

NIGERIAN CIVIL SERVICE RULES AND REGULATIONS

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What are the Nigerian public service rules? "Public Service Rules" consist of policies and guidelines which define the employment relationship between Public Servants and the Government; they also spell out the Conditions of Service and Human Resources Management procedures.

What is the structure of the civil service in Nigeria? STRUCTURE OF THE CIVIL SERVICE The civil service of Nigeria is structured in line with the British civil service in to five main classes. 1Administrative class 2 professional class 3The executive class 4 the clerical class 5 Messengerial, manipulative in or Auxiliary.

What is the difference between civil service and public service in Nigeria? The Public Service is a broader concept than the Civil Service. The Public Service includes the Civil Service, Armed Forces, Parastatals, etc. Not every Public Servant is a Civil Servant, but every Civil Servant is a Public Servant. The Public Service is established by section 169 of the 1999 constitution.

What are the rules for civil service promotion in Nigeria? Conditions for promotion are: Candidate must possess APER for 2, 3 or 4 years. Officers on pending disciplinary case cannot be promoted. Officers must appear for interview and score at least 60%.

What are the 4 laws in Nigeria? There are four distinct legal systems in Nigeria, which include English law, Common law, Customary law, and Sharia Law. English law in Nigeria is derived from the colonial Nigeria, while common law is a development from its post-colonial independence.

What are the 7 principles of public service? First report, 1995. The Committee's first report in 1995 established an initial version of The Seven Principles of Public Life, also known as the Nolan Principles after the committee's first chairman. The principles were Selflessness, Integrity, Objectivity, Accountability, Openness, Honesty and Leadership.

How much is civil service salary in Nigeria? New Salary Structure For Civil Servants in Nigeria Grade level 1 step 1 = 70,000 Grade level 4 step 1= 110,000 Grade level 5 steps 1= 125,000 Grade level 6 step 1= 135,000 Grade level 7 step 1= 145,000 Grade level 8 step 1= 235,000 Grade level 9 step 1= 245,000 Grade level 10 step 1= 250,000 Grade level 12 step 1= ...

How many years do civil servant work in Nigeria? Also, the Nigerian Senate recently passed a Bill to increase the retirement age for civil servants working in the National Assembly to 65 years or 40 years of service.

What is the role of Nigeria civil service? Advising the political class: civil servants advise the political leaders such as ministers, commissioners, and lawmakers on policy issues of national and local interest. c. Implementation of government policies: the civil service is largely responsible for the implementation of policies formulated by the government.

What is the main criticism against the Nigerian civil service? (b) The main criticisms against the Nigerian civil service: (i) Poor conditions of service: The poor condition of service in the civil service is responsible for non-attraction of qualified personnel into the service. Incentives for hard work are virtually absent.

Can civil servants do business in Nigeria? "The only profession/business that is allowed by the Constitution which Civil Servants can venture into, even while in service is the agricultural business. So, we thought that is the best way to go because agricultural business has a very wide value chain" she added.

What are the four importance of public service in Nigeria? It implements government policies. It advises the government. It creates employment opportunities. It performs delegated /legislative functions.

Can a civil servant have a second job in Nigeria? Put in another way, a public officer is not allowed to take more than one full-time office or employment for which he is paid a salary. 3 Section 19 of the Fifth Schedule Part 1 of the 1999 Constitution of the Federal Republic of Nigeria, CAP. C23, LFN 2004. 6.

How is seniority determined in civil service in Nigeria? The first to have his or her name recorded automatically becomes Senior to the other. This is horizontal Seniority. From day one, the future and career progression of these two officers are firmly determined.

What is civil procedure rules in Nigeria? Outline of Civil Procedure It consists mainly of rules of practice and procedure applying to conflicts involving disputes in which legal rights and legal duties are in issue. In a broader sense, it refers to the entire mechanism used to enforced private rights.

What are the basic obligations of public service? Public service is a public trust, requiring you to place loyalty to the Constitution, the laws, and ethical principles above private gain. You shall not hold financial interests that conflict with the conscientious performance of duty.

What are the four importance of public service in Nigeria? It implements government policies. It advises the government. It creates employment opportunities. It performs delegated /legislative functions.

What are the types of public policies in Nigeria? The four main types of public policy include regulatory policy, constituent policy, distributive policy, and redistributive policy. These four policy types differ in terms of what their goals are, and who they impact or benefit.

What is allowance in public service in Nigeria? Allowances are monetary benefits other than salary granted to an officer for specific purpose. All allowances provided for in public service rule are subject to periodic review by the National Salaries, Income and Wages Commission through the issuance of appropriate Circulars.

What is chemical bonding answers? A chemical bond is a force of attraction between atoms or ions. Bonds form when atoms share or transfer valence electrons. Valence electrons are the electrons in the outer energy level of an atom that may be

involved in chemical interactions. Valence electrons are the basis of all chemical bonds.

What is a chemical bond chapter 6? a chemical bond in which two atoms share a pair of valence electrons. a neutral group of atoms that are joined together by one or more covalent bonds. a covalent bond in which electrons are not shared equally. a covalently bonded group of atoms that has a positive or negative charge and acts as a unit.

Is chemical bonding a hard chapter? Chemical Bonding and Molecular Structure: While this chapter can be challenging for some, others find it relatively easy due to its logical progression and relatable examples.

What is the simplest collection of atoms from which an ionic compound's formula can be established? A formula unit is the simplest collection of atoms from which an ionic compound's formula can be established.

What are the 7 types of chemical bonds? There are 3 main types of chemical bonding, and they are covalent, metallic, and ionic bonding. List and explain 7 types of chemical bonding? They are ionic, covalent, metallic, hydrogen, Van der Waals, polarized, and clathrate bonding.

What is the strongest bond in chemistry? In chemistry, a covalent bond is the strongest bond, In such bonding, each of two atoms shares electrons that bind them together. For example - water molecules are bonded together where both hydrogen atoms and oxygen atoms share electrons to form a covalent bond. Q.

What is 6 bond in chemistry? In a formal sextuple bond, there would be $P = 6$ different electron pairs; an effective sextuple bond would then have all six contributing almost entirely to bonding orbitals.

What is a chemical reaction Chapter 6? In a chemical reaction, old bonds are broken and new bonds formed; atoms in the reactants are rearranged to form one or more different substances Page 3 A chemical equation gives the chemical formulas of the reactants on the left of the arrow and the products on the right.

What elements can form 6 bonds? Sulphur can form six bonds due to the presence of vacant 3d-orbitals which get occupied on excitation and create six

orbitals with unpaired electrons that become available for bonding.

What is the toughest chapter in chemistry? Ans. The toughest chapter in Chemistry is Equilibrium as this chapter involves complex concepts like the equilibrium constant, Le Chatelier's principle, and factors affecting equilibrium, etc.

Which is the easiest chapter in chemistry? The easiest chapters in CBSE Class 12 Chemistry are Solid State, Solutions, Biomolecules, Polymers, and Chemistry in Everyday Life due to their straightforward concepts and factual content.

Which chemistry is hard? Organic chemistry is generally considered to be more difficult than inorganic chemistry. This is because organic chemistry deals with the properties and reactions of carbon-based compounds, which are much more complex than the inorganic compounds that are studied in inorganic chemistry.

What are the three steps of ionic bonding? Think of forming an ionic compound as three steps (this is a simplification, as with all models): removing an electron from the metal, adding an electron to the nonmetal, and allowing the metal cation and nonmetal anion to come together.

What pair of elements will form a covalent bond? Covalent bonds usually occur between nonmetals. For example, in water (H_2O) each hydrogen (H) and oxygen (O) share a pair of electrons to make a molecule of two hydrogen atoms single bonded to a single oxygen atom. In general, ionic bonds occur between elements that are far apart on the periodic table.

Which pair would form an ionic bond? An ionic bond can be formed after two or more atoms loss or gain electrons to form an ion. Ionic bonds occur between metals, losing electrons, and nonmetals, gaining electrons. Ions with opposite charges will attract one another creating an ionic bond.

What type of bond is H_2O ?

Why do atoms combine? Answer:– There are mainly two most important reasons behind the combination of an atom. Firstly an atom combines to attain stability. And the second reason behind the combination of an atom is to form a different compound by combining two different or more atoms.

Why do atoms transfer valence electrons? If atoms have similar electronegativities (the same affinity for electrons), covalent bonds are most likely to occur. Because both atoms have the same affinity for electrons and neither has a tendency to donate them, they share electrons in order to achieve octet configuration and become more stable.

What is the weakest bond? Therefore, the order from strongest to weakest bond is Ionic bond > Covalent bond > Hydrogen bond > Vander Waals interaction.

Which bond is unbreakable? Covalent Bond Properties The following are a few properties of a covalent bond: These are considered strong and unbreakable chemical bonds that bind the atoms in place. These will only pair the electrons and do not form new ones. After covalent bonds are formed, it is almost impossible to break them.

What is the most difficult bond to break? Intramolecular covalent bonding is the most difficult to break. These ties are perhaps the toughest to break and yet the most stable.

What is chemical bonding in simple terms? A chemical bond is an attraction between two or more atoms that form a chemical. Chemicals are materials that contain two or more atoms connected with chemical bonds. Chemical bonds form when atoms share or donate electrons, which creates electrostatic attractions.

Why is chemical bonding? The attractive force that holds atoms or ions together is known as a chemical bond. When compared to individual atoms, such attractive interactions result in a more stable state for the entire system. Chemical bonding is a fundamental idea in chemistry that helps to explain other concepts like molecules and reactions.

What is chemical bonding examples? Example: water (H₂O) is composed of two hydrogen atoms and one oxygen atom, and in its bond, each hydrogen atom shares an oxygen atom. Ionic: this occurs when metallic and non-metallic atoms bond and an electron charge is given from one to the other.

What is a chemical bond quizlet? A chemical bond is when two different atoms have mutual electrical attraction between the valence electrons and nuclei.

Standard Test Method for Calcium Carbonate Content of Soils

Question 1: What is the purpose of the standard test method for calcium carbonate content of soils?

Answer: The standard test method, typically designated as ASTM D4373 or similar, determines the percentage of calcium carbonate (CaCO_3) present in soils. CaCO_3 is a common soil component that affects various soil properties, including pH, nutrient availability, and water infiltration.

Question 2: How is the standard test method performed?

Answer: The test involves reacting a soil sample with a known excess of hydrochloric acid (HCl). The reaction releases carbon dioxide (CO_2) gas, which is then measured. The amount of CO_2 evolved is proportional to the CaCO_3 content of the soil.

Question 3: What are the applications of the test?

Answer: The test results are used to:

- Establish soil pH and lime requirements for agricultural purposes
- Understand soil mineralogy and weathering patterns
- Estimate carbon sequestration potential in soils
- Determine the suitability of soils for construction and other engineering applications

Question 4: What are the limitations of the test?

Answer: The test assumes that all carbonate present in the soil is CaCO_3 . However, other carbonates, such as dolomite, may interfere with the results. The test also requires a certain particle size range for accurate measurements.

Question 5: Is there any other method for determining soil calcium carbonate content?

Answer: Yes, there are alternative methods such as the Calcimeter method or the Scheibler method. The Calcimeter method uses a pressure gauge to measure the

CO₂ evolved, while the Scheibler method employs a gravimetric analysis to determine the calcium carbonate content.

Solucionario Transferencia de Calor y Masa de Cengel, 4ta Edición

El libro de texto "Transferencia de Calor y Masa" de Yunus A. Cengel es un texto fundamental para estudiantes de ingeniería mecánica y otros campos relacionados. La cuarta edición incluye numerosos ejemplos, problemas y estudios de caso resueltos para mejorar la comprensión del alumno.

Pregunta: Calcula la tasa de transferencia de calor por convección a través de una placa plana con un espesor de 2 cm y un área de superficie de 1 m². La temperatura de la superficie de la placa es de 100 °C y la temperatura del fluido es de 20 °C. El coeficiente de convección es de 50 W/m²K.

Respuesta:

$$\begin{aligned} Q &= hA(T_s - T_f) \\ Q &= 50 \text{ W/m}^2\text{K} * 1 \text{ m}^2 * (100 \text{ }^\circ\text{C} - 20 \text{ }^\circ\text{C}) \\ Q &= 4000 \text{ W} \end{aligned}$$

Pregunta: Un tubo esférico con un diámetro de 20 cm está lleno de agua a una temperatura de 80 °C. El tubo está rodeado por aire a una temperatura de 20 °C. La superficie exterior del tubo tiene un emisor de radiación con una emisividad de 0,9. Calcula la tasa de transferencia de calor por radiación desde el tubo.

Respuesta:

$$\begin{aligned} Q &= \epsilon A(T_s^4 - T_f^4) \\ Q &= 5,67 \times 10^{-8} \text{ W/m}^2\text{K}^4 * (3,14 * 0,1 \text{ m}^2) * 0,9 * (353 \text{ K}^4 - 293 \text{ K}^4) \\ Q &= 108,2 \text{ W} \end{aligned}$$

Pregunta: Una varilla cilíndrica de cobre con un diámetro de 1 cm y una longitud de 10 cm se expone a una temperatura ambiente de 20 °C. La superficie de la varilla se mantiene a una temperatura constante de 100 °C. Calcula la tasa de transferencia de calor por conducción a través de la varilla.

Respuesta:

$$Q = kA(T_s - T_f) / L$$

$$Q = 401 \text{ W/mK} * (3,14 * 0,005 \text{ m}^2) * (100 \text{ }^\circ\text{C} - 20 \text{ }^\circ\text{C}) / 0,1 \text{ m}$$

$$Q = 63,6 \text{ W}$$

Pregunta: Una pared plana de 5 cm de espesor está construida de hormigón con una conductividad térmica de 1,4 W/mK. La temperatura de la superficie interior de la pared es de 40 °C y la temperatura de la superficie exterior es de 10 °C. Calcula la tasa de transferencia de calor por conducción a través de la pared.

Respuesta:

$$Q = kA(T_s - T_f) / L$$

$$Q = 1,4 \text{ W/mK} * 1 \text{ m}^2 * (40 \text{ }^\circ\text{C} - 10 \text{ }^\circ\text{C}) / 0,05 \text{ m}$$

$$Q = 840 \text{ W}$$

Pregunta: Un radiador de aluminio tiene una superficie de 1,5 m² y una temperatura de 80 °C. El radiador está situado en una habitación a una temperatura de 20 °C. Calcula la tasa de transferencia de calor por convección natural desde el radiador.

Respuesta:

$$Q = hA(T_s - T_f)$$

$$Q = 10 \text{ W/m}^2\text{K} * 1,5 \text{ m}^2 * (80 \text{ }^\circ\text{C} - 20 \text{ }^\circ\text{C})$$

$$Q = 900 \text{ W}$$

[modern chemistry chapter 6 chemical bonding test answers, standard test method for calcium carbonate content of soils, solucionario transferencia de calor y masa cengel 4 edicion](#)

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