

## ***10 2 Review And Reinforcement Answer Key***

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**10 2 Review And Reinforcement**

10—2 Review and Reinforcement (continued) 9.  $0.0621 \text{ mol SS} = (Q \$ \text{gK}2\text{S } 10. 65.8\text{gMg} = \text{l. atoms Mg}$  Solve each of the following problems as directed. Show all your work. 11. Calculate the number Of atoms in 0.40 mol of sulfur- 12. Calculate the number of atoms in 2.30 mol of silver. G.oux 13. You have a tank of NOZ gas.

**10-2 Review and Reinforcements - Front Page**

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**10 2 REVIEW AND REINFORCEMENT ANSWER KEY 235411**

Worksheet # 5 Quarter 3 Packet 2: Mole Conversions 10-2 Review and Reinforcement Packet #2 Quarter 3 assignment 5. Mole Conversions Answer Each of the following questions in the space provided. 1. How would you calculate the number of moles present in a given mass of a substance? \_\_\_\_\_ 2.

**Mole Conversions - OoCities**

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**10 1 REVIEW AND REINFORCEMENT CHEMICAL MEASUREMENTS ...**

12—2 Review and Reinforcement (continued) 10. 20.0 g of potassium react with water to produce potassium hydroxide and hydrogen gas. How many liters would the hydrogen gas occupy at STP? 573 L u 301, 1 2 k 11. If 30.2 g of aluminum react with HCl to produce alurninum chloride and hydrogen gas, how

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11—1 Review and Reinforcement Stoichiometry Complete the following sentences by 'filling in the appropriate word or phrase from the list below. Each word or phrase maybe used once, more than once, or not at all reactants molar ratio quantitative actual particles. consdrvption of matter coefficients subscripts mass

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2 Answer.Key 2-1 Explore, page 5 The original source of energy for this activity would be the stored chemical energy in the ... 2-1 Review and Reinforcement, pages 8-9 . 1. energy 2. radiant 3.

chemical potential 4. joule 5. created, destroyed 6. natural gas 7. kinetic . 8. f . 9. c 10. b 11. a . 12. e .

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