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Class 12 Chemistry. Electrochemistry InText Solution. Question 1: How would you determine the standard electrode potential of the system  $\text{Mg}^{2+} | \text{Mg}$ ?. Answer: We know that it is not possible to measure the electrode potential of a half cell. We can only measure the difference between electrode potentials of two half cells, which gives the cell potential of the cell.

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Recognize that  $\text{Cu}^{2+}$  is the stronger oxidizing agent (see Table P2) and will be more likely to be reduced. Thus, the  $\text{Cu}^{2+}$  half reaction is where copper is reduced and will be the reduction half-cell. The  $\text{Zn}^{2+}$  half reaction will be the oxidation half-cell because Zinc gets oxidized at the anode.

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