Ammonia is produced in the equation below:

 $N_2 + 3H_2 \rightleftharpoons 2NH_3 \qquad \Delta H = -92 \text{ kJ}$

The forward reaction is exothermic.

- 1. Increase in temperature will increase the yield of ammonia at equilibrium.
- 2. Increase in pressure will increase the yield of ammonia at equilibrium.

Which of these statements is false?