

26 saturated and unsaturated solution pogil key

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Saturated and Unsaturated Solutions**

What is a Saturated Solution?

A saturated solution is one in which no more solute (the substance being dissolved) can be dissolved at a given temperature. In other words, the solvent (the liquid doing the dissolving) has reached its maximum capacity for holding the solute.

What is an Unsaturated Solution?

In contrast to a saturated solution, an unsaturated solution contains less solute than its solvent can hold at a given temperature. There is still room for more solute to dissolve.

How to Tell if a Solution is Saturated or Unsaturated

The features of a graph that can help identify saturated solutions include:

- **Constant solute concentration:** The graph will show a horizontal line once the solution becomes saturated.
- **Excess solute:** Any additional solute added to the solution will not dissolve and will appear as a precipitate or solid particles.

Determining if a Liquid is Saturated or Unsaturated

The liquid that drips from a filter is typically unsaturated because the dissolved particles are too small to be trapped by the filter.

Key Characteristics of a Saturated Solution

- Maximum solute concentration for a given temperature
- No more solute can dissolve
- Excess solute forms a precipitate
- Equilibrium is reached between dissolved and undissolved solute

Dissolved Particles in Saturated Solutions

When a small amount of additional solute is added to a saturated solution, the number of dissolved particles does not increase. Instead, the excess solute precipitates out of the solution.

Pure Substances vs. Solutions

A pure substance is a substance that has a fixed composition and consists of only one type of particle. A solution is a mixture of two or more pure substances.

Concentrated vs. Dilute Solutions

Two ways to differentiate a concentrated solution from a dilute solution are:

- **Amount of solute:** A concentrated solution has a higher amount of solute dissolved in it.
- **Appearance:** A concentrated solution may appear cloudy or opaque due to the high concentration of solute particles.

Saturated vs. Unsaturated Simple

- **Saturated:** No more solute can be dissolved.
- **Unsaturated:** There is still room for more solute to dissolve.

Examples of Saturated and Unsaturated Solutions

- **Saturated:** Saltwater with all the salt it can hold
- **Unsaturated:** Sugar water with sugar that can still be dissolved

Solubility and Saturated Solutions

The solubility of a substance depends on the temperature, pressure, and the nature of the substance and solvent.

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