



Cite this: DOI: 10.1039/c6ce01787h

Received 16th August 2016,
Accepted 17th October 2016

DOI: 10.1039/c6ce01787h

www.rsc.org/crystengcomm

Observing non-classical crystallisation processes in gypsum *via* infrared attenuated total reflectance spectroscopy†

R. Stach,^a P. Krebs,^a F. Jones^b and B. Mizaikoff^{*a}

The nature of crystallisation processes is of major interest, as they are among the most frequently occurring reactions associated with a variety of relevant processes in chemistry, biochemistry, and geochemistry. In this study, an innovative approach towards fundamentally understanding crystallisation pathways in a seemingly simple system – gypsum – has been developed *via* infrared spectroscopic techniques. Specifically, infrared attenuated total reflection spectroscopy (IR-ATR) was instrumental in revealing detailed information on inter- and intramolecular interactions during gypsum crystallization *via* subtle changes in the vibrational spectra of the involved reactants. When applying D₂O as an isotope marker, it was shown that isotopically labelled water may serve as a viable spectroscopic probe during mid-infrared (3–15 μm) studies providing unique insight into the crystallization process at molecular-level detail. In addition, it was revealed that H₂O and D₂O give rise to distinctly different reaction kinetics during the crystallization process.

Gypsum is a well-studied molecule with its crystal structure investigated *via* a variety of different methods.^{1–3} In addition, the intra- and intermolecular interactions between water molecules and gypsum have been investigated.^{4,5} Compared to other non-soluble sulfates such as barite, water indeed plays a crucial role during the crystallisation process, as reported by Jones *et al.*⁶ Consequently, in the present study D₂O was applied as an isotope marker for tracing the pathways of crystallisation in molecular detail. While it is anticipated that its molecular behaviour is similar to water, D₂O has a distinct infrared signature facilitating its spectral differentiation from H₂O even in mixtures.⁷ To date, the influence of D₂O on the reaction kinetics was only reported for bio-minerals though.⁸ While the influence of deuterium, and especially of HOD on

the crystal structure of CaSO₄ was the subject of earlier studies, to the best of our knowledge the process of gypsum crystallisation in heavy water has been not examined in satisfactory detail.⁹

Even more so, in view of the recently emerging non-classical nucleation theory (non-CNT), the investigated system is of substantial interest, as the reaction proceeds sufficiently slow for monitoring the associated kinetics and nucleation pathways. Non-CNT mechanisms were observed during the nucleation of a variety of different systems including proteins, organic crystals, and crystals derived from bio-minerals.^{8,10,11} Derived from previous reports,^{6,12} it was anticipated that gypsum should likewise follow a non-CNT pathway. Recently, Stawski *et al.* proposed a four step mechanism including the formation of 3 nm anhydrite particles determined *via* SAXS/WAXS.¹³ Alternatively and complementarily, IR-ATR spectroscopy – as an *in situ* technique – provides chemical information on the occurring processes without the immediate need of synchrotron radiation. Consequently, the

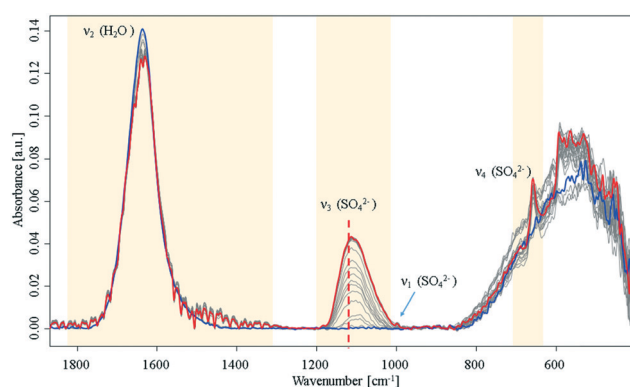


Fig. 1 Infrared spectra of an aqueous Ca²⁺ and SO₄²⁻ solution during crystallisation. Spectra were collected over a period of 16 h in the closed sample cell under nitrogen atmosphere. Blue: start at *t*₀; red: end at *t* = 16 h. For a detailed view of the ν₃ band, see ESI† Fig. S2. The dashed line highlights a shift of the SO₄²⁻ ν₃ mode at 1105 cm⁻¹ to higher wavenumbers.

^a Institute of Analytical and Bioanalytical Chemistry, Ulm University, Albert-Einstein-Allee 11, 89081 Ulm, Germany. E-mail: boris.mizaikoff@uni-ulm.de

^b Department of Chemistry and Nanochemistry Research Institute, Curtin University, GPO Box U1987, Perth, Western Australia 6845, Australia

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c6ce01787h

aim of the present studies was to gain in depth understanding on the crystallization processes of CaSO_4 in various media for unequivocally determining non-classical pathways in this system.

The crystallisation process towards $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$ and $\text{CaSO}_4 \cdot 2\text{D}_2\text{O}$ was monitored as a function of time *via* infrared attenuated total reflection (IR-ATR) spectroscopy using a single-bounce diamond ATR crystal (detailed experimental procedure see ESI†). To prevent evaporation and exchange processes with the ambient environment, a fluidic sample cell was designed and 3D printed (see ESI† Fig. S1), which may be sealed in a controlled nitrogen atmosphere.

During the crystallisation of gypsum in ultrapure water, significant changes of the H_2O bending mode (1640 cm^{-1}), and of the SO_4^{2-} vibrations at 1000 cm^{-1} (ν_1), 1105 cm^{-1} (ν_3), and 645 cm^{-1} (ν_4) were observed.^{14,15} The SO_4^{2-} ν_3 mode minutely increases until it reaches a plateau. Hence, it is concluded that this initial crystallisation reaction eventually (Fig. 1) stops, and eventually saturation is reached as anticipated for a batch system assuming that precursors for later crystalline nuclei formation are established.

During the entire observed period (16 h) the sulfate ν_3 band intensity revealed an exponential approximation towards a plateau along with a shift to higher wavenumbers. After approx. 10 h, the ν_1 and ν_4 bands appeared in the spectrum. This observation is indicative of a transition of sulfate from a T_d to a C_{2v} or C_{3v} symmetry, as the ν_1 band is in fact IR-inactive in the T_d symmetry.^{16,17} Additionally, an increase of the ν_4 vibration located at 645 cm^{-1} was evident during the crystallisation process, which has been reported in literature.^{6,9,14,15} Resulting from group theory, the ν_4 vibration of the calcium sulfate molecule was expected to be observed as a triplet – especially regarding recent studies on formation of anhydrite during the process.^{16–18} However, the remaining two peaks were not observed during these experiments; instead, the signal at 645 cm^{-1} leads to the assumption that an amorphous phase has formed rather than a final crystalline phase, which would reveal all characteristic signals of the crystal lattice. Likewise, the characteristic splitting of the gypsum water band (1685 cm^{-1} and 1623 cm^{-1})⁹ was not observed. The onset of a signal at 1480 cm^{-1} combined with the minute decrease of the water bending band indicates that a new population of H_2O – Ca^{2+} – SO_4^{2-} related modes is present. Direct precipitation of the gypsum monocline structure with C_{2h} symmetry¹⁹ may be precluded though, as the ν_4 doublet is absent.^{15,20}

After reaching equilibrium the system remained stable for at least 16 hours until a perturbation was applied (approx. 10 hours after the spectral behaviour was equilibrated), *e.g.*, *via* forced evaporation. Then, later stages of crystallisation were observed *via* additional spectral changes indicative of the final formation of the crystal lattice. This behaviour matches with a process of forming disordered solids from stable pre-nucleation clusters as the initial step towards a final lattice.^{10,21} The suggested hypothesis of solution clusters is consistent with a recently proposed pathway describing crystalli-

zation *via* an amorphous precursor phase by Y. W. Wang *et al.*¹² and Stawski *et al.*¹³

Following previous reports on crystallization of biomolecules⁸ and preliminary experiments during the present study, it was expected that crystallisation occurs more rapidly in D_2O vs. H_2O . For example, it was reported that D_2O enhances the aggregation speed of lysozyme. Turbidity experiments and kinetic calculations suggested that different kinetic routes for H_2O and D_2O are evident, and thus, different crystallisation mechanisms prevail.⁸ The associated IR spectra recorded during the crystallization process in neat D_2O clearly evidenced distinct differences in crystallization behaviour, as shown in Fig. 2.

In comparison to the same processes in H_2O , the spectra in deuterated water show progressively distinct spectral changes over the observation period, and a more rapid crystallisation process. The sulfate ν_3 vibration rises rapidly with an exponential behaviour (see, ESI† Fig. S3), and approaches maximum intensity after approx. 350 min. Furthermore, the signal intensity is 10-times stronger compared to the same signal in H_2O . The early (*i.e.*, at approx. 60 min) emergence of the sulfate ν_1 and ν_4 vibrations again evidence a change in symmetry. In the spectral region of the D_2O stretching mode, a peak-shift towards higher wavenumbers was observed combined with a peak splitting. Fig. S3 (ESI†) illustrates the progress of the peak splitting. The green curve indicates the increase of a new band at 2590 cm^{-1} , which occurs during a simultaneous decrease and shift of the 2500 cm^{-1} band. Likewise, the D_2O bending vibration splits into two separate peaks. It is hypothesized that this splitting behaviour is similar to the characteristic water splitting observed in $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$ IR spectra, and provides strong evidence that D_2O is incorporated into the crystal lattice instead of H_2O . It should be noted that during these experiments bands of H_2O (*i.e.*, stretching vibrations) were also observed resulting from the usage of hydrated initial constituents. However, minute amounts of remaining water apparently did not interfere with the observed processes. In addition, two bands were arising at 665 cm^{-1} and 597 cm^{-1} , where in fact a

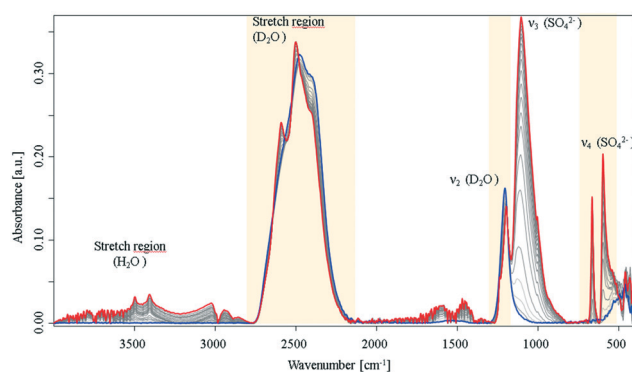


Fig. 2 IR-ATR spectra of a CaSO_4 solution in D_2O during crystallisation collected over a period of 16 h under nitrogen atmosphere. Evolution of the D_2O stretching vibrations and of the sulfate ν_3 mode, see ESI† Fig. S3.

triplet from the ν_4 mode was expected due to the C_{2v} symmetry of calcium sulfate molecules; however, the spectral behaviour indicates direct precipitation of crystalline species with C_{6h} symmetry. This is in distinct contrast to the process in H_2O , which only formed a crystalline product when evaporation was allowed. After the crystallisation process appeared complete, the samples were extensively dried under a nitrogen atmosphere, which resulted in crystalline $CaSO_4 \cdot 2H_2O$ and $CaSO_4 \cdot 2D_2O$ (see, ESI† Fig. S4). Interestingly, the deuterated gypsum samples revealed excellent stability against H-D exchange with ambient water, which diffuses into the solid over a period of ten days. Likewise, treatment with small amounts of water did not lead to a measurable exchange of hydrogen and deuterium. Hence, it may be assumed that D_2O forms very stable deuterium bonds within the crystal lattice. In the IR-ATR spectra of the dried samples (ESI† Fig. S4), the features, which appeared during the crystallisation process appear well resolved, and the generated gypsum spectra compare well with the spectrum of natural gypsum. As expected, the water bending and stretching vibrations split into two vibrational modes, and additionally, a well-resolved SO_4^{2-} ν_4 triplet is apparent (see, ESI† Fig. S4, blue spectrum), as reported elsewhere as a shoulder for natural gypsum samples.²² This indicates that the precipitate does not provide perfect C_{6h} symmetry in contrast to the samples grown in D_2O .

The D_2O stretching and bending modes clearly show splitting as well (red spectrum). In comparison to $CaSO_4 \cdot 2H_2O$, the bands within the H_2O stretching region of $CaSO_4 \cdot 2D_2O$ are shifted to smaller wavenumbers. This implies that this band is resulting from HDO integration into the crystal lattice due to H-D exchange with residual water. Furthermore, a strong sulfate vibration is evident. Interestingly, the ν_4 mode of D_2O -grown gypsum still appears as a pronounced doublet (red spectrum). This indicates that the structural arrangements of both crystalline species do not reveal the same symmetry. The apparent doublet leads to the assumption that sulfate in $CaSO_4 \cdot 2D_2O$, which was crystallized in heavy water established a perfect C_{6h} symmetry. Although the $CaSO_4 \cdot 2H_2O$ spectrum reveals a ν_4 triplet, the vibrational pattern of the spectrum – as reported elsewhere – provides no direct evidence for the presence of anhydrite or basanite.^{15,23} However, the lifetime of these species is too short to be observable at the timescales of the present studies. SEM images of $CaSO_4 \cdot 2D_2O$ (ESI† Fig. S4) reveal distinct monocline crystals, as commonly reported for gypsum, which aligns well with the obtained characteristic IR spectra and a C_{6h} symmetry.²⁰ In comparison, gypsum samples grown in H_2O have a tendency to crystallise in large plates with only a small proportion of monocline structures. Combined with observations derived from the spectral data, it may be assumed that D_2O promotes nucleation and growth during the crystallisation process by following different kinetic routes explained as follows.

According to non-classical nucleation theory, two energy barriers are proposed for the reaction from ion pairs to ordered crystalline nuclei. First, the reactants have to overcome

an energy barrier to form stable solution clusters. This barrier can be insignificant in some systems, for example calcium carbonate (see the work of Gebauer and Cölfen).^{10,24} From this stable state, the system needs to overcome a second energy barrier for establishing ordered crystalline nuclei from the initially disordered yet stable state.^{10,11} In pure D_2O , proof for solution clusters – as evident in the H_2O experiments – could not be observed. For the case of D_2O , if a non-CNT mechanism is operative, the transition of the pre-nucleation clusters into crystalline solids is kinetically faster than the collection of the spectrum. Thus, either the crystallisation process is *via* a classical pathway or in the presence of D_2O there is a very fast pre-nucleation clusters to amorphous solids to crystalline solids transition. Clearly, the presence of D_2O also increases the overall rate of crystallisation (a crystalline product is formed after 60 min). Nonetheless, the reactants precipitated as perfect monocline crystals and a perfect C_{6h} symmetry. The obtained IR-ATR spectra provide evidence for the occurrence of crystalline species, which are different in symmetry compared to the same reaction in water. Solubility effects may be neglected, since the solubility in both solvents only differs by $12.86\% \pm 1.6\%$ measured *via* total reflection X-ray fluorescence (TXRF), see ESI†.

In pure H_2O , Stawski *et al.*,¹³ proposed a four step process that includes the formation of pre-nucleation clusters and nano-sized particles prior to formation of gypsum particles. These solids were proposed based on calculated electron densities, which involved some assumptions, while the work presented here is based only on the infrared signatures of the system as it evolves temporally.

Additional experiments, were performed by spiking D_2O into an aqueous $CaSO_4$ system at deliberately selected intervals. The first spike was injected after approx. 30 min, *i.e.*, sufficient time was provided to facilitate the presence of solution clusters and amorphous solids; the associated spectra are shown in Fig. 3 (experimental details are given in the ESI†).

It is clearly evident that each D_2O spike induces a crystallisation event. After each injection of D_2O , the sulfate band increases with an exponential behaviour until D_2O is consumed by the system. In turn, this indicates if the freely available D_2O concentration drops below a certain limit, precipitation stops and the system resides in a stable equilibrium state (see, ESI† Fig. S5). After nine spikes, the entire ATR crystal surface is covered, and no further changes may be observed. Effects due to different solubility may safely be neglected, as the injected amount of D_2O is too small (*i.e.*, 5% for each spike) to change the overall solubility of gypsum within the system.

Due to the behaviour of the water bending mode, it is suggested that D_2O is consumed by both H-D exchange and crystal growth. Regarding Fig. 3a), all bands of deuterated water show distinct band splitting. Likewise, bands in the H_2O /HDO stretching region shift to higher wavenumbers, get more pronounced, and start to split. Thus, apparently both

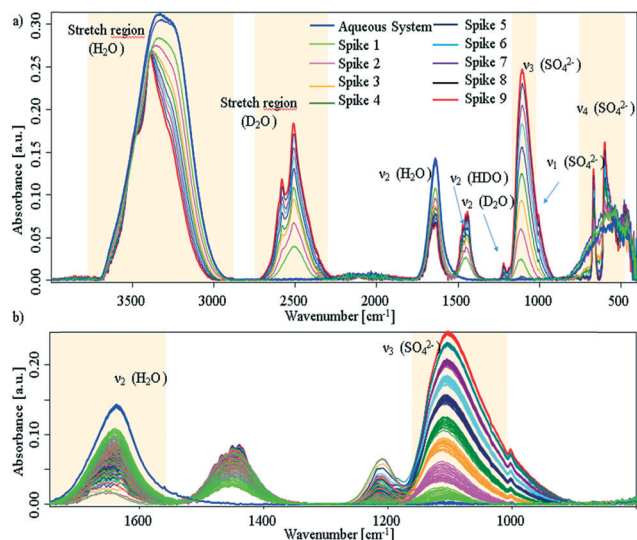


Fig. 3 a) IR-ATR spectra of aqueous CaSO_4 solution, which was spiked with D_2O in intervals of 30 min. Blue spectra: prior to spiking; red spectrum: last D_2O spike. b) Zoom onto the sulfate ν_3 band and water bending region. The spectral behaviour as a function of time is given in Fig. S5 within the ESI†

species are incorporated into the crystal lattice. In lieu of the ν_4 triplet of gypsum seen, a doublet is apparent, which compares well to the D_2O -grown crystals shown in ESI† Fig. S4. It can therefore be assumed that a crystalline structure with C_6^{2h} symmetry is directly established and precipitates. The missing triplet and well resolved splits in the D_2O stretching regime proves that deuterium oxide not only initiates the reaction, but is intimately involved in the crystallisation process within that mixture.

The IR-ATR spectra obtained in the present study clearly supports a non-classical crystallization mechanism for the water-gypsum system, and reveals a substantial impact of D_2O on the crystallisation kinetics. Compared to SAXS/WAXS requiring a synchrotron light source, IR-ATR spectroscopy provides in-depth *in situ* information on the nature and chemistry of such nucleation processes with fewer assumptions being necessary. Evidence for the formation of anhydrite or basanite during the nucleation process, as reported elsewhere¹³ could not be observed, which is probably related to the lifetime of these species.²³ In any case, the obtained results are consistent with previously suggested amorphous precursor phases.^{6,12,13}

Due to the impact of D_2O on the crystallisation process, transformation from solution clusters to crystalline solids – may be induced *via* D_2O spikes. Each D_2O injection into the H_2O system, which is initially in an equilibrium state, instantly triggers precipitation into perfect monocline crystals.

Thus, it is suggested that molecular restructuring from solution clusters to ordered crystalline nuclei is a critical rate-determining step for the formation of gypsum solids. From the perspective of the two step nucleation theory, this indicates that the energy barrier between those states is lower in D_2O vs. H_2O , which explains the accelerated crystal forma-

tion. This data also shows, very importantly, that the kinetic barriers are critical in determining whether the system crystallises through a classical or non-classical pathway.

Acknowledgements

RS thanks D. Rommel at the team of Prof. K. Leopold (IABC, Ulm Univ.) for assistance with TXRF measurements.

Notes and references

- G. A. Lager, T. Armbruster, F. J. Rotella, J. D. Jorgensen and D. G. Hinks, *Am. Mineral.*, 1984, **69**, 910–918.
- H. Takahashi, I. Maehara and N. Kaneko, *Spectrochim. Acta, Part A*, 1983, **39**, 449–455.
- W. F. Cole and C. J. Lancucki, *Acta Crystallogr., Sect. B: Struct. Crystallogr. Cryst. Chem.*, 1974, **30**, 921–929.
- R. Kling, *J. Chem. Phys.*, 1971, **54**, 5331.
- K. Iishi, *Phys. Chem. Miner.*, 1979, **4**, 341–359.
- F. Jones, *CrystEngComm*, 2012, **14**, 8374–8381.
- D. Neubauer, J. Korbacher, M. Frick, J. Kiss, M. Timmler, P. Dietl, O. H. Wittekindt and B. Mizaikoff, *Anal. Chem.*, 2013, **85**, 4247–4250.
- L. Xin Qi and S. Yoh, *J. Protein Chem.*, 1998, **17**, 9–14.
- M. Hass and G. B. B. M. Sutherland, *Proc. R. Soc. London, Ser. A*, 1956, **236**, 427–445.
- D. Gebauer and H. Cölfen, *Nano Today*, 2011, **6**, 564–584.
- D. Erdemir, A. Y. Lee and A. S. Myerson, *Acc. Chem. Res.*, 2009, **42**, 621–629.
- Y.-W. Wang, Y.-Y. Kim, H. K. Christenson and F. C. Meldrum, *Chem. Commun.*, 2012, **48**, 504.
- T. M. Stawski, A. E. S. van Driessche, M. Ossorio, J. Diego Rodriguez-Blanco, R. Besselink and L. G. Benning, *Nat. Commun.*, 2016, **7**, 11177.
- I. R. Moraes, M. C. P. M. Da Cunha and F. C. Nart, *J. Braz. Chem. Soc.*, 1996, **7**, 453–460.
- J. L. Bishop, M. D. Lane, M. D. Dyar, S. J. King, A. J. Brown and G. A. Swayze, *Am. Mineral.*, 2014, **99**, 2105–2115.
- F. A. Cotton, *Chemical applications of group theory*, Wiley, Cambridge, 3rd edn, 1991, vol. 249.
- J. J. Wylde, C. A. Geoffrey and I. R. Collins, *Appl. Spectrosc.*, 2001, **55**, 1155–1160.
- G. Renaudin, R. Segni, F. Leroux and C. Taviot-Gueho, *12th Int. Congr. Chem. Cem.*, 2007.
- J. F. Ferraro, *Low-Frequency Vibrations of Inorganic and Coordination Compounds*, Plenum Press, New York, 1st edn, 1971.
- W. A. Deer, R. A. Howie and J. Zussman, *An introduction to the Rock-Forming Minerals*, Pearson Education Limited, Edinburgh, 2nd edn, 1992.
- A. S. Myerson, *Faraday Discuss.*, 2011, **179**, 11–14.
- M. D. Lane, *Am. Mineral.*, 2007, **92**, 1–18.
- U. Tritschler, A. E. S. Van Driessche, A. Kempter, M. Kellermeier and H. Cölfen, *Angew. Chem., Int. Ed.*, 2015, **54**, 4083–4086.
- D. Gebauer, M. Kellermeier, J. D. Gale, L. Bergström and H. Cölfen, *Chem. Soc. Rev.*, 2014, **43**, 2348–2371.