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THE BEHAVIOR OF GASES SECTION 14.1 PROPERTIES OF GASES (pages 413–417) This section uses kinetic theory to explain the properties of gases. This section also explains how gas pressure is affected by the amount of gas, its volume, and its temperature. Compressibility (pages 413–414) 1. Look at Figure 14.1 on page 413.

SECTION 14.1 PROPERTIES OF GASES(pages 413–417) - MAFIADOC.COM

Objectives & Standards 14.1: Properties of Gases By: Arienne Rancudo & John Molina Gases can expand to fill a container. Compressibility: Measure of how much the volume decreases under pressure. Gases are easily compressed. Particles in a gas are filled together under pressure.

14.1: Properties of Gases by Arienne Rozel Rancudo on Prezi

Properties of Gases: 1. A gas has no definite shape or volume of its own. It acquires the shape of the container. ADVERTISEMENTS: Reason: Intermolecular attraction is the weakest in gases whereas intermolecular separation is the largest. Hence, molecules in a gas move very fast and the gas expands to fill all the space available.

What are the Properties of Gases? - preservearticles.com

What Are Five Properties of Gases? Low Density. Gases contain scattered molecules that are dispersed across a given volume... Indefinite Shape or Volume. Gases have no definite shape or volume.The random movement... Compressibility and Expandability. The low density of gases makes them ...

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properties of gases lead to some unique uses of them. For example, because the particles in a gas are relatively far apart, gases are poor thermal conductors and are used for insulation between the glass panes in double-paned windows. FYI In this section, students look at the relationships between the variables P, V, T,

14.1 Properties of Gases 14 - Henry County School District

CEM 141 – General Chemistry Description: Atoms, molecules, ions; chemical calculations; reactions, energy changes, gases; periodic properties of elements; chemical bonds; states of matter, solutions; acids and bases; aqueous reactions and ionic equations. Credit: 4 hours (3 hours of lecture and 1 hour of recitation)

CEM 141 - General Chemistry

THE PROPERTIES OF GASES 14.1 Section Review Objectives why gases are easier to compress than solids or liquids are Describe the three factors that affect gas pressure Vocabulary compressibility Part A Completion Use this completion exercise to check your understanding of the concepts and

terms that are introduced in this section.

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SECTION 14.1 PROPERTIES OF GASES(pages 413–417) This section uses kinetic theory to explain the properties of gases. This section also explains how gas pressure is affected by the amount of gas, its volume, and its temperature. Compressibility (pages 413–414) 1.

SECTION 14.1 PROPERTIES OF GASES(pages 413–417)

Gases are easily compressed because of the space between the particles in a gas. Compressibility
•The volume of the particles in a gas is small compared to the overall volume of the gas. •The distance between particles in a gas is much greater than the distance between particles in a liquid or solid.

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