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Chapter 16 Solutions 167 SECTION 16.1 PROPERTIES OF SOLUTIONS (pages 471-477) This section identifies the factors that affect the solubility of a substance and determine the rate at which a solute dissolves. Solution Formation (pages 471-472) Look at Figure 16.1 on page 471 to help you answer Questions 1 and 2. 1.

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What is the molarity of a solution that contains 212.5 g of sodium nitrate (NaNO_3) in 3.0 liters of solution? 4. You must prepare 300.0 mL of 0.750M NaBr solution using 2.00M NaBr stock solution. How many milliliters of stock solution should you use? 5. In order to dilute 1.0 L of a 6.00M solution of NaOH to 0.500M solution, how much water ...

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Pearson Accelerated Chemistry Chapter 16: Section 1: Properties of Solutions

Chapter 16 Solutions 403 Section Review Objectives • Identify the three colligative properties of solutions • Describe why the vapor pressure, freezing point, and boiling point of a solution differ from those properties of the pure solvent. Vocabulary • colligative properties • freezing-point depression • boiling-point elevation Part A Completion Use this completion exercise to check ...

05 CTR ch16 7/12/04 8:14 AM Page 399 PROPERTIES OF ...

Chapter 16: Colligative Properties of Solutions 45 16-4. The mole fraction of $(\text{NH}_4)_2\text{SO}_4(\text{aq})$ is given by $x_{(\text{NH}_4)_2\text{SO}_4} = \frac{n_{(\text{NH}_4)_2\text{SO}_4}}{n_{(\text{NH}_4)_2\text{SO}_4} + n_{\text{H}_2\text{O}}}$ Because $(\text{NH}_4)_2\text{SO}_4(\text{aq})$ is a strong electrolyte, it dissociates completely into $\text{NH}_4^+(\text{aq})$ and $\text{SO}_4^{2-}(\text{aq})$ ions. Assume a one kilogram solution. The number of moles of ions in one ...

CHAPTER 16. Colligative Properties of Solutions

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Section 16.1 Properties of Solutions 473 What is happening? Particles move from the solid into the solution. Other dissolved particles move from the solution back to the solid. Because these two processes occur at the same rate, no net change occurs in the overall system. As Figure 16.2 illustrates, a state of dynamic equilibrium

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