Ammonia is produced in the equation below:

$$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g) \quad \Delta H = -92 \text{ kJ}$$

The forward reaction is exothermic.
1 Increase in temperature will reduce the yield of ammonia at equilibrium.
2 Increase in pressure will increase the yield of ammonia at equilibrium.

Which of these statements are correct?