

**MARK SCHEME for the October/November 2010 question paper
for the guidance of teachers**

9701 CHEMISTRY

9701/33

Paper 3 (Advanced Practical Skills), maximum raw mark 40

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

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Question	Sections	Indicative material	Mark	
1 (a)	PDO layout	I Volume given for Rough titre and accurate titre details tabulated.	1	
	MMO Collection	II In the correct spaces, records Initial and final burette readings for Rough titre and ; Initial and final burette readings and , volume of FB 2 added recorded for each accurate titre <i>Headings should match readings.</i> <i>Do not award this mark if:</i> <i>50(.00) is used as an initial burette reading;</i> <i>More than one final burette reading is 50.(00);</i> <i>Any burette reading is greater than 50.(00)</i>	1	
	MMO Decisions	III Has two uncorrected, accurate titres within 0.1 cm^3 <i>Do not award this mark if having performed two titres within 0.1 cm^3 a further titration is performed which is more than 0.10 cm^3 from the closer of the initial two titres, unless a fourth titration, within 0.1 cm^3 of the third titration or of the first two titres has also been carried out.</i>	1	
	PDO Recording	IV All accurate burette readings (initial and final) recorded to nearest 0.05 cm^3 . <i>Assessed on burette readings only.</i>	1	
	MMO Quality	V, VI and VII Round any burette readings to the nearest 0.05 cm^3 Check and correct subtractions in the titre table. Select the “best” titre using the hierarchy: two identical; titres within 0.05 cm^3 , titres within 0.10 cm^3 etc. Award <u>V, VI and VII</u> for a difference to Supervisor within 0.20 cm^3 Award <u>V and VI only</u> for a difference of $0.20+ \text{ cm}^3 - 0.40 \text{ cm}^3$ Award <u>V only</u> for a difference of $0.40+ \text{ cm}^3 - 0.80 \text{ cm}^3$ <i>If the selected “best” titres are $> 0.50 \text{ cm}^3$ apart, cancel one of the Q marks awarded.</i>	3	
				[7]

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(d)	ACE Interpretation	Gives $0.1(0) \text{ cm}^3$ as the maximum error in (i). <i>Ignore any sign</i> and the expression $^{0.1}/_{\text{cand titre in (b)}} \times 100$ in (ii)	1	[2]
		Evaluates $^{0.06}/_{25.0} \times 100$ in step (iii) <i>Accept only 0.240 or 0.24,</i> or <i>rounded to 0.2 provided 0.24 has been seen in the working.</i>	1	
	[Total: 15]			

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2	(a)	<p>PDO Layout</p> <p>PDO Recording</p> <p>MMO Decisions</p> <p>MMO Quality</p>	<p>I Records at least four different balance readings and at least one mass of solid/gas <i>Accept 0.0(0X) g as the mass of the empty tube or a statement that the tube is tared.</i></p> <p>II Gives all appropriate headings and units when recording results. <i>Do not accept mass of empty tube as 0.0(00)g here unless tube is described as tared.</i> <i>(minimum of three pieces of information)</i></p> <p>III All recorded balance readings consistent to at least 1 decimal place. <i>(minimum of three balance readings)</i></p> <p>IV Evidence of reheating to “constant” mass. For balances reading to 1 d.p. two masses must be identical For 2 or 3 d.p. balances, two masses must be within 0.05 g</p> <p>V and VI Check and correct all subtractions in the results table. Calculate $\frac{\text{mass heated}}{\text{mass of residue}}$ to 3 significant figures. Compare to Supervisor standard or standard value of 1.45. Award <u>V and VI</u> for a difference up to 0.15 Award <u>V only</u> for a difference of 0.15+ to 0.30 <i>Where a candidate repeats the experiment use cumulative masses of FA 3 and residue.</i> <i>Where masses of FA 3 and residue cannot be checked, accept candidate values to calculate the ratio.</i></p>	<p>1</p> <p>1</p> <p>1</p> <p>1</p> <p>2</p>	
	(b)	<p>ACE Interpretation</p>	<p>Evaluates $\frac{\text{cand mass loss from (a)}}{\text{cand mass of FA 3}}$ correct to 2–4 significant figures. <i>Where mass loss or mass of FA 3 is not given in (a), check, from balance readings, the values.</i> <i>A candidate who incorrectly describes the mass of the residue as the mass loss in tabulated results in (a) may “correct” the error and use the correct mass loss here.</i></p>	1	

[6]

[1]

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FA 4 is $Al_2(SO_4)_3(aq)$; FA 5 is $ZnSO_4(aq)$; FA 6 is $Pb(NO_3)_2(aq)$; FA 7 is $MgSO_4(aq)$				
3	(a)	MMO Collection	1 mark for correct observations in each of the vertical columns. or 1 mark for correct observations in each of the horizontal rows (i), (ii) and (iii). 3 mark maximum Mark the section by the method which gives the better mark.	4
				[4]

test		observations			
		FA 4	FA 5	FA 6	FA 7
(i)	addition of NaOH	white ppt	white ppt	white ppt	white ppt
	further addition of NaOH	ppt soluble	ppt soluble	ppt soluble	ppt insoluble
(ii)	addition of NH_3	white ppt	white ppt	white ppt	white ppt
	further addition of NH_3	ppt insoluble	ppt soluble	ppt insoluble	ppt insoluble
(iii)	addition of KI	no ppt, no reaction, colourless or yellow solution	no ppt, no reaction, colourless or yellow solution	yellow ppt	no ppt, no reaction, colourless or yellow solution

Minimum evidence required in observations for the ion identity marks **I, II and III** in (b)

In some cases, identification may be allowed from incomplete observations. There must, however, be no observations that are contrary to those expected with any “correctly” identified ion.

The same criteria will be applied to “candidate’s supporting evidence in awarding mark **IV**. Candidates are not permitted to introduce (from the Qualitative Analysis Notes) supporting evidence that is not given in the observations. Precipitate colour need not be mentioned in supporting evidence.

Al^{3+}	(white) precipitate, soluble in (excess) NaOH, if yellow ppt with KI
Zn^{2+}	(white) precipitate, soluble in (excess) $NH_3(aq)$
Pb^{2+}	Yellow precipitate with KI
Mg^{2+}	(white) precipitate, insoluble in (excess) NaOH

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FA 4 is $\text{Al}_2(\text{SO}_4)_3(\text{aq})$; FA 5 is $\text{ZnSO}_4(\text{aq})$; FA 6 is $\text{Pb}(\text{NO}_3)_2(\text{aq})$; FA 7 is $\text{MgSO}_4(\text{aq})$				
(b)	ACE Conclusions	Do not accept any ion other than Al^{3+} , Zn^{2+} , Pb^{2+} or Mg^{2+} in any section. Marks I to III Ions must be correct, including charge, if a symbol has been given. – <u>no ecf in this section.</u>	1	
		Award <u>I only</u> if one ion only is identified from correct observations.	1	
		Award <u>I and II</u> if two ions only are identified from correct observations.	1	
		Award <u>I, II and III</u> if all four cations are identified from correct observations. <i>The 4th cation may be identified by elimination from incomplete supporting evidence.</i>	1	
		Award mark <u>IV</u> if the supporting evidence fits the ion identified and the practical performed for at least three of the four ions. Allow ecf on ion order on mark <u>IV</u> .	1	[4]
(c)	MMO Decisions	Selects sodium or potassium chromate(VI), sulfuric acid or hydrochloric acid soln containing one of the following named ions or formula given followed by (aq): CrO_4^{2-} , SO_4^{2-} , Cl^- , Br^- but not I^- , soln containing CrO_4^{2-} ions, H_2SO_4 , HCl ,		[1]

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FA 8 is CuSO ₄ (aq)				
(d)	MMO Collection	I	Records blue colour of solution fading/disappearing on adding zinc powder in (i) <i>If no reaction with Zn(s) is reported do not allow blue to light blue solution.</i>	1
		II	Records a temperature rise in (i) <i>Accept reaction is exothermic/produces heat</i>	1
		III	Records a red-brown, orange-brown, brown or black solid in (i)	1
		IV	Observes a green, lime green, fluorescent green or yellow-green solution in (ii)	1
		V	Observes solution turning blue, or blue solution in (iii) if solution green in (ii) or solution going towards blue in colour on adding water in (iii) If solution is not mentioned in (ii) or (iii) but colours are correct – award point V only .	1
				[5]
(e)	ACE Conclusions	Completes the equation: → Cu(s) + Zn ²⁺ (aq) State symbols required		1
				[1]
				[Total: 15]