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CHAPTER 3 REVIEW Atoms: The Building Blocks of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Explain the difference between the mass number and the atomic number of a nuclide. Mass number is the total number of protons and neutrons in the nucleus of an isotope.

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26. According to the kinetic-molecular theory, the particles in a liquid can change relative positions but still are influenced by attractive forces.

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REVIEW Chemical Bonding MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. a.

6 Chemical Bonding - Effingham County Schools / Overview

40. 7.05 1016 Hz 41. 1.28 s 42. 4.58 10¹⁹ J The Periodic Law, pp. 36–45 TEST A 1. b 2. d 3. b 4. 5. d 6. a 7. b 8. 9. b 10. a 11. c 12. a 13. c 14. d 15. c 16. b

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Holt McDougal Modern Chemistry 2 Chapter Test Chapter Test A, continued ____ 3. The periodic law states that a. no two electrons with the same spin can be found in the same place in an atom.

b. the physical and chemical properties of the elements are functions of their atomic number. c. wave patterns repeat at regular intervals.

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CHAPTER 10 REVIEW States of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Match description on the right to the correct crystal type on the left. b ionic crystal (a) has mobile electrons in the crystal c covalent molecular crystal (b) is hard, brittle, and nonconducting a metallic crystal (c) typically has the lowest melting point of the four

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