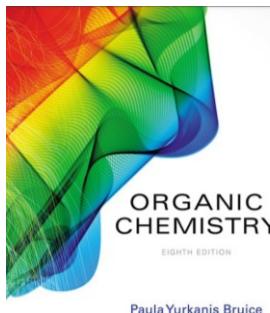




# CHM 2210

# Lecture 14



Paula Yurkanis Bruice

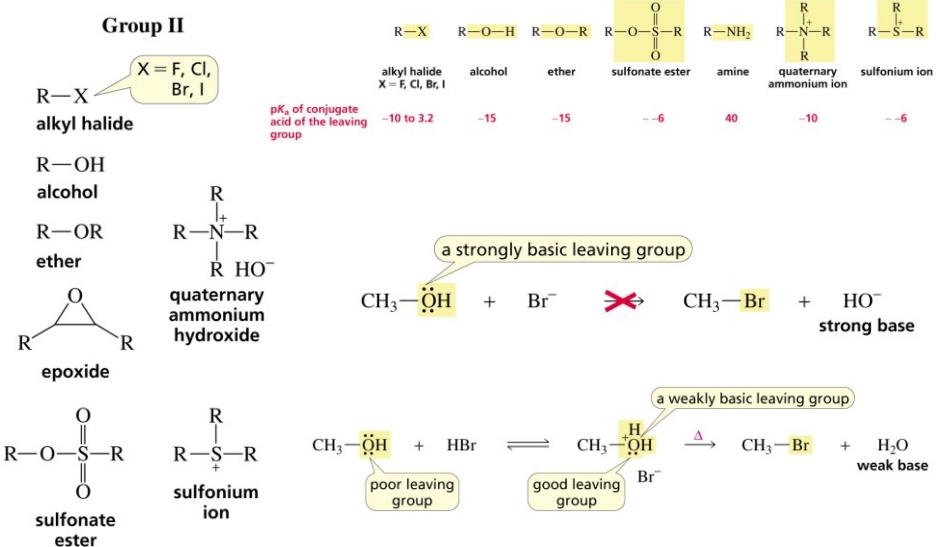
## Outline

- ❖ Alcohols
    - Halogenation
    - Dehydration
    - Oxidation
    - Diol Synthesis
  - ❖ Epoxides
    - Acidic
    - Basic
  - ❖ Amines
  - ❖ Thiols

Soumya (SAM) Bhattacharya

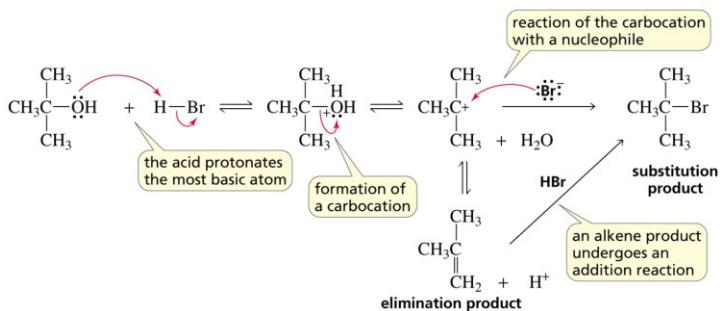
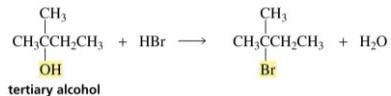
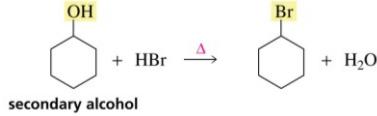
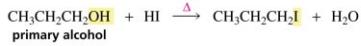


# Leaving Group Saga





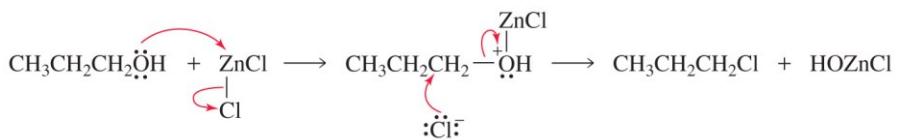
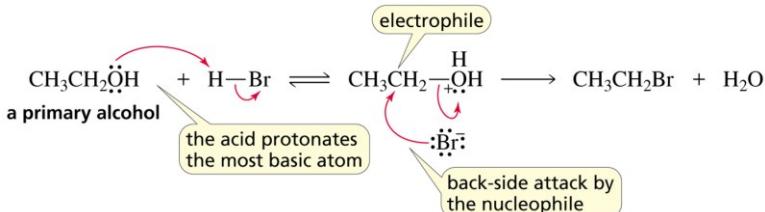
## Alcohol – I: Halogenation



3



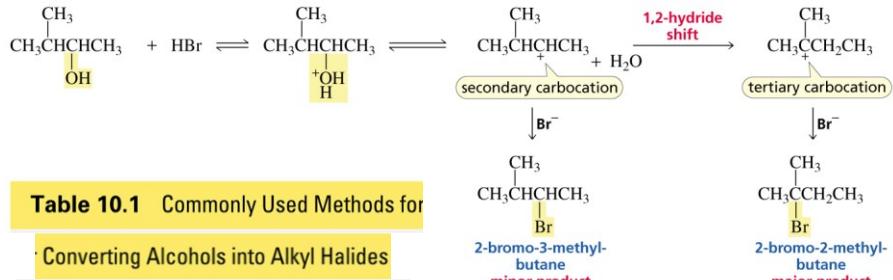
## Alcohol – II: Halogenation



4



## Alcohol – III: Halogenation



**Table 10.1** Commonly Used Methods for

### Converting Alcohols into Alkyl Halides

$\text{ROH} + \text{HBr} \xrightarrow{\Delta}$	$\text{RBr}$
$\text{ROH} + \text{HI} \xrightarrow{\Delta}$	$\text{RI}$
$\text{ROH} + \text{HCl} \xrightarrow{\text{ZnCl}_2}$	$\text{RCl}$
$\text{ROH} + \text{PBr}_3 \xrightarrow{\text{pyridine}}$	$\text{RBr}$
$\text{ROH} + \text{PCl}_3 \xrightarrow{\text{pyridine}}$	$\text{RCl}$
$\text{ROH} + \text{SOCl}_2 \xrightarrow{\text{pyridine}}$	$\text{RCl}$

$\text{R-OH}$   
*alcohol*

$\xrightarrow{\text{HX} \quad \Delta}$

$\text{R-X}$   
*alkyl halide*  
 $\text{X} = \text{Cl}, \text{Br}, \text{I}$

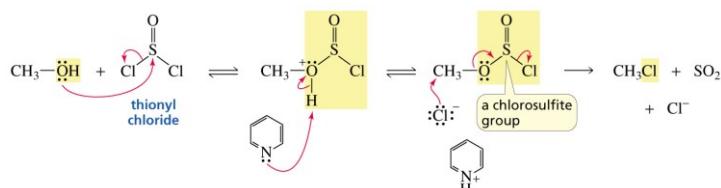
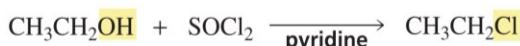
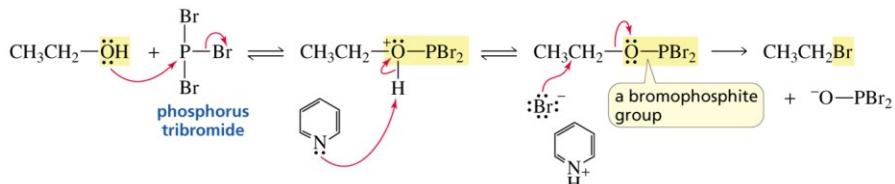
$\begin{array}{c} \text{R}-\text{OCH}_3 \\ \diagup \\ \text{R}-\text{O} \end{array} \xrightarrow{\text{C}\equiv\text{CCH}_3} \text{R}-\text{C}\equiv\text{CCH}_3$

$\begin{array}{c} \text{R}-\text{C}\equiv\text{N} \\ \diagup \\ \text{R}-\text{C} \end{array}$

5



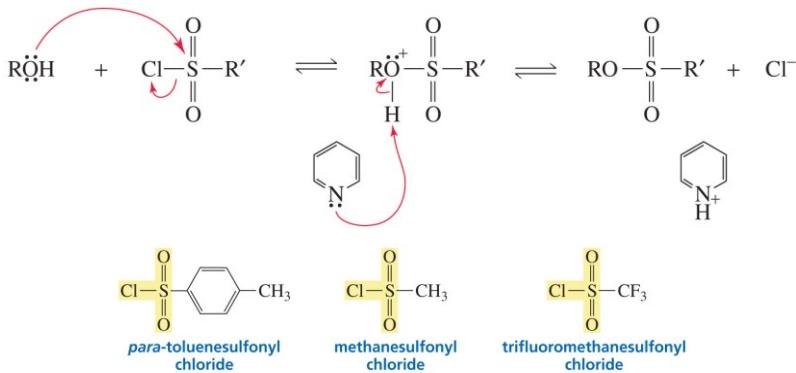
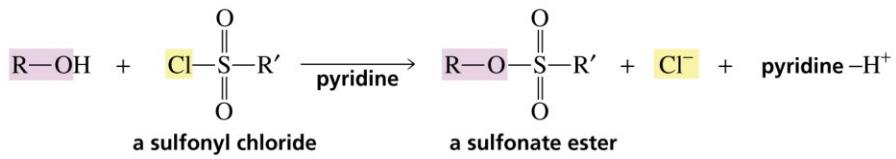
## Alcohol – IV: Halogenation



6



## Alcohol – V: Sulfonation

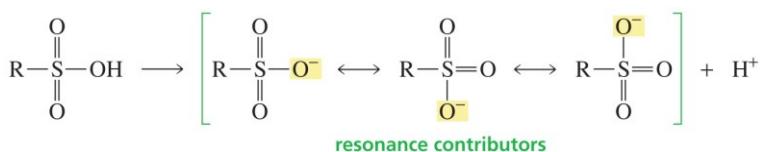
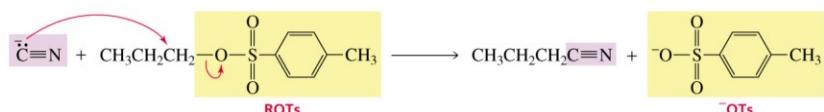
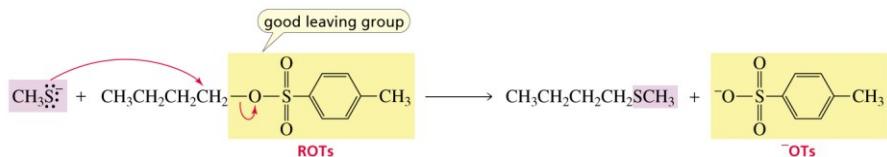


tosyl chloride = TsCl

7



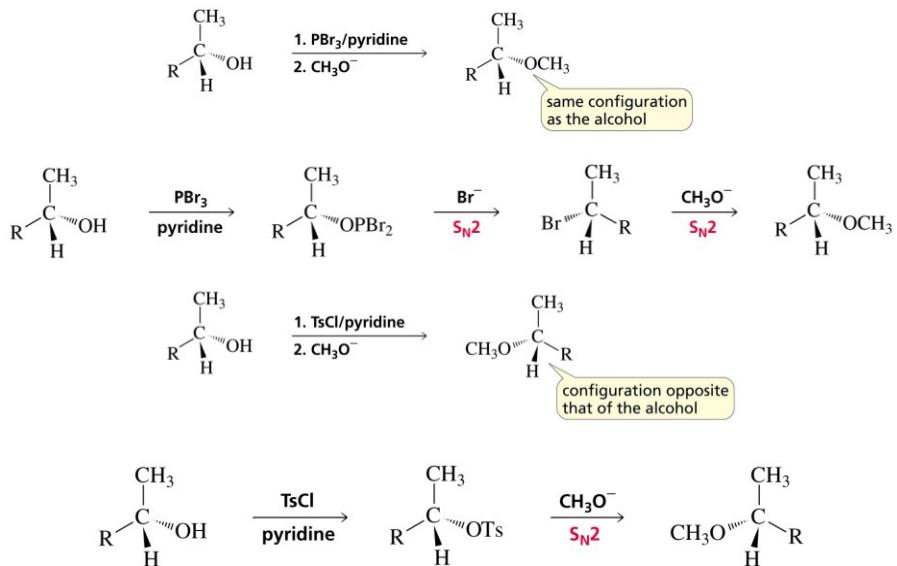
## Alcohol – VI: Sulfonation



8



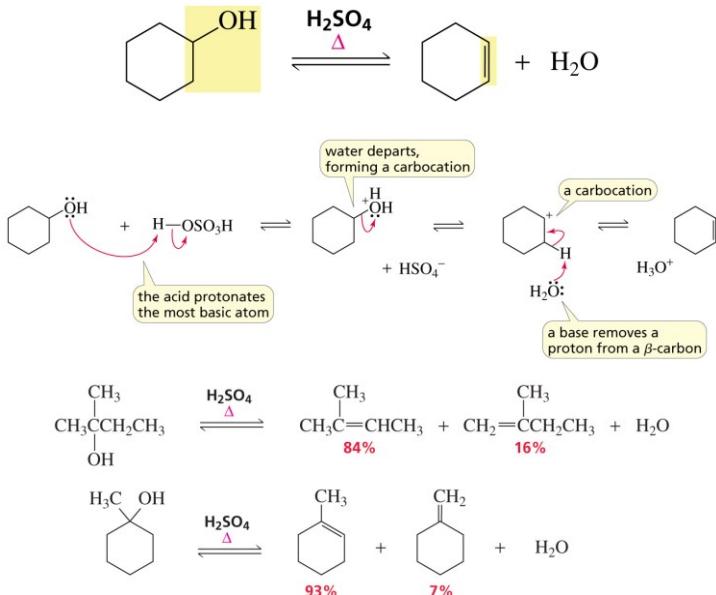
## Alcohol – VII: Configuration Considerations



9



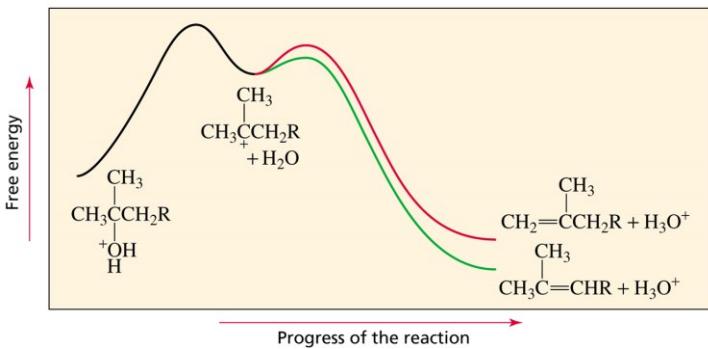
## Alcohol – VIII: Dehydration



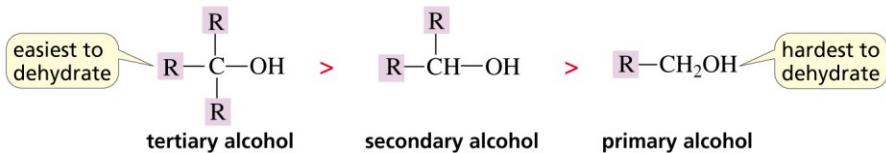
10



## Alcohol – IX: Dehydration



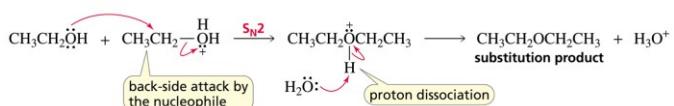
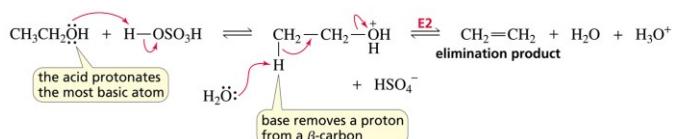
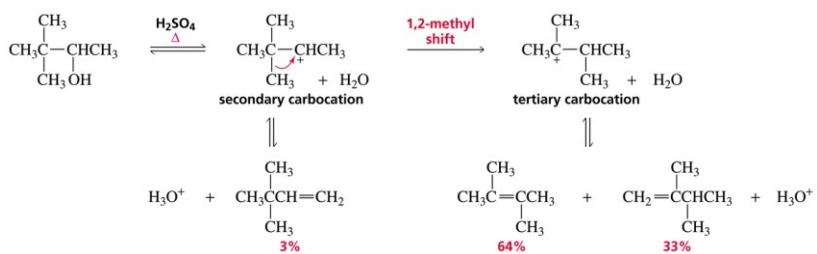
**relative ease of dehydration**



11



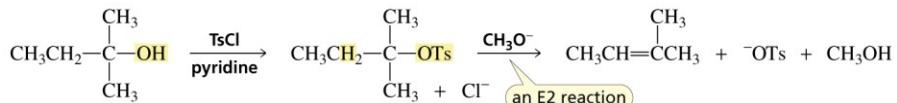
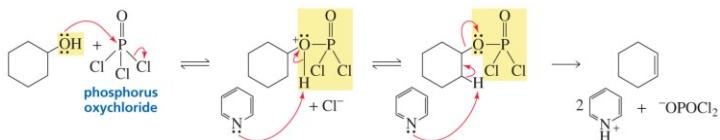
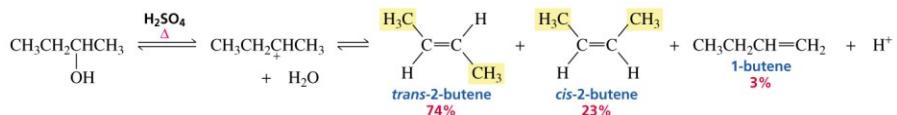
## Alcohol – X: Dehydration



12



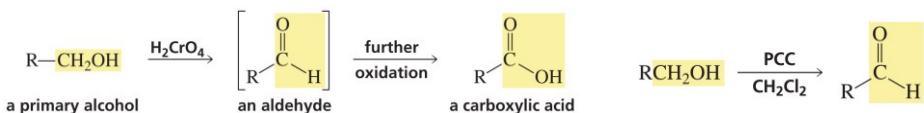
## Alcohol – XI: Dehydration



13



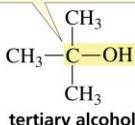
## Alcohol – XII: Oxidation



### ❖ Jones Reagent

- Dichromate Salt
- Concentrated Sulfuric Acid
- Acetone

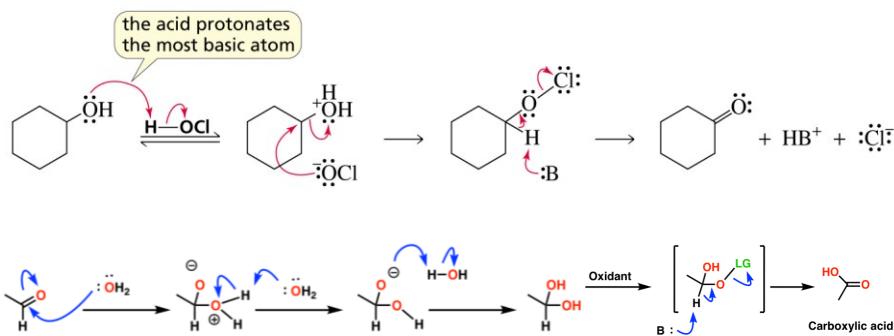
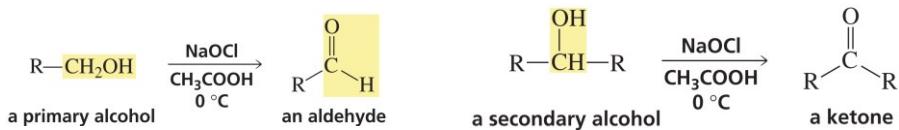
this C is not bonded to an H, so the alcohol cannot be oxidized to a carbonyl compound



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## Alcohol – XIII: Oxidation

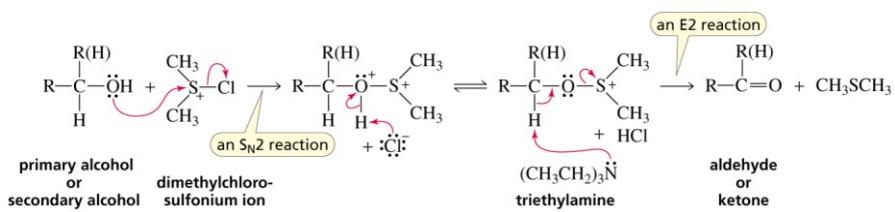
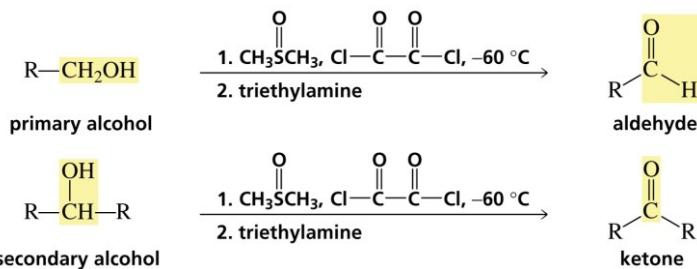


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## Alcohol – XIV: Oxidation

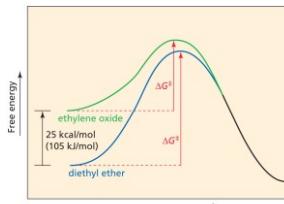
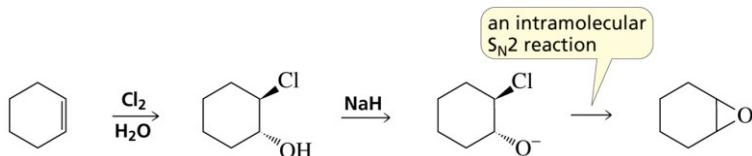
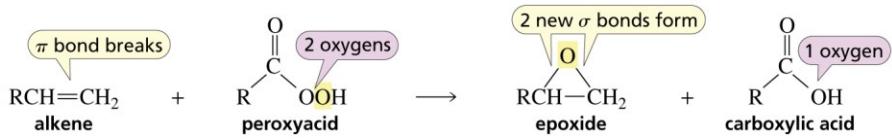
### ❖ Swern Oxidation



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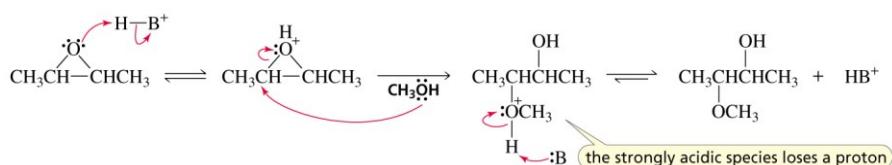
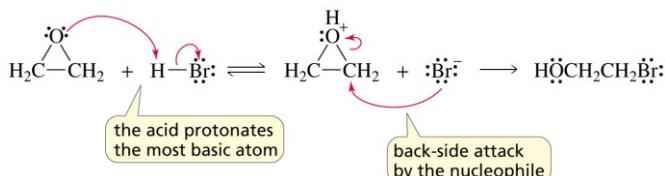
## Epoxides – I: Synthesis



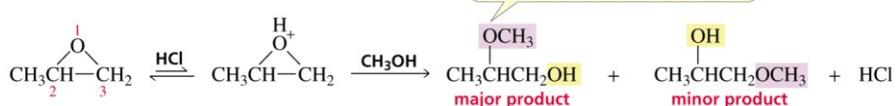
17



## Epoxides – II: Acid Catalyzed Ring Opening



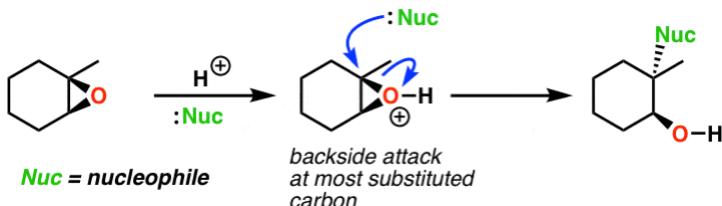
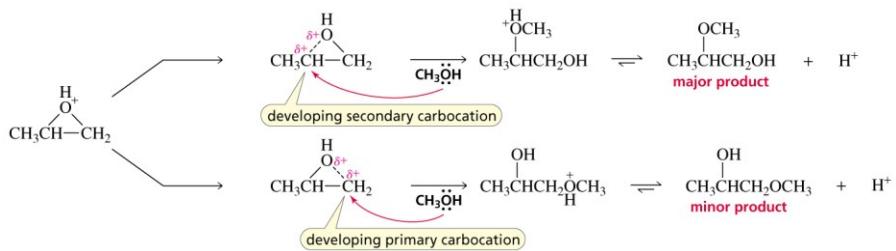
obtained by nucleophilic attack on the more substituted ring carbon



18



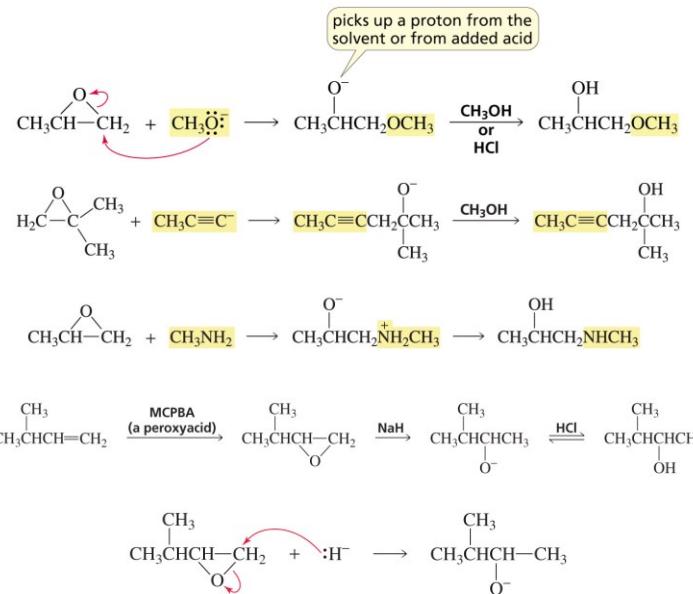
## Epoxides – III: Acid Catalyzed Ring Opening



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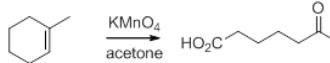
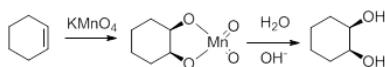
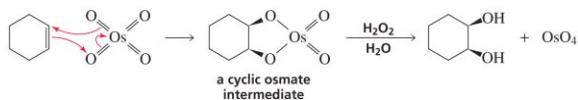
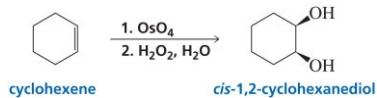
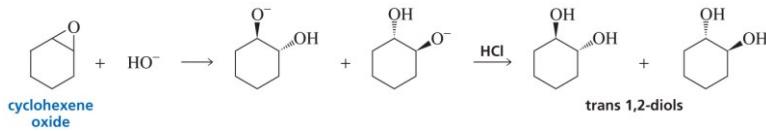
## Epoxides – IV: Base Catalyzed Ring Opening



20



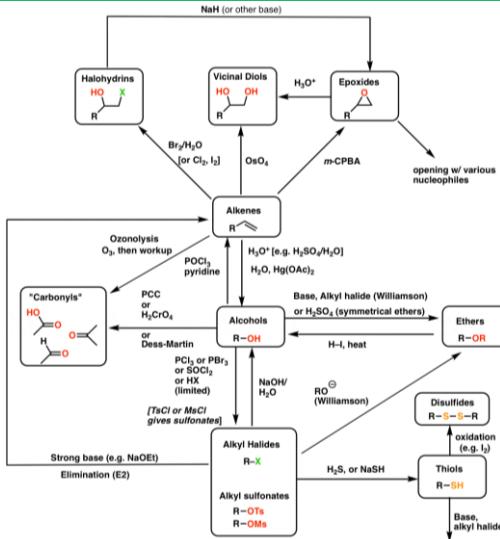
## Diols



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## Reaction Map – I: Alcohols



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