

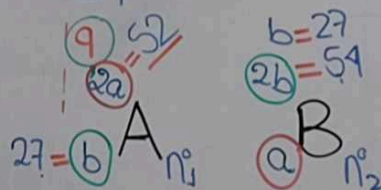
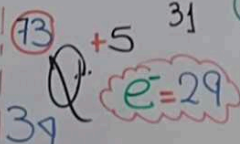
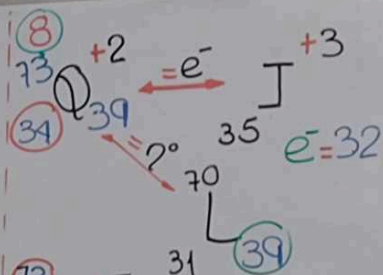
$\begin{matrix} \textcircled{7} & A \\ & \diagdown \quad \diagup \\ z_1 & & n_1 \end{matrix}$
 $\begin{matrix} A \\ | \\ z_2 & n_2 = 32 \end{matrix}$

$$n_1^o + n_2^o = 40$$

$$\underline{Z_1 - Z_2 = n_2^0 - n_1^0}$$

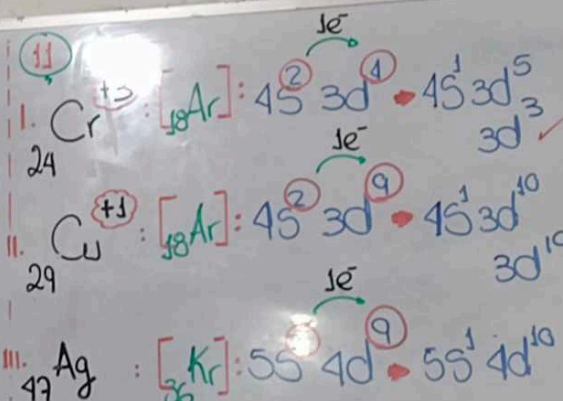
$$\Delta G = n_2^0 + n_3^0$$

$$g = n^\circ$$



$$2a - 27 + 54 - a = 53$$

$$a=26$$



$$(2\sqrt{zn})^2 = (z+n)^2$$

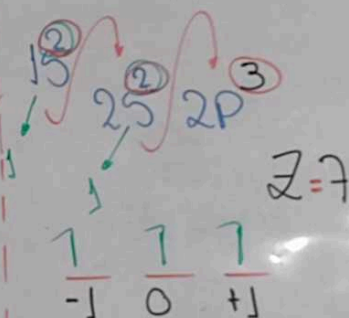
$$4. \mathbb{Z} \eta^{\circ} = \mathbb{Z}^2 + 2\mathbb{Z} \eta^{\circ} + \eta^{\circ 2}$$

$$0 = \cancel{2}^2 - 2\cancel{2}n^{\circ} + n^{\circ 2}$$

$$Q = (z - n^0)^2$$

$$Z = n^0 = 7$$

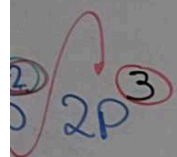
Posee la mínima cantidad de e para tener solamente 5 orbitales.



$2 = 7$

$A = 14$: 4 e⁻ apareados

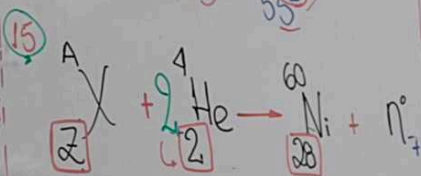
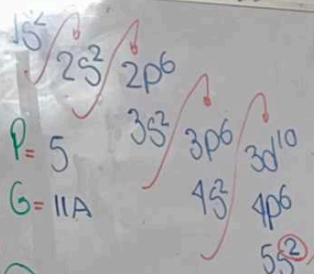
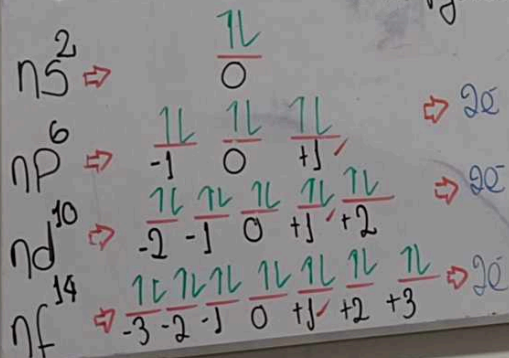
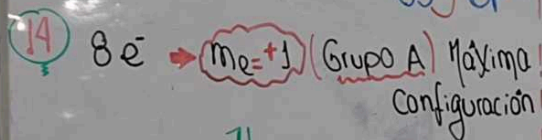
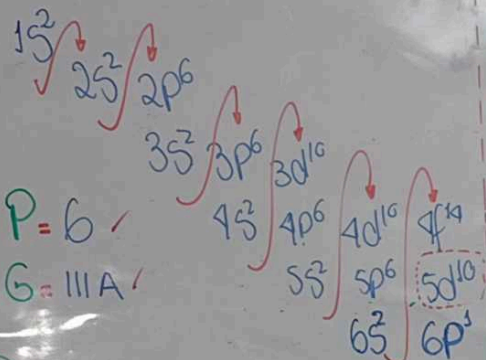
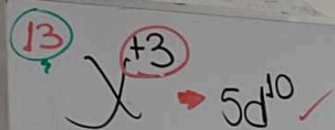
mínima cantidad
a tener solamen-
tales -



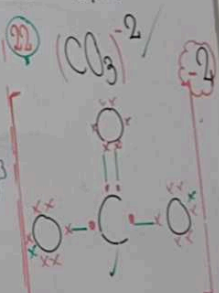
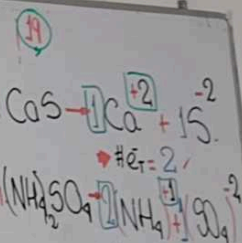
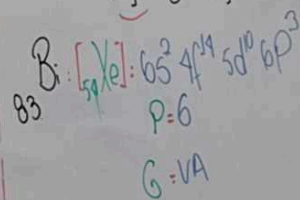
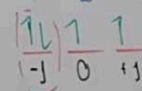
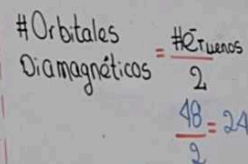
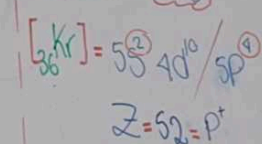
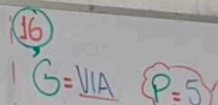
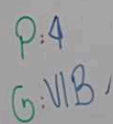
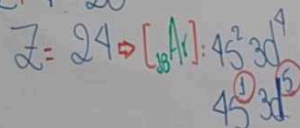
$$Z=7$$

$$\frac{1}{0} \frac{1}{+1}$$

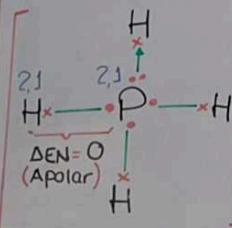
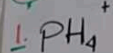
4e apareados



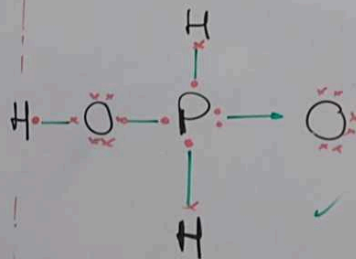
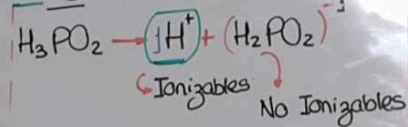
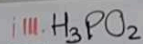
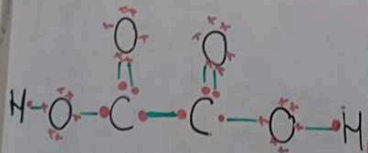
$$Z+4=28$$



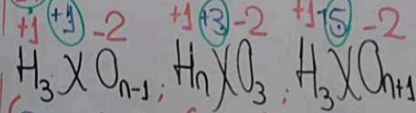
(23)



4 Enlaces Apolares

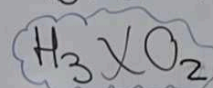
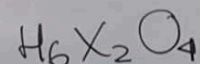
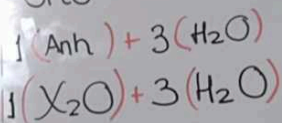


(27)

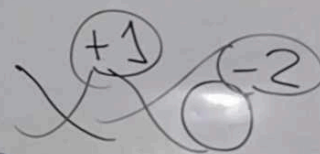
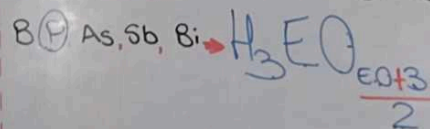


$$(2n-5) + (6-n)(2n-1) = 9$$
$$n=3$$

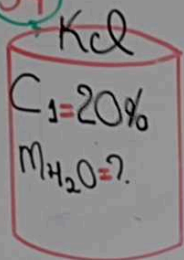
Orto



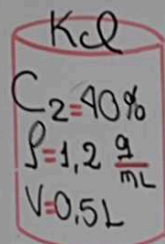
atomicidad



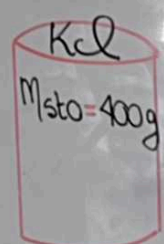
59



+



→



Solución 1

Solución 2

Solución Mezcla

Datos de la Solución 2

$$\rho = \frac{m}{V}$$

$$\Rightarrow m = \rho \times V$$

$$= 1,2 \frac{g}{mL} \times 500 mL$$

$$m_{sol II} = 600g$$

$$\% m_{sto} = \frac{m_{sto}}{m_{sol}} \times 100$$

$$40 = \frac{m_{sto II}}{600} \times 100$$

$$240 = m_{sto II}$$

Datos de la Solución de la Mezcla

$$m_{sto} = m_{sto I} + m_{sto II}$$

$$400 = m_{sto I} + 240$$

$$160g = m_{sto I}$$

Datos de la Solución 1

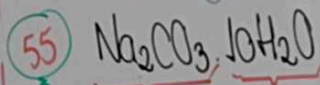
$$20 = \frac{160}{m_{sol I}} \times 100$$

$$m_{sol I} = 800g$$

$$m_{sol} = m_{sto} + m_{ste}$$

$$800 = 160 + m_{H_2O}$$

$$m_{H_2O} = 640g$$



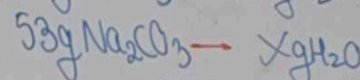
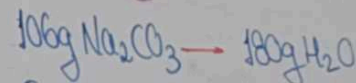
$$PF = 106$$

$$PM = 180$$

$$\Rightarrow PM = 286$$

$$10 = \frac{m_{sto}}{530} \times 100$$

$$m_{sto} = 53g$$



$$x = 90g$$

$$m_{sol} = 53 + 90 = 143$$

68

	$2NO_{2(g)} \rightleftharpoons N_2O_{4(g)}$	
Inicio	1mol	0 moles
Rx	2x	1x
Equilibrio	1-2x	x

$$V_0 = 1L \quad K_c = 1 \frac{L}{mol}$$

$$K_c = \frac{[N_2O_4]}{[NO_2]^2}$$

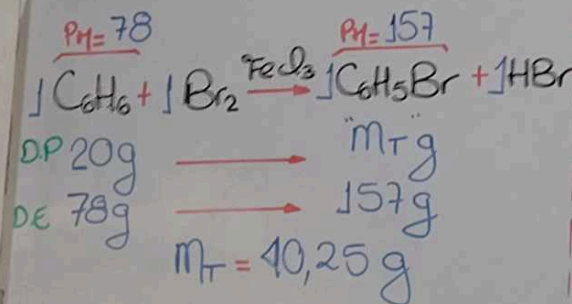
$$\frac{\left[\frac{x}{1}\right]}{\left[\frac{1-2x}{1}\right]^2} = 1$$

$$x = 1 - 4x + 4x^2$$

$$0 = 4x^2 - 5x + 1$$

$$x = \frac{1}{4}$$

52

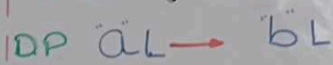
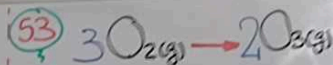


$$\% \text{P} = \frac{m_{\text{P}}}{m_{\text{T}}} \cdot 100$$

$$\% \text{P} = \frac{23}{40,25} \times 100 = 57,2\%$$



53



$$b = \frac{2}{3}a$$

a = Volumen de O_2 consumido

b = Volumen de O_3 formado

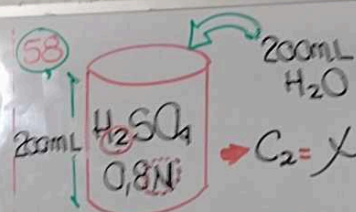
$$V_{\text{R}} = V_{\text{O}_2(\text{queda})} + V_{\text{O}_3}$$

$$10 = (12 - a) + b$$

$$10 = 12 - a + \frac{2}{3}a$$

$$a = 6\text{L}$$

58



$$C_1 V_1 = C_2 V_2$$

$$0,8 \times 200 = X \cdot 400$$

$$0,4\text{N} = X \checkmark$$

$$V_{\text{agregar}} = \frac{400}{4} = 100$$

Nuevo Volumen = 300mL

Vol H_2O agregado = 3(300) = 900mL

$$C_1 V_1 = C_2 V_2$$

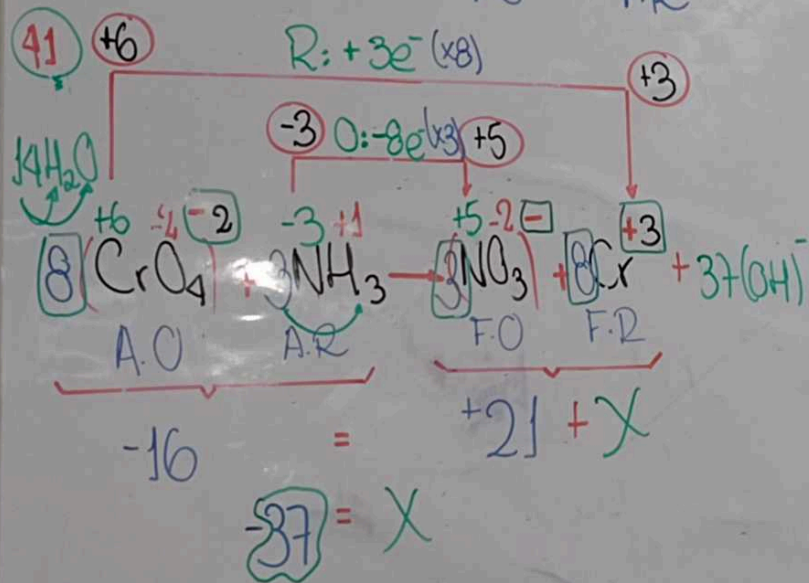
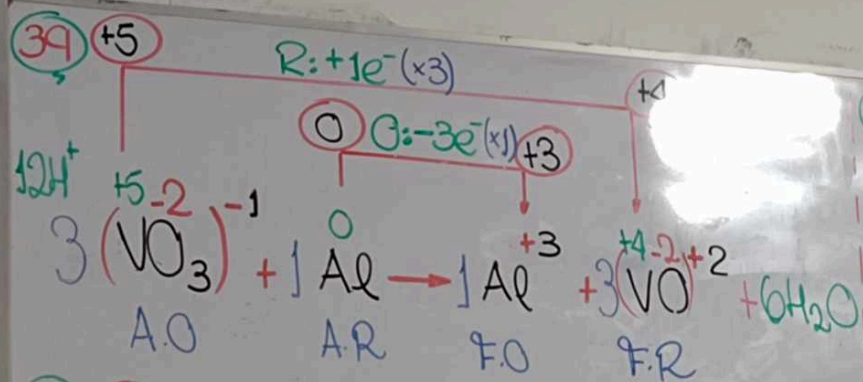
$$M_A V_1 = M'_A V_2$$

$$0,2 \times 300 = M'_A \times 1200$$

$$N = n \times \theta$$

$$0,4 = n_A \times 2$$

$$0,2 = n_A$$



44 18 Ktes $\rightarrow 75\% \text{Au} + 25\% \text{Ag}$
 $M_{\text{Au}} = 52,53\text{g} (0,75) = 39,39\text{g}$

197g Au \rightarrow NA Atomos
 39,39g Au \rightarrow X Atomos
 $X = 0,20 \text{ NA}$

43 $\text{I}_2\text{O}_5 \rightarrow \% \text{O} = 30\% \rightarrow \text{PM} = 2\text{I} + \text{O}$
 $\text{I}_2\text{O}_3 \rightarrow \% \text{I} = ? \rightarrow \text{PM} = 2\text{I} + 3\text{O}$

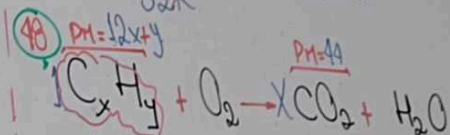
$$\% \text{E} = \frac{\# \text{at-g(e)} \times \text{PA(e)}}{\text{PM O PF}} \times 100$$

$$30 = \frac{1 \times \text{O}}{2\text{I} + \text{O}} \times 100$$

60I + 30O = 100O
 60I = 70O
 $\text{I} = \frac{7\text{K}}{6\text{K}}$

$$\% \text{J} = \frac{2 \times \text{J}}{2\text{J} + 3\text{O}} \times 100$$

$$\% \text{J} = \frac{14\text{K}}{32\text{K}} \times 100 = 43,75\%$$

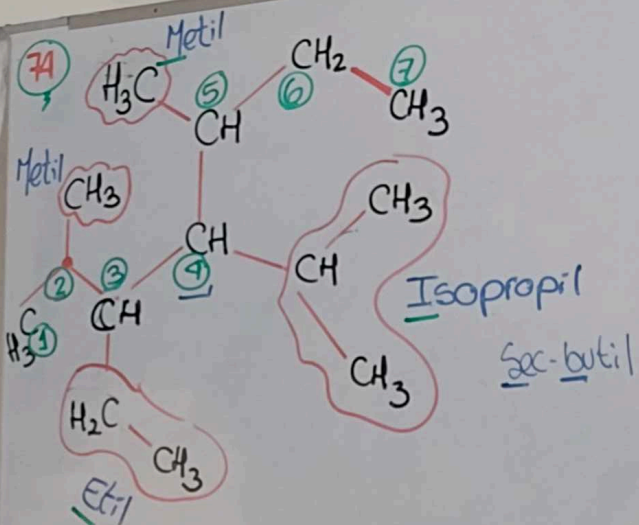


D.P. 1,2g \rightarrow 3,3g
 D.E. (12x+y)g \rightarrow 44x

x=1

y=4





3-Etil-4-Isopropil-2,5-dimetilheptano

