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Rank the following elements from smallest to largest electronegativity based on the trends you have discovered thus far in the periodic table: barium (atomic number 56), bromine (atomic number 35), and iron (atomic number 26).

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15 (6 Oâ(ÜIQ 2. In general, what is the trend in atomic radius as you go down a group in Model 1? Support your answer, using examples from three groups. t,xreaSeS . H K 2 B (33) (3. Using your knowledge of Coulombic attraction and the structure of the atom, explain the trend in atomic radius that you identified in Question 2.

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HS-PS1-1: Use the periodic table as a model to predict the relative properties of elements based on the patterns of electrons in the outermost energy level of atoms. Students will be using experimental data to infer the periodic trends for atomic radius, first ionization energy and electronegativity.

Ninth grade Lesson Periodic Trends- Reciprocal Style Day 1

Chem 115 POGIL Worksheet - Week 10 Periodic Trends Why? The size of an atom or ion and the attraction between the nucleus and the outermost electrons play important roles in determining the chemistry of an element. Knowing the trends in atomic and ionic sizes, ionization energies, and electron affinities aids in understanding chemical behavior

Chem 115 POGIL Worksheet - Week 10 Periodic Trends Why ...

ANSWER KEY Trends on the Periodic Table Part 1: Identifying chemical properties/characteristics of the elements For each of the following elements, draw a caricature or cartoon above the description to show its "chemical personality" based on the description provided. (Note: See the Neon example for ideas.) a "magnetic" personality.

ANSWER KEY Trends on the Periodic Table - rocklin.k12.ca.us

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Periodic Trends 3 6. Locate the numbers in Model 1 that represent the ionization energy. The ionization energy is the amount of energy needed to remove an electron from an atom. a. Using your knowledge of Coulombic attraction, explain why ionization—removing an electron from an atom—takes energy. b.

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Periodic(Trends(Worksheet(((Answer Key 1. Circle)the)element)withthe)largest)atomic(radius)andput)a)square)aroundthe)element)withthe)smallest) atomic(radius:))) Cu) K) Ni) Br))) Largest – K Smallest - Br) Explainwhy)you)madethesechoices.) Atomic radius decreases as you go left to right across a period. Potassium is in the far

Periodic Trends Worksheet Answers

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Periodic trends: Electronegativity answers. Name ____ 1. What is the electronegativity of an element? It is the tendency of a bonded atom to accept or to donate an electron. 2. Is the “electronegativity” the same thing as the “electron affinity”? No, the electronegativity relates to bonded atoms and their tendency to attract ...

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