

5 3 review and reinforcement chemistry answers

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5.3 Review and Reinforcement Chemistry Answers

Question 1: What is covalent bonding?

Answer: Covalent bonding occurs when atoms share electrons to form a stable molecule. The shared electrons are attracted to the nuclei of both atoms, creating a strong bond.

Question 2: Describe the different types of covalent bonds.

Answer: There are two main types of covalent bonds: single bonds, where two atoms share a pair of electrons, and double bonds, where two atoms share two pairs of electrons.

Question 3: How is electronegativity related to covalent bonding?

Answer: Electronegativity is the ability of an atom to attract electrons. The more electronegative an atom, the more strongly it will attract the shared electrons in a covalent bond.

Question 4: What is the difference between a polar and nonpolar covalent bond?

Answer: A polar covalent bond is one in which the electrons are not evenly shared between the two atoms. A nonpolar covalent bond is one in which the electrons are evenly shared.

Question 5: Give an example of a molecule with a polar covalent bond.

Answer: Water is a molecule with a polar covalent bond. The oxygen atom is more electronegative than the hydrogen atoms, so the electrons in the covalent bonds are pulled towards the oxygen atom. This creates a partial negative charge on the oxygen atom and a partial positive charge on the hydrogen atoms.

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