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b) O—H has the stronger bond c) O is smaller than S. ChemActivity 35 1. a) —3T.01 kJ b) —696.56 kJ c) —6T.18 id 3. a) C13 has the strongest bond (all are single bonds) and should have the most negative Maui; .

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The C—O bond order in formaldehyde is two. The C—O bond order in the carbonate ion is 4/3. Therefore, the C—O bond length in the carbonate ion should be closer to 143 pm than to 116 pm. We can calculate a bond length by assuming that bond length vs. bond order is linear between bond orders one and two.

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WHENEVER POSSIBLE, Ms. Myriah Day will post the HOMEWORK KEYS ON THIS PAGE. PLEASE NOTE: If you have a question about these answers, it is your responsibility to come to office hours or ask during class work time.

**HW Keys - Roosevelt High School AP Chemistry 2017-18**

Based on your answers to CTQs 4 and 5, ... 32 ChemActivity 5 The Shell Model (II) Information Related eBooks: Wpf Tutorial For Dependency Properties Soil Testing Laboratory Manual December 8 2009 Memorandum For Division Dodea Bca Ist Year ... Answers To Chemactivity 32 Molarity

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ChemActivity 1 The Nuclear Atom 3 Critical Thinking Questions 11. —How many protons are found in 2C? 13C? C? 122. 13How many neutrons are found in C? 13C? C—? 133. 13How many electrons are found in 12C? C? C—? 4. Based on your answers to CTQs 1-3, what do all carbon atoms (and ions) have in common? 5.

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**Chem 124 ChemActivity Homework - Wofford College**

Use the Ideal Gas Law and the answers to questions 12, 13, and 14 to calculate the number of moles of hydrogen gas present.  $R = 0.0821 \text{ L} \cdot \text{atm/mol} \cdot \text{K}$  (watch sig figs.) In a Hardware Model Lab, a student counted out a dozen large pins symbolizing Mg und found th.il they

weighed 24.501 g.

**Use The Ideal Gas Law And The Answers To Questions ...**

ChemActivity 33 Problems 1. The Ideal Gas Law 189 2. 3. Consider the three flasks in the figure below. Assume that the connecting tubes have no volume and the temperature is held constant. a) Calculate the partial pressure of each gas when all stopcocks are open.

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