## 1 Fluorination

$$CH_4$$
 +  $F_2$   $\xrightarrow{h\nu}$   $C_{black}$  +  $4HF$ 

F is highly reactive, hence the reaction is exothermic.

## 2 Iodination

$$CH_4$$
 +  $I_2$   $\stackrel{h\nu}{\rightleftharpoons}$   $CH_3I$  +  $HI$ 

As the reaction is reversible, an oxidizing agent such as  $HIO_3, HNO_3,$  etc. is used.

## 3 Chlorination

$$CH_4$$
 +  $Cl_2$   $\xrightarrow{h\nu}$   $CH_3Cl$  +  $HCl$ 

- $\bullet$  Cl is more reactive, hence being less selective.

## 4 Bromination

$$CH_4$$
 +  $Br_2$   $\xrightarrow{h\nu}$   $CH_3Br$  +  $HBr$ 

- $\bullet$  Br being less reactive, hence more selective.
- Reactivity ratio  $1^{\circ}: 2^{\circ}: 3^{\circ}:: 1:80:1600$