

Light Waves Chem Worksheet 5 1 Answer Key

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Light Waves Chem Worksheet 5

waves determines the type of light. All of the various light waves move with the same speed, a value abbreviated (c) equal to 3.00×10^8 m/s. The frequency that light waves pass a given point is ... Chem Worksheet 5-1 $1 \text{ m} = 1 \times 10^3 \text{ mm}$ $1 \text{ m} = 1 \times 10^6 \mu\text{m}$ useful equations

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Answer to Section 4.1a Light Waves Name ____ Chem Worksheet 5-1 The behavior of light indicates that it is comprised of waves. The distance between...

Section 4.1a Light Waves Name _____ Chem ...

The behavior of light indicates that it is comprised of waves. The distance between successive waves is called the wavelength (λ) and the wavelength determines the type of light. The size of the waves determines the type of light. All of the various light waves move with the same speed, a value abbreviated (c) equal to 3.00×10^8 m/s. The ...

CHEMISTRY LEE UNIT FOUR - CHEMSTEM - Home

A visible light wave has a frequency of 7.5×10^{14} 1/s. Find the wavelength in nanometers (nm) and determine the color of the light. ... Chem Worksheet 5-2 $h\nu = E = hc/\lambda$ $E = hc/\lambda$ $\lambda = hc/E$ $\lambda = (6.63 \times 10^{-34} \text{ J}\cdot\text{s})(3.00 \times 10^8 \text{ m/s}) / (4.0 \times 10^{-19} \text{ J})$ $\lambda = 4.97 \times 10^{-7} \text{ m}$ $\lambda = 497 \text{ nm}$ Spectrum of Hydrogen and the Bohr Model 1. What does the ground state of an atom represent?

1. Draw a wave below and label the following parts: peak ...

6. A ruby laser produces red light that has a wavelength of 500 nm. Calculate its energy in joules. 7. What is the frequency of UV light that has an energy of 2.39×10^{-18} J? 8. What is the wavelength and frequency of photons with an energy of 1.4×10^{-21} J? Planck's Equation Name _____ Chem Worksheet 5-2 $h\nu = E = hc/\lambda$ $E = hc/\lambda$ $\lambda = hc/E$ $\lambda = (6.63 \times 10^{-34} \text{ J}\cdot\text{s})(3.00 \times 10^8 \text{ m/s}) / (1.4 \times 10^{-21} \text{ J})$ $\lambda = 1.42 \times 10^{-4} \text{ m}$ $\lambda = 142 \text{ nm}$

Planck's Equation Name Chem Worksheet 5-2

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