

Ch En HW 3

1.) $T_1 = 300\text{K}$, $P_1 = 101\text{ kPa}$, $V_1 = 0.5\text{ m}^3$
 $P_2 = 900\text{ kPa}$

$$W = \int F dx = \int P dV$$

$$P = \frac{nRT}{V} \quad \therefore W = - \int_{V_1}^{V_2} \frac{nRT}{V} dV = -nRT [\ln V] = -nRT (\ln V_2 - \ln V_1)$$

$$nRT = PV \rightarrow W = PV \ln \left| \frac{V_2}{V_1} \right|$$

$$V_2 = \frac{nRT}{P_2} \quad V_1 = \frac{nRT}{P_1} \rightarrow V_2 P_2 = V_1 P_1$$

$$\therefore W = P_1 V_1 \ln \left| \frac{P_1}{P_2} \right| \rightarrow W = 101\text{ kPa} \cdot 0.5\text{ m}^3 \cdot \ln \left| \frac{101\text{ kPa}}{900\text{ kPa}} \right|$$

$$\frac{V_2}{V_1} = \frac{P_1}{P_2}$$

$$W = nRT \ln \left| \frac{P_1}{P_2} \right|$$

$$n = \frac{m}{M}$$

$$m = \rho \cdot V$$

$$M = 28.0134\text{ g/mol}$$

$$m = 1.25\text{ g} = 500\text{ L}$$

$$n = 22.32$$

$$m = 625.3\text{ g}$$

$$\therefore W = 22.32\text{ mol} \cdot 8.314\text{ J/mol}\cdot\text{K} \cdot 300\text{K} \cdot \ln \left| \frac{101000}{900000} \right|$$

$$= 121569.7\text{ J} = 122\text{ kJ}$$

b) Heat removed from the gas $\Delta U = Q - W$, $\Delta U = 0$ (isothermal)

$$Q = W$$

$$\therefore Q = +122 \rightarrow -122\text{ kJ removed from gas}$$

2.) $W = \int_a^b P dV \rightarrow W = P_x \int_a^b dV = P_x [V]_a^b = P_x [b - a] = W$

$$P_1 = 750\text{ kPa}$$

$$P_{avg} = \frac{3550}{6} = 591.67\text{ kPa}$$

$$P_2 = 500$$

$$P_3 = 400$$

$$P_4 = 400$$

$$P_5 = 700$$

$$P_6 = 800$$

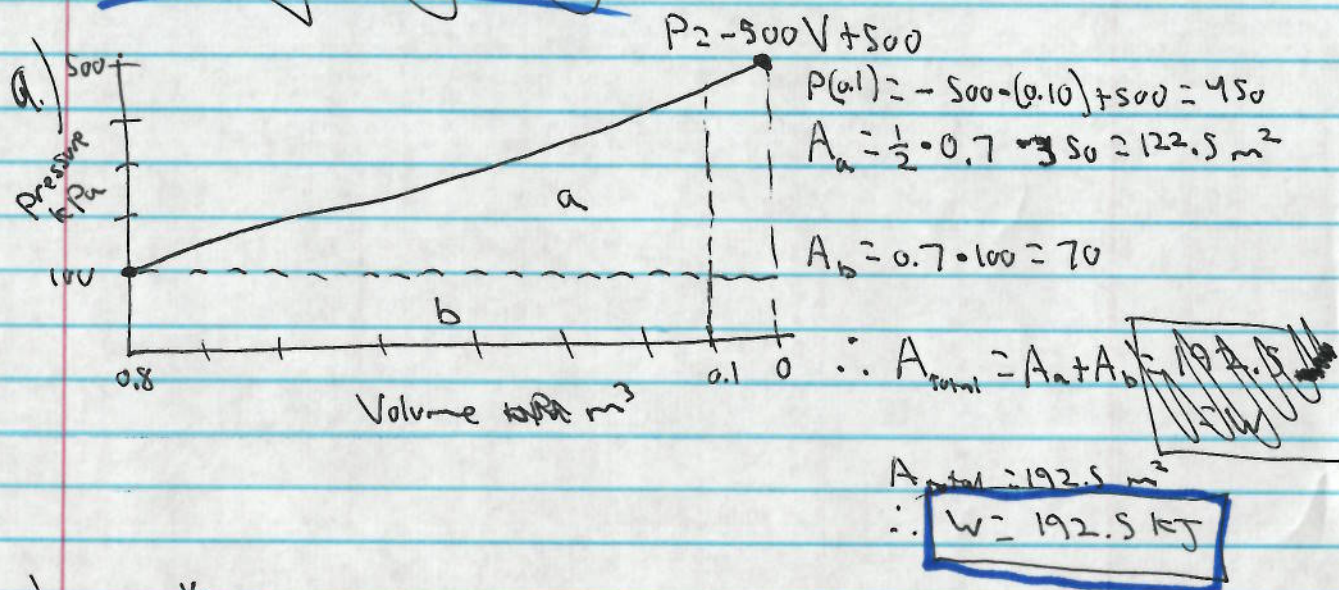
$$a = 0.01\text{ m}^3$$

$$\therefore W = P_{avg} [0.02 - 0.01] = 5.9\text{ kJ}$$

Ch E HW3

3) $V_1 = 0.8 \text{ m}^3$ $V_2 = 0.10 \text{ m}^3$, $P = aV + b$
 $a = -500 \text{ kPa/m}^3$ $b = 500 \text{ kPa}$

~~Plot on next page~~



b) $W = - \int_{V_1}^{V_2} P dV$ $P = P_{\text{avg}} = 275 \text{ kPa}$

$\therefore W = -275 [V_2 - V_1] = -275 (-0.7) = 192.5 \text{ kJ}$

Compression of a Gas

