

Acid base or salt physical science if8767

answers

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What is Acid, Base, and Salt in Science?

Acid: A substance that releases hydrogen ions (H^+) in aqueous solutions, has a pH less than 7, and has a sour taste.

Base: A substance that releases hydroxide ions (OH^-) in aqueous solutions, has a pH greater than 7, and tastes bitter and slippery.

Salt: An ionic compound formed by the reaction of an acid and a base, which is usually neutral ($pH = 7$) and has a salty taste.

Is $FeBr_3$ an Acid, Base, or Salt?

$FeBr_3$ is a salt.

Which is More Chemically Active, Li or Na?

Na is more chemically active than Li.

20 Examples of Bases

1. NaOH (Sodium hydroxide)
2. KOH (Potassium hydroxide)
3. $Ca(OH)_2$ (Calcium hydroxide)
4. $Mg(OH)_2$ (Magnesium hydroxide)

5. NH_3 (Ammonia)
6. CH_3NH_2 (Methylamine)
7. $(\text{CH}_3)_2\text{NH}$ (Dimethylamine)
8. $(\text{CH}_3)_3\text{N}$ (Trimethylamine)
9. NaHCO_3 (Sodium bicarbonate)
10. K_2CO_3 (Potassium carbonate)
11. Na_2CO_3 (Sodium carbonate)
12. CaCO_3 (Calcium carbonate)
13. MgCO_3 (Magnesium carbonate)
14. $\text{Fe}(\text{OH})_2$ (Iron(II) hydroxide)
15. $\text{Fe}(\text{OH})_3$ (Iron(III) hydroxide)
16. $\text{Al}(\text{OH})_3$ (Aluminum hydroxide)
17. $\text{Zn}(\text{OH})_2$ (Zinc hydroxide)
18. $\text{Cu}(\text{OH})_2$ (Copper(II) hydroxide)
19. AgOH (Silver hydroxide)
20. AuOH (Gold hydroxide)

10 Examples of Acids

1. HCl (Hydrochloric acid)
2. H_2SO_4 (Sulfuric acid)
3. HNO_3 (Nitric acid)
4. CH_3COOH (Acetic acid)
5. HCOOH (Formic acid)
6. $\text{C}_6\text{H}_8\text{O}_7$ (Citric acid)
7. H_2CO_3 (Carbonic acid)
8. H_3PO_4 (Phosphoric acid)
9. H_2SO_3 (Sulfurous acid)
10. HNO_2 (Nitrous acid)

Is Fe_2O_3 an Acid, Base, or Salt?

Fe_2O_3 is an acidic salt.

Is FeCl_3 Acid or Salt?

FeCl₃ is a salt.

Is FeSO₄ an Acid, Base, or Salt?

FeSO₄ is a salt.

Is Sodium More Reactive than Li?

Yes, sodium is more reactive than lithium.

Is Li or Na More Hydrated with Water?

Lithium is more hydrated with water than sodium.

Is Potassium or Sodium More Chemically Active?

Potassium is more chemically active than sodium.

Are Organic Acids Strong or Weak?

Organic acids are generally weak acids.

Why are Strong Acids Also Strong Electrolytes?

Strong acids completely dissociate in water, releasing all their hydrogen ions and making them highly conductive.

Is Pure Water Acidic, Basic, or Neither?

Pure water is neutral (pH = 7).

What is the Difference Between Strong and Weak Bases?

Strong bases completely dissociate in water, releasing all their hydroxide ions, while weak bases only partially dissociate.

What are the 7 Strong and Weak Acids?

Strong Acids:

1. HCl
2. H₂SO₄
3. HNO₃
4. HClO₄
5. HBr
6. HI
7. CH₃COOH (strong only in concentrated solutions)

Weak Acids:

1. H₂CO₃
2. HF
3. HNO₂
4. H₃PO₄
5. CH₃COOH (weak in dilute solutions)
6. H₂SO₃
7. HCN

What are Six Strong Bases?

1. NaOH
2. KOH
3. Ca(OH)₂
4. Sr(OH)₂
5. Ba(OH)₂
6. LiOH

Is Na₂O an Acidic Salt?

No, Na_2O is a basic salt.

Is FeCl_3 an Acid, Base, or Salt?

FeCl_3 is a salt.

Is O_2 an Acid, Base, or Salt?

O_2 is neither an acid nor a base.

Is K_2SO_4 a Base or Acid?

K_2SO_4 is a salt.

Is KCl a Base or Acid?

KCl is a salt.

Is Na_2CO_3 an Acid or Base?

Na_2CO_3 is a base.

What is an Acid and Base in Science?

See the definitions provided earlier.

What is Acidic and Basic Salt?

Acidic salt: A salt formed by the partial neutralization of a strong base and a weak acid. **Basic salt:** A salt formed by the partial neutralization of a weak base and a strong acid.

What is Salt in a Short Answer?

Salt is an ionic compound formed by the combination of a positively charged ion (cation) and a negatively charged ion (anion).

What is the Definition of Acid Salt?

An acid salt is a partially neutralized salt that exhibits acidic behavior due to the presence of a weak base and a strong acid.

Is Water an Acid or Base?

Pure water is neutral (neither acidic nor basic).

Is Milk an Acid or Base?

Milk is slightly acidic, with a pH around 6.5.

Do You Put Acid in Water or Water in Acid?

Always add acid to water, not vice versa, to avoid splashing and dangerous reactions.

What is Acid Base and Salt Answer?

See the definitions provided earlier.

What are 10 Examples of Basic Salts?

1. Na_2CO_3
2. K_2CO_3
3. CaCO_3
4. MgCO_3
5. NaHCO_3
6. KHCO_3
7. $\text{Ca}(\text{OH})_2$
8. $\text{Mg}(\text{OH})_2$

9. $\text{Fe}(\text{OH})_2$

10. $\text{Al}(\text{OH})_3$

How is Basic Salt Formed?

Basic salts are formed by the reaction of a weak acid with a strong base.

Is Salt a Mixture?

No, salt is not a mixture. It is a pure substance composed of positive and negative ions.

Is NaOH a Basic Salt?

No, NaOH is a strong base, not a salt.

How to Prepare Salts in Chemistry?

Salts can be prepared by various methods, including:

1. Acid-base reactions
2. Precipitation reactions
3. Oxidation-reduction reactions
4. Neutralization reactions
5. Electrolysis

What is the Definition of a Salt?

See the definition provided earlier.

What is an Acid Base?

See the definitions provided earlier.

Is Salt an Acid or Base?

Salt is neither an acid nor a base. It is neutral.

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