Ammonia is produced in the equation below:

$$N_{2 (g)} + 3H_{2(g)} \Rightarrow 2NH_{3(g)} \Delta H = -92 \text{ kJ}$$

The forward reaction is exothermic.

- 1 Increase in temperature will reduce the yield of ammonia at equilibrium.
- 2 Increase in pressure will increase the yield of ammonia at equilibrium.

Which of these statements are correct?