

Analysis Of Commercial Bleach Lab Answers

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Analysis Of Commercial Bleach Lab

Partner: Judy 29 March 2012. Analysis of a Commercial Bleach. Purpose: The purpose of this lab is to determine the amount of sodium hypochlorite (NaClO) in commercial bleach. accurate, starch was added to help determine the endpoint of the solution.

Analysis of a Commercial Bleach Purpose

A solution of sodium thiosulfate of a known concentration will be added to the bleach using a buret in a titration. An average of 17.2 mL of sodium thiosulfate was used in the titration, resulting in a calculation of 4.7% of sodium hypochlorite in the commercial bleach.

Analysis of a Commercial Bleach Laboratory #10

Haziq Rauf Jacky Chan Analysis of Commercial Bleach Lab I. 2. A solution of household vinegar (a mixture of acetic acid and water) is to be analyzed. A pipet is used to measure out 10.0 mL of the vinegar, which is placed in a 250 mL volumetric flask. Distilled water is added until the total volume of solution is 250 mL.

Analysis of Commercial Bleach Lab | Redox | Titration

Analysis of a Commercial Bleach Purpose: The purpose of this lab is to determine the amount of sodium hypochlorite (NaClO) in commercial bleach. This can be done by forming triiodide ions. To make the measurement more accurate, starch was added to help determine the endpoint of the solution.

Analysis of Commercial Bleach Lab Essay - 1708 Words ...

Analysis of a Commercial Bleach Lab. Introduction Many commercial products are effective because they contain oxidizing agents. Some products which contain oxidizing agents are bleaches, hair coloring agents, scouring powders, and bathroom cleaners. The most common oxidizing agent in bleaches is sodium hypochlorite, NaClO (also written NaOCl).

Analysis of a Commercial Bleach Lab - FHS AP Chemistry

Analysis of Commercial Bleach Lab I. Purpose In this experiment, the amount of sodium hypochlorite in a commercial bleach will be determined by reacting it with sodium thiosulfate in the presence of iodide ions and starch. A solution of sodium thiosulfate of known concentration will be added to the bleach using a buret in a titration procedure.

Analysis of Commercial Bleach Lab Essay - 1681 Words ...

Analysis of a Commercial Bleach Lab - Analysis of a... Pipet 25.0 mL of the dilute bleach into an Erlenmeyer flask. Add the solid KI and about 25 mL of distilled or deionized water. Swirl to dissolve the KI. Working in a fume hood, add approximately 2 mL of 3 M HCL while stirring the solution.

Analysis of a Commercial Bleach Lab - Course Hero

Procedure 2. Titration of the Liquid Bleach a) To a clean 125 mL or 250 mL Erlenmeyer flask add 30 mL deionized water and 0.5 g KI (about 1 small scoop). Swirl until the KI has dissolved. Add 10 mL 1M HCl using your small graduated cylinder and swirl to mix.

Experiment 2: Analysis of Commercial Bleach Solutions

Experiment 8 - Analysis of Hypochlorite in Bleach. 8-2. The endpoint of the titration is when the last trace of yellow color from the iodine disappears. Because this change from pale yellow to colorless is not very distinct, a small quantity of starch solution is added near the endpoint.

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