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The first law of thermodynamics connects the two energy units and allows one to relate heat and work energy or to relate calories and Joules. ‡ Extensive and Intensive Properties. Properties (or state variables) are extensive or intensive . Extensive variables depend on the size of the system such as volume or mass.

Introduction to the Thermodynamics of Materials

The first law of thermodynamics, also known as Law of Conservation of Energy, states that energy can neither be created nor destroyed; energy can only be transferred or changed from one form to another. For example, turning on a light would seem to produce energy; however, it is electrical energy that is converted.

The Three Laws of Thermodynamics | Introduction to Chemistry

Work: Work is the amount of energy transferred to or from a body or system as a result of forces acting upon the body, causing displacement of the body or parts of it. More specifically, the work done by a particular force is the product of the displacement of the body and the component of the force in the direction of the displacement.

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