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!\\jame - - - - - Date _ Class _ 11-2 Practice Problems (continued)13. Find the mass of sodium required to 20. What volumes of H₂S gas and oxygen are produce 5.68 L of hydrogen gas at STP necessary to produce 14.2 L of sulfur from the reaction described by the dioxide gas?

Name 11-2 Practice Problems - Harrison High ... - FlipHTML5

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0.001934 moles NaI x 2 moles NaCl = 0.001934 mole NaCl. 2 mole NaI. 0.001934 mole NaCl x 58.45 g = 0.11 g NaCl. 1 mole NaCl. Determine the mass of carbon dioxide, CO₂, produced when 0.85 g of butane, C₄H₁₀, reacts with oxygen according to the following equation: (Hint to balance you will have to multiply by 2) ... 11-2 Practice Problems ...

11-2 Practice Problems - Mr. Fischer.com

2. Identify the limiting reactant when 4.68 g of Fe reacts with 2.88 g of sulfur to produce FeS. 3. ... 11-3 Practice Problems ... Fraction Review Sheet - Quia

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11-3 Practice Problems 1. Identify the limiting reactant when 1.22 g of O₂ reacts with 1.05 g of H₂ to produce water. 2. Identify the limiting reactant when 4.68 g of Fe reacts with 2.88 g of sulfur to produce FeS. 3. Identify the limiting reactant when 5.87 g of Mg(OH)₂ reacts with 12.84 g of HCl to form MgCl₂ and water. 4.

11-3-Problems-2 - 11-3 Practice Problems 1 Identify the ...

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Jefferson Township High School Science Department. 11-3 Practice Problems Answer key. February 1st, 2013 | Author: kkula

11-3 Practice Problems Answer key

Name Date _ Class _ ----- 11-2 Practice Problems 1.Determine the mass of lithium hydroxide produced when 0.38 g of lithium nitride reacts with water according to the following equation: Li₃N + 3H₂O → NH₃ + 3LiOH 2. What mass of sodium chloride is produced when chlorine reacts with 0.29 g of sodium iodide? 3.

Stoich practice prob 11-2 with key.pdf - Yumpu

Chapter 11 Section 11.2 (continued) CONCEPTUAL PROBLEM 11.5 Answers 15. 2HI → H₂ + I₂ 16. HBr Use Visuals Conceptual Problem 11.5 Have stu-dents study the photograph in the problem. Explain that the explosive properties of dynamite are due to the rapid production of large amounts of gases. A related reaction is the decom-position of ...

11.2 Types of Chemical Reactions - bleiker.weebly.com

Practice Problems: Stoichiometry (Answer Key) Balance the following chemical reactions: a. 2 CO + O₂ → 2 CO₂ b. 2 KNO₃ → 2 KNO₂ + O₂ c. 2 O₃ → 3 O₂ d. NH₄NO₃ → N₂O + 2 H₂O e. 4 CH₃NH₂ + 9 O₂ → 4 CO₂ + 10 H₂O + 2 N₂ f. Cr(OH)₃ + 3 HClO₄ → Cr(ClO₄)₃ + 3 H₂O Write the balanced chemical equations of each reaction:

Practice Problems: Stoichiometry

AE 11-1 Practice Problems Answers. AE 11-2 Practice Problems Answers. AE Extra Stoichiometry Problems Answers. AE Stoich Practice Sheet Answers. AE 11-3 Practice Problems Answers. AE Percent Yield and Limiting Reactant 2.0 Answers. AE Extra Stoichiometry Problems 2 Answers.

AE Stoichiometry Answer Keys - Baumritter - Google Sites

15-4 Practice Problems (continued) 11. How many grams of silver would have to be dissolved in 1120 g of ethanol to lower the freezing point by 0.25 °C? K_f for ethanol is 1.99 °C/m. 12. What is the freezing point depression when 85.3 g of oxygen is dissolved in 1500 g of water? K_f for water is 1.86 °C/m.

15-4 Practice Problems - Weebly

Unit Practice Problem Answers Additional Practice Problem Answers 1) tenths 2) hundredths 3) thousandths 1) tenths 2) hundredths 3) thousandths 4) ten thousandths 5) hundred thousandths 4) ten thousandths 5) hundred thousandths 6) millionths 7) four-tenths 8) forty-three hundredths 6) millionths 7) six-tenths 8) sixty-three hundredths

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11. If 18.6 g of methanol is used to dissolve 2.68 g of $\text{Hg}(\text{CN})_2$ what is the molality of the solution? 6. ... 15-2 Practice Problems (continued) 13. What is the molality of a solution made by dissolving : 8.11 g of potassium sulfide (K_2S) in : 47.6 g of ethanol? 14. What is the molality of a solution

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