

# Atoms atomic structure questions and answers

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**How to answer atomic structure questions?**

**What questions do you have about atoms?**

**What is the structure of an atom Class 9 very short answer?** Solution : According to J.J. Thomson's model of an atom, an atom consists of a positively charged sphere with electrons embedded in it. However, it was later found that the positively charged particles reside at the centre of the atom called the nucleus, and the electrons revolve around the nucleus.

**What are the fundamental particles of an atom Mcq?** Atoms are made of three smallest particles- electrons, protons and neutrons. The protons and neutrons remain in the nucleus of the atom and electron revolve around the nucleus. The electrons are negatively charged, protons are positively charged and neutrons are neutral.

**Does an atom have a color?** atoms (as opposed to molecules) do not have colors - they are clear except under special conditions.. you could not see the color of one atom or molecule - not because it is too small - but because the color of one atom would be too faint.

**What are the 3 rules of atomic structure?** That is, we follow the three important rules: Aufbau Principle, Pauli-exclusion Principle, and Hund's Rule. The electronic configuration of cations is assigned by removing electrons first in the outermost p orbital, followed by the s orbital and finally the d orbitals (if any more electrons need to be removed).

**What are atoms mostly made of?** Protons & electrons make up most of the mass of an atom.

**How are atoms structured?** Atoms consist of an extremely small, positively charged nucleus surrounded by a cloud of negatively charged electrons. Although typically the nucleus is less than one ten-thousandth the size of the atom, the nucleus contains more than 99.9% of the mass of the atom.

**Are atoms made of particles?** Atoms are constructed of two types of elementary particles: electrons and quarks. Electrons occupy a space that surrounds an atom's nucleus. Each electron has an electrical charge of -1. Quarks make up protons and neutrons, which, in turn, make up an atom's nucleus.

**How are neutrons formed?** A neutron is formed by an electron and a proton combining together. Therefore, it is neutral. Canal rays are positively charged radiations which led to the discovery of protons. Electrons are negatively charged sub-atomic particles having negligible mass.

**How to find valence electrons?** For neutral atoms, the number of valence electrons is equal to the atom's main group number. The main group number for an element can be found from its column on the periodic table. For example, carbon is in group 4 and has 4 valence electrons. Oxygen is in group 6 and has 6 valence electrons.

**How to calculate valency?** The electrons in the outermost shell of an atom are called valence electrons. For elements having one to four valence electrons, valency = valence electrons. For elements having five to seven valence electrons, valency = (8-valence electrons).

**What are the two main parts of an atom?** The two parts of an atom are the nucleus and the electron cloud. The nucleus is composed of protons and neutrons, densely packed in the nucleus and held together by nuclear forces. The electron cloud is an area around the nucleus where the electrons orbit the nucleus in different shells.

**Are all the atoms the same?** Yes, all things are made of atoms, and all atoms are made of the same three basic particles - protons, neutrons, and electrons. But, all

atoms are not the same. You know that the number of protons in an atom determines what element you have. For instance hydrogen has one proton, carbon has six.

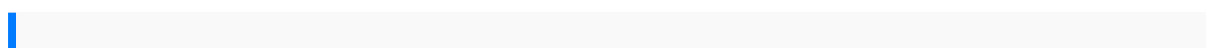
**Is an atom a chemical?** An atom is the basic building block of chemistry. It is the smallest unit into which matter can be divided without the release of electrically charged particles. It also is the smallest unit of matter that has the characteristic properties of a chemical element.

**How do you solve for atomic structure?**

**Is atomic structure a difficult chapter?** 5 Easiest Chapters in CBSE Class 11 Chemistry Structure of Atom: Understanding the atomic structure, isotopes, and electronic configuration of elements is comparatively easier.

**What do  $^{72}\text{Zn}$ ,  $^{75}\text{As}$ , and  $^{74}\text{Ge}$  have in common?** Hence, all three given atoms have the different number of electrons, protons and mass numbers. Hence, all three atoms have the same number of neutrons.

**How do you answer the atomic number?** The atomic number of an atom is equal to the number of protons in the nucleus of an atom or the number of electrons in an electrically neutral atom. For example, in a sodium atom, there are 11 electrons and 11 protons. Thus the atomic number of Na atom = number of electrons = number of protons = 11.



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