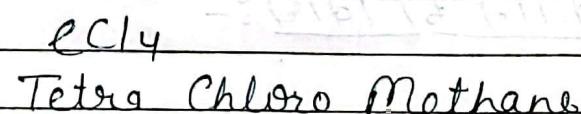
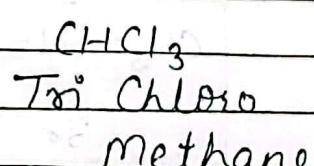
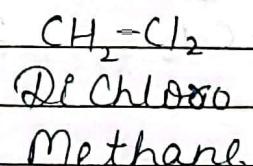
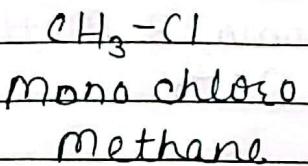


⑥ Alkyl Halides

Date: / /
Page: 1

Alkyl Halide:-

संतृप्त दाईरी कार्बन के एक या अधिक दाईरीजन परमाणु का हलोजन परमाणु कारा प्रतिस्थापन करने पर प्राप्त युग्मन की Alkyl Halide कहा जाता है। इनका सामान्य सूत्र $C_nH_{2n+1}X$ होता है। इनके सामान्यतः RX में भी प्रदर्शित किया जाता है।



क्रियकरण:-

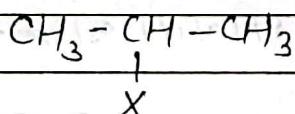
Alkyl Halide

मैन प्रकार के होते हैं।

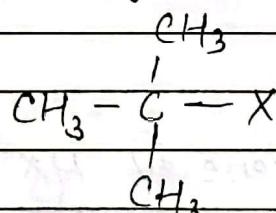
① Primary Alkyl Halide



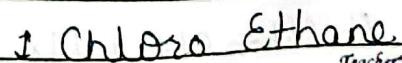
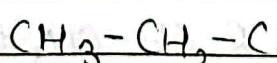
② Secondary Alkyl Halide



③ Tertiary Alkyl Halide



उदाहरण:-



Teacher's Signature.....



Ques 1, 2 वाले

~~Alkene का रेग्युलर तरीका है -~~

(2) Alkene का :-

लोम्प.

Bromo

F > Cl > Br > I

- का रेग्युलर तरीका है -

Ques 1, 2 वाले का रेग्युलर तरीका है -

CH₃-CH₃ + X₂ \longrightarrow CH₃-CH₂-X

Ques 1, 2 वाले का रेग्युलर तरीका है -

(1) Alkane का :-

Mono Halide का रेग्युलर तरीका है -

Br > Cl

Butane

CH₃-CH-CH₃

1

2

3

4

Butane

2-Bromo 3-Methyl

CH₃

Propane

CH₃-CH-CH₂-Cl

1

2

3

4

1-Chloro 2-methyl

Br

9-Bromo propane

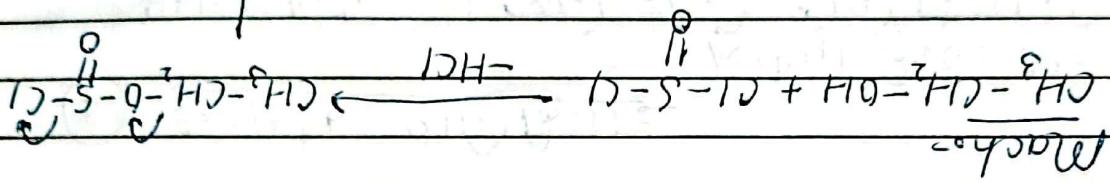
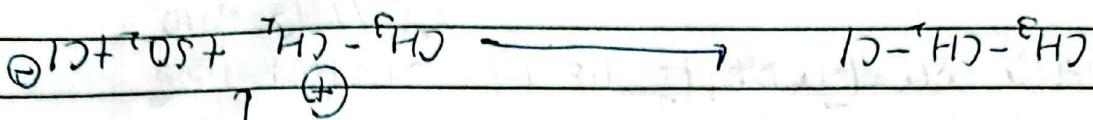
CH₃-CH-CH₃

1

2

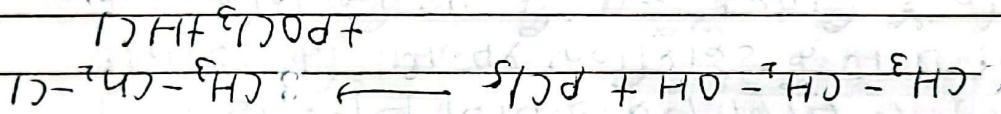
3

4

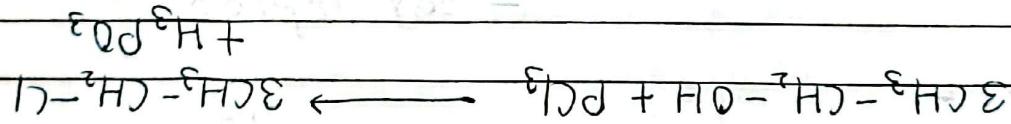


④ Alc. & SOCl_2 (alkyl halide + AlCl_3) $\xrightarrow{\text{HgO}}$ $\text{R}_3\text{P}(\text{O})(\text{OR})_2$

~~alkyl halide + AlCl_3~~



② Alc. & Hg(OAc)_2 $\xrightarrow{\text{HgO}}$ $\text{R}_3\text{P}(\text{O})(\text{OR})_2$



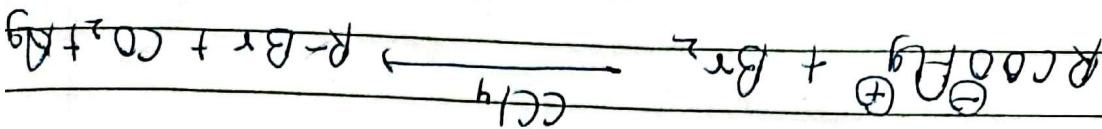
⑥ Alc. & Hg(OAc)_2 $\xrightarrow{\text{HgO}}$ $\text{R}_3\text{P}(\text{O})(\text{OR})_2$



⑨ Alcohol & HX $\xrightarrow{\text{HgO}}$ $\text{R}_3\text{P}(\text{O})(\text{OR})_2$

③ Alcohol & HX

Nitrone & Hg(OAc)_2 $\xrightarrow{\text{HgO}}$ $\text{R}_3\text{P}(\text{O})(\text{OR})_2$

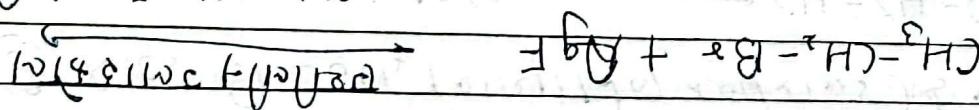
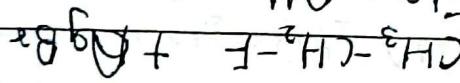


Q11 N/2

~~Q11 N/2~~ 1. ~~Alky / Alkene~~ ~~Alkyl halide~~ ~~Alkyl halide~~

(5) ~~Q11 N/2~~

Hence Alkane



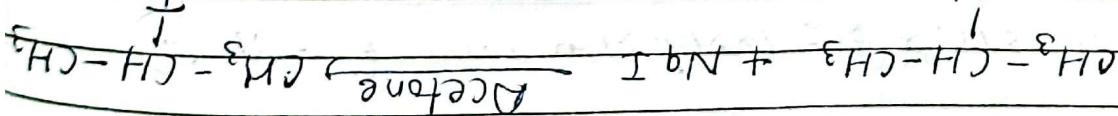
Q11 N/2

~~Q11 N/2~~ 2. ~~Alkyl halide~~ ~~Alkyl halide~~ ~~Alkyl halide~~

(6) ~~Q11 N/2~~ Rego.

+ NaCl

Q11 N/2



Q11 N/2

~~Q11 N/2~~ 3. ~~Alkyl halide~~ ~~Alkyl halide~~ ~~Alkyl halide~~

~~Q11 N/2~~ 4. ~~Alkyl halide~~ ~~Alkyl halide~~ ~~Alkyl halide~~

~~Q11 N/2~~ 5. ~~Alkyl halide~~ ~~Alkyl halide~~ ~~Alkyl halide~~

~~Q11 N/2~~ 6. ~~Alkyl halide~~ ~~Alkyl halide~~ ~~Alkyl halide~~

~~Q11 N/2~~ 7. ~~Alkyl halide~~ ~~Alkyl halide~~ ~~Alkyl halide~~

(7) ~~Q11 N/2~~ Rego.

Q11 N/2

Teacher's Signature

1. Hydrogen atom (H) has one proton and one electron.

2. Hydrogen atom (H) has one proton and one electron.

3. Hydrogen atom (H) has one proton and one electron.

4. Hydrogen atom (H) has one proton and one electron.

5. Hydrogen atom (H) has one proton and one electron.

6. Hydrogen atom (H) has one proton and one electron.

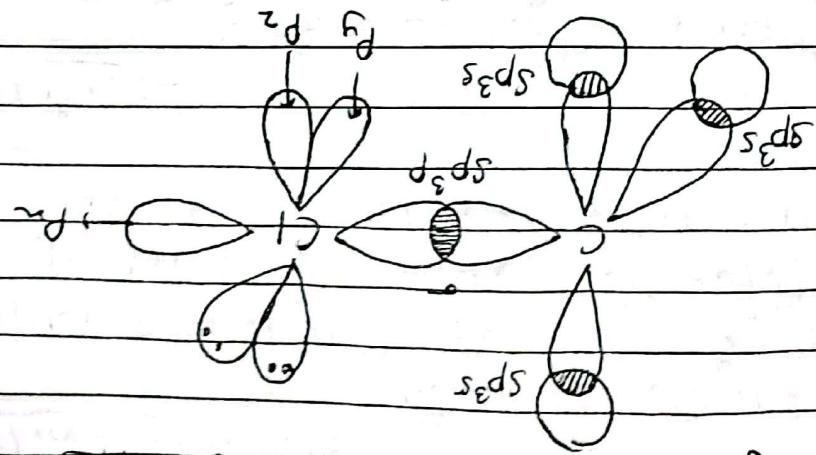
7. Hydrogen atom (H) has one proton and one electron.

8. Hydrogen atom (H) has one proton and one electron.

9. Hydrogen atom (H) has one proton and one electron.

10. Hydrogen atom (H) has one proton and one electron.

-: Hydrogen atom (H) has one proton and one electron.

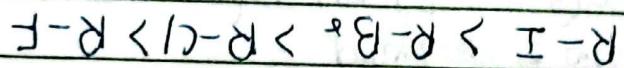


AI 튜닝/Hyperparameter Tuning 알고리즘

(2)

1) $\text{CH}_3\text{I} > \text{CH}_2\text{I} > \text{CH}_3\text{Cl}$
 2) $\text{CH}_3\text{I} > \text{CH}_2\text{I} > \text{CH}_3\text{Br}$

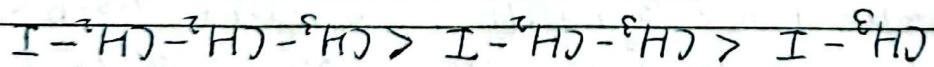
$\text{CH}_3\text{I} > \text{CH}_2\text{I} > \text{CH}_3\text{Cl}$



(1)

$\text{CH}_3-\text{CH}_2-\text{CH}_2-\text{I} > \text{CH}_3-\text{CH}_2-\text{CH}_2-\text{I} > \text{CH}_3-\text{CH}_2-\text{I}$

$\boxed{\text{B.P. } \alpha = 117^\circ \text{ C}}$



$\boxed{\text{B.P. } \alpha = 117^\circ \text{ C}}$

(2)

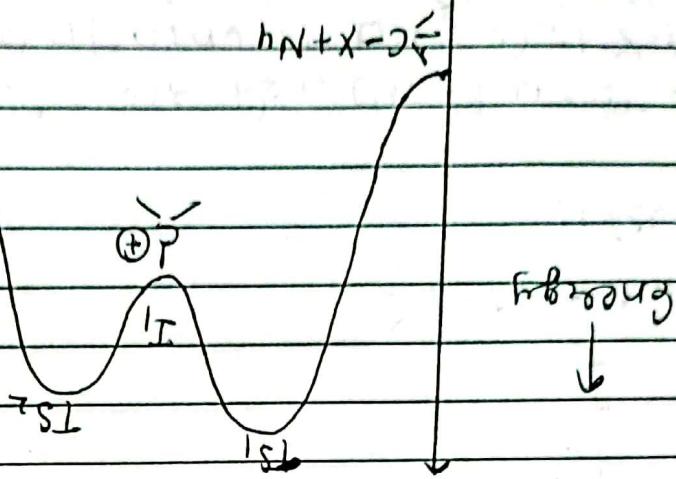
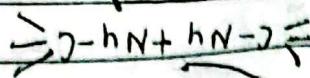
$\text{CH}_3-\text{I} > \text{CH}_3-\text{CH}_2-\text{I} > \text{CH}_3-\text{CH}_2-\text{CH}_2-\text{I}$

(1)

$\text{AlE}_3 \text{ Halide} > \text{Al}_2\text{I}_3 \text{ Halide} > \text{Al}_3\text{I}_5 \text{ Halide}$

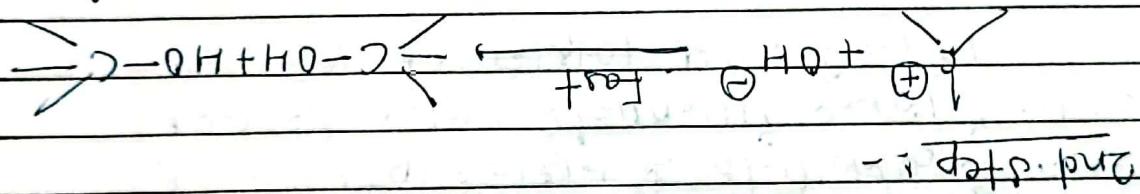
$\boxed{\text{B.P. } \alpha = 117^\circ \text{ C}}$

← 4/10/2018 8:51 AM

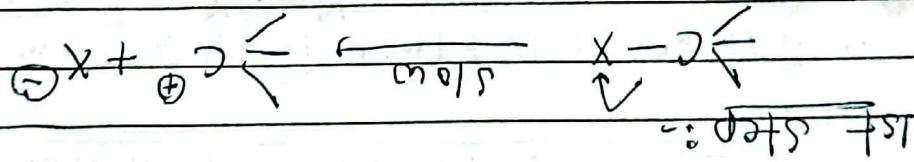


Racemic Nithium

all (50% - 50%)



d-Reactant Carbocation



SN₁ Reactions

for example SN₁ reactions

minimizing energy

SN₁ Reactions

SN₁ Reactions

SN₁ Reactions

SN₁ Reactions

SN₁ Reactions

SN₁ Reactions

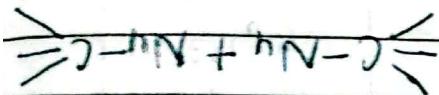
- First step :-

(A)

<input type="checkbox"/>	t =
1	2000
2	4000

SN₂ Reaction:

~~(108-08) np~~



$$\Theta^{hN+1}$$



12 IN 13 MILS. 10151-14 14% OS-OS

Suppose INTL will buy 2000 units

101+5112. 1+42N. 11-12 4+121241014 1m. 2

Reo. 42012

ԱՐԵՎՈՒԹՅՈՒՆԻՑ ԽՈՐԵ ՀԱՅՈՒԹՅՈՒՆ ՀԱՅ ՀԱՅՈՒԹՅՈՒՆ

~~2008 (M 2 INFO HARAKATI. 6 26 47 AM 105~~

~~6-9.9.9~~, NS

~~- 2011 R 4 10112 1101013~~ ③

~~2024 (NP) to fit in with the other types of film that I'm doing.~~

~~14 INSTITUTE FOR THE STUDY OF~~

~~It's important to have a good understanding of the concepts involved.~~

~~دستورات اخراجی~~ دستورات اخراجی

וְעַתָּה תִּשְׁמַח בְּנֵי יִשְׂרָאֵל אֶת-בְּנֵי עֲמֹק, נָ

$$\boxed{[(x-c)\leq] \geq = \text{log} \cdot (\textcircled{3})}$$

-: alter topic ①

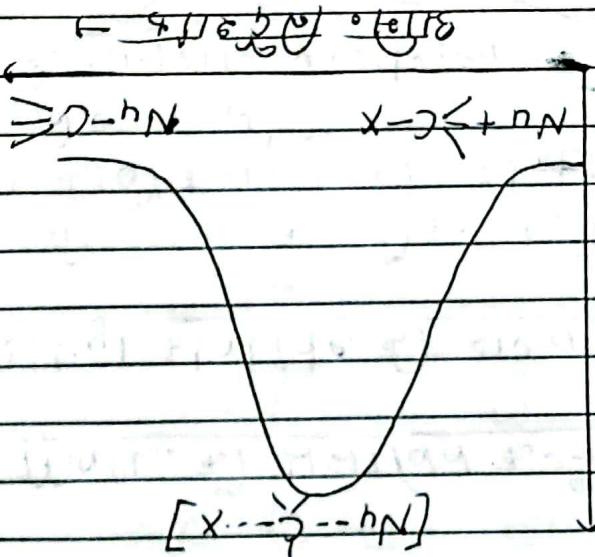
- : TOLK. F. 125 831811251, NS

12 (Nitrates)

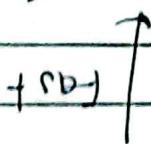
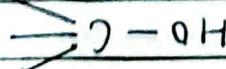
~~Ammonium is a strong base. It has a low solubility product constant. It dissociates completely in water.~~

$$[x - \text{NH}_4^+] [x - \text{NO}_3^-] = K_{\text{sp}}$$

$$- \frac{K_{\text{sp}}}{[x]^2} = \frac{\partial \ln K_{\text{sp}}}{\partial x} \quad \text{---} \quad \text{SN}^2 \text{ Pog. 34. 4. 101: -}$$



Stability - $\text{CH}_3\text{B}_6^{2-} > \text{Li}^+ > \text{Na}^+$

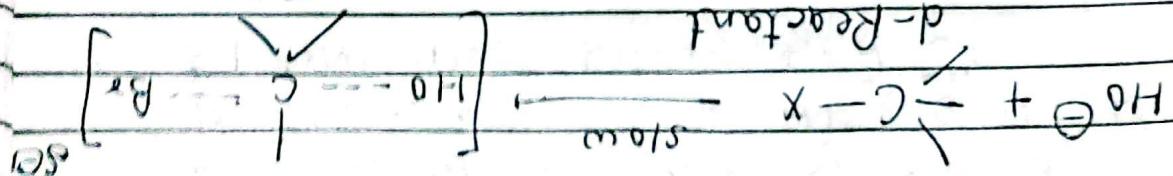


$$= \frac{\partial \ln K_{\text{sp}}}{\partial x}$$

$$K_{\text{sp}} [\text{OH}] [\text{C} - x]$$

$$= \frac{\partial \ln K_{\text{sp}}}{\partial x}$$

T.S.



~~Ammonium is a strong base. It has a low solubility product constant. It dissociates completely in water.~~

~~Ammonium is a strong base. It has a low solubility product constant. It dissociates completely in water.~~

~~Ammonium is a strong base. It has a low solubility product constant. It dissociates completely in water.~~

~~Ammonium is a strong base. It has a low solubility product constant. It dissociates completely in water.~~

1	6
1	1

Halide & Halogen के बीच का सम्पर्क SN₂Po₃, NS

(2) 319 वें दिन का कागज़ :-

3.

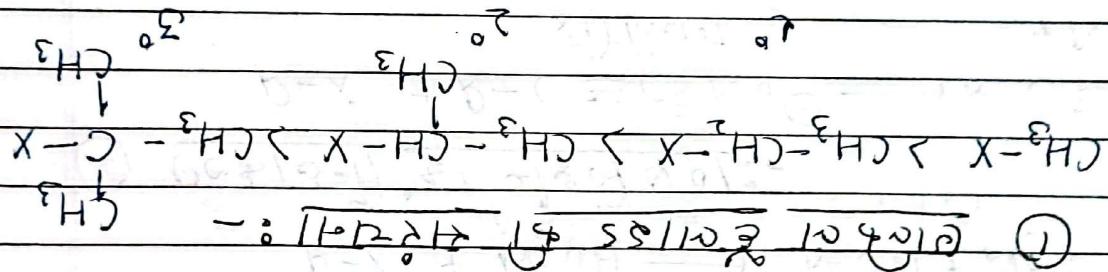
2.

1.

CH₃CH

Lycium

SN² Po₂. It is a yellow solid. It is very soluble in water.



δN^2 Macho $\sqrt{4}$ $\sqrt{4} \times 10^6$ $\sqrt{2} \times 10^6$ $\sqrt{2} \times 10^6$ $\sqrt{2} \times 10^6$ $\sqrt{2} \times 10^6$



12 214K. 1-101 14 14 b2b2113
418.2 12 b2b2113 → 16 h6 yd14K 431h2E
14 12b1102B 1m 2 12 1-14 3 10 3 14
2010b1E 12 1024 2 14 2 10 5 17.924S 12
1-12 4 1014 1-110 2 4.1040 14 2 14 2010b1E
12 1N1-12 1-12, 1014 1-1013H 1-13 1-5007P

Alcohol

Sodium

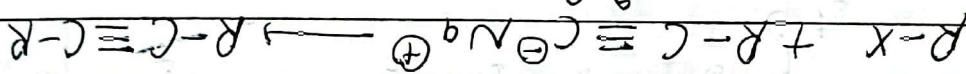


Alkyl Halide + Metal Oxide $\xrightarrow{\text{Heat}}$ Alkyl Alcohol + Metal Chloride
 Alkyl Halide + Potassium Hydroxide $\xrightarrow{\text{Heat}}$ Alkyl Alcohol + Potassium Chloride

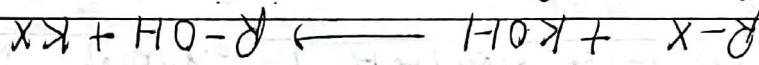
(3) $\text{R-OH} + \text{NaCl} \xrightarrow{\Delta}$

$\text{R-OH} + \text{NaCl} \xrightarrow{\Delta}$

+ NaX



(2) $\text{R-OH} + \text{AgNO}_3 \xrightarrow{\Delta}$



(1) $\text{R-OH} + \text{AgNO}_3 \xrightarrow{\Delta}$

$\text{R-OH} + \text{AgNO}_3 \xrightarrow{\Delta}$

(1) $\text{R-OH} + \text{AgNO}_3 \xrightarrow{\Delta}$

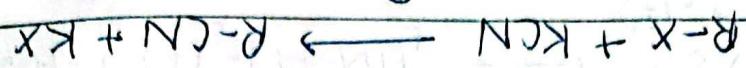
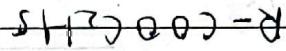
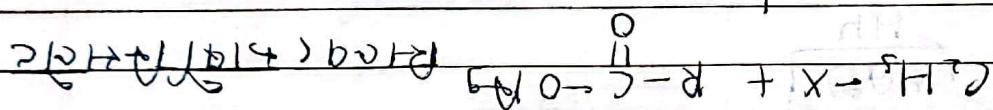
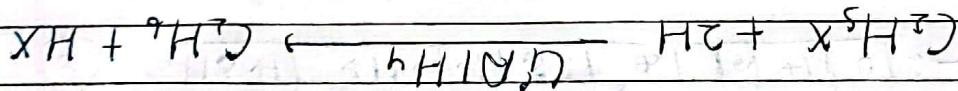
(2) $\text{R-OH} + \text{AgNO}_3 \xrightarrow{\Delta}$

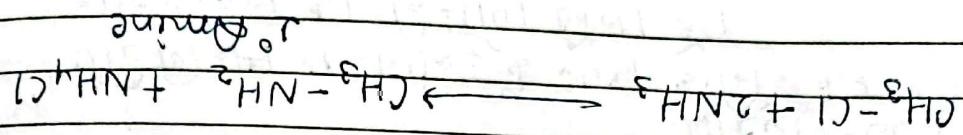
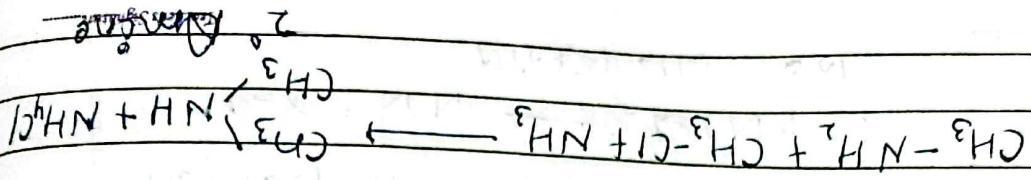
Alkyl Halide + Potassium Nitrosochloride

(3) $\text{R-OH} + \text{AgNO}_3 \xrightarrow{\Delta}$

$R-I < R-Br < R-Cl < R-F$

(2) $\text{R-OH} + \text{AgNO}_3 \xrightarrow{\Delta}$

ChloroformAlicyclic HalideReactionEffect $-X$ ReactionAlicyclic HalideProductAlkaneAlkeneReactionProductButaneAlkeneReactionAlicyclic HalideProduct

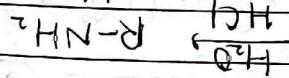
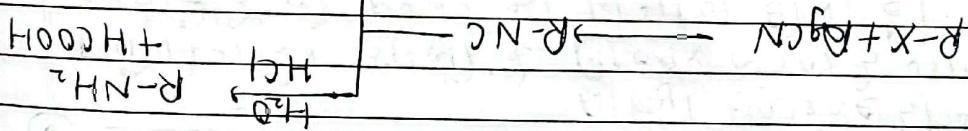
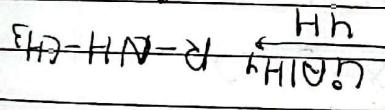


~~3. 2. 3. Formation of salts~~

~~Alky! Halide~~

(g) ~~3. 2. 3. Formation of salts~~

3. ~~Amine~~

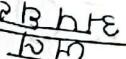
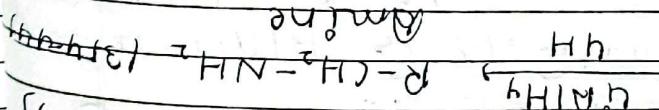


~~3. 2. 3. Formation of salts~~

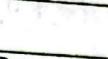
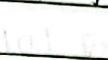
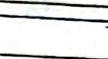
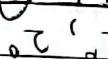
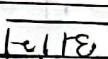
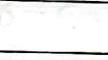
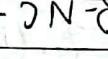
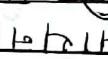
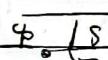
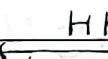
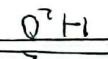
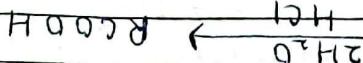
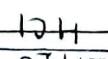
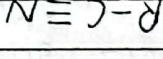
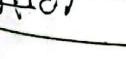
~~Alky! Halide~~

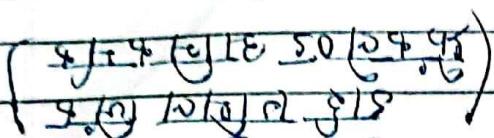
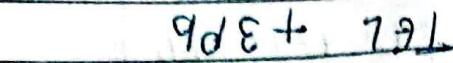
(8) ~~3. 2. 3. Formation of salts~~

Amide

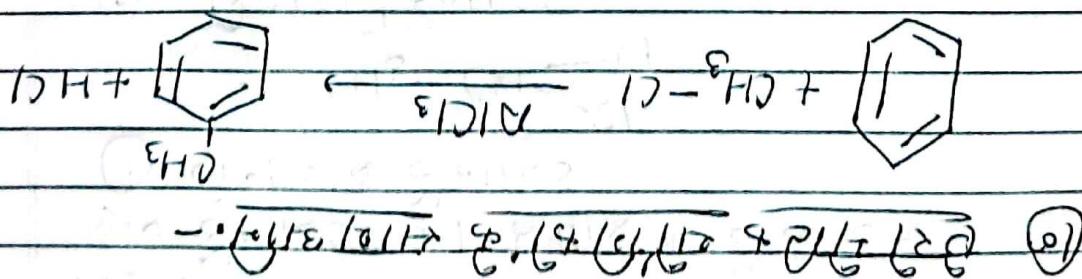
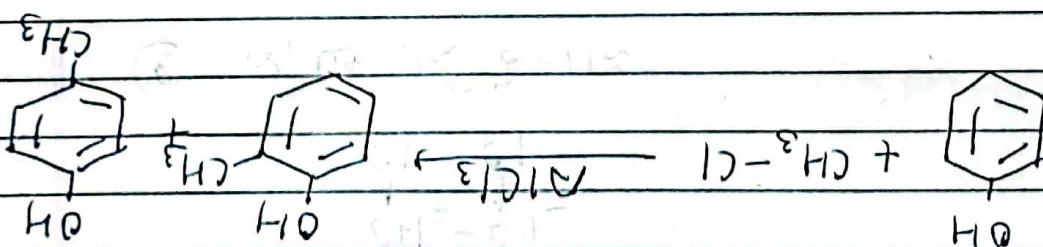


Amide

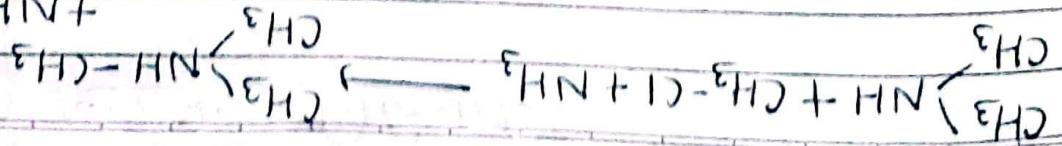




~~Alkyli Halide + AlCl₃ → Alkyl Chloride + HCl~~



3° Amine



47. $\text{CH}_3\text{CHO} + \text{PCl}_3 \rightarrow \text{CH}_3\text{COCl} + \text{HCl}$

Ethyl chloride

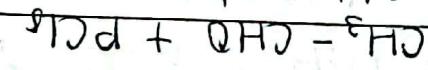
H₁ at Cl atom

+ POCl₃

Cl

Cl

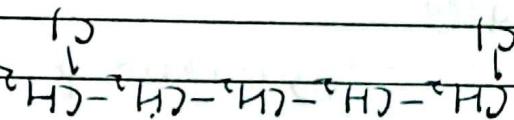
Amino chloride



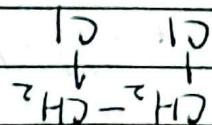
48. $\text{CH}_3\text{CH}_2\text{OH} + \text{PCl}_3 \rightarrow \text{CH}_3\text{CH}_2\text{OCl} + \text{HCl}$

①

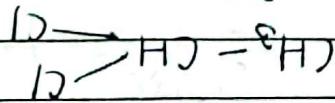
$\text{CH}_3\text{CH}_2\text{OH} + \text{PCl}_3 \rightarrow \text{CH}_3\text{CH}_2\text{OCl} + \text{HCl}$



② $\text{CH}_3\text{CH}_2\text{OCl} + \text{HCl}$



② $\text{CH}_3\text{CH}_2\text{OCl} + \text{HCl} \rightarrow \text{CH}_3\text{CH}_2\text{OH} + \text{Cl}_2$



① $\text{CH}_3\text{CH}_2\text{OCl} + \text{HCl} \rightarrow \text{CH}_3\text{CH}_2\text{OH} + \text{Cl}_2$

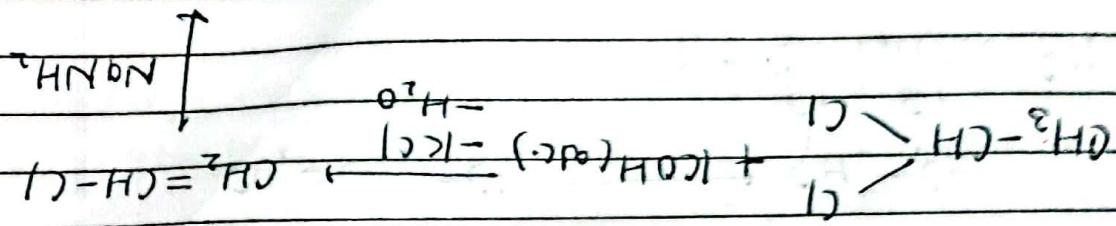
② $\text{CH}_3\text{CH}_2\text{OCl} + \text{HCl} \rightarrow \text{CH}_3\text{CH}_2\text{OH} + \text{Cl}_2$

③ $\text{CH}_3\text{CH}_2\text{OCl} + \text{HCl} \rightarrow \text{CH}_3\text{CH}_2\text{OH} + \text{Cl}_2$

④ $\text{CH}_3\text{CH}_2\text{OCl} + \text{HCl} \rightarrow \text{CH}_3\text{CH}_2\text{OH} + \text{Cl}_2$

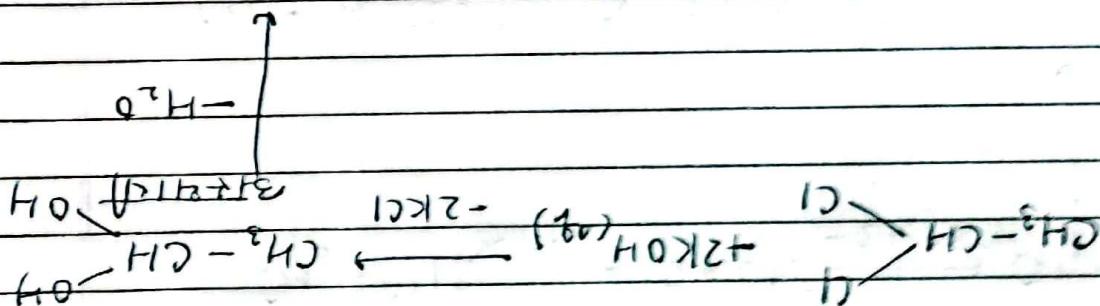
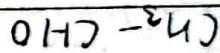
⑤ $\text{CH}_3\text{CH}_2\text{OCl} + \text{HCl} \rightarrow \text{CH}_3\text{CH}_2\text{OH} + \text{Cl}_2$

Alfonso CH3CH + NH3 + NaCl



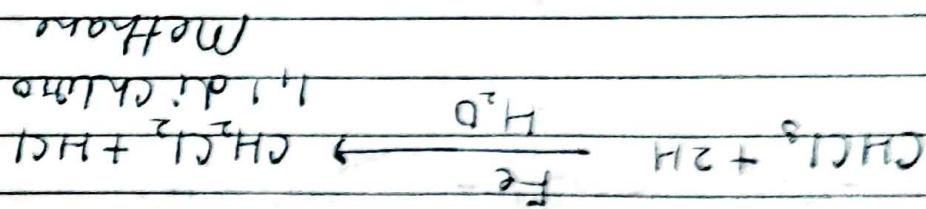
$\text{AlC}_6\text{KOH} \xrightarrow{\text{Heat}} \text{Al}_2\text{O}_3 + 3\text{H}_2\text{O}$ (Alumina)

Aldophydr



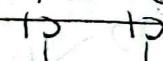
لهم إني أنت عبدي فلا جندي لك ولا ملك لك ولا سلطانا لأنك أنت عبدي

~~استعفیت از مطالعه این سایت~~



$$\text{CH}_3\text{Cl} + \text{HBr} \longrightarrow \text{CH}_2 = \text{CH}-\text{Br}$$

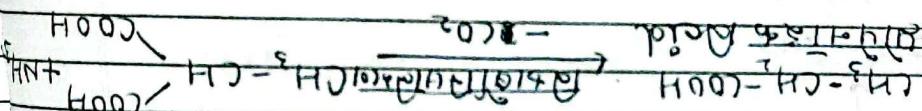
1. 2 ③ CH₂O EtH₂O



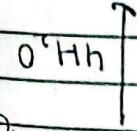
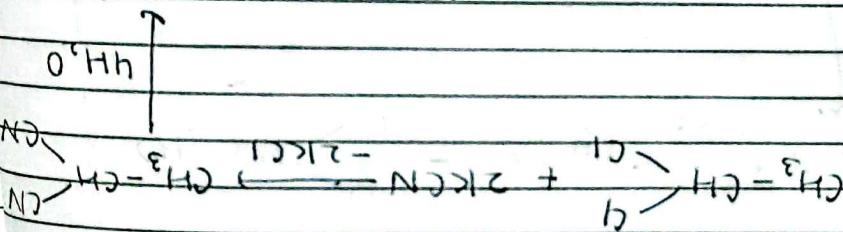
③

2. ④ ~~CH₃-COOH + H₂N-CH₂-COOH → CH₃-CH₂-CONH-CH₂-COOH~~

~~CH₃-CH₂-CONH-CH₂-COOH~~

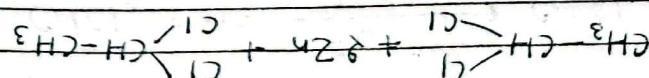
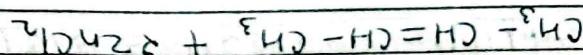


②



④ KCN + P_og.

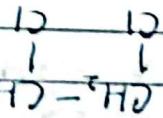
Bu₂F-O-ene



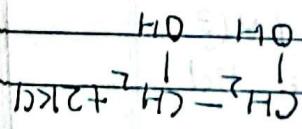
⑤ CH₃-CH₂-CH₂-CH₂-P_og.

⑥ CH₃-CH₂-CH₂-CH₂-P_og.

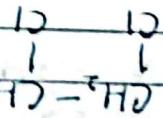
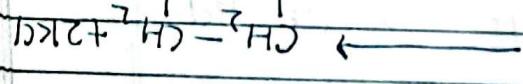
कार्बन डायामांड

KOH + CH₄ → C₆₀Alc. KOH + CH₄ → C₆₀.

+ 2KOH (aq.)



→

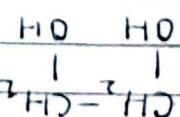
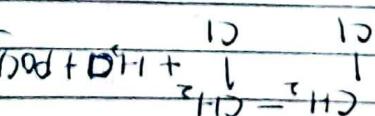
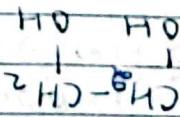
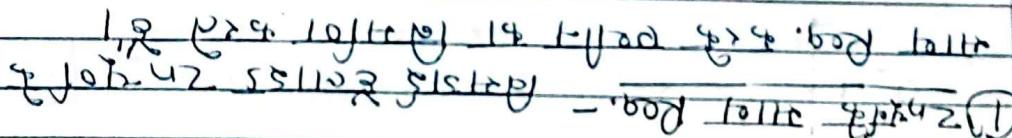
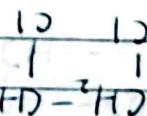
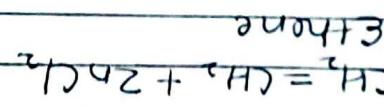


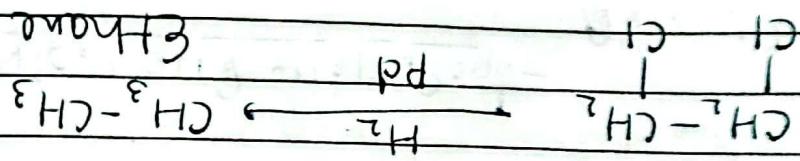
कार्बन डायामांड

C₆₀ + KOH + CH₄ →

-

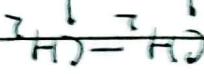
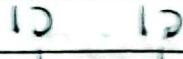
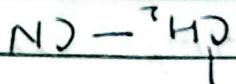
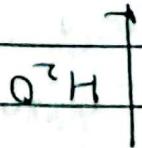
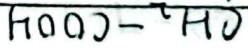
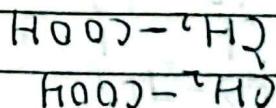
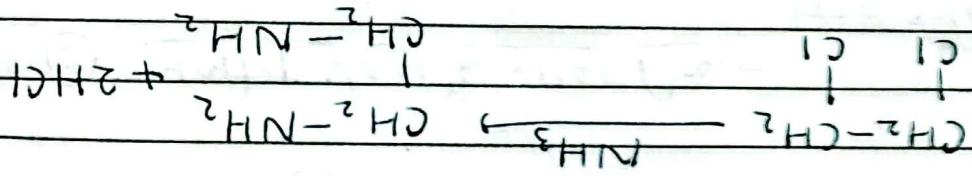
-





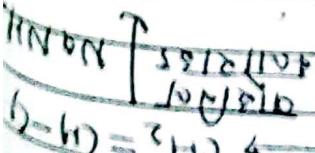
12 1N13

- Fritz ⑧



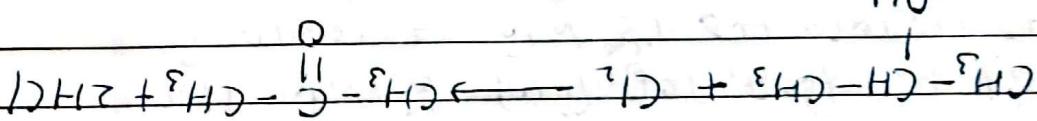
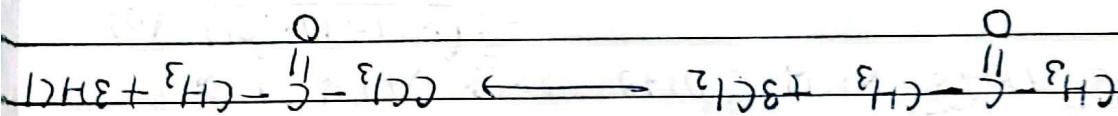
12 1N13 → 1N14-13 12 13
Fritz 12 13 14-13 12 13

- ⑨



200gm Al_2O_3 + $2\text{Na}_2\text{O} \rightarrow 2\text{NaAlO}_2$ (Alumina + Soda ash \rightarrow Sodium aluminate)

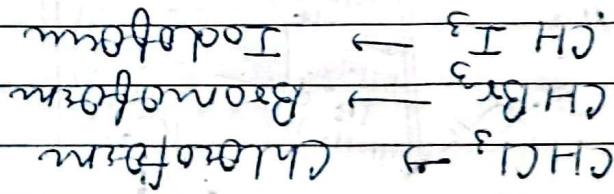
- (2) NaAlO_2



Charge separation

- (1) NaAlO_2

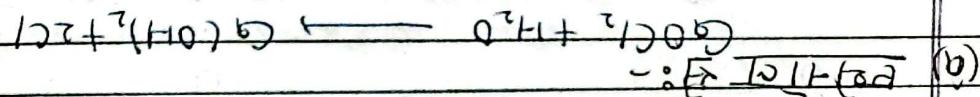
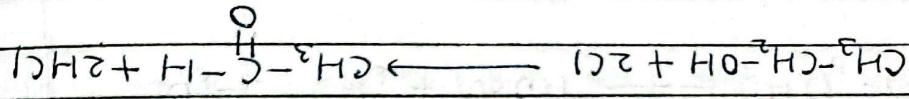
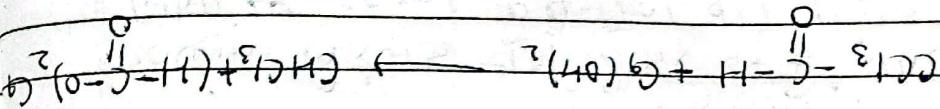
- (2) NaAlO_2



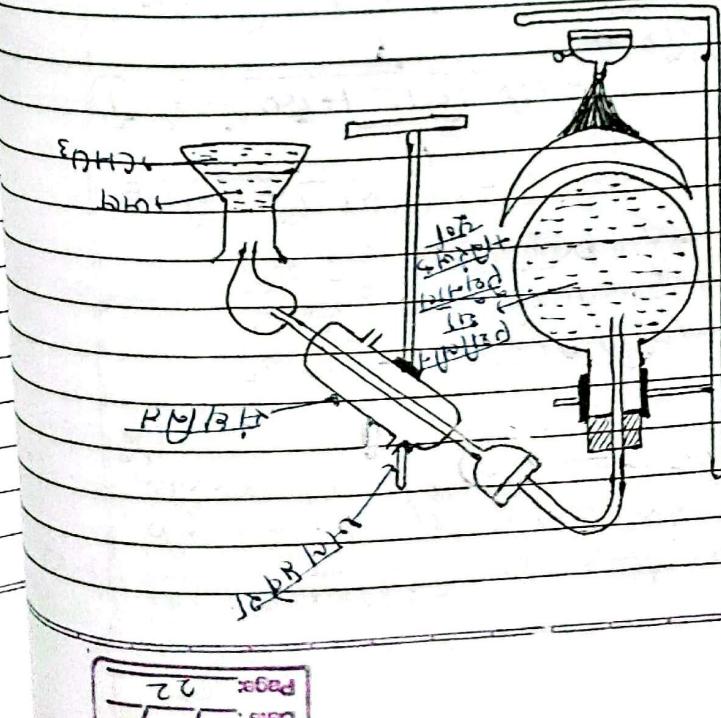
13

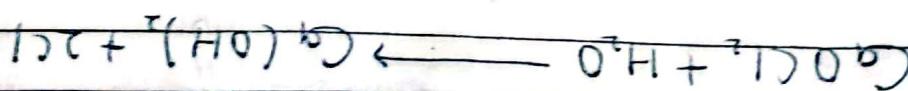
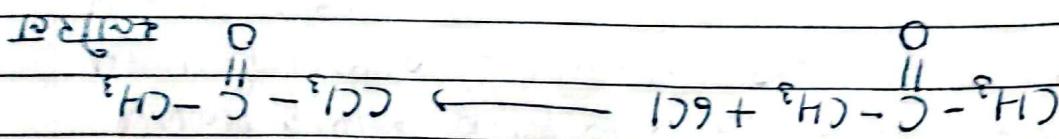
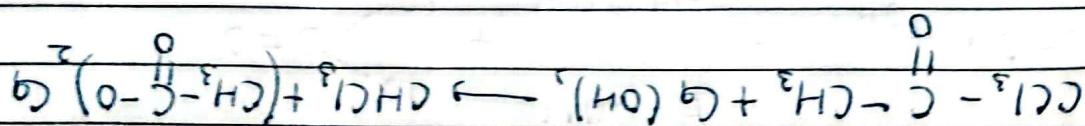
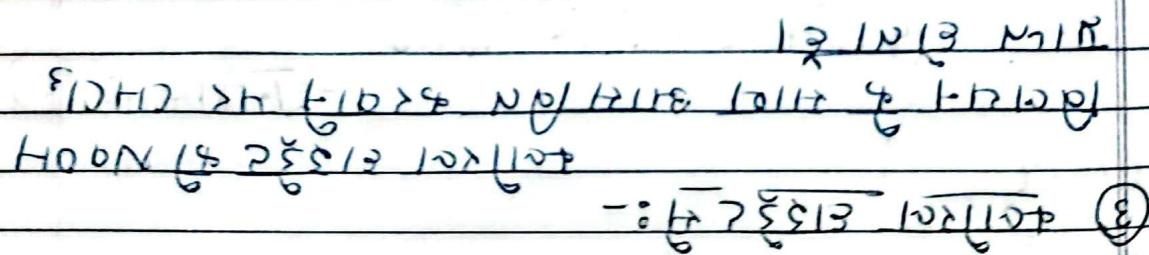
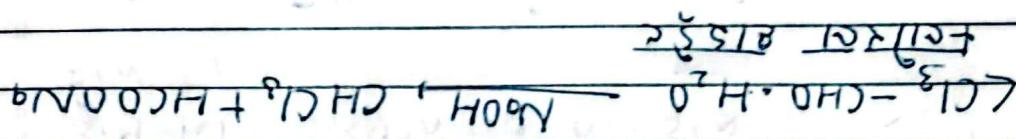
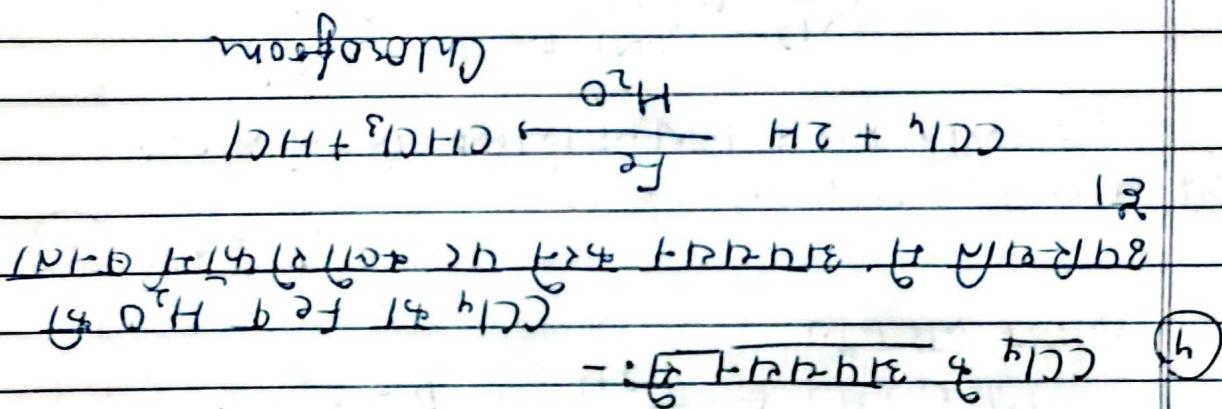
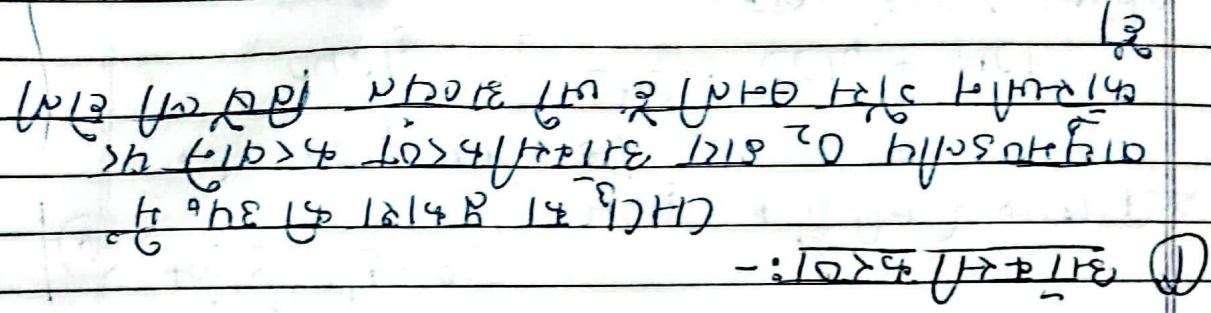
Hydrogen atom with Halogen atom

Tellurides:-



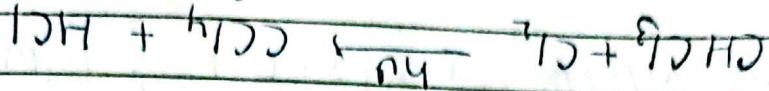
NaOH + HCl → NaCl + H₂O
 FeCl₃ + 3HCl → 3FeCl₂ + 2H₂O
 CH₃COCl + NaOH → CH₃COONa + HCl
 CH₃COONa + HCl → CH₃COCl + NaCl
 CH₃COCl + NaCl → CH₃COONa + HCl





(5) 31/1/2021

~~Chloride~~



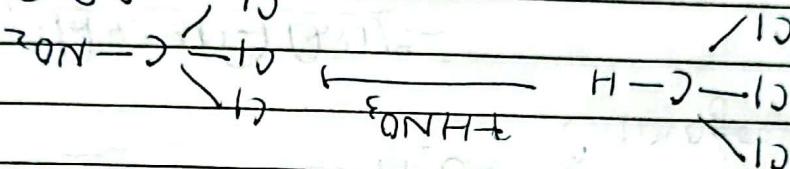
~~CCl₄ + HCl~~

~~CHCl₃ + HNO₃~~

~~-OCl > Cl / HCl / HCl~~

(D)

~~CHCl₃ + HNO₃~~

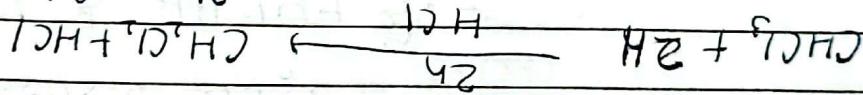


~~CHCl₃ + HNO₃~~

~~-OCl > Cl / HCl / HCl~~

(E)

~~CHCl₃ + HNO₃~~



~~CHCl₃ + HNO₃~~

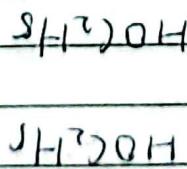
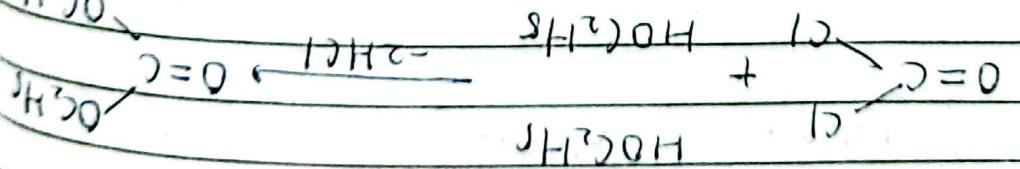
~~CHCl₃ + HNO₃~~

~~-OCl > Cl / HCl / HCl~~

(F)

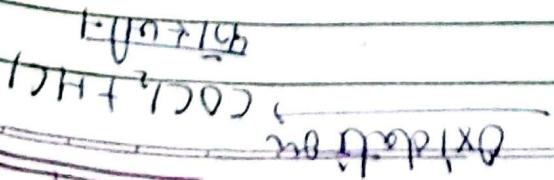
~~CHCl₃ + HNO₃~~

~~CHCl₃ + HNO₃~~

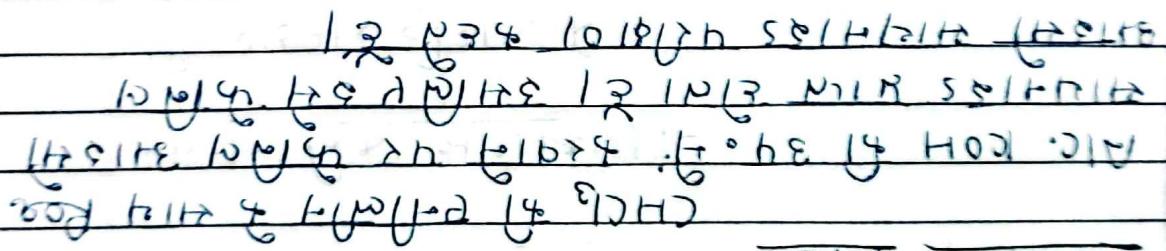
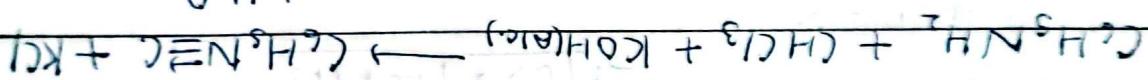
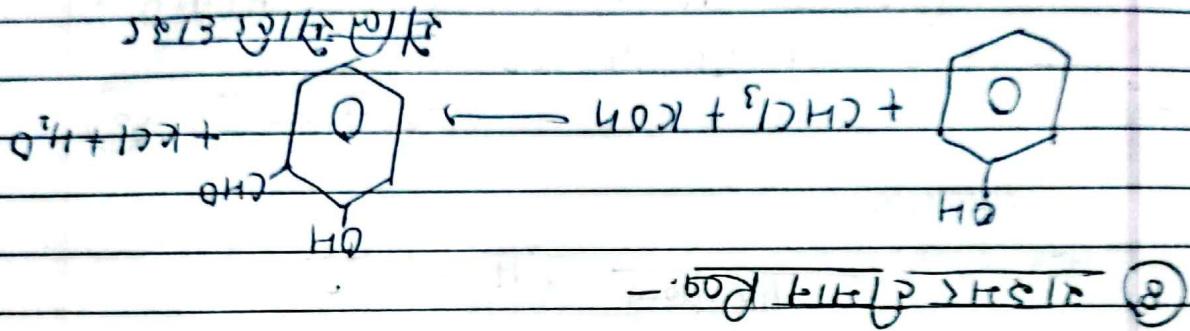
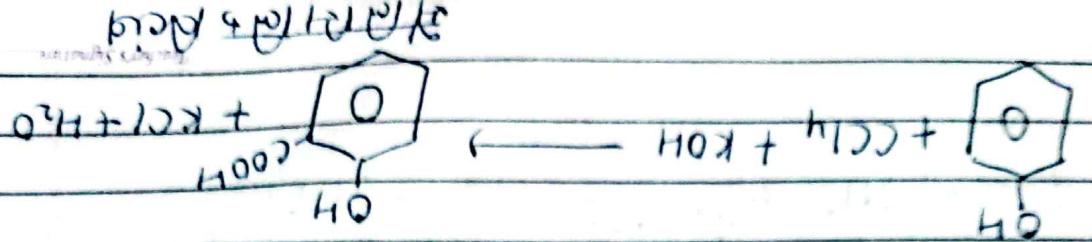


~~CHCl₃ + HNO₃~~

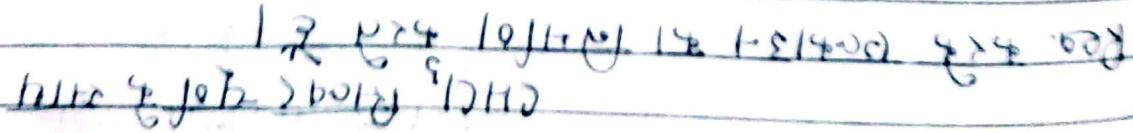
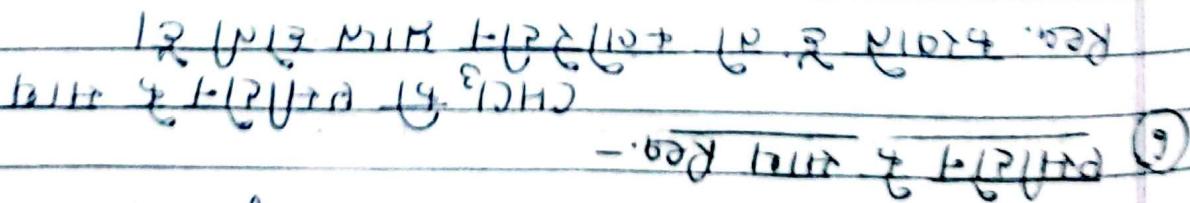
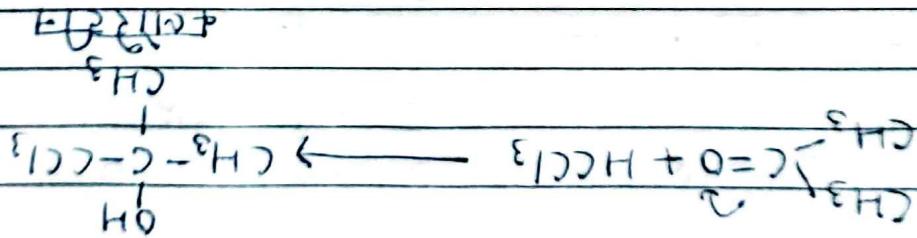
~~CHCl₃ + HNO₃~~



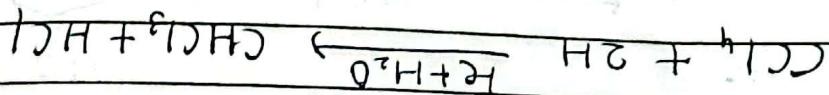
~~CHCl₃ + HNO₃~~



(9) ~~Alkalies का अध्ययन~~



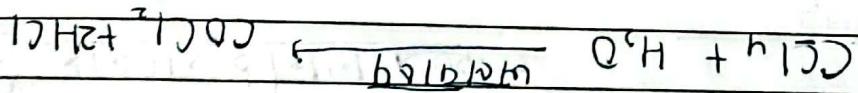
(14) ~~Alkalies का अध्ययन~~



~~34.0 g CCl₄ reacts with 34.0 g H₂O to form 34.0 g CHCl₃~~

3) ~~34.0 g Fe reacts with 34.0 g H₂O to form 34.0 g HCl~~

~~34.0 g Fe reacts with 34.0 g H₂O to form 34.0 g HCl~~

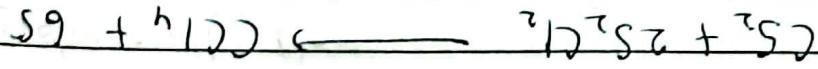


~~34.0 g CCl₄ reacts with 34.0 g H₂O to form 34.0 g COCl₂ and 34.0 g HCl~~

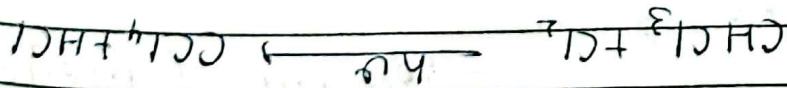
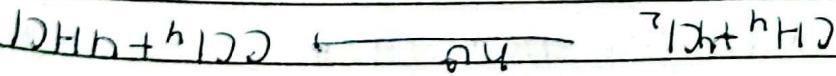
2) ~~34.0 g COCl₂ reacts with 34.0 g NaOH to form 34.0 g NaCl~~

① ~~34.0 g NaCl reacts with 34.0 g Ba(OH)₂ to form 34.0 g BaCl₂~~

~~34.0 g BaCl₂ reacts with 34.0 g Na₂SO₄ to form 34.0 g BaSO₄~~



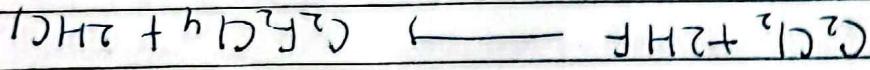
~~34.0 g CCl₄ reacts with 34.0 g Cl₂ to form 34.0 g $\text{CCl}_4 \cdot \text{Cl}_2$~~



~~34.0 g CCl₄ reacts with 34.0 g Cl₂ to form 34.0 g $\text{CCl}_4 \cdot \text{Cl}_2$~~

~~34.0 g $\text{CCl}_4 \cdot \text{Cl}_2$ decomposes to form 34.0 g CCl₄ and 34.0 g Cl₂~~

Teflon Handmade -

~~Page 311 - 112~~~~Page 311 - 113~~

~~Page 311 - 114~~
~~CFCl₃ + HF → CF₃Cl₂ + HCl Reg. No. 42417~~

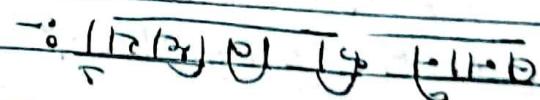
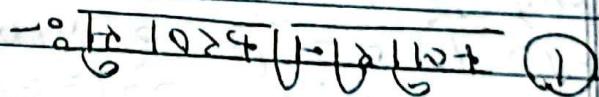
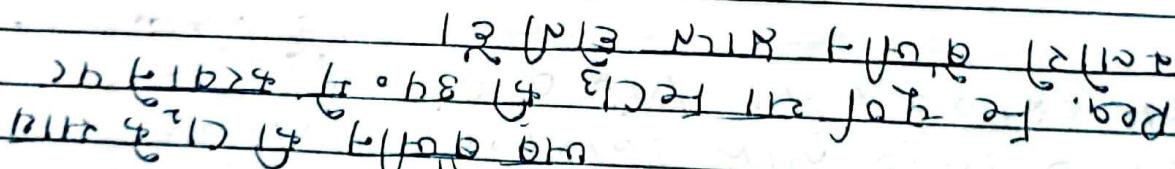
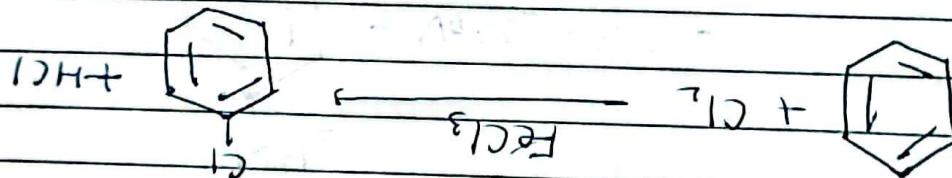
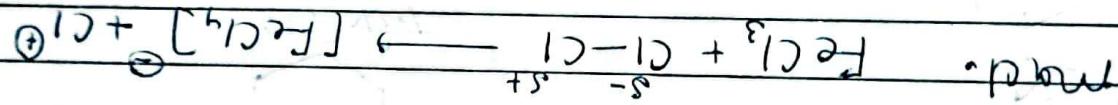
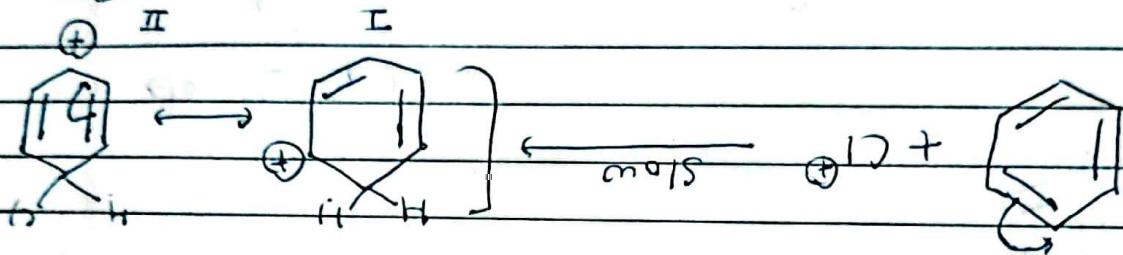
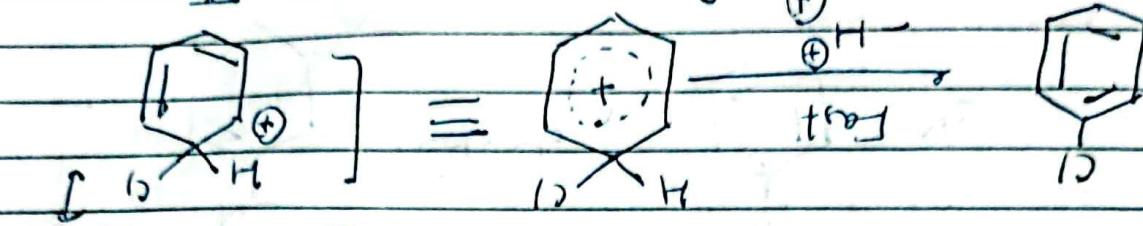
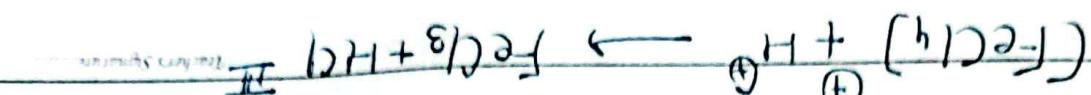
~~Page 311 - 114~~~~Page 311 - 115~~

~~HF + CCl₄ Reg. No. 42417~~
~~CFCl₃ + HCl → CF₃Cl₂ + HF~~

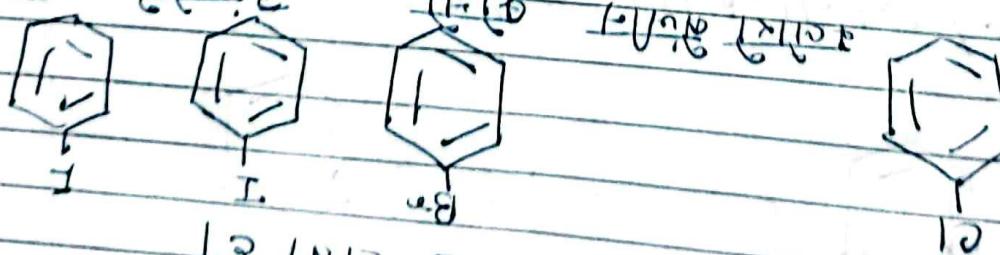
~~Page 311 - 116~~

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Page:	27

$$\left. \begin{array}{l} E = Z \quad (N.o.) \\ H + I = Y \quad (N.o.) \\ C - L = X \quad (N.o.) \end{array} \right\}$$



Benzene Benzene Benzene



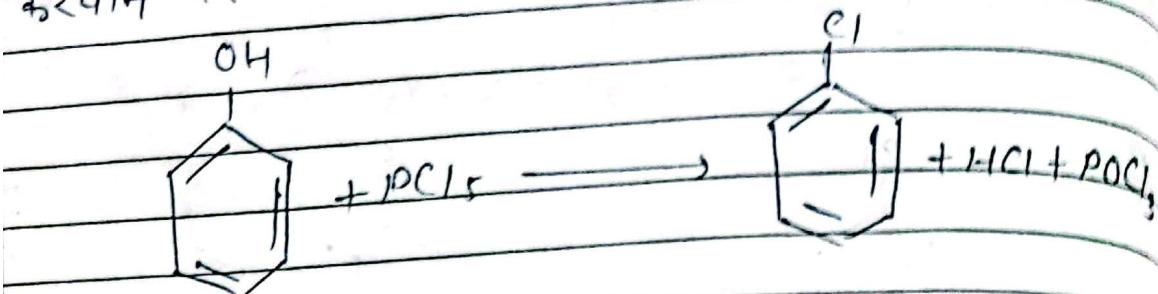
Compound. Alkyl Halide

Alkyl Halide:

Date:	1/1/2023
Page:	1

② फीनोल से:-

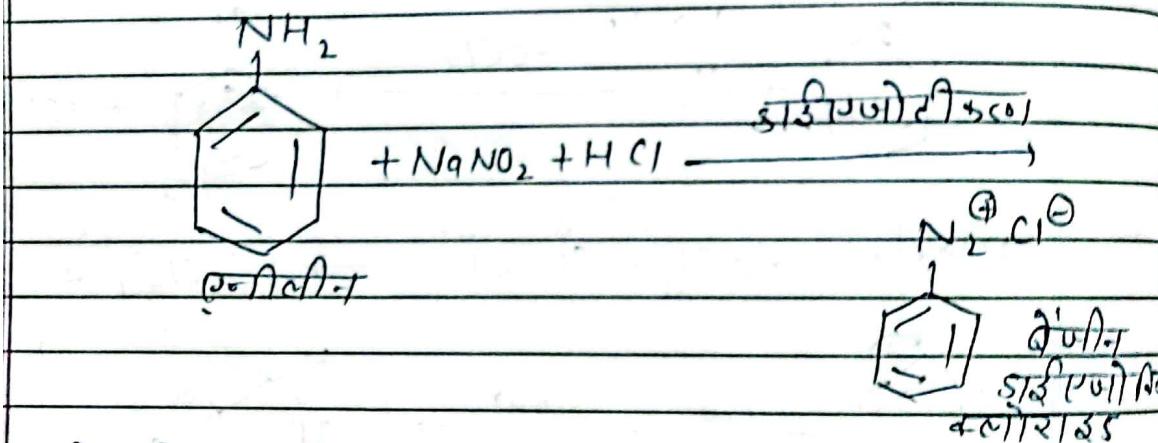
फीनोल की PCl_5 से रासायनिक प्रक्रिया करपाने पर क्लोरी बैंगन माप होता है।



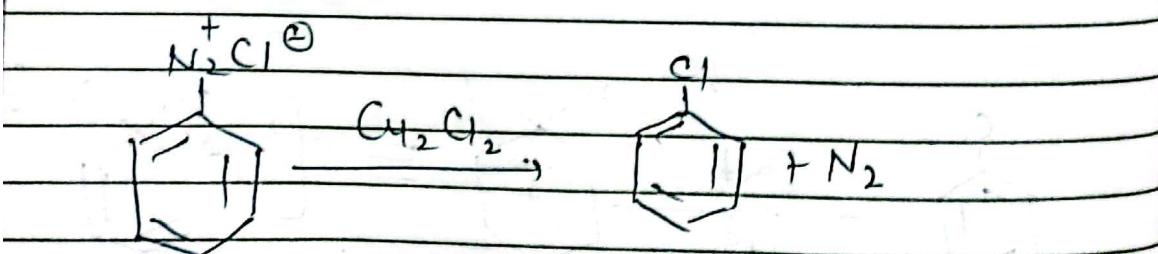
③ बैंगन दाई ऐजो विधम क्लोराइड (सेंट्रल पर)

गात्रमान अभि. -

इस विधि में प्राप्यमान एथिल
एमिन का सीध्यमनावद्वारा उत्पन्न HCl की ३५% में
दाई ऐजोटीकरण करपाते हैं तो बैंगन दाई
ऐजो विधम क्लोराइड माप होता है जिससे
इसी Reactions द्वारा एरिल हैलाइड बना
सकते हैं।



सेंट्रल पर



Teacher's Signature.....

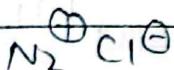


$C_4, Br_2,$



+ N_2

गाटर प्र० रेग.



C_4



+ N_2



KI

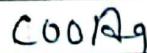


+ $N_2 + KCl$

(4)

काबोरिसलिक अम्ली के रिएक्टलेशनों की श्रीमीटा
हुन्सडी कर अधिक्रिया:-

काबोरिसलिक अम्ली के
सिल्वलेशनों की श्रीमीटा के साथ $R_o g.$ करवाने पर
एरिल हेलाइ (श्रोणी वेंजीन) बनता है।



+ $Br_2.$



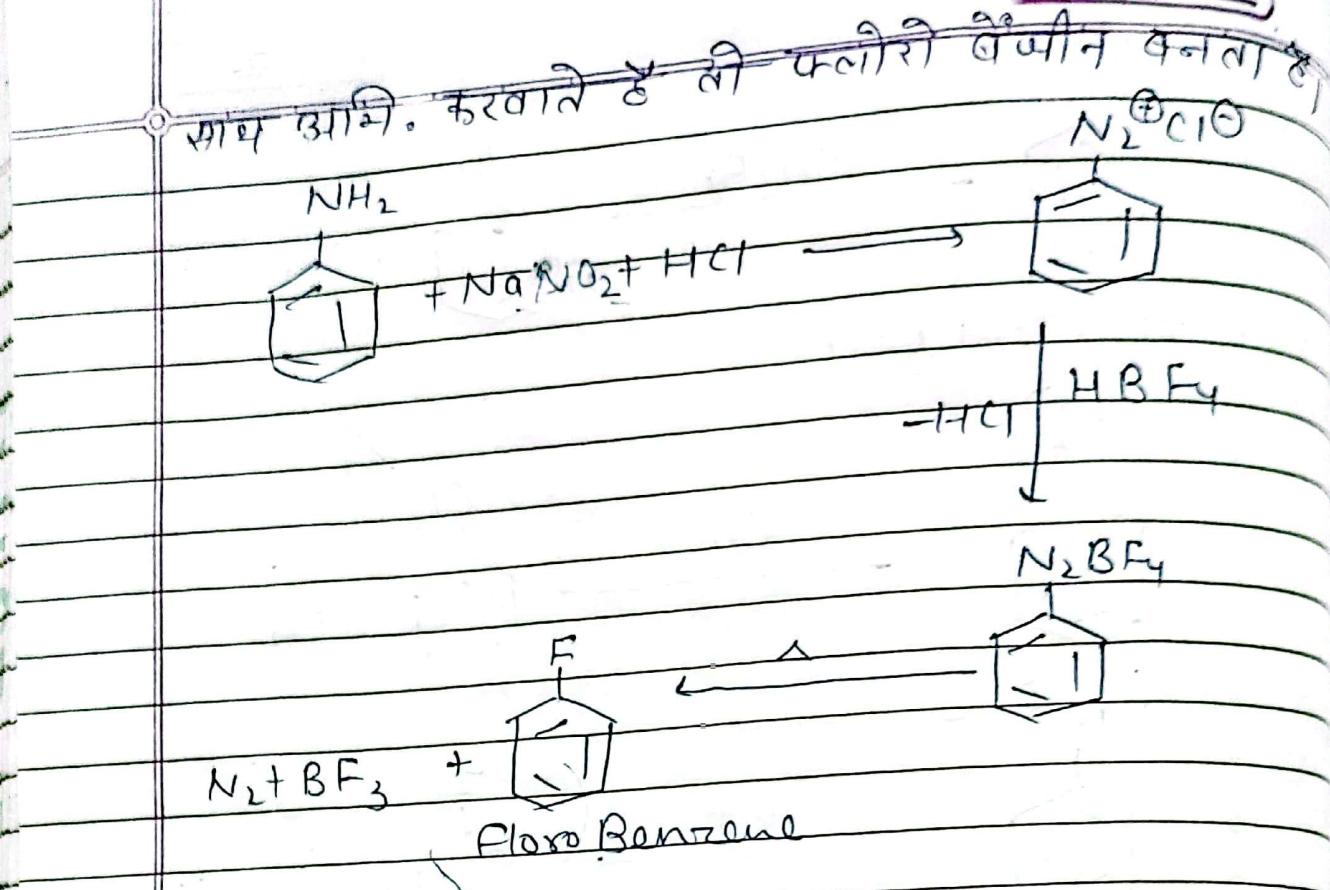
+ $CO_2 + AgBr$

श्रोणी वेंजीन

(5) बाल्जरिमान $R_o g.$ -

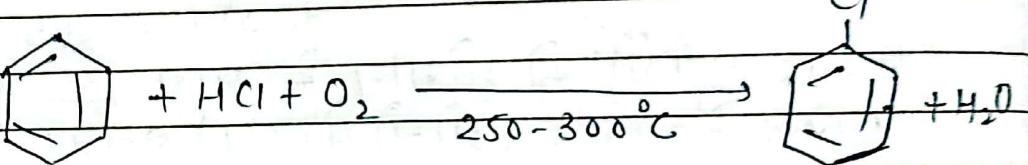
इस विधि का प्रयोग ऐल
फलोराइड बनाने में किया जाता है इसमें श्रीमीटा
डाइएजीबियम लेपा की फलोरी श्रीरं आइके

Teacher's Signature.....

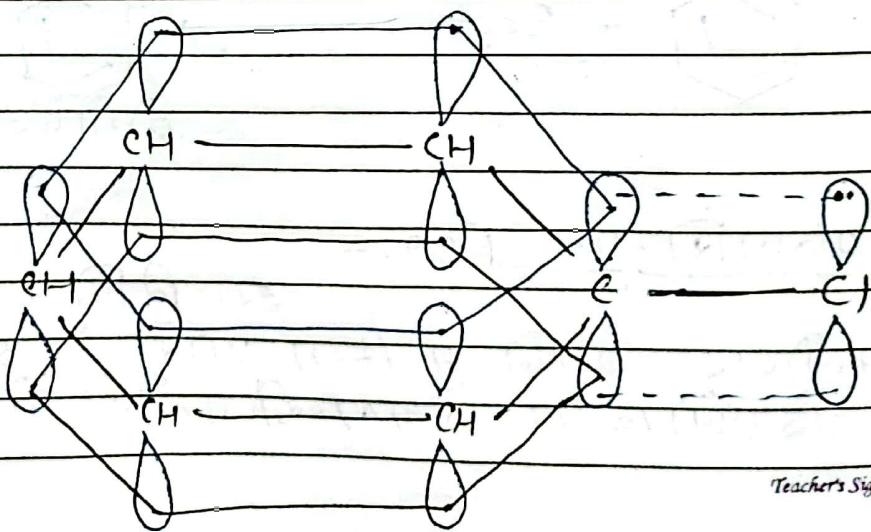


6

रेशिंग मूल्य द्वारा :- बैनिं की वायु व मूला के साध ($250 - 300^\circ\text{C}$) Temp पर गर्ज करने पर कलोरो बैनिं ब्राप्ट होता है।



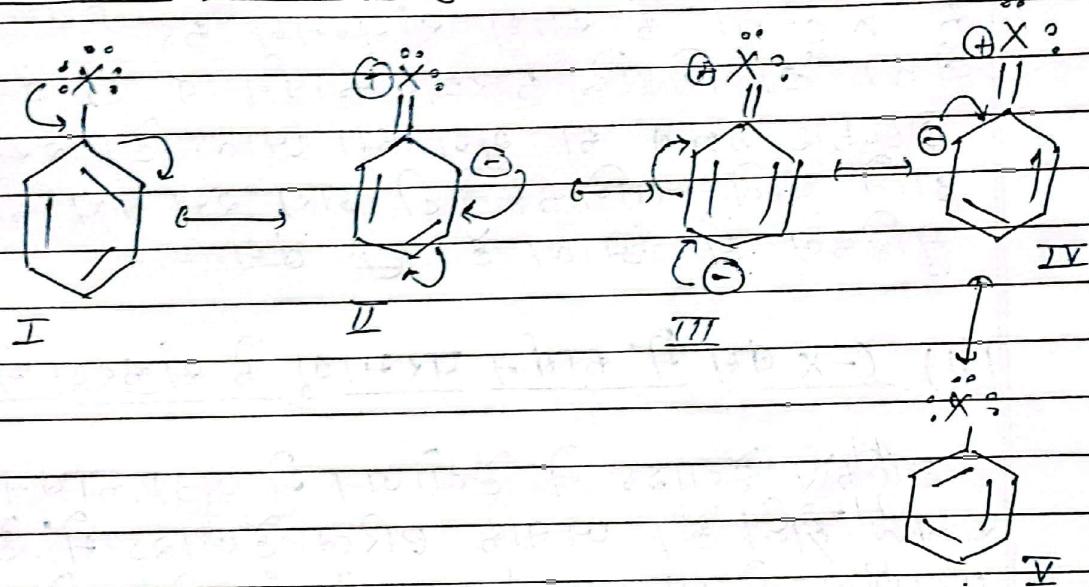
ईरो हैलाइड की MO सर्चना :-



Teacher's Signature

एरिल हैलाइड की संरचना बैंबीन की संरचना के समान ही ही होती है परंतु एरिल हैलाइड में हैलोजन परमाणु बैंबीन Ring के साथ जुड़े रहते हैं जिन पर p Orbital में lone pair 34. दीते हैं। बैंबीन Ring के कार्बन परमाणुओं के p कक्षाक और Halogen के lone pair के p कक्षाक 14. दीते हैं। जैससे ये सम्पादवीय अतिव्यापन करके विस्तारित पाइआट्रिक बंध बनाते हैं जिसके कारण कार्बन व हैलोजन के मध्य Double Bond का गुण आ जाता है।

एरिल हैलाइड की अनुनादी संरचनाएँ:-



एरिल हैलाइड में हैलोजन के l.p. व बैंबीन Ring के गे-आपस में संयुगमन करते हैं जिसके कारण कार्बन व हैलोजन के मध्य Double bond का गुण आ जाता है।

रासायनिक उत्पाद की यात्रा:-

(i) नाभि के स्वेच्छी प्रतिरक्षण आविष्कार:-

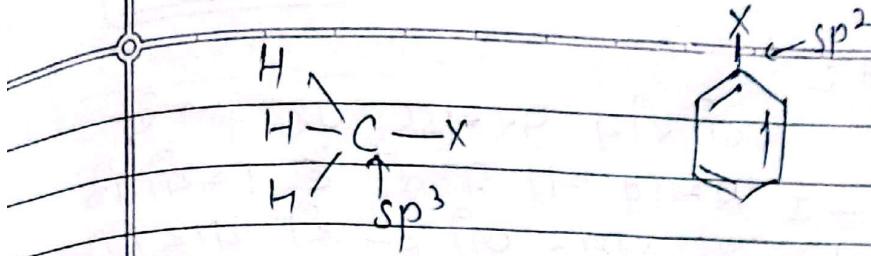
नाभि के स्वेच्छी प्रतिरक्षण आविष्कार की एलिल हैलाइट ही प्रदर्शित कर सकते हैं। एरिल हैलाइट ने आविष्कार की प्रदर्शित नहीं करता। उसके बिना (i) कारण हैं—

(ii) अनुनाद प्रभाव:-

एरिल हैलाइट में हैलोजन परमाणु पर उपरिधात Mope Ryde e- Benzene ring के C=C -नीके साथ संयुक्त हो जाता है। इस कारण Double bond का गुण आ जाता है। इस कारण आने वाले नाभिक स्वेच्छी होते हारा इस बंध की तीव्रता मुश्किल ही जाता है। Ex :-

(iii) $C-X$ बंध में कार्बन परमाणु के संकरण में अंतरः

एलिल हैलाइट में हैलोजन से जुड़ा कार्बन SP^3 संकरित होता है। जबकि एरिल हैलाइट में हैलोजन परमाणु से जुड़ा कार्बन SP^2 संकरित होता है। इस कारण एरिल हैलाइट में C का गुण उत्तराधारा जाता है। एलिल हैलाइट में $C-C$ बंध लंबाई 177 pm होती है। जबकि एरिल हैलाइट में $C-C$ बंध लंबाई 169 pm होती है। नाभिक स्वेच्छी प्रतिरक्षण द्वारा एलिल हैलाइट की तुलना में एरिल हैलाइट का क्रियाशील होता है।



③ फैबिल धनायन का अस्थायित्वः-

आयनन के कारण

PRIM हेलाइड में हेलीजन परमाणु छाएव बिकल जाता है जिसके कारण फैबिल धनायन बनता हो केविल धनायन अनुनाद के द्वारा स्थायी नहीं हो पाएगा और HN⁺ नामिक स्नेही प्रतिस्थापन Reg. Show नहीं करेगा।

② अन्य अभिक्रियाएँ:-

ब्लॉरी बैंधीन दी सकार की Reg.

Show करता है।

(i) बैंधीन वलय की Reg.

(ii) ब्लॉरीन परमाणु की Reg.

① बैंधीन वलय की Reg.-

ऐरोमेटिक Ring ये e- स्नेही प्रतिस्थापन अभिक्रियाएँ देते हैं। ब्लॉरीन परमाणु के बैंधीन पर दी प्रभाव पड़ते हैं।

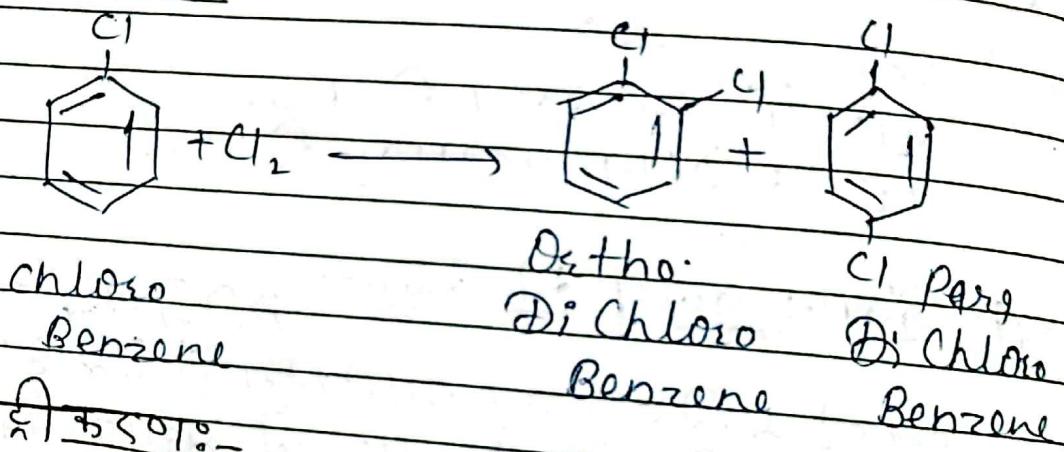
② +M प्रभावः-

जब Ring पर lone pair होता समूह जुड़ा होती अनुनाद के द्वारा Ring में e- धनत्व बढ़ जाता है जिससे वलय की Ortho & para position पर -ive charge उपलब्ध होता है। अतः आने वाला e- स्नेही Ortho & para स्थिति पर जुड़ता है। इसलिए इन्हें आशो-पैरा विद्युत कारी कहा जाती है।

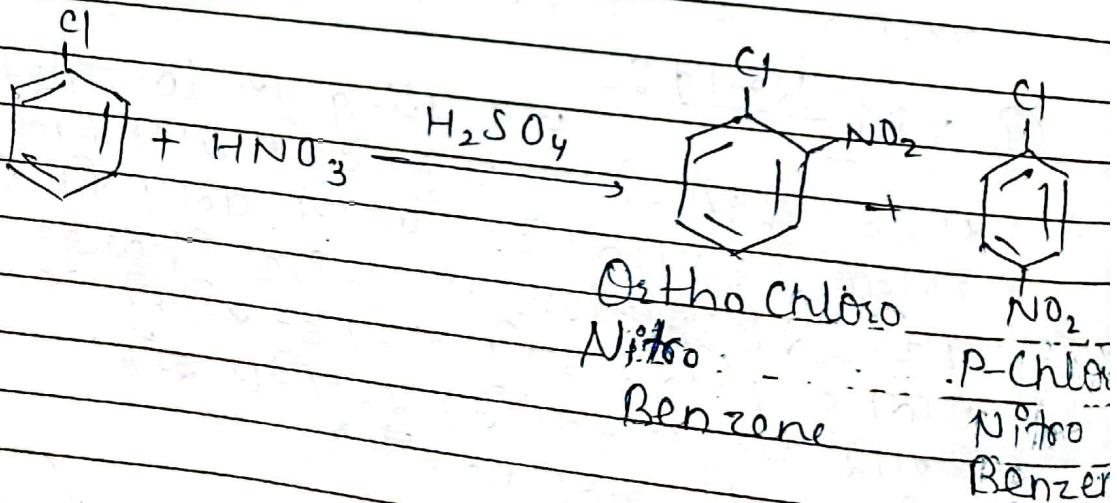
(b) प्रभाव

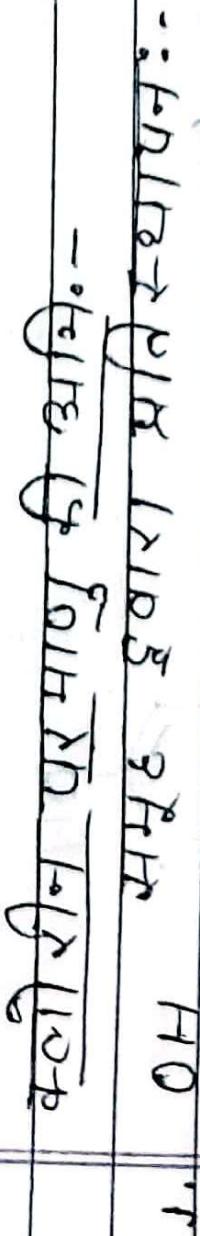
6) - I
 क्लोरीन परमाणु पर +1 के साथ-साथ -I प्रभाव भी पड़ता है। यहीं निम्नलिखित ए-आक्षर्णी समूह जो e- की ओर डिस्ट्रिब्यूट करता है जिसकी विद्युत असंतुष्टि उद्यादा होती है। इस पर -I प्रभाव उद्यादा होता है और जिसकी वजह से बेंजीन वलय e-न्यून हो जाती है। अतः e-र्जी ही स्थिरस्थापन PbCl_3 के लिए क्लोरो बेंजीन, बेंजीन की तुलना में कम उत्थाशील हो जाएगा। फिर भी Chlorobenzene कुछ Reactions show करता है जो निम्न प्रकार हैं।

① → बलोरीन-१ कृपा:-

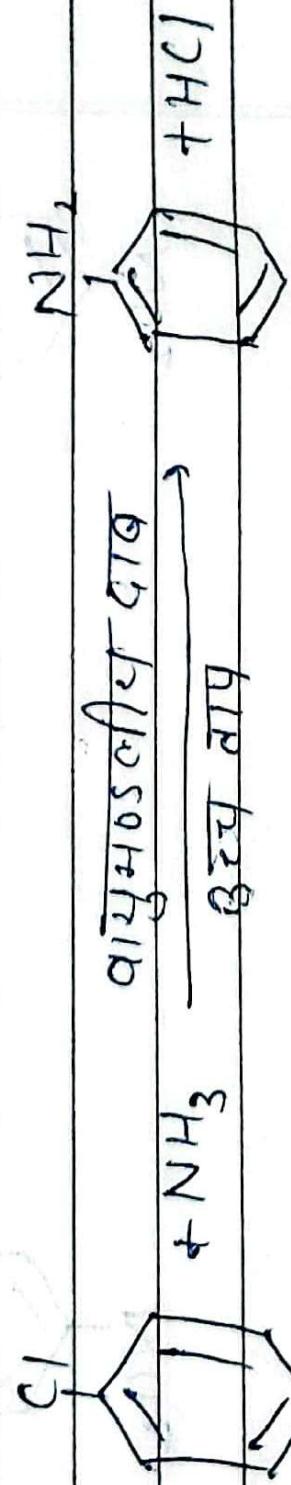


(2) \rightarrow नाइट्रो क्लोरो:





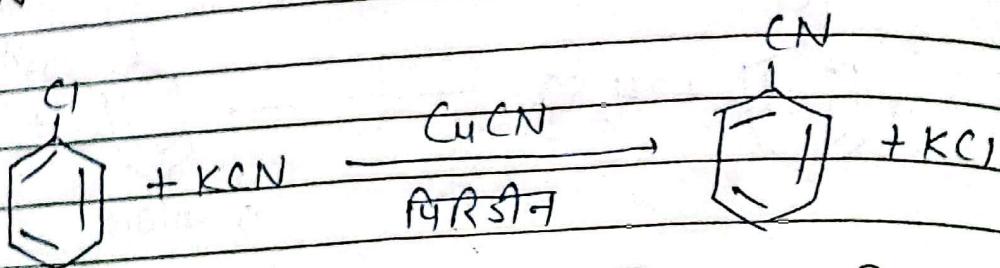
2.



Teacher's Signature.....

CN समृद्ध द्वारा प्रतिस्थापनः -

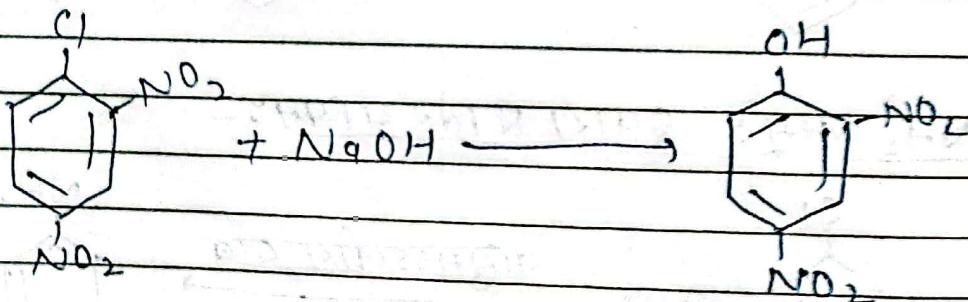
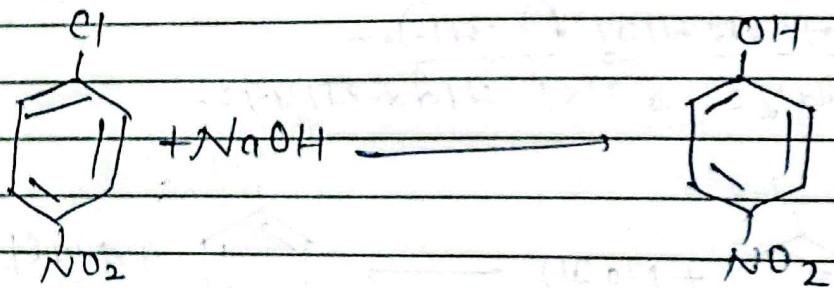
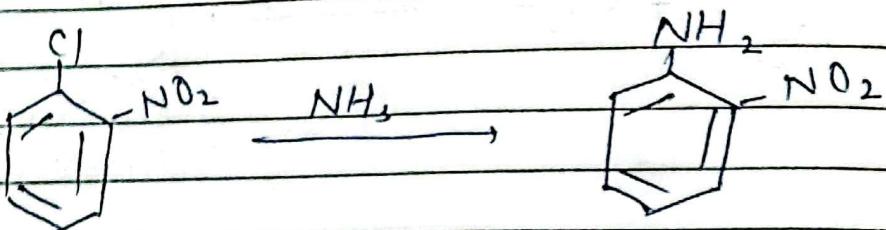
3.



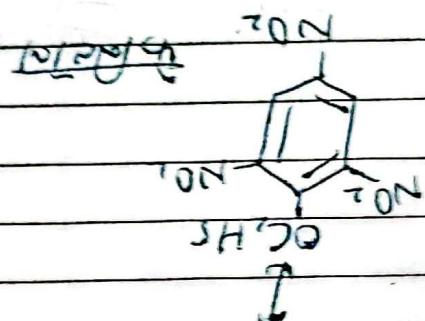
नानिक सनेही प्रतिस्थापन आणि पर प्रतिस्थापने

का प्रभावः -

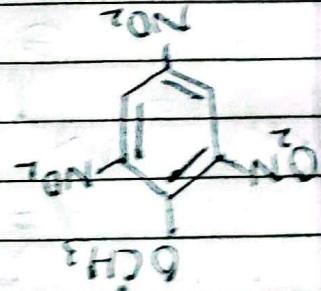
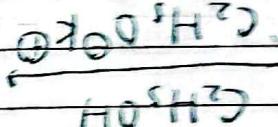
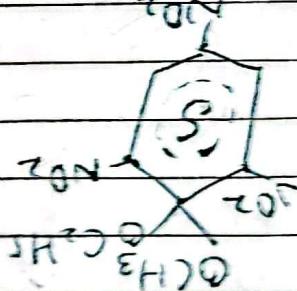
ब्लॉरी डेंबीन में आधीरे तर पैरा रेखा पर कोई इलैंग आकर्षित करने वाला समृद्ध भुजा हो तो क्रियाशीलता बढ़ जाती है।



وَهُوَ مُكَبَّلٌ بِالْمُنْدَرِ وَالْمُنْدَرِ
 وَهُوَ مُكَبَّلٌ بِالْمُنْدَرِ وَالْمُنْدَرِ
 وَهُوَ مُكَبَّلٌ بِالْمُنْدَرِ وَالْمُنْدَرِ
 وَهُوَ مُكَبَّلٌ بِالْمُنْدَرِ وَالْمُنْدَرِ



وَهُوَ مُكَبَّلٌ بِالْمُنْدَرِ



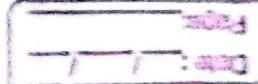
CH_3

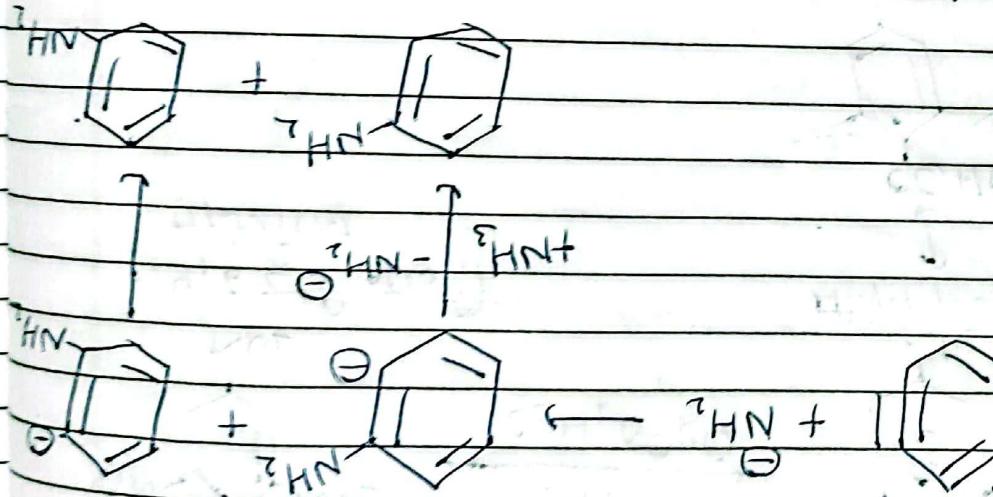
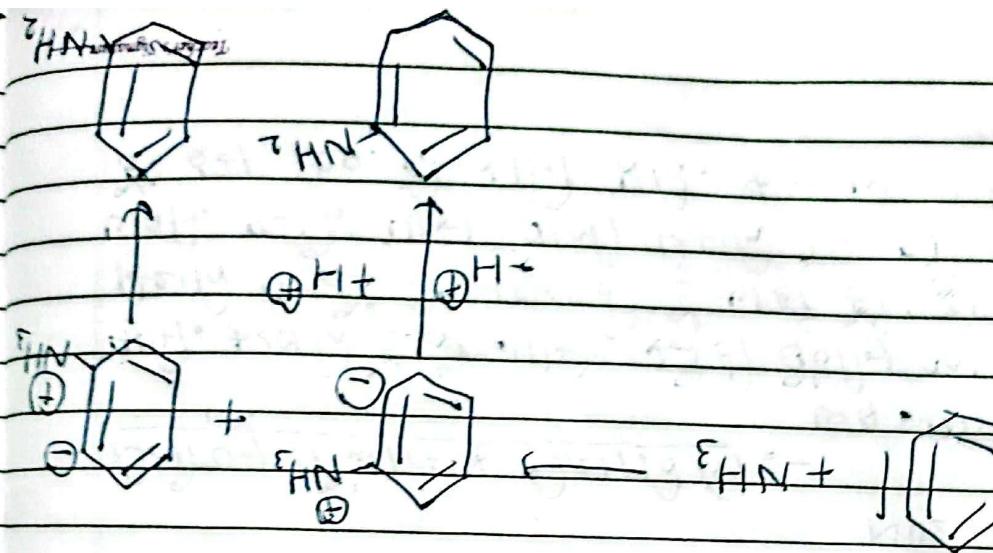
وَهُوَ مُكَبَّلٌ بِالْمُنْدَرِ وَالْمُنْدَرِ
 وَهُوَ مُكَبَّلٌ بِالْمُنْدَرِ وَالْمُنْدَرِ
 وَهُوَ مُكَبَّلٌ بِالْمُنْدَرِ وَالْمُنْدَرِ
 وَهُوَ مُكَبَّلٌ بِالْمُنْدَرِ وَالْمُنْدَرِ

Exampole:

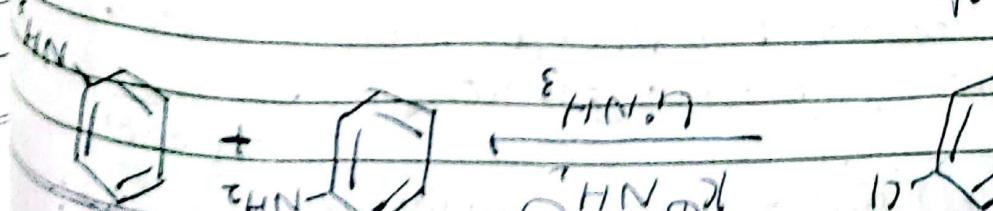
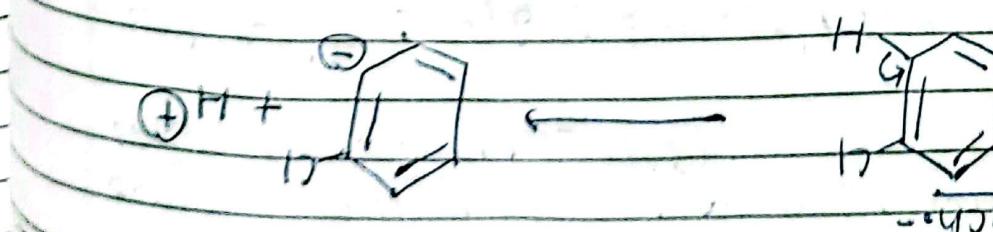
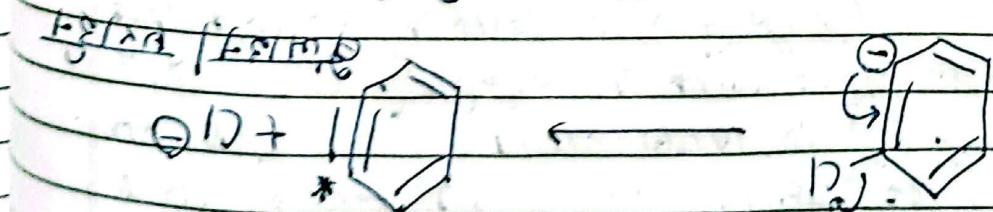
وَهُوَ مُكَبَّلٌ بِالْمُنْدَرِ وَالْمُنْدَرِ
 وَهُوَ مُكَبَّلٌ بِالْمُنْدَرِ وَالْمُنْدَرِ
 وَهُوَ مُكَبَّلٌ بِالْمُنْدَرِ وَالْمُنْدَرِ

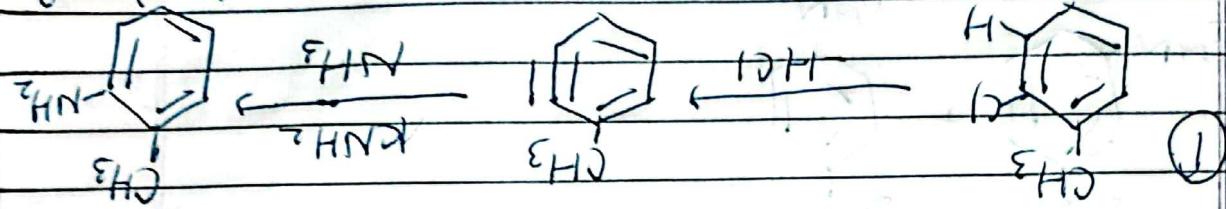
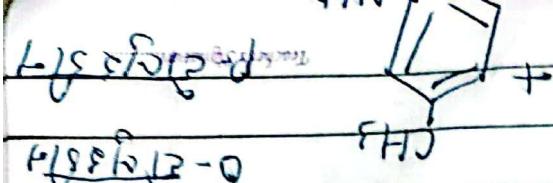
وَهُوَ مُكَبَّلٌ بِالْمُنْدَرِ وَالْمُنْدَرِ





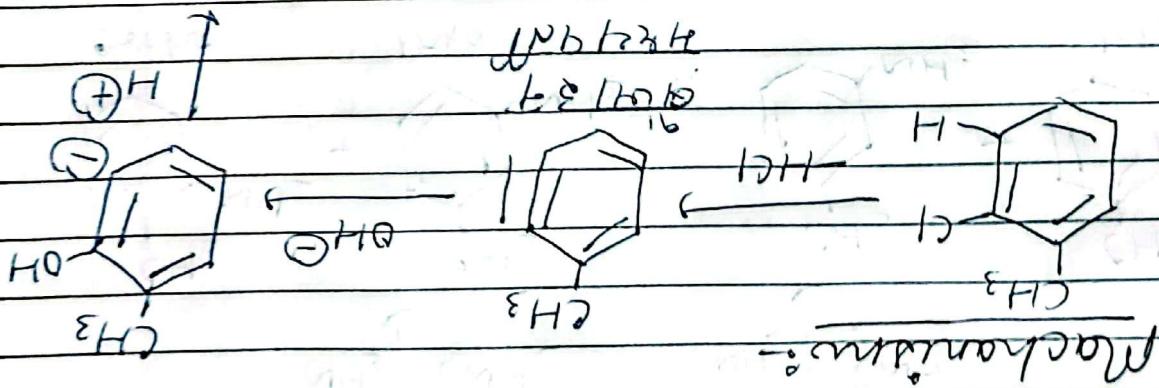
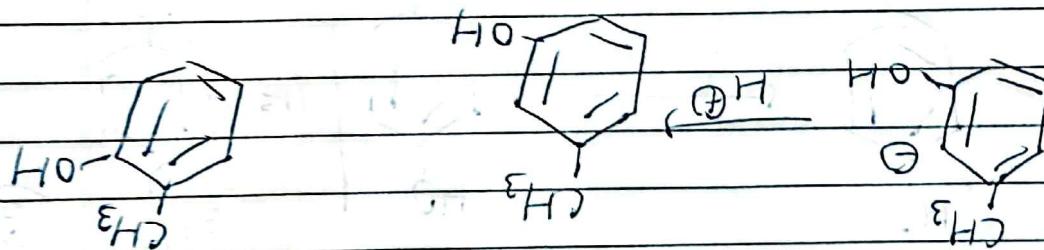
12. $\text{C}_6\text{H}_5\text{NH}_2 + \text{H}_3\text{O}^+ \rightleftharpoons \text{C}_6\text{H}_5\text{NH}_3^+$





~~Q-2/3/4/5/6~~

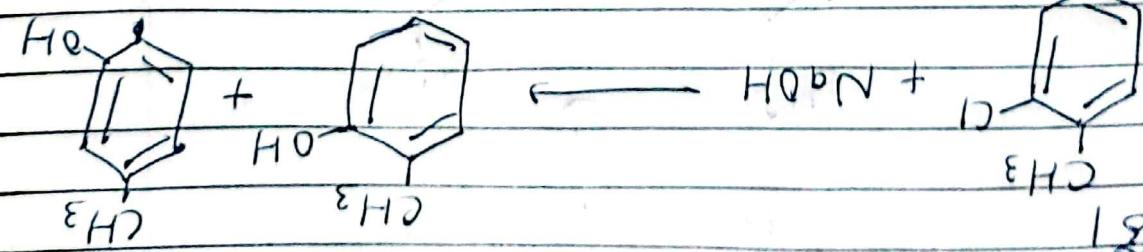
Chloro Toluene $\xrightarrow{\text{NaOH}}$ $\text{C}_6\text{H}_5\text{NH}_2(\text{CH}_3)_2$



~~Q-2/3/4/5/6~~

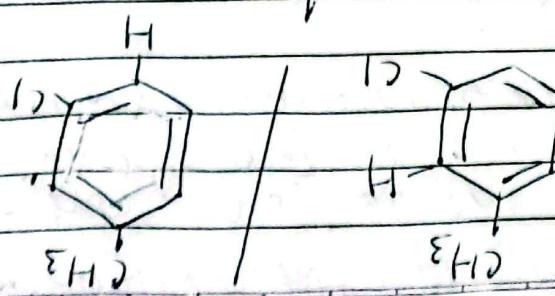
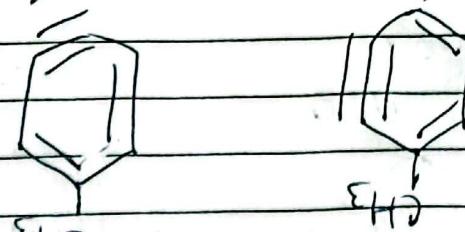
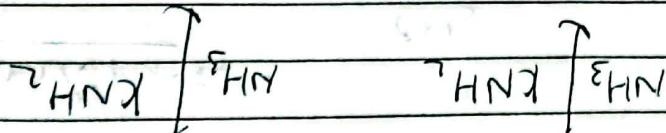
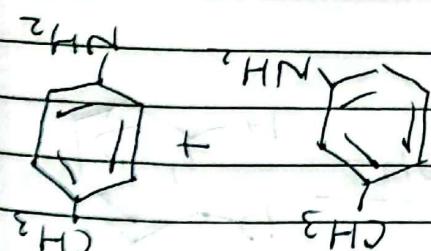
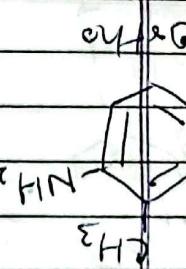
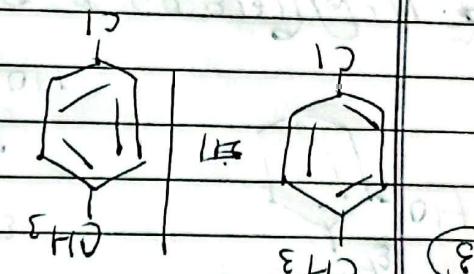
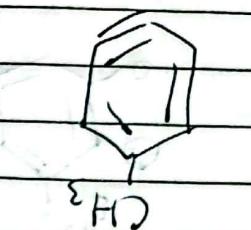
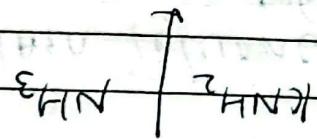
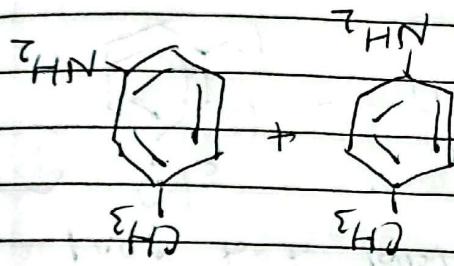
Phenol
2-Methyl-3-methylphenol

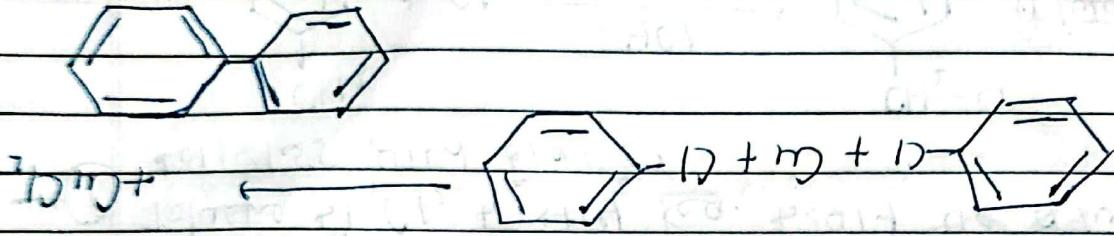
Toluene
 $\text{C}_6\text{H}_5\text{Cl}$



~~Q-2/3/4/5/6~~

2-Methyl-3-methylphenol
NaOH
Toluene + $\text{C}_6\text{H}_5\text{Cl}$





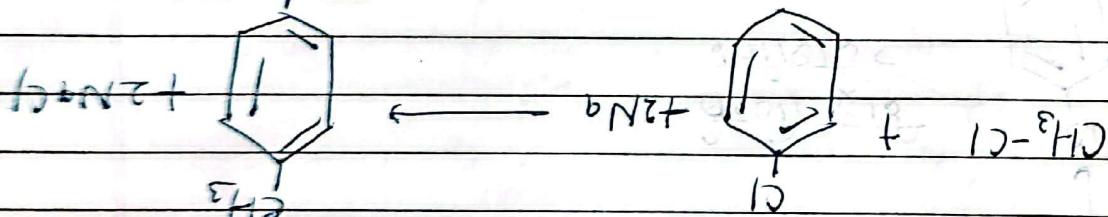
12 1/2 10/14 14

Phenyl Cu+1 + Cl- → C₆H₅CuCl

-: 1/2 10/14 14

(h)

Toluene



12 1/2 10/14 14

Hydrochloride of Phenyl Chloride

-: 1/2 10/14 14

(3)

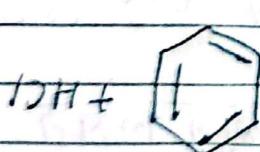
Chloride

Phenyl Magnesium



-: 1/2 10/14 14

(7)



NaOH

2H

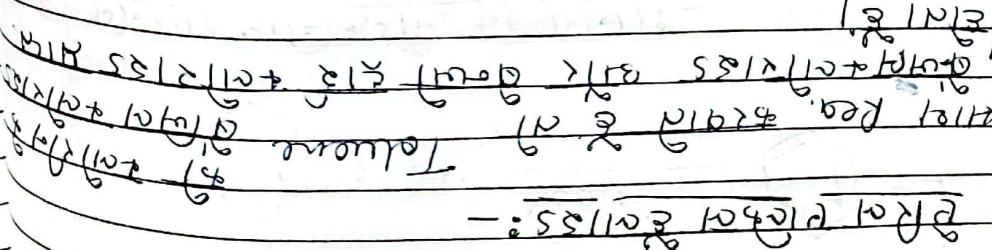
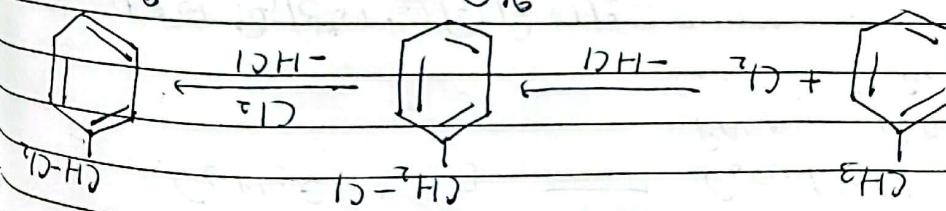
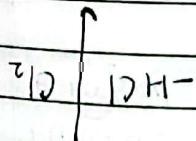
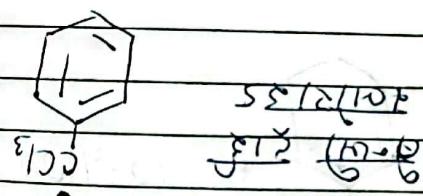
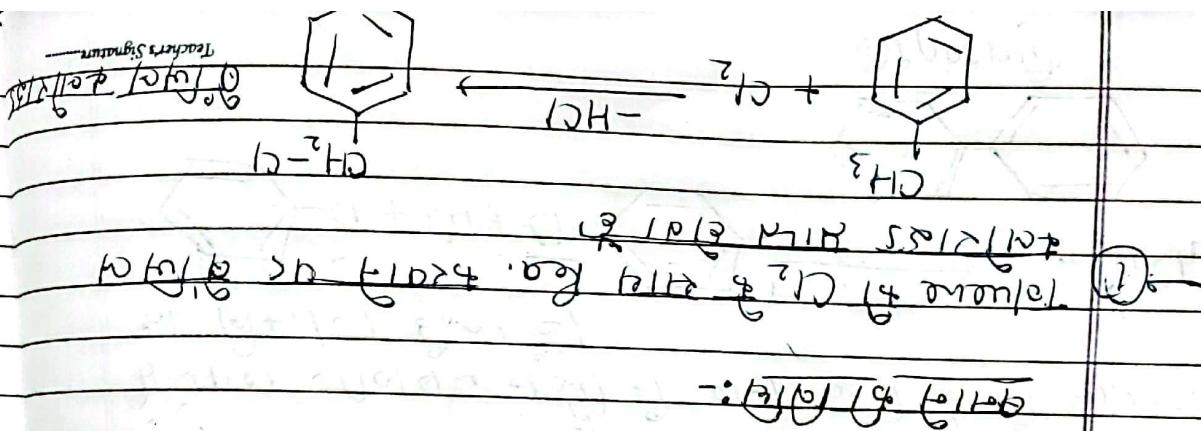


12 1/2 10/14 14

-: 1/2 10/14 14

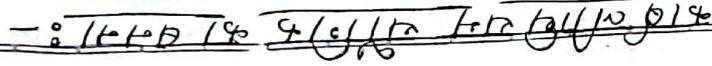
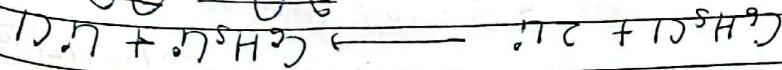
(1)

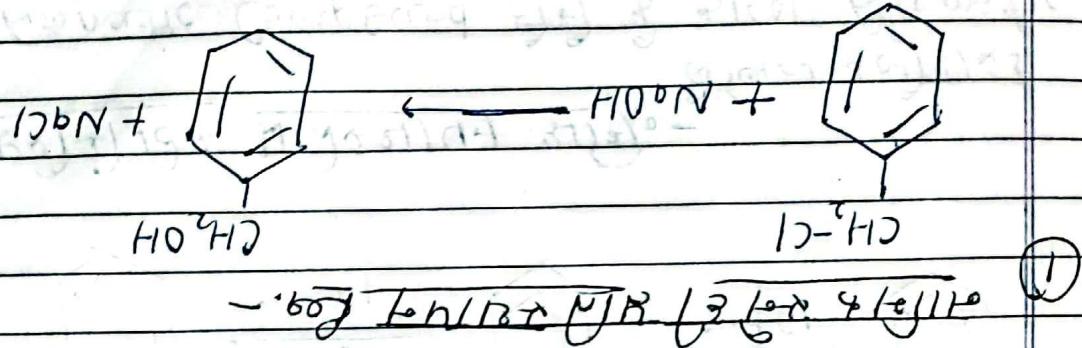
13 4 14 5 10/14 14



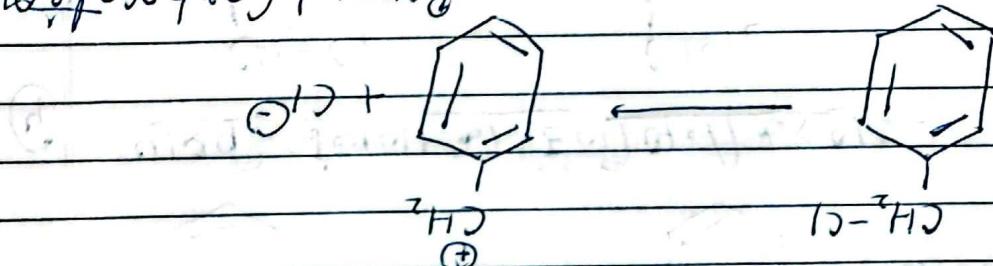
#

Chloroform

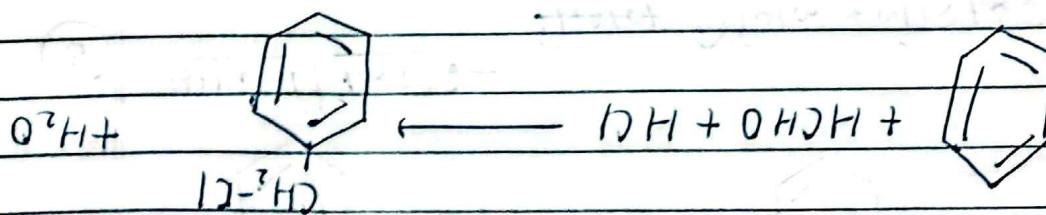




Benny's Collection

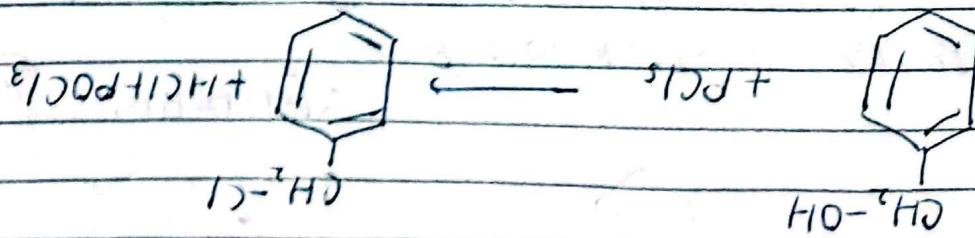


Chromatic Floating -



12 IN 13 MIN SS/100 + 104/10 8 BEST OVER 800000
13 IN 14 MIN SS/100 + 104/10 + 14

Benzyl Alcohol



BRITISH ALCOHOL POLICY

L2 S14N

~~55/12/10 2010 10:11:12~~

$$\text{CH}_2-\text{Cl} \xrightarrow{\text{HCl}} \text{CH}_2=\text{Cl} \xrightarrow{\text{HCl}} \text{CH}_2=\text{C}(\text{Cl})=\text{CH}_2$$

$\text{Benzene} + \text{H}_2\text{O} \xrightarrow{\text{acid}} \text{C}_6\text{H}_5\text{OH}$

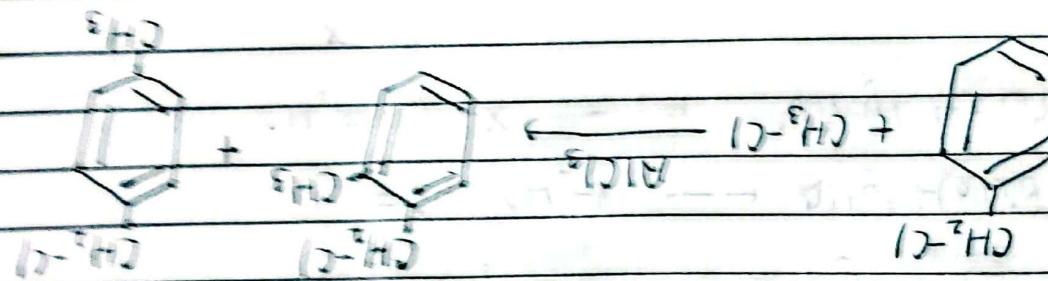
~~12/101-12 PROB 45 10.0 > 12. 10 > 4 10 > 4 U > 2118
14 55 12 110 > 10118 total~~

$$\text{C}_6\text{H}_5\text{CH}_3 + \text{HCl} \longrightarrow \text{C}_6\text{H}_5\text{CH}_2\text{Cl}$$

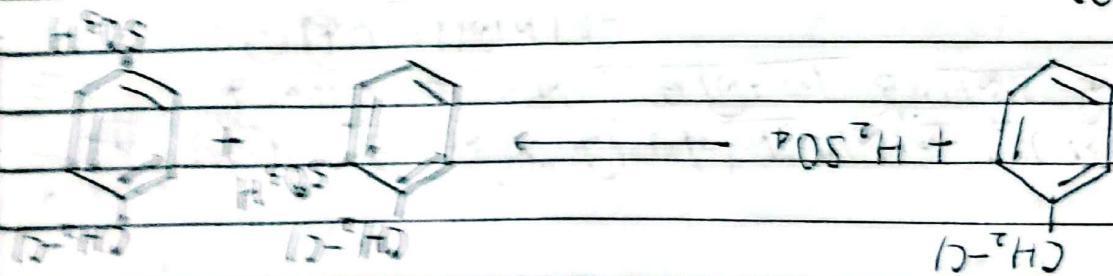
~~12 1N/3 10/1+14 14 Tolumec 24 24
10/12-11/18 14 8.5/10/10+10/10
- 10/12-11/18~~

$$\text{C}_6\text{H}_5\text{CH}_2\text{Cl} \xrightarrow[\text{KCN}, \text{HgCl}_2]{\text{CuI}} \text{C}_6\text{H}_5\text{CH}_2\text{CN}$$

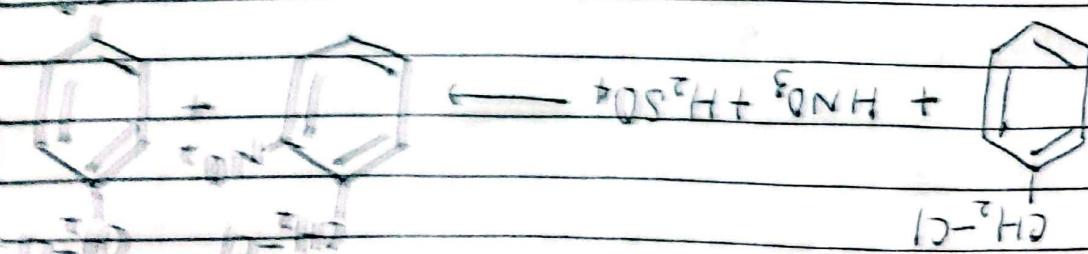
Ques 10. Aromatic substitution reactions



(3) SN1 Reactions



(2) SN1 Reactions



(1) SN1 Reactions



~~Alkyl halide + H₂O → Alkene + HX~~

~~Alkyl halide + NaOH → Alkene + NaX~~

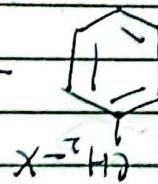
~~Alkyl halide + NaOEt → Alkene + EtX~~

~~Alkyl halide + NaHCO₃ → Alkene + HCOONa~~

~~Alkyl halide + Na₂CO₃ → Alkene + CO₂ + NaX~~

~~Alkyl halide + Na₂SO₃ → Alkene + SO₃²⁻ + NaX~~

Alkyl Alkyl Halide



Alkyl Halide



Vinyl Halide



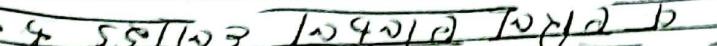
Alkyl Halide



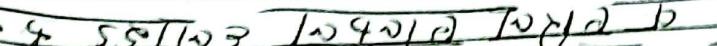
Alkyl Halide



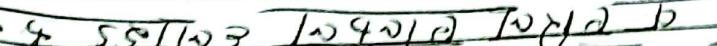
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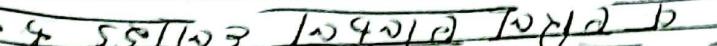
Alkyl Halide



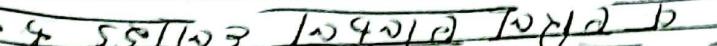
Alkyl Halide



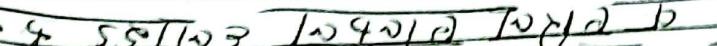
Alkyl Halide



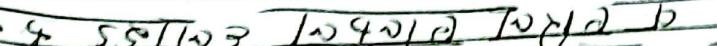
Alkyl Halide



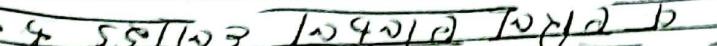
Alkyl Halide



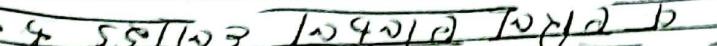
Alkyl Halide



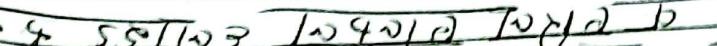
Alkyl Halide



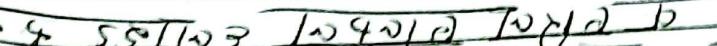
Alkyl Halide



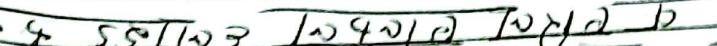
Alkyl Halide



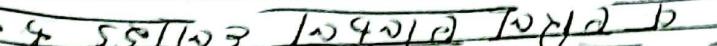
Alkyl Halide



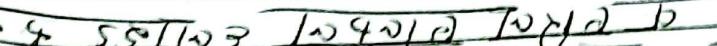
Alkyl Halide



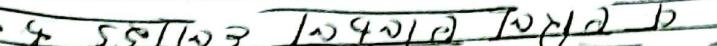
Alkyl Halide



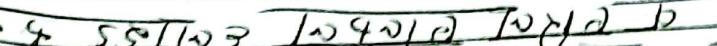
Alkyl Halide



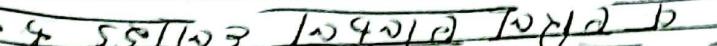
Alkyl Halide



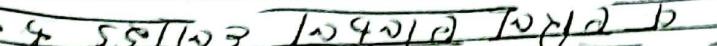
Alkyl Halide



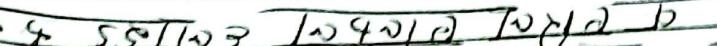
Alkyl Halide



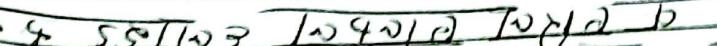
Alkyl Halide



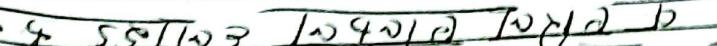
Alkyl Halide



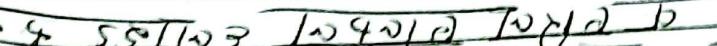
Alkyl Halide



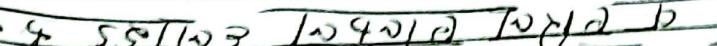
Alkyl Halide



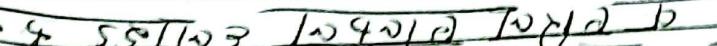
Alkyl Halide



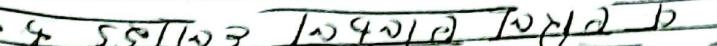
Alkyl Halide



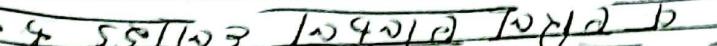
Alkyl Halide



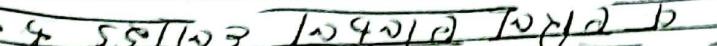
Alkyl Halide



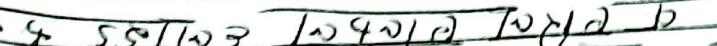
Alkyl Halide



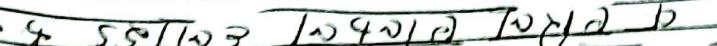
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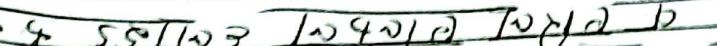
Alkyl Halide



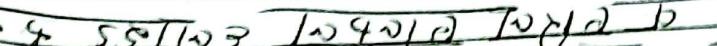
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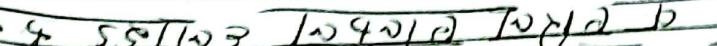
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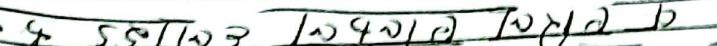
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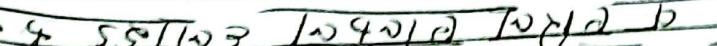
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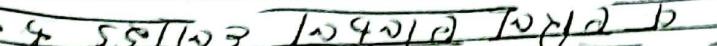
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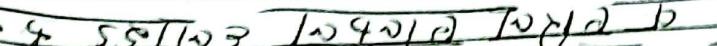
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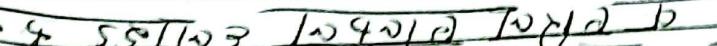
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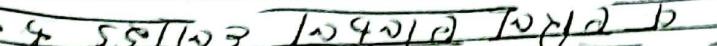
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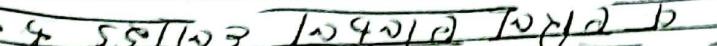
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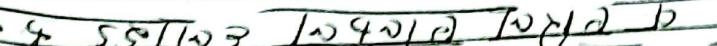
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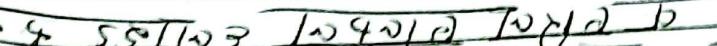
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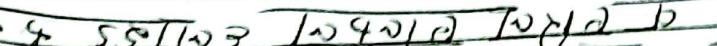
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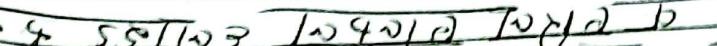
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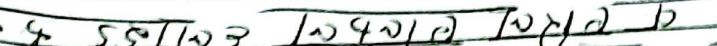
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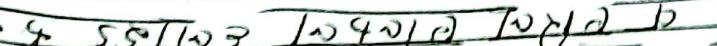
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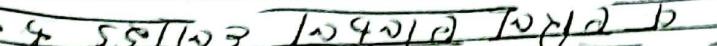
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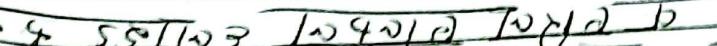
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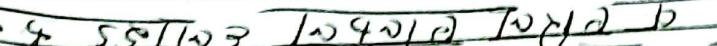
Alkyl Halide



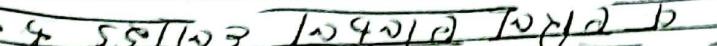
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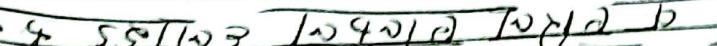
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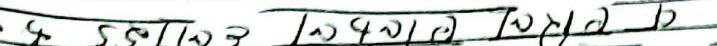
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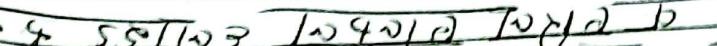
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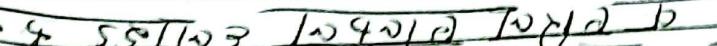
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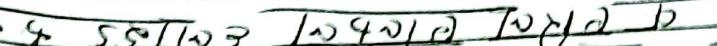
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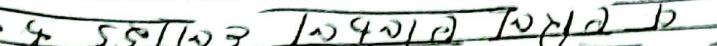
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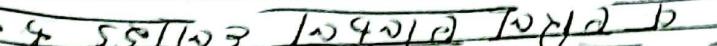
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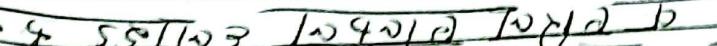
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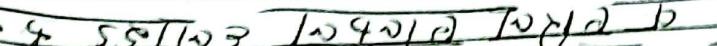
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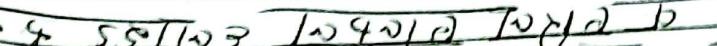
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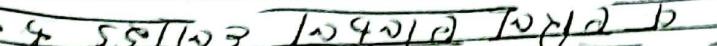
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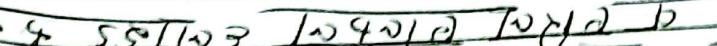
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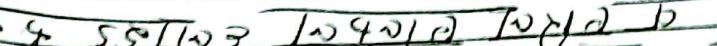
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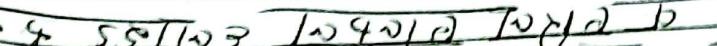
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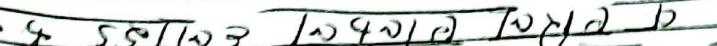
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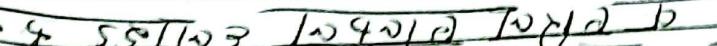
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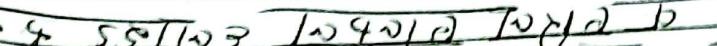
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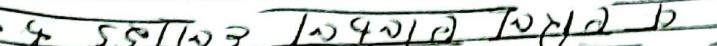
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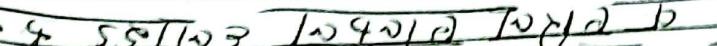
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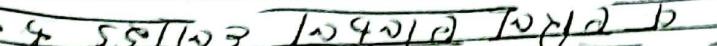
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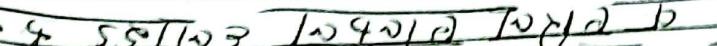
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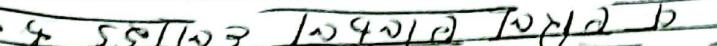
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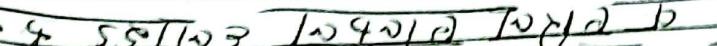
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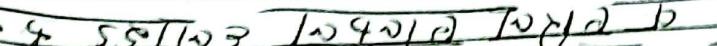
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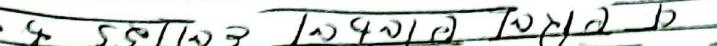
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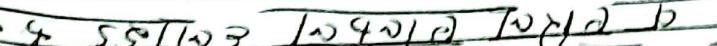
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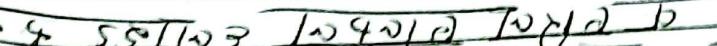
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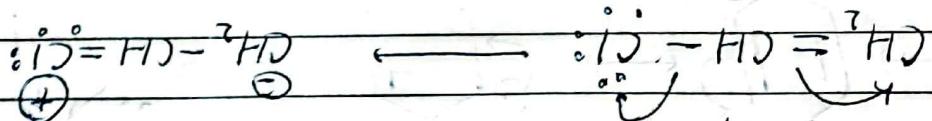


Alkyl Halide



12 IN 12

It's part of a larger topic called multiple inheritance.
Multiple inheritance is when a class inherits from multiple parents.
Multiple inheritance is often used to implement composition.
Composition is when a class contains another class as a member variable.

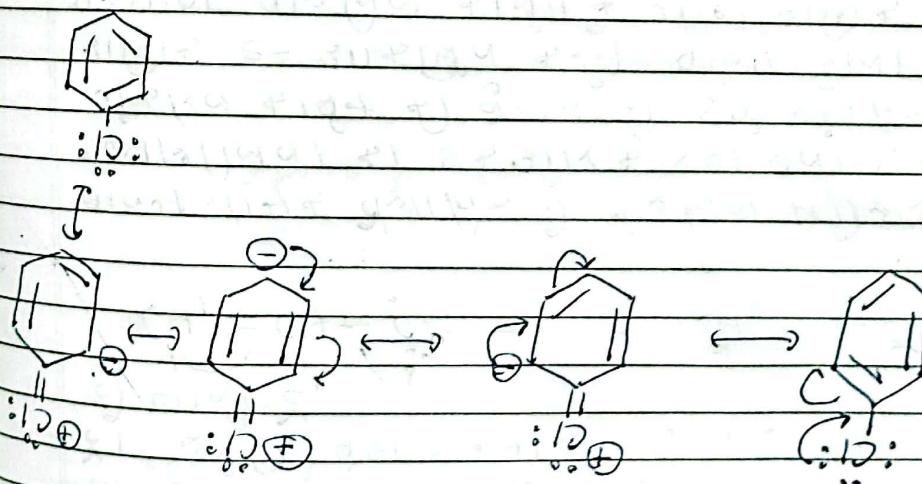
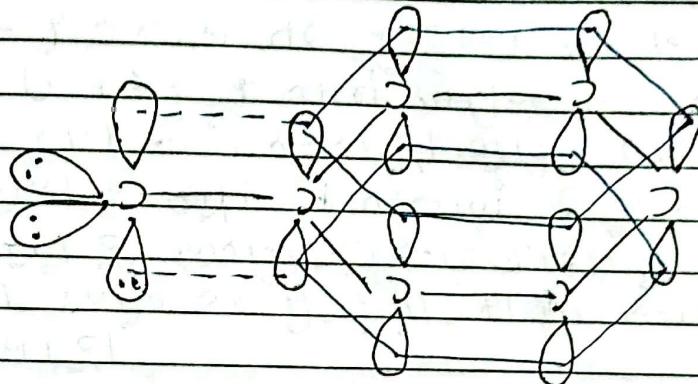


12 (n) b (?)

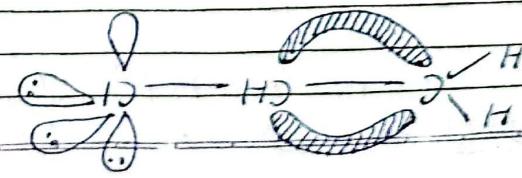
Double Bond a lone pair of electrons
top/bottom followed by 1014-1014

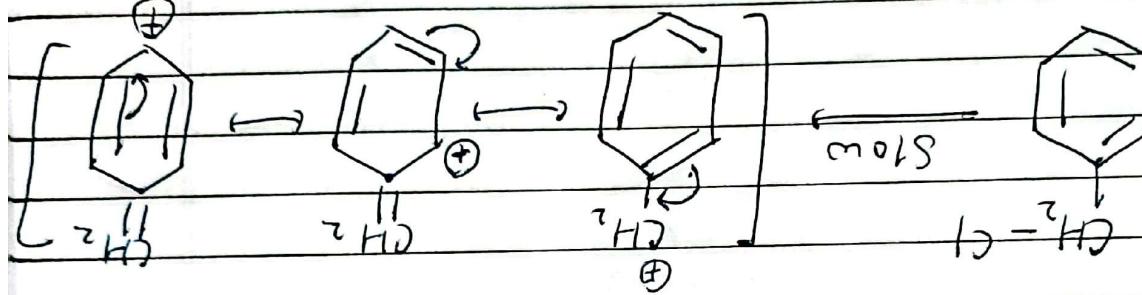
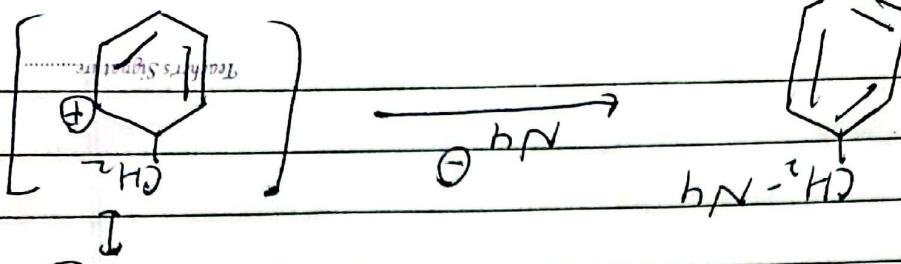
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~~1,2 N12 10 U5 1~~



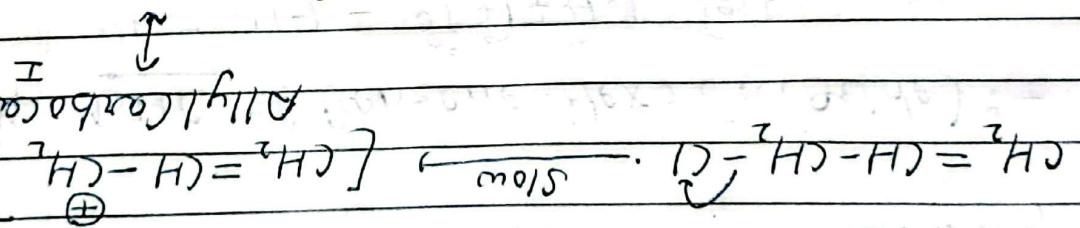
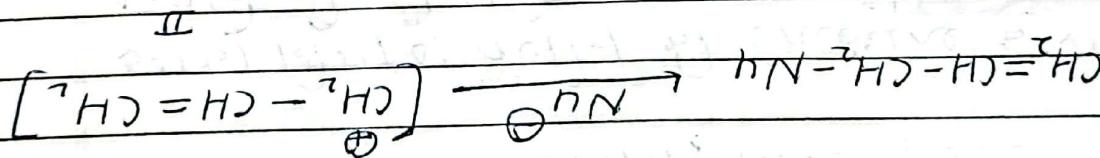
- Shows 14 π for 6 C atoms.





1. $\text{CH}_2=\text{CH}-\text{CH}_2-\text{NH}_2 \xrightarrow{\text{NH}} \text{CH}_2=\text{CH}-\text{CH}_2-\text{NH}^+$
 2. $\text{CH}_2=\text{CH}-\text{CH}_2-\text{NH}^+ + \text{CH}_2=\text{CH}-\text{Cl} \xrightarrow{\text{Slow}} \text{CH}_2=\text{CH}-\text{CH}_2-\text{NH}-\text{CH}_2-\text{Cl}$

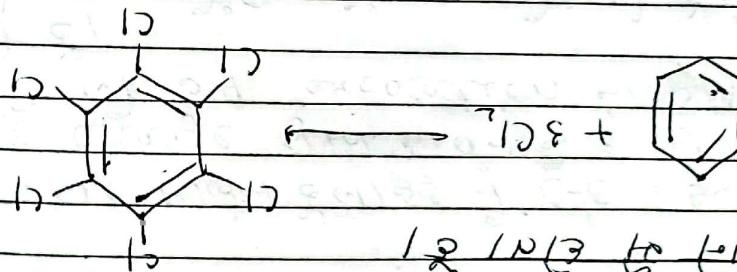
+ive Charge \Rightarrow N^+ + H_2 \Rightarrow NH_2 \Rightarrow NH_2 + H_2 \Rightarrow NH_2
 $\text{C-C Double Bond} \Rightarrow \text{C-C Double Bond}$



1. $\text{CH}_2=\text{CH}-\text{NH}_2 \xrightarrow{\text{NH}_2^-} \text{CH}_2=\text{CH}-\text{NH}_2$
 2. $\text{CH}_2=\text{CH}-\text{NH}_2 + \text{CH}_2=\text{CH}-\text{Cl} \xrightarrow{\text{Slow}} \text{CH}_2=\text{CH}-\text{NH}-\text{CH}_2-\text{Cl}$

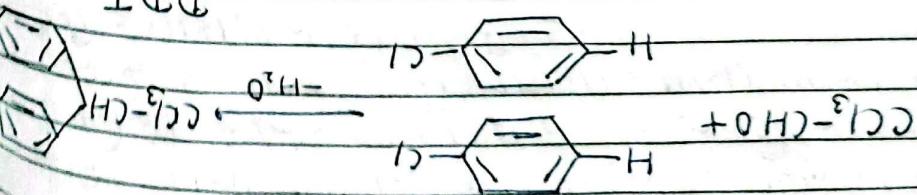
1. $\text{CH}_2=\text{CH}-\text{NH}_2 \xrightarrow{\text{NH}_2^-} \text{CH}_2=\text{CH}-\text{NH}_2$
 2. $\text{CH}_2=\text{CH}-\text{NH}_2 + \text{CH}_2=\text{CH}-\text{Cl} \xrightarrow{\text{Slow}} \text{CH}_2=\text{CH}-\text{NH}-\text{CH}_2-\text{Cl}$

BHC



BHC (Benzene Hexa Chloride) :-

DDT



DDT (DDT) :-