



Successive reduction-oxidation activity of FeO_x/TiO₂ for dehydrogenation of ethane and subsequent CO₂ activation

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ABSTRACT

Prototype FeO_x/TiO₂ was applied for dehydrogenation of ethane to ethylene and successive CO₂ activation to CO for further chemical looping (CL) application. The Fe(5)/TiO₂ exhibited a facile redox cyclic activity with an insignificant coke formation and comparable CO₂ activation through reduction-oxidation reaction cycles. During the oxidative dehydrogenation of chemical looping (CL-ODH), the Fe nanoparticles (Fe₂O₃) on the rutile TiO₂ were partially transformed to the thermally stable FeTiO₃ phases with its lower oxidation state (Fe²⁺) below 700 °C. However, the formations of Fe₂TiO₅ phases were found to be less active than the smaller sizes of iron oxides (Fe³⁺). The thermally stable iron phases with their reversible redox natures between FeTiO₃ and Fe₂O₃ phases were responsible for a stable reduction-oxidation activity. The robust preservation of the partially reduced surface Fe²⁺ sites with the copresence of the larger Fe₂O₃ hematite crystallites on the TiO₂ was responsible for improving the redox activity and stability.

1. Introduction

A fast increase of anthropogenic CO₂ in an atmosphere has been mainly considered as a major chemical component for a recent global warming problem [1–3], and it largely encourages to develop effective utilization methods of CO₂ to value-added chemicals. Among the proposed chemical processes, one potential way to utilize CO₂ seems to be a chemical looping (CL) process, which is based on redox cycles of oxygen carriers such as partially reducible transition metal oxides [4–7]. The CL process for typical combustion reaction is generally composed of two interconnected reactors such as fuel reactor (reduction) and air reactor (oxidation) [8–10]. The oxygen carriers can release the active oxygen atoms from crystalline lattices in the fuel reactor through a reduction of metal oxides, and the partially reduced oxides can be regenerated by the oxidation reaction with air or CO₂ to simultaneously generate CO in an air reactor [11,12]. In terms of the CL scheme, the combination of CL with an oxidative dehydrogenation of light paraffins (CL-ODH) has a lot of potentials for a direct CO₂ capture/transformation as well as simultaneous light olefin production. For example, ethane (C₂H₆) dehydrogenation to ethylene (C₂H₄) can be carried out at a fuel reactor, whereas an air regenerator of the partially reduced metal oxides can be operated under CO₂ reduction to CO. One of the most important building blocks in petrochemical industries is ethylene, which is a starting material for the production of various

petrochemical intermediates [13,14]. Generally, the ethylene feedstock has been produced through conventional steam cracking and/or fluidized catalytic cracking reaction [15–17]. However, the pyrolysis processes have many drawbacks such as highly endothermic reaction natures requiring higher operation temperatures up to 1100 °C with substantial energy consumption by producing 1.0–1.2 ton(CO₂)/ton (C₂H₄) with a large amount of coke depositions [18–20]. Although direct oxidative dehydrogenation (ODH) of paraffins have many potential advantages for olefin production, there are a number of unsolved critical challenges for its commercial implementation such as the requirement of air separation unit (ASU) and safety issues due to the copresence of oxygen and ethane with a comparable separation cost of the final gaseous products [21,22]. The proposed CL-ODH processes in the present study can be described with the following two main reaction steps;

- (1) Reduction step of metal oxides (fuel reactor): C₂H₆ + M_xO_y → C₂H₄ + M_xO_{y-1} + H₂O
- (2) Oxidation step of metal oxides (air reactor): CO₂ + M_xO_{y-1} → CO + M_xO_y

Compared to the general co-feeding of ethane and oxygen for ODH reaction, the CL-ODH by using the active lattice oxygens of the heterogeneous catalysts has been less studied. Several vanadium-based

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