

# PHY64 Experiment 2: The Millikan Oil-drop Experiment

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## I. INTRODUCTION

In 1910, American physicist Robert Millikan performed the first experiment to determine the elusive value of the electron charge. Using a microscope, Millikan observed small oil droplets as they fell through two horizontal metal electrodes. He first calculated their terminal velocity with no voltage applied to determine the radius of the droplets. Then, using the known density of the oil, Millikan was able to calculate their mass and therefore the gravitational force on each droplet. Next Millikan applied a voltage between the plates to induce an electric field. By observing the motion of a the droplets in several charge states, Millikan showed that the charges were all small integer multiples of a certain base value. Millikan calculated this value to be  $1.5924 \times 10^{-19}$  C, within 1% of the currently accepted value of  $e$ .

In this experiment, we measure the electron charge using the same method that Millikan used in 1910. The main device used is a cylindrical chamber manufactured by Pasco Scientific. The bottom half of the device consists of a capacitor with plate separation set by a 7.6 mm thick spacer that serves as a droplet viewing chamber. A small

hole at the top of this chamber allows for oil droplets to enter the viewing region, where they are illuminated by an LED and viewed through a microscope. Using an atomizer, oil droplets are deposited into the upper chamber. In the process, the droplets acquire an electric charge. The droplets diffuse from the upper chamber through the small hole at the top of the lower chamber and enter the viewing region. By applying a voltage and observing the velocities of the oil droplets, we may deduce the concentration of charge on each droplet. Finally, we may calculate the difference between these charge states and conclude that they are integer multiples of some constant value. By calculating this value, we will obtain the charge of a single electron.

## II. THEORY

## III. RESULTS

## IV. ERROR

## V. CONCLUSION