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2)  $4.75\text{ g} \times 1\text{ kg} = 0.00475\text{ kg}$ . 1000g. 3)  $0.000278\text{ mole} = 0.0582\text{ m}$ .  $0.00475\text{ kg}$ . The molality of a solution of chlorine and water is  $0.0362\text{ m}$ . This solution contains 3500g of water. How much chlorine, in grams, was used to prepare this solution? ... 15-2 Review and Reinforcement ...

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10—2 Review and Reinforcement (continued) 9.  $0.0621\text{ mol SS} = (Q \$ \text{ gK}2\text{S } 10. 65.8\text{ gMg} = \text{l. atoms Mg}$  Solve each of the following problems as directed. Show all your work. 11. Calculate the number Of atoms in  $0.40\text{ mol}$  of sulfur- 12. Calculate the number of atoms in  $2.30\text{ mol}$  of silver. G.oux 13. You have a tank of NO<sub>2</sub> gas.

### 10-2 Review and Reinforcements - Front Page

2 Answer.Key 2-1 Explore, page 5 ... 2-1 Review and Reinforcement, pages 8-9 . 1. energy 2. radiant 3. chemical potential 4. joule 5. created, destroyed 6. natural gas ... 15. chromatography 16. electrolysis 17. distillation 18. A magnet could be used to separate them

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'16—1A Review and Reinforcement, pages 9-10 6. Answers Will Vary. When the rates Of forward and reverse reactions are equal, equilibrium is established. For the example of ... 15. The forward and reverse reactions are both occurring, but with no net change in concentration. -16—2 Review Reinforcement, pages 15-16

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15. Why is water called the universal solvent? \_\_\_\_ 16. How could you determine experimentally whether a given substance is an electrolyte? ... Title: 13-1 Review and Reinforcement Author: NPSD Last modified by: NPSD Created Date: 9/27/2005 1:14:00 PM Company: NPSD

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