

## Molecular Dynamics Lab

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### Molecular Dynamics (MD) Module Problems

During this exercise we will get hands on experience running various molecular dynamics simulations using the Large-scale Atomic/Molecular Massively Parallel Simulator (LAMMPS) Molecular Dynamics code. In doing so we will explore issues of equilibration, estimating error, and determining thermodynamic quantities such as melting temperature, latent heat, and thermal expansion. This lab will focus on **Aluminum (Al)** as described by an Embedded Atom Method (EAM) interatomic potential. The specific parts of the lab are:

<b>1</b>	<b><i>Equilibration</i></b>	<b>1</b>
1.1	<i>Equilibration Time</i>	1
1.2	<i>Error Estimation</i>	2
1.3	<i>Extrapolating Error Estimation</i>	3
<b>2</b>	<b><i>Heating and Cooling</i></b>	<b>3</b>
2.1	<i>Solidification/Melting</i>	4
2.2	<i>Latent Heat</i>	4
2.3	<i>Specific Heat</i>	4
2.4	<i>Thermal Expansion</i>	4
2.5	<i>Confirming the Clausius-Clapeyron Equation (More challenging)</i>	5

The first concept we will address is that of equilibration. Equilibration is the process of letting the systems reach thermodynamic equilibrium. Equilibration is needed at the start of a Molecular Dynamics simulation when initial atom positions and velocities must be input by the user and may not represent the final thermodynamic conditions. Equilibration is also needed when one changes parameters in a system during or between runs, e.g., taking a step in temperature. In general, it is usually necessary to let the system equilibrate from the start of a run before extracting any materials information related to equilibrium properties. During the first exercise we will explore how long this equilibration can take, and how to assess when a system has reached equilibrium.

## 1 Equilibration

Here we explore how long it takes to reach equilibrium conditions of volume, temperature, pressure, and total energy when we take a temperature step. We use solid fcc aluminum as the example.

### 1.1 Equilibration Time

Using the script file EquilRun.txt perform an **equilibration run** for **10000 steps** with a **step size of 0.001 picoseconds**. Set the **temperature to 700 K** and the **initial lattice parameter to 4.04525 Angstroms**. To set these parameters look for the “**Parameters for Exercise 1:**” section near the top of the file. After “equal” for each variable enter the number for the variable. For example: variable Allattice equal ###

To execute the run, use the following command from the terminal (note: for mac and linux you will omit the .exe and % portions)

```
% run_lammps
```

Do this directly with the \*.in file. See the README in the home directory on the cluster for more information and updates on this.

Before proceeding with the data analysis below I recommend you skip Problem 1.2. Setup and start the run there before proceeding with the analysis here as that run will take several minutes to run and you won't have to wait if you start it now.

Several output files will be generated from the run. The key one we're interested in right now is `npqfile.txt`. This is a space delimited file with columns for Time Step, Temperature (K), Pressure (bar), Volume ( $\text{\AA}^3$ ), and Total Energy (eV/atom). Use Excel's text to columns feature if you are unsure of how to get the data imported properly.

Using the output `npqfile.txt` file plot the system volume, temperature, pressure, and total energy against the timestep. Based on each of these parameters, estimate where the system appears to enter equilibrium. **Record the timestep and total simulation time for each property in Table 1. Also copy your four plots into the space below Table 1. Remember to include axis titles and units and make sure the plot shows the data points and not just a fit line.**

## 1.2 Error Estimation

Next, we'll look at how to extract system properties and precision estimates (error bars) from our MD equilibration run. Note that this is precision, not accurate, and a precise value can still be incorrect due to, e.g., errors in the potential.

Using the same input files and settings from **Problem 1.1** set the number of steps to run to one million instead of ten thousand. Note that this run could take several minutes so I recommend setting this up to run in a separate directory while you work on the plotting and analysis in Problem 1.1.

Using the file `npqfile_averages.txt` generated during the same MD run we have the same system properties from part 1.1 in 1000 step block averages over the simulation time. Starting after equilibration steps (let us all start from block 3 (i.e., ignore the first 2 block averages) to be consistent) calculate the Standard Error in the Mean for Volume per atom and Energy per atom as a function of how many blocks we include in the average. **Fill in these values in Table 2 for number of blocks included (5, 10, 100, 500, 900).** Remember that the equation for Standard Error in the Mean for uncorrelated data is:

$$\sigma_m = \frac{\sigma}{\sqrt{N}}$$

Where  $\sigma_m$  is the Standard Error in the Mean,  $\sigma$  is the standard deviation in the block averages, and  $N$  is the number of blocks used to generate the standard deviation.

**Also include a plot of the Standard Error in the Mean for both Volume per atom and Energy per atom for every block number (it should range 1-998). Remember to include axis titles and units and make sure the plot shows the data points and not just a fit line.**

Note that the energy output in the npteq files is already given in units eV/atom. Convert these to meV/atom to give more readable values. Also, the volume is given in cubic Angstroms ( $\text{\AA}^3$ ) per cell. Please convert it to  $\text{\AA}^3/\text{atom}$  by dividing by the number of atoms in the simulation (32). By representing values per atom, you can easily compare different system sizes.

NOTE: An easy way to set up the standard deviations for these various numbers of blocks in excel is to sort the data by TimeStep from largest to smallest, then generate the standard deviation for the full TimeStep range. Next extend that standard deviation formula down to the bottom of the dataset. This works because excel automatically ignores empty cells, so as the formula includes more and more empty cells toward the bottom of the dataset the standard deviation is being calculated based only on the first few blocks, ignoring everything below.

### 1.3 Extrapolating Error Estimation

Now we have an idea about how precise our values of Energy per atom and Volume per atom are in our system. However, what if we wanted to know how precise we could be after 100, 200, or maybe 1,000 blocks? For our current system it may be trivially easy to just increase the step count in the run and do the analysis again. However, for more complex systems and calculations that take more time this may not be practical. Let's take advantage of what we know about the functional form of the standard error in the mean and see how well we can extrapolate. Recall that the standard error in the mean takes the form:

$$\sigma_m = \frac{\sigma}{\sqrt{N}}$$

Taking advantage of logarithms, we could also write this equation as

$$\log(\sigma_m) = \log(\sigma) - 0.5 * \log(N)$$

Now we have a linear equation with respect to the number of blocks included in analysis.

Take the log of the Standard Error in the Mean for Volume and Energy that you calculated previously. **For block numbers 100-998 plot these values against each other and fit a line through the data. Fill in Table 3 with the slope values that you find for Volume and Energy. Also copy below your log plot(s). Please plot these two data sets in one figure. Remember to include axis titles and units and make sure the plot shows the data points and not just a fit line. Make sure you include at least 2 significant figures in the slope.**

## 2 Heating and Cooling

In this part of the lab, we will use MD to calculate properties for a series of temperatures for Al and show how we can extract a number of very useful thermodynamics properties.

## 2.1 Solidification/Melting

We will now explore solidification/melting of our aluminum solid. To do this we'll perform an MD run in which we ramp the temperature from 900K to 1400K in 25K intervals. Then we will ramp the temperature back down using the same process. At each temperature we will perform an equilibration run using the equilibration settings determined earlier in the lab (holding for 2,000-time steps), and then a production run to extract the system properties. Solidification/melting is a first-order phase transition which requires nucleation of the new phase. As this process is stochastic and takes time, the simulation will typically not actually solidify (melt) until below (above) the true melting temperature, and effect called hysteresis. In some cases, the system may not go through the phase transition at all during the run time of the simulation.

Open the script file `Ramp.txt` and set the initial and final temperature, and the temperature step in the same way you set the variables in Problem 1. Then Run the script file using the usual submit command.

Open the `prall.txt` file using excel to see the overall output of the run. You may also notice additional files generated for each temperature step. We won't look in detail at these, but they have all the detailed results for each temperature step. **Plot system volume per atom vs temperature and system total energy per atom vs temperature for the ramp up and ramp down and copy those plots below Table 4.** Again, remember to normalize your data per atom. Also split the ramp up and ramp down into different series on your plots. Use these plots to identify the melting temperature from the heating and cooling data and enter the results in **Table 4**.

## 2.2 Latent Heat

The latent heat is the enthalpy change in the system at a fixed temperature during a phase transition. For a liquid to solid transition this is called the "latent heat of fusion". Use the data determined in **Problem 2.1 to estimate** the latent heat of fusion under both the heating and cooling conditions and enter them in **Table 5**.

## 2.3 Specific Heat

The specific heat of a material captures the way internal energy changes with temperature and is defined as

$$c_P = \left( \frac{\partial U}{\partial T} \right)_{P,N}$$

A simple expression for the specific heat is the formula  $c_P = A + BT$ . Determine this expression from your solid phase heating data by fitting to a quadratic and differentiating and enter the values in **Table 6**.

## 2.4 Thermal Expansion

The thermal expansion coefficient of a material captures the way volume changes with temperature and is defined as

$$\alpha = \frac{1}{V} \left( \frac{\partial V}{\partial T} \right)_{P,N}$$

Using the volume vs. temperature data you obtained in Problem 2.1 calculate  $\alpha$  for the solid phase of the Al from the heating curve. To find  $\alpha$  fit  $V(T)$  to a linear function of  $T$ , differentiate, and then divide by your average  $V$  over the temperature range as  $V(T)$  changes a small amount. Enter the data in **Table 7**.

## 2.5 Confirming the Clausius-Clapeyron Equation (More challenging)

The Clausius-Clapeyron equation states

$$\left( \frac{\partial P}{\partial T} \right)_{CC} = \left( \frac{\Delta S}{\Delta V} \right)$$

The left side is the partial derivative of function  $P(T)$  along the coexistence curve of two phases, i.e., how much you need to change pressure for a small change in temperature to keep the two phases in equilibrium. The right-side is the change in entropy ( $S$ ) and volume ( $V$ ) between the two phases. All values are evaluated at the same  $P, T$  in the equation. Rearranging this equation gives

$$\Delta S = \Delta V \left( \frac{\partial P}{\partial T} \right)_{CC}$$

This  $\Delta S$  value is called the entropy of melting (or freezing, depending on the sign) when the coexistence is between the solid and liquid. To explore this approach to getting  $\Delta S(P_0, T_m)$  please calculate  $\Delta V(P_0, T_m)$  and  $\left( \frac{\partial P}{\partial T} \right)_{CC, (P_0, T_m)}$ , where  $(P_0, T_m)$  denotes that they are all to be evaluated at standard pressure  $P_0 = 1 \text{ atm}$  and  $T_m$  = the predicted melting temperature of Al. To calculate  $\left( \frac{\partial P}{\partial T} \right)_{CC}$  please determine  $T_m(P)$  for at least 6 values of  $P$  from 0 to 100 GPa (1 GPa =  $10^9$  Pa) and fit it and differentiate. Note that you will need to extend the range of temperatures, e.g. to 2000K. Focus only on melting from heating as kinetics of freezing is slow and stochastic and has a lot of hysteresis. You will observe clearer trends from a larger system size, e.g., 500 atoms, so consider trying that (although not required). Then compare this value of  $\Delta S$  to the value found by calculating  $\Delta S$  from the condition of thermal equilibrium between the two phases on the coexistence curve, which requires that  $\Delta S(P_0, T_m) = \frac{\Delta H(P_0, T_m)}{T_m}$ . Please enter the relevant data in Table 8.

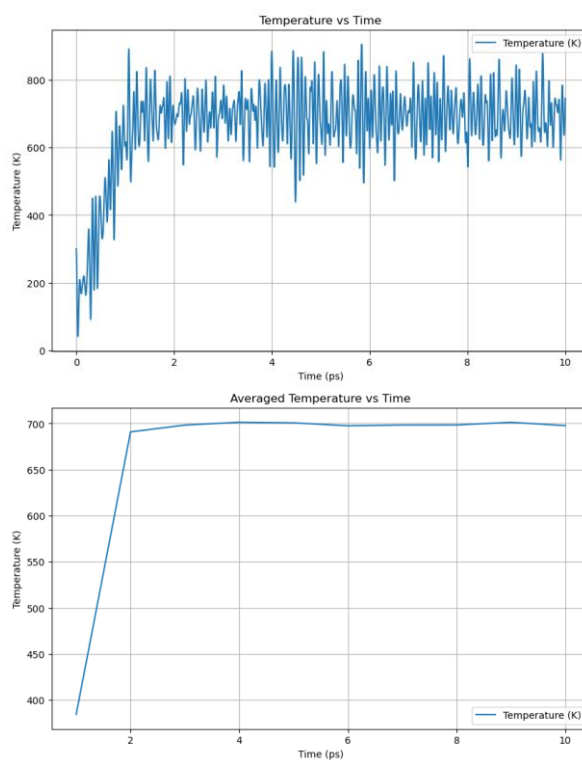
### Problem 1

#### Problem 1.1

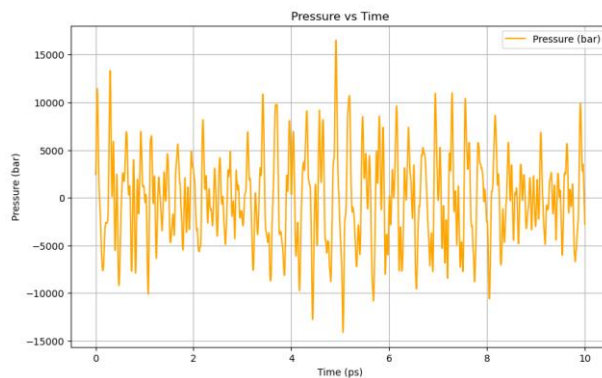
**Table 1:** Estimate the time steps and total simulation time to equilibrate the following system properties for Aluminum. Fill in after running the EquilRun.txt script using LAMMPS to complete Problem 1.1. Include below your plots of Temperature, Pressure, Volume, and Total Energy vs Time Step. On the volume plot normalize the total volume to volume per atom by dividing by the number of atoms in our simulation (32).

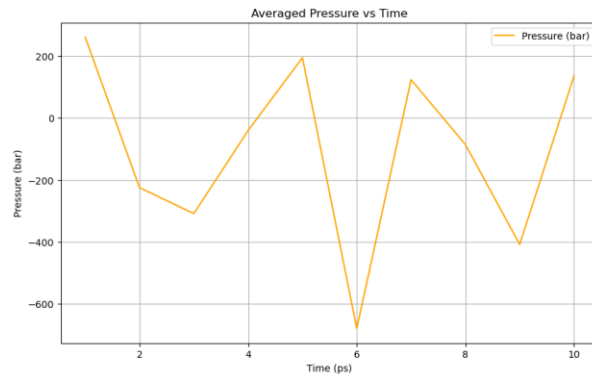
System Property	Time Steps to Equil	Simulation Time to Equil (picosecond)
Temperature (K)	3000	3
Pressure (bar)	N/A	N/A
Volume (Angstrom <sup>3</sup> )	3000	3
Total Energy (eV/atom)	3000	3

## Temperature vs Timestep

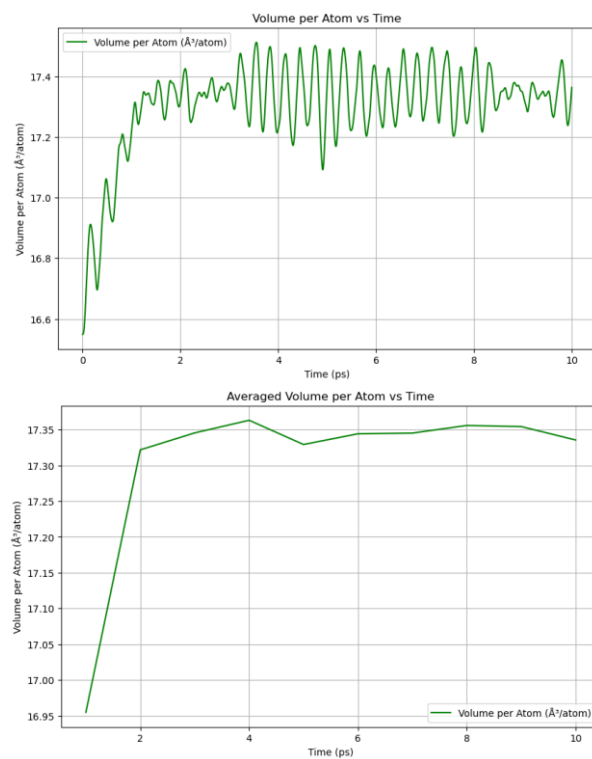


## Pressure vs Timestep

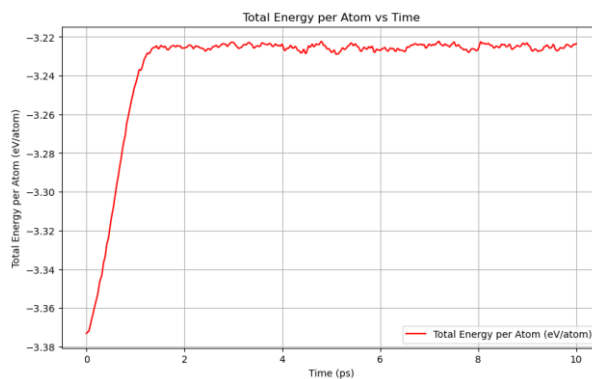


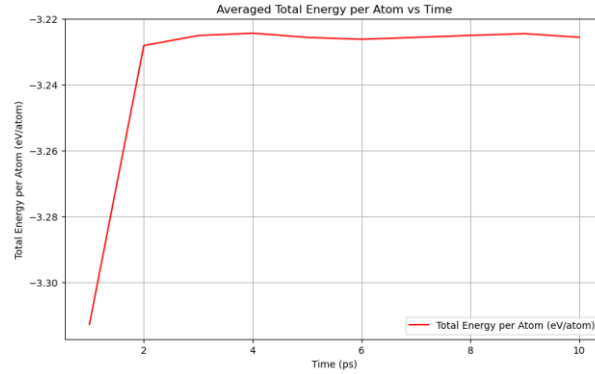


## Volume per atom vs Timestep



## Total Energy per atom vs Timestep





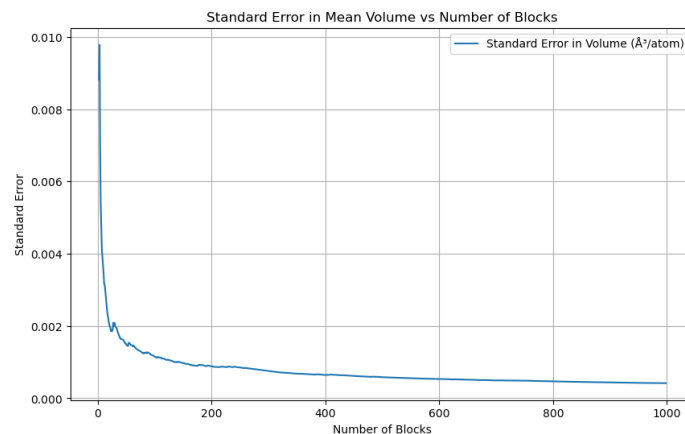
Data from `npteqfile.txt` and `npteqfile_averages.txt` is plotted using code from `problem_1-1.ipynb`.

## Problem 1.2

**Table 2:** Calculate the Standard Error in the Mean for Temperature and Energy of the system as a function of number of blocks included in the calculation. Also calculate the total simulation time for each corresponding number of blocks. Fill in the individual values for when the number of blocks is 5, 10, 100, 500, and 900 to complete Problem 1.2.

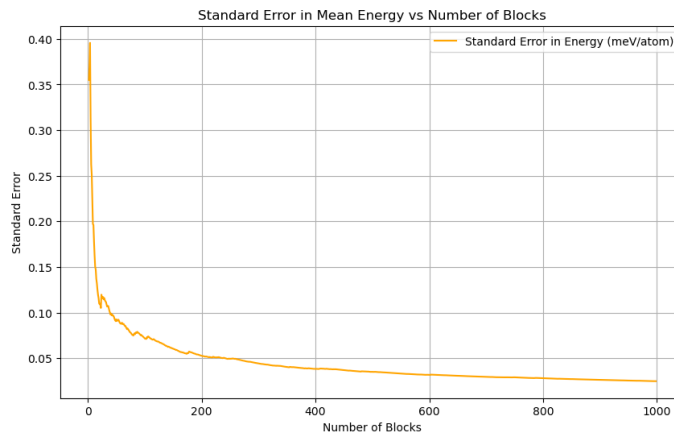
Number of Blocks	Simulation Time (picosecond)	Standard Error in the Mean Volume ( $\text{\AA}^3/\text{atom}$ )	Standard Error in the Mean Energy (meV/atom)
5	5	0.005368	0.313091
10	10	0.003468	0.196124
100	100	0.001152	0.072015
500	500	0.000583	0.035138
900	900	0.000438	0.026359

## Standard Error in Mean Volume vs Number of Blocks





## Standard Error in Mean Energy vs Number of Blocks



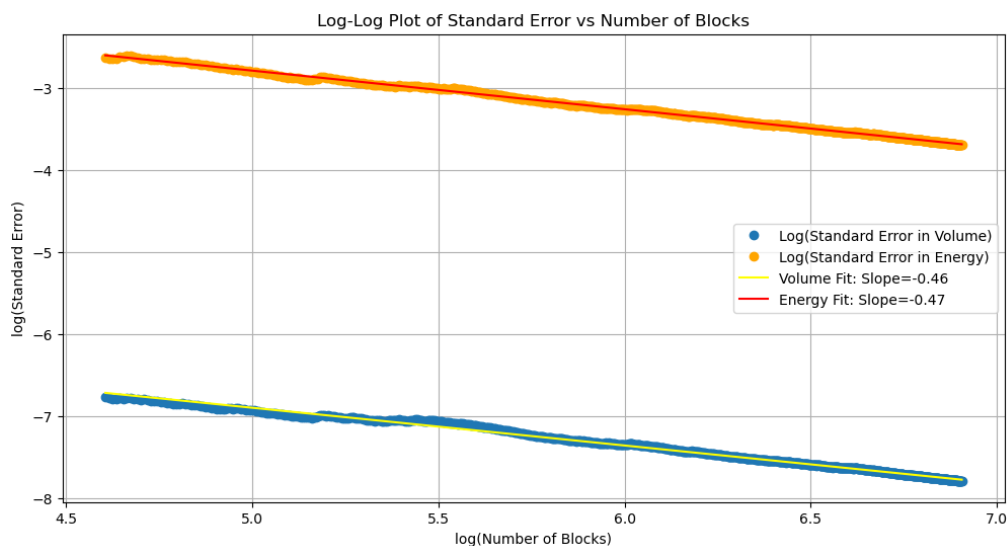
Data from `npteqfile_averages.txt` is plotted using code from `problem_1-2.ipynb`.

### Problem 1.3

**Table 3:** Plot the log of Standard Error in the Mean for Volume and Energy against the number of blocks included in the calculation. Do this for the range 100-998 under which we might expect to see a linear relationship emerge. Record the slopes of the linear fit for Volume and Energy to complete Problem 1.3.

	Volume	Energy
Linear fit slope	-0.46	-0.47

## Standard Error in the Mean for Volume and Energy vs Number of Blocks



Data from `npteqfile_averages.txt` is plotted using code from `problem_1-3.ipynb`.

Based on these values have we done enough steps in this equilibration run to be able to accurately extrapolate our volume precision? How about Energy?

**Answer:**

Based on the slope values in the log-log plot for standard error versus number of blocks, which are **-0.46 for volume** and **-0.47 for energy**, it appears that the standard error does not follow the **ideal -0.5 scaling** expected from an inverse square root relationship. This suggests that our system has **not reached a sufficient level of equilibration** for accurate error extrapolation.

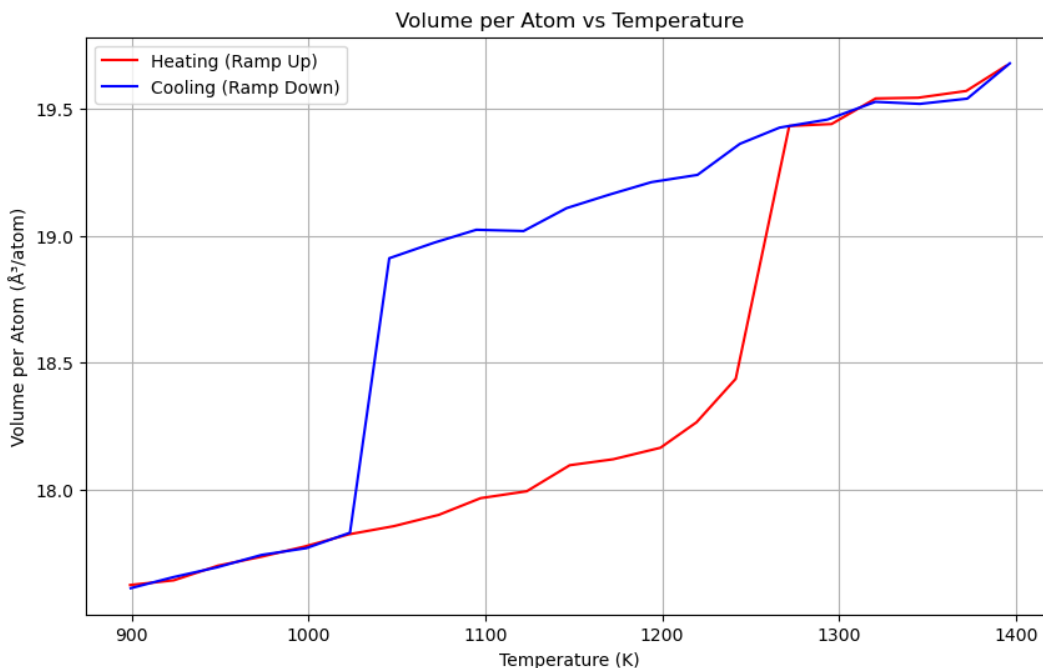
**Problem 2**

**Problem 2.1**

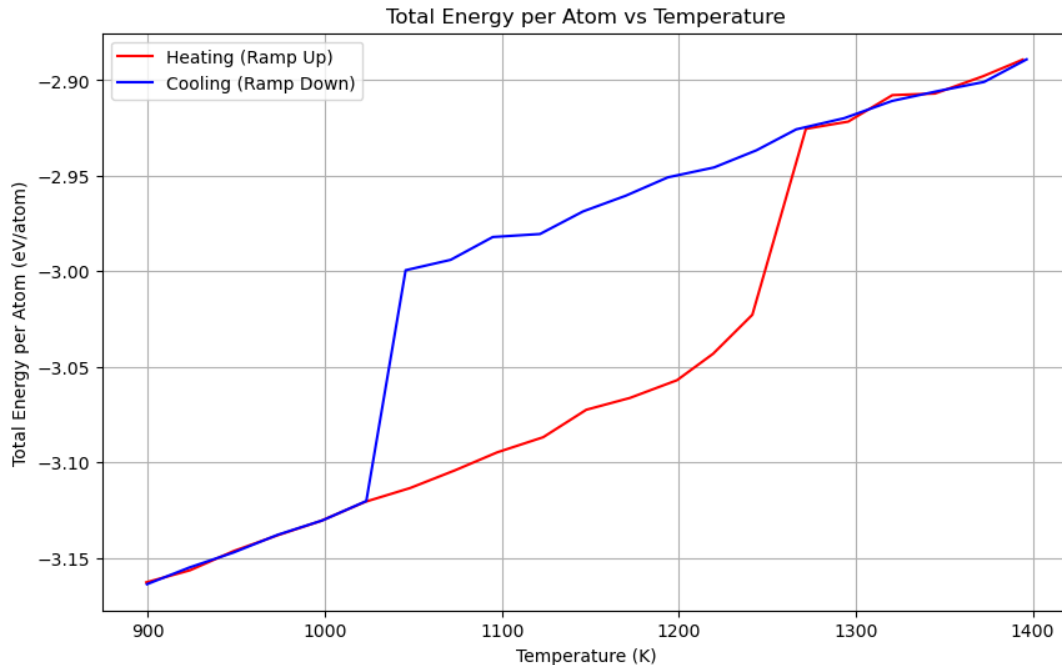
**Table 4:** Plot **average energy per atom** against temperature and identify the melting temperature from the heating and cooling data to complete Problem 2.1. If the system did not go through a transition just give a bound (e.g., on cooling  $T_{\text{melt}} < 900\text{K}$ )

Ramp	$T_{\text{melt}}$ (K)
Heating	1241.5
Cooling	1045.6

**Volume vs Temperature**



**Energy vs Temperature**



Data from `prall.txt` is plotted using code from `problem_2.ipynb`.

1. How large is the **hysteresis between ramping up and cooling down** and how might we **reduce** this? If you have a bound in the above Table 4, just take the bound values to estimate the bound on the hysteresis (e.g., hysteresis is  $> 180\text{K}$ ) and discuss ways to reduce this.

**Answer:**

The hysteresis between the ramp-up (heating) and ramp-down (cooling) processes can be calculated as the difference in melting temperatures:

$$\text{Hysteresis} = T_{\text{melt (heating)}} - T_{\text{melt (cooling)}}$$

From the results:

Melting temperature during heating: 1241.5K

Melting temperature during cooling: 1045.6K

So, the hysteresis is:

$$1241.5\text{K} - 1045.6\text{K} = 195.9\text{K}$$

The hysteresis between ramping up and cooling down is approximately 196K.

To reduce hysteresis between heating and cooling:

1. **Slow Heating/Cooling Rates:** Reducing the rate of temperature change allows more time for equilibration at each step, lowering kinetic barriers.

2. **Increase Equilibration Time:** Spending more time at each temperature point helps the system reach true equilibrium.
  3. **Larger System Size:** Increasing particle count can reduce finite-size effects, leading to more accurate phase transitions.
  4. **Enhanced Sampling Methods:** Techniques like Metadynamics can help overcome energy barriers, reducing hysteresis.
2. **Compare our calculated melting temperatures values** (if you only have one value then use that one) to experimental values. Let's use an **experimental value of 933K** [[https://www.engineeringtoolbox.com/melting-temperature-metals-d\\_860.html](https://www.engineeringtoolbox.com/melting-temperature-metals-d_860.html)]. What are our **percentage errors** vs. experiment? **What might lead to these errors?** For the melting temperature, use the average of the heating and cooling values.

**Answer:**

To calculate the percentage error compared to the experimental melting temperature, we first find the average of our calculated melting temperatures from heating and cooling:

$$T_{\text{melt, average}} = \frac{1241.5 + 1045.6}{2} = 1143.55 \text{ K}$$

The experimental melting temperature is given as 933K.

The percentage error relative to the experimental value is:

$$\text{Percent Error} = \left| \frac{T_{\text{melt (average)}} - T_{\text{melt (experimental)}}}{T_{\text{melt (experimental)}}} \right| \times 100\%$$

Substitute the values:

$$\text{Percentage Error} = \frac{|1143.55 - 933|}{933} \times 100\% \approx 22.57\%$$

Sources of Error:

1. **Simulation Rate:** Rapid heating/cooling rates can lead to inaccurate phase transition temperatures due to insufficient equilibration.
2. **Finite System Size:** Small system sizes can introduce finite-size effects that shift transition temperatures.
3. **Interatomic Potential:** The force field used in simulations may not perfectly replicate real material properties, leading to discrepancies.

## Problem 2.2

**Table 5:** From the energy curves in Problem 2.1 estimate the latent heat of fusion for the heating and cooling runs to complete Problem 2.2 (if you only have a transition on heating or on cooling just use that value and say “Not Available” for the other).

Ramp	Latent Heat (eV/atom)
Heating	0.097387
Cooling	0.120865

Heating:  $-2.92542 - (-3.022807) = 0.097387$

Cooling:  $-2.99947 - (-3.120335) = 0.120865$

3. Compare our calculated latent heat values to the experimental value of 10.79 kJ/mol. [<https://www.nuclear-power.net/aluminum-specific-heat-latent-heat-vaporization-fusion/>] Use the **heating run** values to be concrete. What are our percentage errors vs. experiment?

**Answer:**

The experimental latent heat of fusion for aluminum is 10.79 kJ/mol (approximately 0.1118 eV/atom). Calculate the percentage error for the calculated latent heat during heating based on this experimental value:

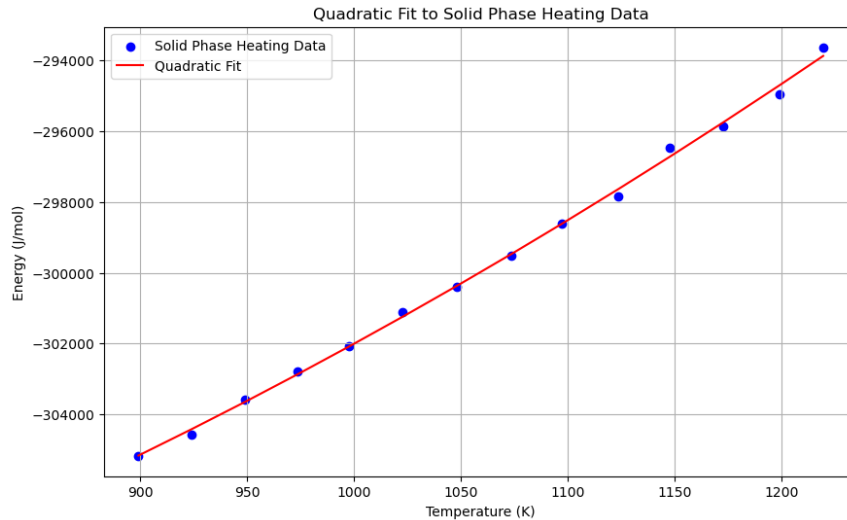
$$\text{Percentage Error} = \frac{|0.097387 - 0.1118|}{0.1118} \times 100\%$$

$$\text{Percentage Error} \approx 12.88\%$$

**Problem 2.3**

**Table 6:** From the heating energy curve in Problem 2.1 estimate a specific heat of the form  $A + BT$  by fitting a quadratic equation and taking the derivative *across the range of solid temperatures* (do not include the liquid as the fit must include only one phase). Enter the values here to complete Problem 2.3

Specific Heat Coefficient Label	Value
A	-2.6535 J/mol/K
B	0.0358 J/mol/K <sup>2</sup>



Data from `prall.txt` is plotted using code from `problem_2.ipynb`.

- Use our calculated values of A and B to predict the specific heat at 900K. Compare this specific heat prediction, not the values of A and B, with experimental values from Dehoff. What is our percentage error in predicting specific heat at this temperature? Note that Dehoff fits a more complete equation of the form  $C_p = A + BT + C/T^2 + DT^2$  with values  $A=31.38$ ,  $B=-1.64E-2$ ,  $C=-3.60E5$ ,  $D=2.08E-5$ . Units are in J/mol. [Dehoff, Thermodynamics in Materials Science].

**Answer:**

#### Step 1: Calculate Specific Heat at 900 K Using A and B

We use the following formula to calculate the specific heat:

$$c_p = A + B \times T$$

where:

- $A = -2.6535 \text{ J/mol/K}$
- $B = 0.0358 \text{ J/mol/K}^2$
- $T = 900 \text{ K}$

So, at 900 K, the specific heat is:

$$c_p(900 \text{ K}) = -2.6535 + 0.0358 \times 900$$

$$c_p(900 \text{ K}) \approx 29.5665 \text{ J/mol/K}$$

#### Step 2: Calculate Specific Heat at 900 K Using Dehoff's Model

Dehoff's model uses a more complex equation:

$$c_p = A + BT + \frac{C}{T^2} + DT^2$$

where:

- $A = 31.38 \text{ J/mol}$
- $B = -1.64 \times 10^{-2} \text{ J/mol/K}$
- $C = -3.60 \times 10^5 \text{ J/mol}$
- $D = 2.08 \times 10^{-5} \text{ J/mol/K}^2$
- $T = 900 \text{ K}$

Calculating the specific heat at 900 K using Dehoff's model:

$$c_p(900 \text{ K}) = 31.38 + (-1.64 \times 10^{-2}) \times 900 + \frac{-3.60 \times 10^5}{900^2} + (2.08 \times 10^{-5}) \times 900^2$$

Solving this gives:

$$c_p(900 \text{ K}) \approx 33.0236 \text{ J/mol/K}$$

### Step 3: Calculate the Percent Error

The percent error between the predicted and experimental values is given by:

$$\text{Percentage Error} = \frac{|29.5665 - 33.0236|}{33.0236} \times 100\%$$

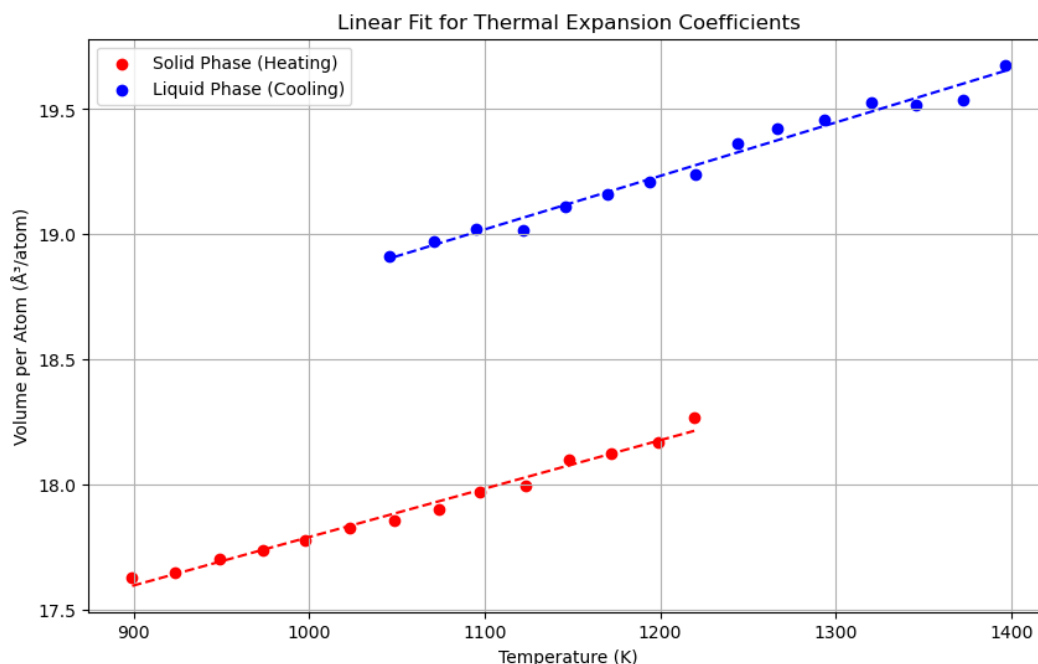
$$\text{Percentage Error} \approx 10.48\%$$

### Problem 2.4

**Table 7:** Plot the volume per atom against temperature and use the LINEST function in excel to fit lines to estimate thermal expansion coefficients. We'll use the heating curve to estimate the solid linear thermal expansion coefficient, and the cooling curve to estimate the liquid linear thermal expansion coefficient. Make sure to fit only the solid phase region on heating and the liquid phase region on cooling.

Note that the values obtained will be volumetric thermal expansion. To approximate linear values, we can divide our values by 3. We're using the LINEST function because it provides a nice way to also get error bars on these numbers. Using the array functionality the second row of each fit will give the error. Include these errors in your answers. These provide the answer to Problem 2.4.

Ramp	Solid Linear Thermal Expansion coefficient (m/(m-K)=1/K)	Liquid Linear Thermal Expansion coefficient (m/(m-K)=1/K)
Up	$1.20179\text{e-}05/3 = 4.00597\text{e-}06$	Leave Blank
Down	Leave Blank	$1.23325\text{e-}05/3 = 4.11083\text{e-}06$



Data from prall.txt is plotted using code from problem\_2.ipynb.

5. Compare our thermal expansion coefficients with experimental data for solid phase ( $2.23 \times 10^{-5}$  (m/(m-K)=1/K)) [[https://www.engineeringtoolbox.com/linear-expansion-coefficients-d\\_95.html](https://www.engineeringtoolbox.com/linear-expansion-coefficients-d_95.html)]. What scale error are we getting? 5%, 10%, Factor of 2, Factor of 10?

**Answer:**

Experimental Solid Linear Thermal Expansion Coefficient:  $2.23 \times 10^{-5}$  1/K

Calculated Solid Linear Thermal Expansion Coefficient:  $4.00597 \times 10^{-6}$  1/K

The percentage error can be calculated using the formula:

$$\text{Percent Error} = \frac{|\text{Calculated Value} - \text{Experimental Value}|}{\text{Experimental Value}} \times 100\%$$

$$\text{Percent Error} = \frac{|4.00597 \times 10^{-6} - 2.23 \times 10^{-5}|}{2.23 \times 10^{-5}} \times 100\%$$

$$\text{Percent Error} = \frac{|4.00597 \times 10^{-6} - 2.23 \times 10^{-5}|}{2.23 \times 10^{-5}} \times 100\%$$

$$\text{Percent Error} = \frac{1.829403 \times 10^{-5}}{2.23 \times 10^{-5}} \times 100\%$$

$$\text{Percent Error} \approx 82.04\%$$

$$\text{Scale Error} = \frac{\text{Experimental Value}}{\text{Calculated Value}}$$

$$\text{Scale Error} = \frac{2.23 \times 10^{-5}}{4.00597 \times 10^{-6}} \approx 5.57$$

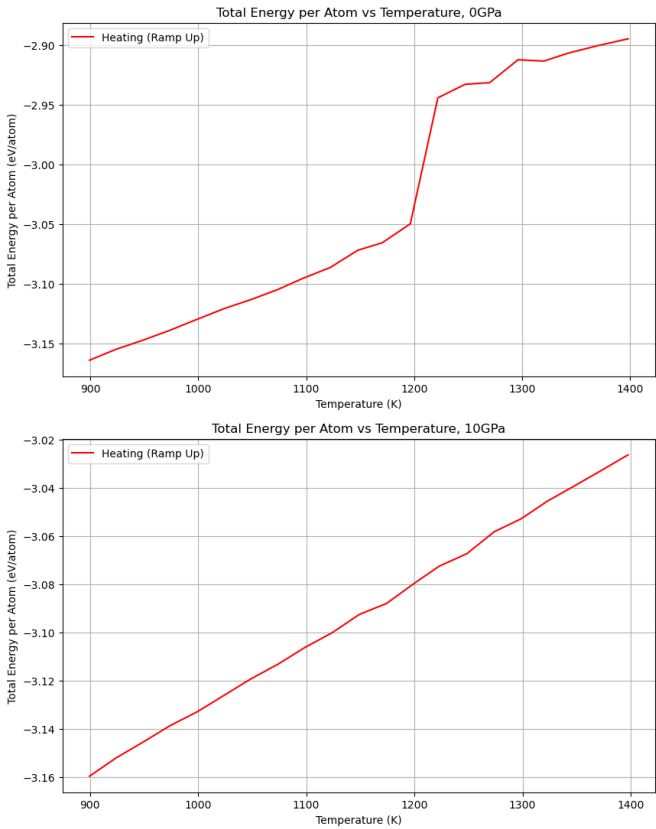
### Problem 2.5

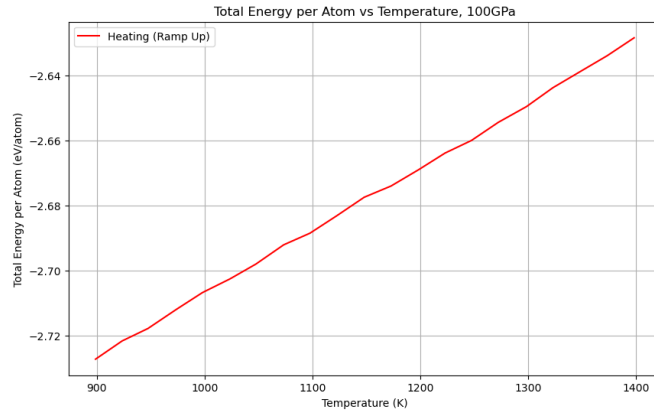
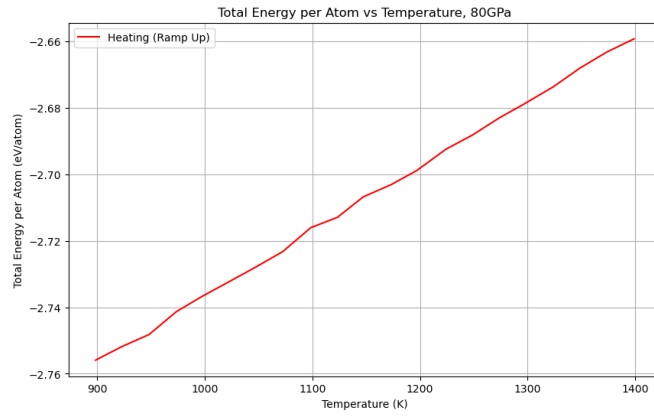
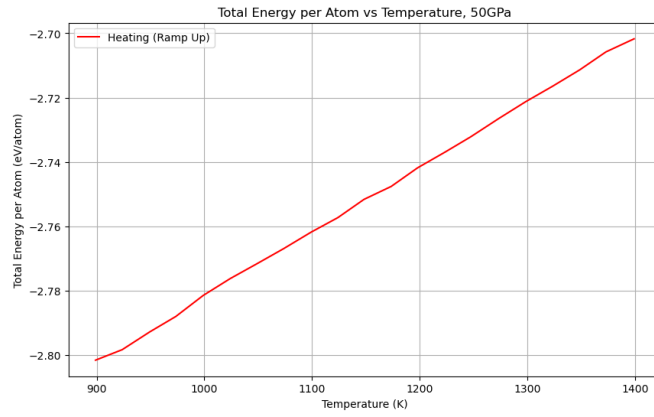
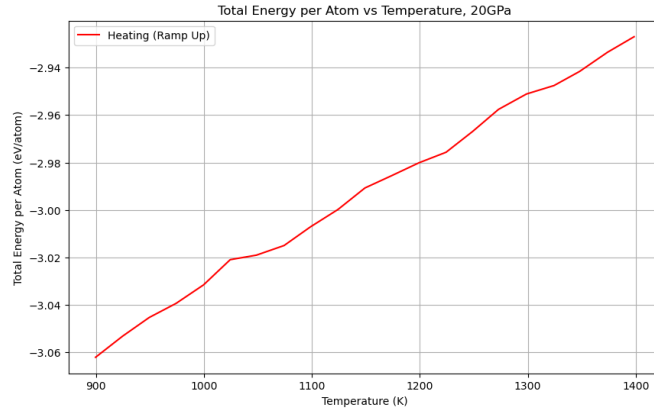
**Table 8:** Values of relevant terms in the determination of  $\Delta S(P_0, T_m)$ .  
Please use the SI units provided.

<b>T<sub>m</sub> (K)</b>	1196.45
<b>P,T values for finding <math>\left(\frac{\partial P}{\partial T}\right)_{CC,(P_0,T_m)}</math></b>	
<b>P (Pa)</b>	<b>T<sub>m</sub> (K)</b>
0 * 10 <sup>9</sup>	1196.45
10 * 10 <sup>9</sup>	1273.79
20 * 10 <sup>9</sup>	1024.35
50 * 10 <sup>9</sup>	998.96



$80 \cdot 10^9$	1098.53
$100 \cdot 10^9$	1323.04
<b>Key Values</b>	
$\left(\frac{\partial P}{\partial T}\right)_{CC,(P_0,T_m)}$ (Pa/K)	4.18e+07
$\Delta S = \Delta V \left(\frac{\partial P}{\partial T}\right)_{CC}$ (eV/(K atom))	2.61e-04
$T_m$ (K)	1196.45
$\Delta H(P_0, T_m)$ (eV/atom)	3.01e-01
$\frac{\Delta S(P_0, T_m)}{\Delta H(P_0, T_m)} = \frac{1}{T_m}$ (eV/(K atom))	2.61e-04





Data is plotted using code from `problem_2.ipynb`.

6. Compare  $\Delta S = \Delta V \left( \frac{\partial P}{\partial T} \right)_{CC}$  and  $\Delta S(P_0, T_m) = \frac{\Delta H(P_0, T_m)}{T_m}$  ? What is the percentage error difference and what might give rise to this difference? How do these values compare to experimental value of 9.28 J/(mol-K) and what might be the source of the errors?

**Answer:**

```
# Given values
S_calculated = 2.61e-4 # eV/(K atom)
S_experimental = 9.62e-4 # eV/(K atom), converted from 9.28 J/(mol-K)

# Calculate the percentage error
percent_error = abs((S_calculated - S_experimental) / S_experimental) * 100
print(f"Percentage Error: {percent_error:.2f}%")
```

✓ 0.0s

Percentage Error: 72.87%

Explanation of Possible Sources of Error:

1. **Finite-Size Effects:** The small system size in simulations can lead to deviations from bulk behavior, affecting the entropy values.
2. **Interatomic Potential:** The potential model used may not perfectly replicate aluminum's true melting properties, affecting entropy and enthalpy values.
3. **Equilibration Time:** Insufficient equilibration, especially for phase transition points, can lead to discrepancies.
4. **Hysteresis Effects:** Differences in melting and freezing points due to hysteresis impact the calculated  $T_m$ , leading to errors.