Discussion

Other Potential Models for Molecular Dynamics Study

Morse interaction model

Morse Interaction model is described as a combination of Brownian motion and improved OU process. Under this model, X_t is a Markov process satisfying the stochastic differential equation (SDE):

$$dX_t = -U'(X_t)dt + \sigma dB_t$$

And U'(x) is the derivative of the Morse potetial energy function:

$$U'(x) = \gamma \cdot (1 - e^{-\alpha \cdot (x-u)})$$

The distance X_t is set to be strictly positive, $X_t > 0$. When X_t is too large, repulsive forces allow bond breaking to occur - the molecules again resemble brownian motion. When the donor and acceptor again get close to each other, the bond is reformed and exchange of photons takes place. The dynamics then resembles OU process.

Morse interaction model is a comprehensive approach towards dynamics study of molecular interaction.

S.Schelstraete, H. Verschelde (1999). Molecular Dynamics. Retrieved from https://www.sciencedirect.com/topics/chemistry/morse-potential

Lennard-Jones Potential

The interaction between two non-bonded and un-charged atoms, known as Van der Waals interaction, has been expressed in terms of potential energy. Lennard-Jones potential is probably the most famous pair potential descriping interatomic Van der Waals forces. It consists of two parts:

- 1) A steep repulsive term; and
- 2) A smooth attractive term

$$V(r) = 4\epsilon \left[\left(\frac{\sigma}{r} \right)^{12} - \left(\frac{\sigma}{r} \right)^{6} \right]$$

In particular, V(x) is the intermolecular potential between the two molecules, and r is the distance of separation between both molecules.

Apart from being a widely used model itself, Lennard-Jones potential also sometimes forms one of 'building blocks' of many force fields, due to its computational expediency.

However, Lennard-Jones potential is not designated to model donor-acceptor interactions.

Libretexts. (2019, June 5). Lennard-Jones Potential. Retrieved from https://chem.libretexts. org/Bookshelves/Physical_and_Theoretical_Chemistry_Textbook_Maps/Supplemental_Modules_(Physical_and_Theoretical_Chemistry)/Physical_Properties_of_Matter/Atomic_and_Molecular_Properties/Intermolecular_Forces/Specific_Interactions/Lennard-Jones_Potential

Research Direction 2

For experiments simulated from OU, we can see a constant improvement in picking the correct model when γ goes up. As $\gamma \to 0$, $\omega_{\Delta t} \to 1$ and τ is still a constant, (todo: equation) becomes $X_{t+\Delta t}|X_t \sim N(X_t,0)$, which is in the same form as (todo:equation). Thus, it is difficult to distinguish the OU simulation from the BM

model. As γ goes up, $\omega_{\Delta t} \to 0$, two model equations differ in mean and variance, and hence becomes easier to distinguish.

For experiments simulated from BM, we can see a general improvement in picking the correct model when σ_{BM} goes up, however, the improvement stops at $\sigma_{BM}=3$ for most of the experiments. We analyze in the following two aspects. First of all, when $\sigma_{BM}\to 0$, (todo:equation) becomes $X_{t+\Delta t}|X_t\sim N(X_t,0)$. From the discussion in the previous paragraph, when $\gamma\to 0$, (todo:equation) also becomes $X_{t+\Delta t}|X_t\sim N(X_t,0)$. Thus, if we do parameter inferences relatively well, it will be difficult to tell apart two models. Furthermore, when σ becomes too large with respect to β_0 and β_1 , for example, when $\sigma=3$ with respect to $\beta_0=5$ and $\beta_1=1$, the probability of X_t being large becomes higher. In this case, $\exp(\beta_0-\beta_1X_t)$ becomes close to 0, and the generated Y_t from Poission (todo:equation) is more likely to be 0. Consequently, Y_t does not provide enough information to the underlying generated process, and hence difficult to tell apart BM model from OU model.