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2 8. FkBr 9. PH 3 O 4 10. SO 3 11. e 2O 3 12. Na 2O Convert the following measurements. Show all work, including units that cancel. 13. 18.2 g ? mol 14. 98.5×10^{24} mol ecu sNO 2 mol 15. 82.6 L of ne at STP mol 16. 4.1 g of Na 2O mol 17. .3 mol SO 3 it r 18. 1.4×10^{24} atoms f K l Chemistry Calculation Review Name _____ Chem Worksheet 12-1

Chemistry Calculation Review Name Chem Worksheet 12-1

What volume is needed to make 2.00 L of 1.00 M solution? 4. Concentrated HCl is 12.0 M. What volume is needed to make 2.00 L of 1.00 M solution? 5. A 0.500 M solution is to be diluted to 500.0 mL of a 0.150 M solution. How many mL of the 0.500 M solution are required? 6. A stock solution of 10.0 M NaOH is prepared.

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2O 2 + O 2 + H 2O 4. How many liters of oxygen at STP are needed to react with 5.2×10^2 molecules of hydrogen sulfide? H 2S + O 2 SO 2 + H 2O 5. What mass of chlorine gas is necessary to synthesize 258 L of hydrogen chloride at STP? H 2 + Cl 2 HCl 6. If there are 6.2×10^2 molecules of calcium carbide, CaC 2, what mass ...

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