{T}$$

Clausius inequality: $dS \geq \delta Q/T$. For an isolated system entropy tends to increase as the system spontaneously finds its equilibrium state $dS0$.

$$C_V = \left(\frac{\delta Q}{d}\right)_V = \left(\frac{\partial U}{\partial T}\right) \rightarrow dU = C_V dT$$

Isolated system at equilibrium

Multiplicity of a macrostate $\Omega(U, V, N)$ counts all equally likely accessible microstates, if the particles are indistinguishable the total number of microstates is reduced by $N!$. Boltzmann's formula $S = k \ln \Omega = -k \sum_s P(s) \ln P(s)$, for equilibrium state at fixed U , entropy is maximized at equilibrium, $dS = 0$.

$$TdS = \frac{1}{T}dU + \frac{P}{T}dV - \frac{\mu}{T}dN \tag{4}$$

For a composite system

$$\frac{\partial S_A}{\partial V_A} = \frac{\partial S_B}{\partial V_B} \quad (\text{at equilibrium})$$

$$\frac{\partial S_A}{\partial N_A} = \frac{\partial S_B}{\partial N_B} \quad (\text{at equilibrium})$$

Extensive and Intensive Quantities

An extensive quantity doubles when you double a system. The quantities which are unchanged when you double the system are intensive.

- Extensive**: V, N, S, U, H, F, G , mass

- Intensive**: T, P, μ , density

extensive \times intensive = extensive

$$\frac{\text{extensive}}{\text{extensive}} = \text{intensive}$$

extensive \times extensive = neither, you did something wrong

extensive + extensive = extensive

extensive + intensive = not allowed

Van der Waals Equation

$$\left(P + \frac{aN^2}{V^2}\right)(V - Nb) = NkT$$

Nb makes it impossible to compress to zero volume. the a term accounts for short range attractive forces. Low density gas $PV = NkT$, ignore particle interactions. $\langle v^2 \rangle = 3kT/m$. Compression $W = NkT \ln V_i/V_f$.

Maxwell Construction

Makes it possible to find the pressure at the phase transition from the PV -diagram, create straight line which makes even areas in the PV -line. Gives us the point of phase transition between liquid and gas. Gas at high volume, liquid at low volume. The intersected regions are unstable. The phase transition happens at constant PRESSURE! We find the pressure on the P-V diagram by the equal areas construction of a given isotherm. Phase transition from liquid to vapor at a constant Gibbs free energy $G(P, T, N)$, $dG_{liquid} = dG_{gas}$.
Clausius Clapeyron relation

$$-S_l dT + V_l dP = -S_g dT + V_g dP$$

$$\frac{dP}{dT} = \frac{S_g - S_l}{V_g - V_l} = \frac{\Delta S}{\Delta V} = \frac{L}{T\Delta V}$$

Entropy jumps going from a liquid to a gas. Volume expansion going from a liquid to a gas. The Clausius Clapeyron relation tells us how much the phase transition pressure changes with changing temperature.

Sterlings approximation

$$N! \approx N^N e^{-N} \sqrt{2\pi N}$$

$$N! \approx N^N e^{-N}$$

$$\ln(N!) \approx N \ln N - N$$

$$\ln \binom{N}{n} \approx N \ln N - n \ln n - (N - n) \ln (N - n)$$

Thermodynamic identities

$$dU = TdS - PdV - \mu dN \tag{5}$$

Maxwell relations

Functions are usually well behaved, so order of differential does not matter, can then use

$$\frac{\partial}{\partial V} \left(\frac{\partial U}{\partial S} \right) = \frac{\partial}{\partial S} \left(\frac{\partial U}{\partial V} \right)$$

By using thermodynamic identities for dU one can derive the maxwell relations, here they are

$$\frac{\partial^2 U}{\partial S \partial V} \rightarrow \left(\frac{\partial T}{\partial V}\right)_S = -\left(\frac{\partial P}{\partial S}\right)_V$$
$$\frac{\partial^2 H}{\partial S \partial P} \rightarrow \left(\frac{\partial T}{\partial P}\right)_S = +\left(\frac{\partial V}{\partial S}\right)_P$$
$$-\frac{\partial^2 F}{\partial T \partial V} \rightarrow \left(\frac{\partial S}{\partial V}\right)_T = +\left(\frac{\partial P}{\partial T}\right)_V$$
$$\frac{\partial^2 G}{\partial T \partial P} \rightarrow \left(\frac{\partial S}{\partial P}\right)_T = -\left(\frac{\partial V}{\partial T}\right)_P$$

Maxwell distribution

Derived from multiplying the probability of having velocity v , and the number of vectors corresponding to velocity v . Probability is $e^{-mv^2/2kT}$. Number of vectors corresponding to v is $4\pi v^2$. Find $D(v)$ by normalizing, $\int_0^\infty dv D^3D(v) = 1$. get

$$D(v) = \left(\frac{m}{2\pi kT}\right)^{3/2} 4\pi v^2 e^{-mv^2/2kT}$$

Boltzmann statistics

Thermal Bath

Can only exchange energy, system + bath is isolated. Probability for the system to be in a specific microstate at fixed T : $P_i = \frac{1}{Z} e^{-E_i/kT}$, with the partition function: $Z = \sum_i e^{-E_i/kT}$, counts all accessible microstates weighted by the Boltzmann factor. The partition function determines the thermodynamic potential which is minimized at given T , V and N . Due to energy exchange with the thermal bath the equilibrium macrostate is an average

$$\langle E \rangle = -\frac{1}{Z} \frac{\partial Z}{\partial \beta} = -\frac{\partial}{\partial \beta} \ln Z \quad \text{equilibrium, reservoir}$$

$$\langle E^2 \rangle = -\frac{1}{Z} \frac{\partial^2 Z}{\partial \beta^2} \quad \text{equilibrium, reservoir}$$

Boltzmann distribution for the average number of particles (occupation number) in a given energy state

$$\langle N_s \rangle = NP(s) = e^{-\beta(E_s - \mu)} \tag{6}$$

We also have the following proerties

$$\langle A^n \rangle = \frac{1}{Z} \sum_i A^n e^{-E_i/kT}$$

$$U = N \langle E \rangle$$

$$F = -kT \ln Z$$

$$Z = e^{-F/kT}$$

$Z_{tot} = Z_1 Z_2 ... Z_N$ distinguishable, identical and independent

$$Z_{tot} = \frac{Z^N}{N!} \text{ indistinguishable, identical and independent}$$

Thermal Bath and particle reservoir

System can exchange energy and particles with a reservoir. Equilibrium at a fixed T and μ . Probability of the system in a specific microstate at fixed T and μ

$$P(s) = \frac{1}{Z_g} e^{-\beta(E_s - \mu N_s)}, \quad Z_g = \sum_s e^{-\beta(E_s - \mu N_s)}$$

Can be showed that the probability of the occupation number N of the given state is

$$P(s) = \frac{e^{-\beta N(E_s - \mu)}}{\sum_N e^{-\beta N(E_s - \mu)}}$$

From this one derives the quantum statistics:

Quantum Statistics

Gibbs factor = $e^{-(E_s-\mu_s N_s)/kT}$

Grand partition function $Z = \sum_s e^{-(E_s-\mu_s N_s)/kT}$

$$v_Q = l_Q^3 = \left(\frac{h}{\sqrt{2\pi m kT}}\right)^3$$

quantum condition $Z_t = Z^N/N!$: $\frac{V}{N} \gg v_Q$

FERMIONS:

$$Z = 1 + e^{-(\epsilon-\mu)/kT}$$
$$\langle n_{FD} \rangle = \frac{1}{1 + e^{(\epsilon-\mu)/kT}}$$

BOSONS:

$$Z = 1 + e^{-(\epsilon-\mu)/kT} + e^{-2(\epsilon-\mu)/kT} + \dots$$
$$\langle n_{BE} \rangle = \frac{1}{e^{(\epsilon-\mu)/kT} - 1}$$

The distribution becomes the boltzmann distribution in the high T limit ($\beta 0$)

Density of state

Average energy:

$$U = \sum_{n_x, n_y, n_z} \langle N(\epsilon) \rangle \epsilon(n) \approx \int_0^\infty \mathrm{d}n_{x,y,z} \epsilon \langle N \rangle = \int_0^\infty \mathrm{d}\epsilon \, g(\epsilon) \cdot \epsilon \cdot \langle N \rangle$$

Average number of particles

$$N(T, V, \mu) = \int_0^\infty \mathrm{d}\epsilon \, g(\epsilon) \langle N \rangle$$

Density of state $g(\epsilon)$ to count all the quantum states at a given energy ϵ . The number of states with energy between ϵ and $\epsilon + \mathrm{d}\epsilon$ = Number of states with state number between n and $n + \mathrm{d}n$ (positive quadrant). Different for different dimensions

$g(\epsilon)\mathrm{d}\epsilon =$

$(3D)\frac{1}{8}4\pi n^2\mathrm{d}n,$ $(2D)\frac{1}{4}2\pi n\mathrm{d}n,$ $(1D)\mathrm{d}n$

(7)

the energy is determined by quantum mechanics

- particle in a box: $\epsilon(n) = h^2 n^2 / (8mL^2)$
- harmonic oscillator: $\epsilon(n) = n \hbar \omega$
- relativistic particles: $\epsilon(n) = \hbar f = \hbar c n / (2L)$

$$g(\epsilon)\mathrm{d}\epsilon = D(n)\mathrm{d}n \rightarrow g(\epsilon) = \frac{D(n)}{\frac{\mathrm{d}\epsilon}{\mathrm{d}n}}$$

For **FERMIONS** multiply by 2, spin up and spin down.
For **PHOTONS** multiply by 2, for the two transverse polarization of EM waves, $\mu = 0$.
For **PHONONS** multiply by 3, for all three polarizations of the sound waves.

Quantum Ideal Gas

$$\epsilon_n = \frac{\vec{p} \cdot \vec{p}}{2m} = \frac{h^2}{8mL^2} \left(n_x^2 + n_y^2 + n_z^2 \right)$$
$$Z_1(T, V) = \sum_{n_x} \sum_{n_y} \sum_{n_z} e^{\frac{-h^2 (n_x^2 + n_y^2 + n_z^2)}{8mL^2 kT}} = \left(\sum_n e^{\frac{-h^2 n^2}{8mL^2 kT}} \right)^3$$
$$Z_1(T, V) \approx_{highT} \frac{1}{2} \int_{-\infty}^\infty \mathrm{d}n e^{\frac{-h^2 n^2}{8mL^2 kT}} = \frac{V}{\Lambda^3(T)}, \Lambda = \sqrt{\frac{h}{2\pi m kT}}$$
$$Z_N(T, V) = \frac{1}{N!} \left(\frac{V}{\Lambda^3(T)} \right)^N$$
$$F_N = -kT \ln Z_N(T, V) = -NkT \left(\ln \left(\frac{V}{N\Lambda^3(T)} \right) - 1 \right)$$
$$\mu(T, V) = \left(\frac{\partial F}{\partial N} \right)_{T, V} = -kT \ln \left(\frac{V}{N\Lambda^3(T)} \right)$$
$$P = - \left(\frac{\partial F}{\partial V} \right)_{T, N} = \frac{kT}{V}$$
$$U = - \frac{\partial}{\partial \beta} \left(\ln Z_N(T, V) \right) = 3NkT/2$$
$$S = \frac{U - F}{T} = Nk \left[\ln \left(\frac{V}{N\Lambda^3(T)} \right) + \frac{5}{2} \right]$$

Integrals

$$\int_0^\infty \frac{x^3}{e^{bx} - 1} = \frac{\pi^4}{15b^4}$$
$$\int_0^\infty \frac{x^3}{e^{bx} + 1} = \frac{7\pi^4}{120b^4}$$

Taylor expansions

$$\ln 1 + x \approx x - x^2/2$$

Tricks

$$\sum_n^N \frac{N!}{n!(N-n)!} a^N b^{N-n} = (a+b)^N$$
$$\sum_n^N n \frac{N!}{n!(N-n)!} a^N b^{N-n} = a \frac{\partial}{\partial a} \sum_n^N \frac{N!}{n!(N-n)!} a^N b^{N-n}$$
$$Rp^R = p \frac{\partial}{\partial p} R^p$$
$$\frac{1}{1-x} = 1 + x + x^2 + x^3 + \dots$$