## Geochemical Thermodynamics - EAS 6211 - Fall 2013

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**Lectures:** TTh 12:05-1:25 pm, EST L1118

Office Hours: Ellery Ingall T, 10:00am-11:00am, W 10:30-11:30am

Grading: Problem Sets 20%

Project(difficulty, paper, presentation) 20%
Participation/Preparation 10%
Exams 50%

**Text:** Physical Chemistry for the Chemical and Biological Sciences

By Raymond Chang, ISBN 1-891389-06-8

Web Page: Class materials will be posted on t-square

#### **Topics**

## **Relevance of Thermodynamics**

What is it?

**Example Applications** 

#### Setting up the Basics

Thermodynamic Terminology

Useful Math

#### **Fundamental Relationships**

Zeroth Law First Law Second Law

Combined First and Second Laws

Third Law

Statistical Interpretation of Entropy

Equations of State

# **Simple Systems**

Free Energy of Formation Clapeyron Equation Phase Rule

**Phase Diagrams** 

## Thermodynamics of Solutions and Gases

Raoults and Henry's Laws Partial Molar Properties Gibbs-Duhem Relation Activity, Fugacity Standard States Chemical Potential

## Dealing with High P and T

The Easy Way
The Correct Way

#### Mineral Equilibria

**Equilibrium Constants** 

Pressure, Temperature Dependence

**Chemical Potential Diagrams** 

## **Solid Solutions**

Order disorder Non-ideal Solutions

## Real Calculations in Aqueous Systems

Activity-Composition Relations Debye-Huckel, Davies Equations

Ion pairs

## **Redox Systems**

Nernst Equation pe-pH diagrams

## Thermodynamic Software

Phreeac

Application to your work