

Acid base indicators

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Acid-Base Indicators: A Guide**

What are the 5 acid-base indicators?

1. Phenolphthalein
2. Methyl orange
3. Bromthymol blue
4. Litmus
5. Alizarin yellow

What are the indicators used in acid-base titration?

- Phenolphthalein
- Methyl orange
- Thymol blue

What are the indicators for strong acid and strong base?

- Methyl orange
- Phenolphthalein

How to choose an acid-base indicator?

- Consider the equivalence point pH of the titration
- Select an indicator with a color change within 1 pH unit of the equivalence point

What are the 3 types of pH indicators?

1. Acid-base indicators
2. Redox indicators
3. Metal-ion indicators

Why does phenolphthalein turn pink?

Phenolphthalein is colorless in acidic solutions and turns pink in basic solutions because it undergoes a structural change that results in the formation of a colored product.

What is the most common acid-base indicator?

Phenolphthalein

What is a natural indicator of acid and base?

Litmus

What is an example of an indicator in acid-base titration?

Phenolphthalein is used in the titration of a strong acid with a strong base.

What are the 7 strong acids and bases?

Strong acids:

- Hydrochloric acid (HCl)
- Sulfuric acid (H₂SO₄)
- Nitric acid (HNO₃)
- Hydrofluoric acid (HF)
- Perchloric acid (HClO₄)
- Hydrobromic acid (HBr)
- Hydroiodic acid (HI)

Strong bases:

- Sodium hydroxide (NaOH)

- Potassium hydroxide (KOH)
- Calcium hydroxide (Ca(OH)₂)
- Barium hydroxide (Ba(OH)₂)
- Strontium hydroxide (Sr(OH)₂)
- Lithium hydroxide (LiOH)
- Rubidium hydroxide (RbOH)

What are the three acid-base indicators?

- Phenolphthalein
- Methyl red
- Litmus

Why is phenolphthalein used as an indicator?

- It provides a clear and sharp endpoint with a color change from colorless to pink.
- It is suitable for strong acid-strong base titrations with equivalence points around pH 8.3.

What are the indicators of acid-base titration?

- Phenolphthalein
- Methyl orange
- Bromothymol blue

Which indicator is best for a base?

Phenolphthalein

Why is phenolphthalein a good indicator for NaOH and HCl?

Because the equivalence point of NaOH and HCl titration is around pH 8.3, which is within the color change range of phenolphthalein.

What is the indicator for strong acid and strong base?

Phenolphthalein

How to choose an indicator?

Select an indicator whose color change range is within 1 pH unit of the equivalence point.

Why do acid-base indicators change color?

Acid-base indicators change color because they undergo structural changes in response to changes in pH.

Is NaOH an acid or base?

NaOH is a strong base.

Is bleach an acid or base?

Bleach is a base.

Is Na₂CO₃ an acid or base?

Na₂CO₃ is a weak base.

Why does phenolphthalein turn pink then clear?

Phenolphthalein turns pink in basic solutions but turns colorless in strongly basic solutions because it undergoes a second structural change at higher pH values.

What is a substitute for phenolphthalein?

Thymolphthalein

Is vinegar an acid or base?

Vinegar is an acidic substance.

Is turmeric an acid-base indicator?

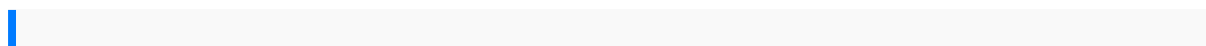
Yes, turmeric can be used as an acid-base indicator. It turns yellow in acidic solutions and brown in basic solutions.

Is onion base or acid?

Onion is an acid.

What is an acid-base indicator at home?

Red cabbage juice can be used as an acid-base indicator. It turns red in acidic solutions, purple in neutral solutions, and green in basic solutions.



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