

# Modeling Molecular Systems

MSSE Bootcamp

August 12, 2020

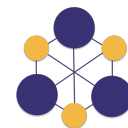
# Monte Carlo Connection to Molecular Systems

According to statistical mechanics

We can use MC to evaluate this integral!

$$\langle Q \rangle = \int_V Q(r^N) \rho(r^N) dr^N$$

- $Q$  quantity which depends on atomic coordinates ( $r^N$ )
- $\langle Q \rangle$  average value of quantity  $Q$  (square brackets denote average)
- $r^N$  atomic coordinates with  $N$  dimensions (x, y, z is three dimensions)
- $\rho(r^N)$  probability density based on thermodynamic properties (beyond scope of this course)



# Monte Carlo Connection to Molecular Systems

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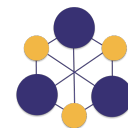
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This integral gets complicated very quickly. Consider a system of 10 atoms in 3 dimensions.

3 dimensions x 10 atoms = 30 dimensional integral!

This integral cannot be solved analytically, but we can use Monte Carlo integration to evaluate it.

Today, we will build our model for our thermodynamic quantity,  $Q$ .



# The Lennard Jones Potential

The Lennard Jones Potential is an equation that is often used to model the interaction energy of nonbonded atoms:

$$Q = U(r) = 4\varepsilon \left[ \left( \frac{\sigma}{r} \right)^{12} - \left( \frac{\sigma}{r} \right)^6 \right]$$

This interaction is **pairwise**, meaning it occurs between two particles.

$r$  – distance between two particles

$\varepsilon$  – strength of particle interaction

$\sigma$  – particle size

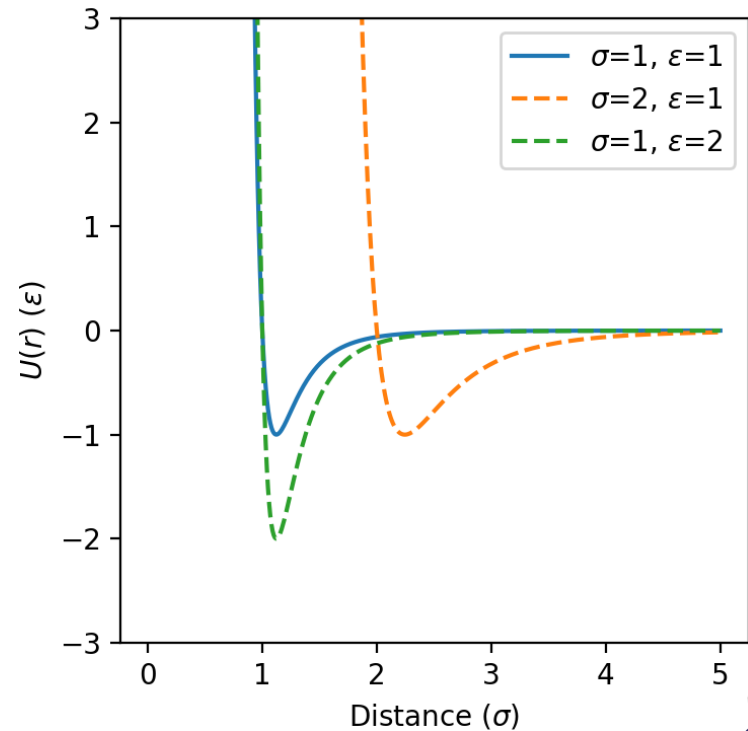
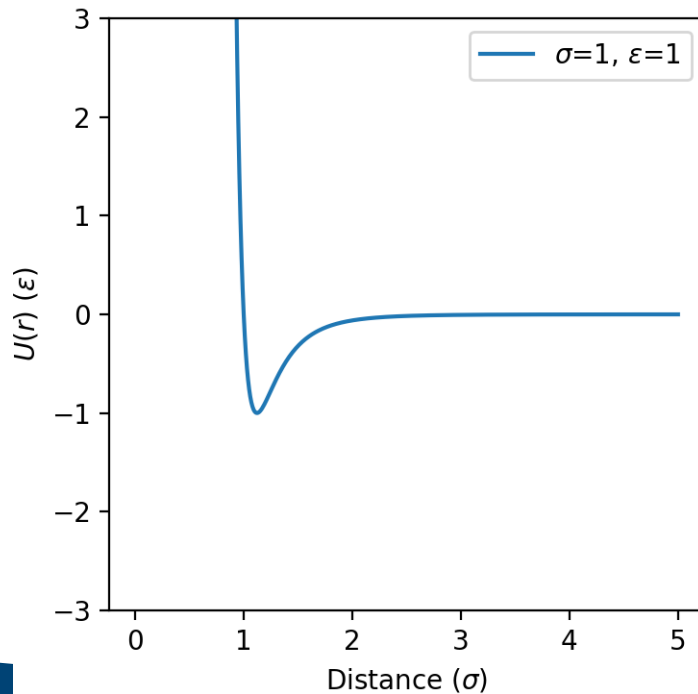
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identity



# The Lennard Jones Potential

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# Reduced Units

For Argon,

$$\varepsilon = 120 K (k_B) = 1.68 \times 10^{-21} J \text{ and } \sigma = 3.4 \times 10^{-10} \text{ meters}$$

These are really inconvenient numbers!

We will normalize our energy by  $\varepsilon$  and our distances by  $\sigma$ .

$$U^*(r) = \frac{U(r)}{\varepsilon}$$

$$r^* = \frac{r}{\sigma}$$

$$U^*(r^*) = 4 \left[ \left( \frac{1}{r^*} \right)^{12} - \left( \frac{1}{r^*} \right)^6 \right]$$

This will make  
 $U^*(r^*)$  be on  
the order of 1.

