

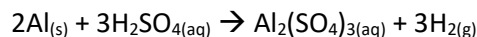
### Extra practice on Stoichiometry

How many moles of  $\text{NH}_3$  form given 65.3 grams of  $\text{N}_2\text{H}_4$ ?

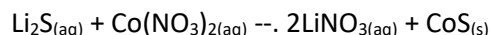


Iron (III) oxide reacts with carbon monoxide. If a reaction mixture contains 22.55 grams of  $\text{Fe}_2\text{O}_3$  and 14.78 grams of  $\text{CO}$ . Once the reaction is complete, how much excess reactant remains?

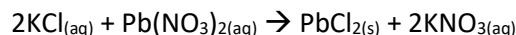
What is the minimum amount of 6.0 M  $\text{H}_2\text{SO}_4$  necessary to produce a 25.0 gram sample of  $\text{H}_2$  according to the balanced reaction?



What volume of 0.150 M  $\text{Li}_2\text{S}$  solution is required to completely react with 125 mL of 0.150 M  $\text{Co}(\text{NO}_3)_2$



A 25 mL sample of 1.20 M potassium chloride solution is mixed with 14.0 mL of 0.900 M lead (II) nitrate solution and this precipitation reaction occurs:



Determine the theoretical and percent yield if 2.45 grams of precipitate is collected.

A 55 mL sample of 0.102 M potassium sulfate solution is mixed with 35.0 mL of 0.114 M lead (II) acetate solution. The solid is collected, dried, and found to have a mass of 1.01 g. Determine the limiting reactant, theoretical yield, and the percent yield.