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2 16.2 Concentrations of Solutions > 7 Copyright © Pearson Education, Inc., or its affiliates. All Rights Reserved Molarity To calculate the molarity of a solution ...

How can you describe the concentration Chapter 16 of a ...
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Solutions can be described by their concen-tration, which is the amount of solute in a specific amount of that solution. These relationships, which include mass % 1m>m2, volume % 1v>v2, mass >volume % 1m>v2, and molarity 1M2, can be used to convert between the amount of a solute and the quantity of its solution. Solutions are also diluted by

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SECTION 16.2 CONCENTRATIONS OF SOLUTIONS (pages 480–486) This section explains how to solve problems involving molarity of a solution, how to prepare dilute solutions from more concentrated solutions, and what is meant by percent by volume and percent by mass. Molarity (pages 480–482) 1.

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Chapter 16
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Chapter 16
In addition to molarity, a number of other solution concentration units are used in various applications. Percentage concentrations based on the solution components' masses, volumes, or

both are useful for expressing relatively high concentrations, whereas lower concentrations are conveniently expressed using ppm or ppb units.

3.4 Other Units for Solution Concentrations – Chemistry
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16.2 Concentrations of Solutions
Chapter 16 - Solutions - 16.2 Concentrations of Solutions - 16.2 Lesson Check - Page 531: 19 Answer To calculate the molarity of a solution, divide the moles of solute by the volume of the solution.

Chapter 16 - Solutions - 16.2 Concentrations of Solutions ...
it is read as "molar." For example, a solution labeled 3M NaCl is read as three molar sodium chloride solution. Figure 16.8 demonstrates how to make a 0.5-molar solution, that is, a solution with a molarity of 0.5. Molarity 1M2 moles of solute liters of solution 16.2 Concentrations of Solutions A supply of clean drinking

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A saturated solution of silver nitrate is prepared in 100.0 g of water at 20 C. The solution is then heated to 50.0 C. How much more silver nitrate must now be added to obtain a saturated solution? (Use Table 16.1.) SECTION 16.2 CONCENTRATIONS OF SOLUTIONS 1. Calculate the molarity of each of the following solutions.

SECTION 16.1 PROPERTIES OF SOLUTIONS
3, solution 3. What is the molarity of a solution that contains 212.5 g of sodium nitrate (NaNO 3) in 3.0 liters of solution? 4. You must prepare 300.0 mL of 0.750M NaBr solution using 2.00M NaBr stock solution. How many milliliters of stock solution should you use? 5. In order to dilute 1.0 L of a 6.00M solution of NaOH to 0.500M solution, how

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