

## *113 Reactions In Aqueous Solution*

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In writing balanced net ionic equations, you must make sure that the ionic charge is balanced. Key Concept By using the general rule for solubility of ionic compounds, you can predict the formation of a precipitate. The equation above displays how to write equations involving

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11.3 Reactions in Aqueous Solution 19 > A net ionic equation shows only those particles involved in the reaction and is balanced with respect to both mass and charge. •Of the five types of reactions identified in this chapter, both single- and double-replacement reactions can be written as net ionic equations.

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Section 11.3 Reactions in Aqueous Solution 343 An ion that appears on both sides of an equation and is not directly involved in the reaction is called a spectator ion.When you rewrite an equation leaving out the spectator ions, you have the net ionic equation. The net ionic equation is an equation for a reaction in solution that shows

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11.3 Reactions in Aqueous Solution Net Ionic Equation. Net Ionic Equations. Sample Problem. An ion that appears unchanged on both sides of the reaction is a spectator ion. You can rewrite the equation leaving out the spectator ions and then you have the net ionic equation.

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11.3 Reactions in Aqueous Solution Reactions that occur in aqueous solutions are double-replacement reactions. The products are precipitates, water, or gases. Lesson Summary Net Ionic Equations Net ionic equations show what species present in solution actually are part of the chemical reaction.

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The reaction of aqueous solutions of silver nitrate with sodium chloride to form solid silver chloride and aqueous sodium nitrate is a double-replacement reaction. The reaction is shown in Figure 11.11.  $\text{AgNO}_3(\text{aq}) + \text{NaCl}(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{NaNO}_3(\text{aq})$  This is the way you have been writing equations involving aqueous solutions of ionic compounds.

### **11.3 Reactions in Aqueous Solution**

Several types of reactions occur in water. When water is the solvent for a reaction, the reaction is said to occur in aqueous solution, which is denoted by the abbreviation (aq) following the name of a chemical species in a reaction. Three important types of reactions in water are precipitation, acid-base, and oxidation-reduction reactions.

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The reaction of aqueous solutions of silver nitrate and sodium chloride to form solid silver chloride and aqueous sodium ... 11.3 Reactions in Aqueous Solution 27 > By using the general rules for solubility of ionic compounds, you can predict the formation of a precipitate.

### 11.3 Reactions in Aqueous Solution - Useful Advice

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