

22 2 review and reinforcement the reaction process answers chemistry

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22.2 Review and Reinforcement: The Reaction Process

1. What is the rate law for a reaction?

The rate law is an equation that expresses the relationship between the rate of a reaction and the concentrations of the reactants. It is typically written in the form $\text{rate} = k[A]^m[B]^n$, where k is the rate constant, $[A]$ and $[B]$ are the concentrations of the reactants A and B, and m and n are the orders of the reaction with respect to A and B, respectively.

2. How can the order of a reaction be determined?

The order of a reaction can be determined experimentally by measuring the rate of the reaction at different concentrations of the reactants. By plotting the rate of the reaction versus the concentration of each reactant, the order of the reaction with respect to that reactant can be determined from the slope of the line.

3. What is the half-life of a reaction?

The half-life of a reaction is the time it takes for the concentration of a reactant to decrease to half its initial value. It is typically expressed in seconds, minutes, or hours. The half-life of a reaction can be used to determine the rate of the reaction.

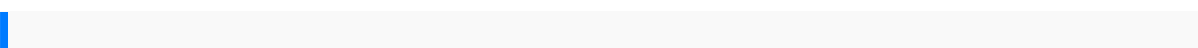
4. How can the rate of a reaction be increased?

The rate of a reaction can be increased by increasing the concentration of the reactants, increasing the temperature of the reaction, or adding a catalyst. A catalyst

is a substance that speeds up the rate of a reaction without being consumed in the reaction.

5. How can the rate of a reaction be decreased?

The rate of a reaction can be decreased by decreasing the concentration of the reactants, decreasing the temperature of the reaction, or adding an inhibitor. An inhibitor is a substance that slows down the rate of a reaction without being consumed in the reaction.



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