

Ryan Jaipersaud  
Final Project Report  
ChE 352  
May 11<sup>th</sup>, 2017  
Hours Spent: 60 hr

Ryan Jaipersaud  
ChE 352  
Final Project Part I

Memorandum

TO: Mr. Sean F. Askebom

FROM: Ryan Jaipersaud

DATE: May 11, 2017

SUBJECT: Fitting vapor pressure to an explicit form

A numerical method was needed to produce reliable results for vapor pressure over a wide range of temperature values without the need for finding the vapor pressure at each temperature through experimentation which is time consuming and costly. The vapor pressure of four species in bars was fitted using the cftool window in matlab. Polyfit was initially used but, it returned negative pressures at temperatures approaching 170 K for some species as well as larger relative errors compared to cftool. Lagrange polynomials produced large relative errors for temperatures that were not used to construct the polynomial and were time consuming to derive. Cftool uses regression to fit the data to an explicit equation which was faster and gave lower error than the two methods previously mentioned. The equation used was the Antoine equation as shown below.

$$\ln P = A - \frac{B}{T + C}$$

The natural log of vapor pressure for each species at given temperatures from Perry's Handbook was calculated and given to the cftool. From there the cftool used regression to solve for parameters A, B and C such that error is minimized.  $R^2$  equaled 1 for all four fits. The equations for species A (ethane), C (vinyl chloride monomer), D (isobutene), and E (1-butene) are shown below.

[VaporPressure] = exp(9.568-(1706/(T-6.064))), for species A between 170 and 305.4 K

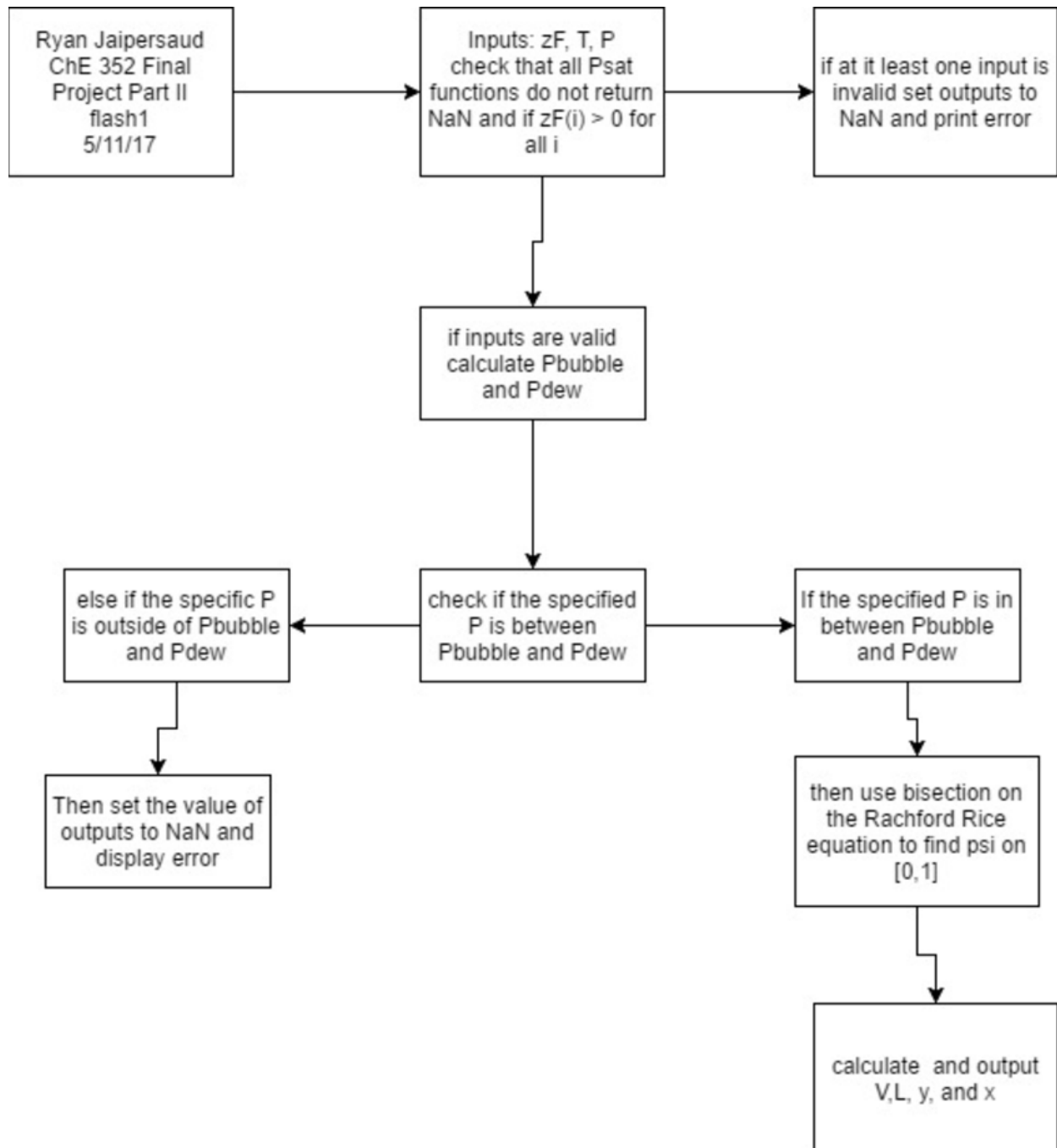
[VaporPressure] = exp(9.001 -(1946/(T-43.15))), for species C between 170 and 432 K

[VaporPressure] = exp(9.186 -(2128/(T-29.53))), for species D between 170 and 408 K

[VaporPressure] = exp(9.331 -(2190/(T-32.18))), for species E between 170 and 418.6 K

The vapor pressure equations for each species returned lower relative errors over the temperature used to construct each fit than those of polyfit. The equations for species A, C, D and E can be accessed by making functions calls to Psat1(T), Psat2(T), Psat3(T), and Psat4(T), respectively. If the input temperature is not within the ranges specified above the functions will return NaN. The number of points that went into constructing each fit were 8, 13, 12, and 11, respectively. The relative error at each point was calculated. The average percent relative error over the number of points to make each curve was calculated to be 0.624, 0.776, 0.827 and 0.68 %, respectively.

Ryan Jaipersaud  
ChE 352  
Final Project Part II  
flash1

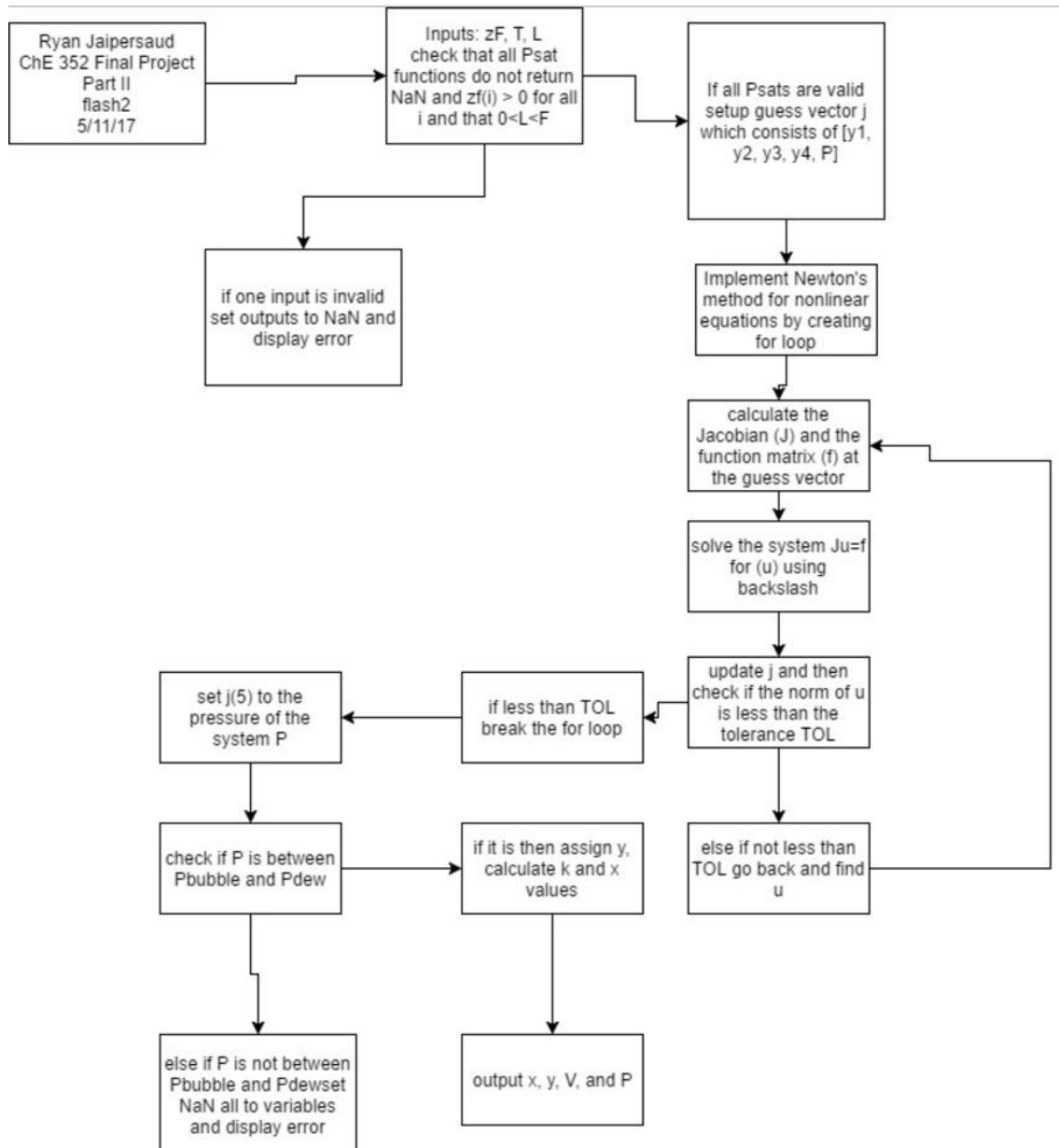


Ryan Jaipersaud  
ChE 352  
Final Project Part II  
flash1

When implementing flash1, the user must input the feed, operating temperature and pressure of the flash unit. The program needs to check if a feed exists and that each element in  $z_F$  is positive which it does by taking the summation of  $z_F$  and by directly checking each element. To make sure that all vapor pressures are valid function calls to  $Psat\#$  are made. Should the specified temperature not be within the bounds for each species the functions will return NaN. The “isnan()” function takes in an arguments that returns 1 if NaN is passed in and this is used to make sure all  $Psat$  functions are valid and don't fall through the code which would give error messages.

To determine whether or not the conditions specified lead to flashing the bubble point and dew point pressures referred to as  $P_{bubble}$  and  $P_{dew}$ , respectively, are calculated. A check is made to make sure the pressure input ( $P$ ) falls in between these values. If ( $P$ ) is between  $P_{bubble}$  and  $P_{dew}$  then the code will proceed to build the Rachford Rice equation [1]. This is an equation of one unknown,  $\psi$ , and therefore bisection is used to determine  $\psi$ . The bounds of  $\psi$  are fixed from 0 to 1 since  $\psi$  is defined as vapor fraction. The code then calculates  $x$ ,  $y$ ,  $V$ , and  $L$ . If  $P$  is not between  $P_{bubble}$  and  $P_{dew}$  an error message will be displayed and all outputs are set to NaN. The code outputs  $x$ ,  $y$ ,  $V$ , and  $L$ .

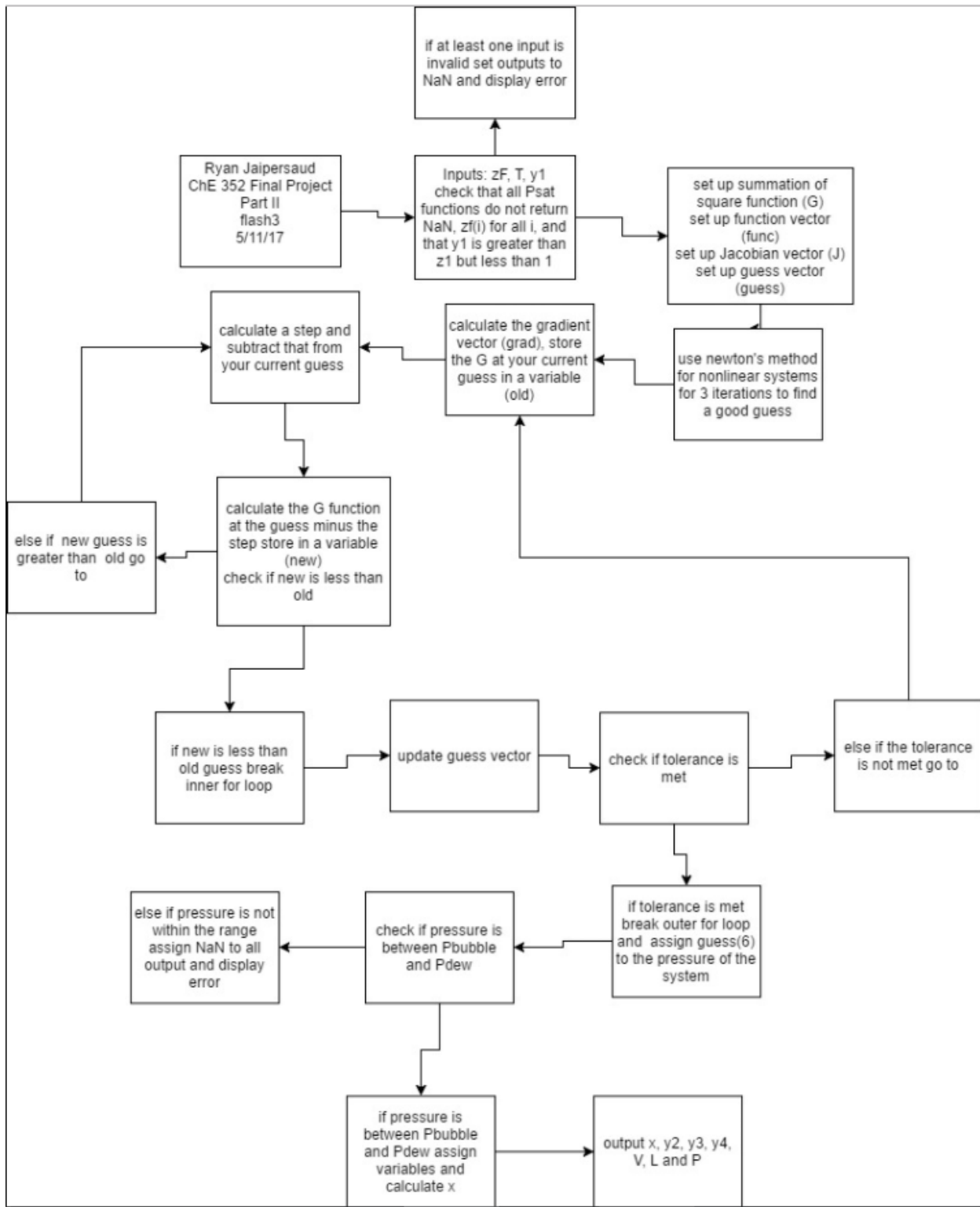
Ryan Jaipersaud  
ChE 352  
Final Project Part II  
flash2



Ryan Jaipersaud  
ChE 352  
Final Project Part II  
flash2

When implementing flash2, the user must input the feed, operating temperature and liquid exit stream of the flash unit. Flash2 starts off the same way as flash1 making the same checks for feed and Psat# functions. But, now it also checks to see if L is between 0 and F. V can be calculated without the need to do anything. This method uses Newton's method for nonlinear equations to solve the system of equations. In order to do this a guess vector must be set up which consists of,  $j = [y_1, y_2, y_3, y_4, P]$ . As a result the Jacobian matrix (J) for the problem is a  $5 \times 5$  and the function vector (f) is  $5 \times 1$ . The system  $Ju=f$  is solved repeatedly for (u) using the backslash operator until the two norm of (u) is less than the  $10^{-11}$ . The guess vector is repeatedly updated by subtracting (u) from the original guess. Once the tolerance for (u) is met the code breaks out of the for loop and assigns j(5) to the pressure of the system. The code then checks if the pressure is between the bubble and dew point pressures. If it is then y is assigned appropriate values and k and x values are calculated. If it is not between the points then set all outputs to NaN. The code outputs x, y, V, P.

Ryan Jaipersaud  
ChE 352  
Final Project Part II  
flash3



Ryan Jaipersaud  
ChE 352  
Final Project Part II  
flash3

When implementing flash3, the user must input the feed, operating temperature and the light key of the flash unit. In addition to checking Psat functions and the feed, y1 must also be checked to see if it is greater than z(1) [explanation in Appendix 1]. A (guess) vector is first made consisting of guess = [y2; y3; y4; V; L; P]. Once all checks are met, a 6 x 6 Jacobian matrix function (J), a 6 x 1 function vector (func), and a summation of squares of the elements in (func) function (G) are created. Newton's method for three iterations is first used to create a decent starting guess for steepest decent. The gradient (grad) and (G) are both calculated at the current guess according to [2]. (G) at the current guess is referred to as (old). A nested for loop is then created to find a step that will decrease (G). (G) is calculated at the guess vector minus the step (new). Once (new) is less than (old) a break statement will get the user out of the nested loop and the (guess) vector will be updated by the step that decreased G. A check is then made to see if the step compared to the guess is lower than the tolerance ( $10^{-10}$ ). If it is lower, a break statement will take the user out of the outer for loop. Guess(6) is assigned to the pressure and the pressure is checked to see if it falls between the bubble and dew point pressures. If the pressure is between the bubble and dew point the code will assign variables and then calculate x. The code outputs x, y2, y3, y4, V, L, and P.



Ryan Jaipersaud  
ChE 352  
Final Project Part III

- 1.) The detector must be 5.9 cm from the wall.
- 2.) This is a second order parabolic differential equation since the determinant of the equation is zero. The Neumann boundary condition in this problem changes the A and B matrixes which did not occur with the differential equations covered in class.

$$D_{MA} \frac{d^2 C}{dx^2} = \frac{dC}{dt}$$

$$\det(X) = \det \left( \begin{vmatrix} D_{MA} & 0 \\ 0 & 0 \end{vmatrix} \right) = 0$$

- 3.) Three boundary conditions are needed, one for time and two for x. They are as follows:  
 $C(x,0) = 0$ ,  $C(0,t) = 18$  and  $\frac{dC}{dx} = 0$  at the wall.
- 4.) Since the detector needs to be within 5.9 cm to detect a leak of  $0.8 \text{ mol/m}^3$  after 10 seconds of methane, which is a fast diffusing gas, diffusion is deemed slow.

Ryan Jaipersaud  
ChE 352  
Final Project

References

- 1.) Seader, J; Henley, E; Roper, D *Separation Process Principles Chemical and Biochemical Operations*, 3<sup>rd</sup> edition; John Wiley & Sons: Hoboken, NJ, 2011; pp 139 – 170
- 2.) Faires, J; Burden, R; *Numerical Methods*, 4<sup>th</sup> edition; Cengage Learning, 2012; pp 427 – 431

Ryan Jaipersaud  
ChE 352  
Final Project

Appendix 1

The equation below can be derived from mass balance.

$$psi = \frac{z1 - x1}{y1 - x1}$$

Since the bounds on psi are from 0 to 1, the only time psi will equal 1 is when y1 is equal to z1. Therefore this serves as a lower bound for what y1 can be inserted as. If y1 is less than z1, psi becomes physically useless.