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Calculating pH 3 Read This! The value $1.0 \times 10-14$ is the equilibrium constant for the autoionization of water (K w). H 2 O + H 2 O @A H 3 O+(aq) + OH-(aq) K w + $1.0 \times 10-14$ This equilibrium occurs in all aqueous solutions (acidic, basic, and neutral).

34 Calculating pH-S - Skyline Chemistry

Unit 6 - Chemical Reactions, Stoichiometry and Acids and Bases Essential Questions: What is the evidence that a chemical reaction has occurred?

UNIT 6 - Welcome to chemistry

Homework due: Enjoy the vacation! Optional: start your lab write-up. Final draft will be due on Friday April 26. Catalyst - Complete the 8.1 try it on your own conjugate acid-base pairs. Class activities - Review Questions on acids and bases that you completed the week before spring break. (Presentation - 8.1 Theories of acids and bases Amphoteric and amphiprotic (slides 17-20)

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