

MARK SCHEME for the May/June 2010 question paper
for the guidance of teachers

9701 CHEMISTRY

9701/34

Paper 32 (Advanced Practical Skills), maximum raw mark 40

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes must be read in conjunction with the question papers and the report on the examination.

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Question 1 Round all thermometer readings to the nearest 0.5°C.

Question	Sections	Indicative material	Mark	
1 (a)	PDO Recording	(i) Presents data in single table of results – <i>to include masses of tube with solid and residue, (mass of solid), initial and final temperatures and temperature change.</i>	1	
		(ii) All columns correctly labelled with appropriate unit shown. <i>Must use solidus, brackets or describe unit fully in words. There must be at least one mass and one temperature. If units not included in column headings every entry must have the correct unit shown.</i>	1	
	MMO Collection	(iii) All balance readings recorded to 1 or more decimal places and showing same precision of measurement	1	
		(iv) All thermometer readings recorded to 0.5°C (must be at least one at 0.5°C)	1	
	MMO Decisions	(v) Two additional masses of NaHCO ₃ evenly spaced between the other readings or one or both extending the plot (Not two between same pair) (Not within 0.5 g of any other) (Not >3.0 g away from any other)	1	
	MMO Quality	(vi) and (vii) Check and correct ΔT where necessary. Compare temperature fall with that obtained by the Supervisor for FB 2 Award (vi) and (vii) for a temp fall difference of 0.0°, 0.5° or 1.0°C Award (vi) only for a difference of 1.5°C	2	
		(viii) and (ix) Check and correct ΔT where necessary. Compare temperature fall with that obtained by the Supervisor for FB 3 Award (viii) and (ix) for a temp fall difference of 0.0°, 0.5° or 1.0°C Award (viii) only for a difference of 1.5°C	2	[9]

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Question	Sections	Indicative material	Mark	
(b)	PDO Layout	(i) Temperature (fall) plotted on <i>y-axis</i> against mass of sodium hydrogencarbonate on <i>x-axis</i> . Clearly labelled axes (ignore units unless $T/\delta T/\theta$ or m as labels) Award even if final temp plotted	1	
		(ii) Uniform and sensible scales that allow points to be plotted in at least half of the squares on each axis. (6 × 4 big squares)	1	
		(iii) check the “sweep” that points plotted for all experiments recorded. Check the plotting of points for FB 1 , FB 2 and FB 3 (and any other “suspect” point). <i>Points should be within ½ of a small square, in the correct half of the small square.</i> Not awarded if final T is plotted Not awarded if only FB 1 , FB 2 and FB 3 are recorded	1	
		(iv) Straight line drawn, passing within ½ small square of the origin and close to the majority of the points	1	[4]
(c)	PDO Layout	Explains that the mass of sodium hydrogencarbonate is the independent (controlled) variable or Temperature change is the dependent variable (<i>or words to that effect</i>)	1	[1]
(d)	ACE Conclusion	Gives correct equation for the reaction $\text{NaHCO}_3 + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O} + \text{CO}_2$ (Not H_2CO_3) (<i>Ignore state symbols</i>)	1	[1]
(e)	PDO Display	Construction on graph for determining the gradient clearly shown (<i>Must span at least 3 large squares in each direction</i>) and working shown for calculation (<i>could be found on graph page</i>)	1	
	ACE Interpretation	Reads intercepts or selects two points <u>on the line</u> to within ½ small square in either direction.	1	
		Evaluates $\Delta y/\Delta x$ from candidate's intercept figures to find gradient correctly to sf shown	1	[3]
(f)	ACE Interpretation	Multiplies answer to (e) by 84	1	[1]
(g)	ACE Interpretation	Some use of $30 \times 4.3 \times$ (answer to (f))	1	
	PDO Display	Answer, in kJ mol^{-1} , correct to 2 or 3 significant figures and showing +ve sign <i>Correct answer is given by $0.129 \times$ answer to (f)</i> (No ecf from first part)	1	[2]

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Question 2

Question	Sections	Indicative material	Mark	
FB 7 is $(\text{NH}_4)_2\text{SO}_4 \cdot \text{FeSO}_4(\text{aq})$; FB 8 is $\text{NH}_4\text{Cl}(\text{aq})$				
2 (a)	MMO Decisions	Chooses barium chloride (or nitrate)/ $\text{Ba}^{2+}(\text{aq})$ /solution containing Ba^{2+} (ions) followed by a <u>specified</u> mineral acid other than sulfuric acid or vice versa. (Allow if acid name specified in (b)) Do not accept lead nitrate/lead ions	1	[1]
(b)	PDO Recording	Results for two solutions and two reagents (ecf from (a)) recorded in a single table with distinct separation between reagents (no repetition of solutions or reagents)	1	[2]
	MMO Collection	Give one mark for correct observations with FB 7 and FB 8 Expected results: FB 7 – white ppt with Ba^{2+} , insoluble in acid (<i>do not award if H_2SO_4 added</i>) FB 8 – No reaction or no precipitate If Pb^{2+} used then white ppt with both or obs as for Ba^{2+} <i>Do not award if any obs are dashes except for FB 8 adding acid after Ba^{2+}</i>	1	
(c)	ACE Conclusion	Mark consequentially on observations in (b) Expected conclusion Identifies FB 7 as solution containing SO_4^{2-} from “white ppt with Ba^{2+} , insoluble in acid given in evidence <i>ecf allowed here. Allow deduction if H_2SO_4 has been added after Ba^{2+} (not with Pb^{2+})</i>	1	[1]
(d)	MMO Collection	One mark for two correct observations with $\text{NaOH}(\text{aq})$ – before heating FB 7 – green/dirty green/muddy green ppt FB 8 – no ppt/no change/no reaction The mark from (e) may be awarded here if the green ppt with FB 7 is recorded as <u>turning brown</u>	1	[3]
		One mark for correct observations with $\text{NaOH}(\text{aq})$ – after heating gas evolved turns red litmus paper blue for both FB 7 and FB 8	1	
		One mark for correct observations with $\text{NH}_3(\text{aq})$ FB 7 – green ppt (as above) FB 8 – no ppt/no change/no reaction	1	
(e)	MMO Collection	Records brown precipitate/residue provided <u>green ppt</u> in first box of (d)	1	[1]

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Question	Sections	Indicative material	Mark	
(f)	ACE Conclusions	Mark consequentially on observations in (d) and (e) Expected conclusions: (i) Common cation is ammonium/ NH_4^+ , from evolution of ammonia or alkaline gas (minimum) or clear ref back to relevant obs (ii) Second cation is Fe^{2+} in FB 7 from green ppt with NaOH and NH_3 or green ppt, turning brown, with either. If Cr^{3+} then grey-green ppt with both alkalis or grey-green ppt with NaOH soluble in excess giving (dark) green solution or grey-green ppt insol in excess NH_3 (no con obs)	1	[1]
FB 9 is a ketone; FB 10 is a primary alcohol; FB 11 is a secondary alcohol; FB 12 is an aldehyde				
(g)	MMO Collection	One mark for correct observations with FB 9 and FB 12 One mark for two correct observations with FB 10 One mark for two correct observations with FB 11 See table below for expected observations	1 1 1	[3]

reagent	observations			
	FA 9	FA 10	FA 11	FA 12
acidified dichromate	no reaction/no change/yellow or orange soln (not no ppt*)	(colour change to) green/blue-green/cyan/turquoise (solution not ppt*)	(colour change to) green/blue-green/cyan/turquoise (solution not ppt*)	*penalise ppt error once in this row If FB9 colour wrong penalise 1 st of additional ppts
2,4-DNPH	yellow ppt		no reaction/ no change	yellow ppt
Tollens' reagent		no reaction/ no change/no ppt	no reaction/ no change/no ppt	silver mirror or black/grey solution or ppt

(h)	ACE Conclusions	Mark consequentially on observations in (g) FB 10 and FB 11 contain the alcohols from both oxidised by dichromate(VI) and give no ppt with Tollens' FB 9 contains the ketone from No reaction with dichromate ($\text{Cr}_2\text{O}_7^{2-}$ obs all correct) or no reaction with $\text{Cr}_2\text{O}_7^{2-}$ and yellow ppt with 2,4-DNPH. (If FB 12 selected evidence must have yellow ppt with 2,4-DNPH and no reaction with Tollens')	1 1	[2]
	Total			[14]