

14 The Properties Of Gases Answers

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14 The Properties Of Gases

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ST AT 14. small compared to the total volume of the gas. Air will rush into a sealed container when the container is opened. Gas flows from a region of lower pressure to a region of higher pressure. Adding air to an object will cause the object to inflate.

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Chapter 14 The Behavior of Gases 147 SECTION 14.1 PROPERTIES OF GASES(pages 413–417) This section uses kinetic theory to explain the properties of gases. This section also explains how gas pressure is affected by the amount of gas, its volume, and its temperature. Compressibility (pages 413–414) 1. Look at Figure 14.1 on page 413.

SECTION 14.1 PROPERTIES OF GASES(pages 413–417)

Gases are easily compressed because of. the space between particles in a gas. temperature, pressure, and amount of gas. Because the gas in the inflated air bag can. be compressed, the bag absorbs some of. If the temperature decreases, the. pressure will also decrease. The volume would need to ...

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THE BEHAVIOR OF GASES SECTION 14.1 PROPERTIES OF GASES (pages 413–417) This section uses kinetic theory to explain the properties of gases. This section also explains how gas pressure is affected by the amount of gas, its volume, and its temperature. Compressibility (pages 413–414) 1. Look at Figure 14.1 on page 413.

SECTION 14.1 PROPERTIES OF GASES(pages 413–417) - MAFIADOC.COM

14 The Properties Of Gases Gas is a state of matter that has no fixed shape and no fixed volume. Gases have lower density than other states of matter, such as solids and liquids. There is a great deal of empty space between ... Properties of Matter: Gases - Live Science Properties and

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SECTION 14.1 THE PROPERTIES OF GASES 1. Using kinetic theory, explain why a tire is more likely to blow out during a trip in the summer than during one in the winter. 2. Use kinetic theory to explain why on a cold autumn morning a camper's air

SECTION 14.1 THE PROPERTIES OF GASES

Gases are easily , or squeezed into a smaller volume 1. because of the between particles in a gas. The four variables 2. used to describe a gas are pressure, (P), (V), (T), 3. and number of (n). 4. You can use theory to predict and explain how gases 5. will respond to a change in conditions. Doubling the amount of 6.

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14.1 Properties of Gases

Upon closer study, they began observing consistent properties that defined gases. The single distinction that initially baffled scientists -- that of gas particles having more space to move freely

than particles of solids or liquids -- informs each of the properties that all gases have in common.

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Gases have three characteristic properties: (1) they are easy to compress, (2) they expand to fill their containers, and (3) they occupy far more space than the liquids or solids from which they form. Compressibility. An internal combustion engine provides a good example of the ease with which gases can be compressed.

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- Relate the total pressure of a mixture of gases to the partial pressures of the component gases • Explain how the molar mass of a gas affects the rate at which

05 CTR ch14 7/12/04 8:13 AM Page 347 THE PROPERTIES OF ...

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Gases are also easily compressed or squeezed into a smaller volume. Compressibility is the measure of how much the volume of matter decreases under pressure. Gases are easily compressed because of the space between the particles in a gas. The volume of the particles of a gas is small compared to the overall volume of the gas. ... Chapter 14.1 ...

Chapter 14.1 Properties of Gases

CDS 14 Physical Properties of Gases I. One of the powers of chemistry is the ability to relate the properties of individual molecules to the physical and chemical properties of the compounds of these molecules. In other words, we want to relate the atomic molecular world to the macroscopic world of materials. We begin this study by observing ...

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