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Section 16.1 Properties of Solutions 473 What is happening? Particles move from the solid into the solution. Other dissolved particles move from the solution back to the solid. Because these two processes occur at the same rate, no net change occurs in the overall system. As Figure 16.2 illustrates, a state of dynamic equilibrium

16.1 Properties of Solutions 16

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Chapter 16 Solutions 167 SECTION 16.1 PROPERTIES OF SOLUTIONS (pages 471–477) This section identifies the factors that affect the solubility of a substance and determine the rate at which a solute dissolves. Solution Formation (pages 471–472) Look at Figure 16.1 on page 471 to help you answer Questions 1 and 2.1.

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Chapter 16: Colligative Properties of Solutions 45 16-4. The mole fraction of (NH 4) 2SO 4(aq) is given by x (NH 4) 2SO 4 = n (NH 4) 2SO 4 n (NH 4) 2SO 4 + n H 2O Because (NH 4) 2SO 4(aq) is a strong electrolyte, it dissociates completely into NH + 4 (aq) and SO2- 4 (aq) ions. Assume a one kilogram solution. The number of moles of ions in one ...

CHAPTER 16. Colligative Properties of Solutions

3) The properties of a solution that depend only on the number of solute particle present, without regard to their identity, are called colligative properties. 4) Freezing point depression and boiling point elevation are colligative properties that are directly proportional to the molal concentration of any solute. These

Chapter 16

Chapter 16 Solutions 403 Section Review Objectives • Identify the three colligative properties of solutions • Describe why the vapor pressure, freezing point, and boiling point of a solution differ from those properties of the pure solvent. Vocabulary • colligative properties • freezing-point depression • boiling-point elevation Part ACompletion Use this completion exercise to check ...

05 CTR ch16 7/12/04 8:14 AM Page 399 PROPERTIES OF ...

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