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Date: **Skills: ORR & AI**

Topic: Redox Reactions

Aim: To compare the reactions of oxidising and reducing agents

Instructions: Conduct the following tests, record observations, and write appropriate inferences.

Test	**Observations**	**Inferences**
I. Add a few drops of KI to $KMnO_4$	Colour changed from purple to orange/yellowish brown solution, with some small dark-coloured precipitate in there	Oxidant: $KMnO_4$ Reductant: KI
II. Add a few drops of $FeSO_4$ to $KMnO_4$	Changed colour from purple to colourless	Oxidant: $KMnO_4$ Reductant: $FeSO_4$
III. Add a few drops of H_2O_2 to $KMnO_4$	Effervescence, changed from purple to coloutless with a brown precipitate on the surface. There was heat produced	Oxidant: $KMnO_4$ Reductant: H_2O_2
IV. Add a few drops of H_2O_2 to KI	Colour changed from colourless to yellow. There is effervescenece with very small bubbles	Oxidant: KI Reductant: H_2O_2
V. Add a few drops of $FeSO_4$ to KI	A little thingy at the top, no visible change	Oxidant: N/A Reductant: N/A
VI. Add a few drops of H_2O_2 to $FeSO_4$	Effervescence, changed from colourless to pale yellow-green	Oxidant: H_2o_2 Reductant: $FeSO_4$

Conclusion: What is the order of the reagents in **increasing** oxidising power?

KI → $FeSO_4$→$KMno_4$ → H_2o_2