GAS IDEALA KONTZEPTUAREN DEFINIZIOAREN BILA

* $U = U(T,V)$ $\quad dU = \left(\frac{\partial U}{\partial T}\right)_V dT + \left(\frac{\partial U}{\partial V}\right)_T dV$

ESPANTSIO ASKEA $dU = 0$ ETA ONARTUZ $dT = 0 \Rightarrow \left(\frac{\partial U}{\partial V}\right)_T = 0$

* $U = U(T,P)$ $\quad dU = \left(\frac{\partial U}{\partial T}\right)_P dT + \left(\frac{\partial U}{\partial P}\right)_T dP$

ESPANTSIO ASKEA $dU = 0$ ETA ONARTUZ $dT = 0 \Rightarrow \left(\frac{\partial U}{\partial P}\right)_T = 0$

dU + PROZESU ISOTERMIKOA $\boxed{U = U(T)}$

ESPERIENTZIAREN ARABERA: $\left(\frac{\partial U}{\partial P}\right)_T \neq 0 \Rightarrow U = U(T,P)$

$\left(\frac{\partial U}{\partial P}\right)_T = f(T) \Rightarrow U = f(T)P + F(T)$

GAS IDEALA

$P \cdot V = nRT$

$\left(\frac{\partial U}{\partial P}\right)_T = 0$

$\left[\left(\frac{\partial U}{\partial V}\right)_T = 0\right]$

$\Rightarrow U = U(T)$

EGOERA-EKUAZIOAK

$P \cdot V = nRT$ MEKANIKOA

$U = U(T)$ TERMIKOA