UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS

GCE Advanced Level

MARK SCHEME for the November 2004 question paper

9701 CHEMISTRY

9701/04 **Paper 4 (Structured Questions A2 Core), maximum raw mark 60**

This mark scheme is published as an aid to teachers and students, to indicate the requirements of the examination. It shows the basis on which Examiners were initially instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began. Any substantial changes to the mark scheme that arose from these discussions will be recorded in the published *Report on the Examination*.

All Examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes must be read in conjunction with the question papers and the *Report on the Examination*.

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CIE is publishing the mark schemes for the November 2004 question papers for most IGCSE and GCE Advanced Level syllabuses.

Grade thresholds taken for Syllabus 9701 (Chemistry) in the November 2004 examination.

	maximum mark available	minimum mark required for grade:		
		A	B	E
Component 4	60	44	39	22

The thresholds (minimum marks) for Grades C and D are normally set by dividing the mark range between the B and the E thresholds into three. For example, if the difference between the B and the E threshold is 24 marks, the C threshold is set 8 marks below the B threshold and the D threshold is set another 8 marks down. If dividing the interval by three results in a fraction of a mark, then the threshold is normally rounded down.

November 2004

GCE A LEVEL

MARK SCHEME
MAXIMUM MARK: 60
SYLLABUS/COMPONENT: 9701/04 **CHEMISTRY** **Paper 4 (Structured Questions A2 Core)**

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1 **(a)** **(i)** strong, because final pH is about 14 **[1]**

(ii) (pH = 0.70) ⇒ $[H^+] = 10^{-0.7}$ = 0.20 (mol dm^{-3}) **[1]**

∴ $[H_2SO_4]$ = (0.10 mol dm^{-3}) ecf **[1]**

(iii) (end point is at 34.0 cm^3 ($\pm$ 0.5 cm^3), so)

amount of H^+ used = 0.2 x 25/1000 = 0.0050 mol ecf from **(ii)** **[1]**

moles of guanidine = moles of H^+ = 0.0050 mol

[guanidine] = 0.005 x 1000/34.0 = **0.147** (mol dm^{-3}) **[1]**

allow range: 0.145 – 0.149 ecf in 0.005 or 34.0

(iv) M_r = 8.68/0.147 = **59** (*allow range* 58 – 60) ecf from **(iii)** **[1]** **6**

(b) **(i)** $\longrightarrow$ **7** $CaSO_4$ + **3** $Ca(H_2PO_4)_2$ + **2** HF **[1]**

(ii) M_r values: $Ca(H_2PO_4)_2$ = 234.1, H_2SO_4 = 98.0 **[1]**

234.1 x 3 = **702.3** 98 x 7 = **686** both **[1]**

ecf from ratios in equation, and from M_r values

∴ mass of H_2SO_4 needed = 1.0 x 686/702.3 = **0.98** kg **[1]**

(correct answer = [3] marks. accurate value is: 0.977 kg.
Allow ecf from incorrect M_r or incorrect multipliers) **4**

(c) **(i)** A solution that **resists** changes in pH [NOT: results in **no** pH change] **[1]**

when **small amounts** of H^+ or OH^- are added **[1]**

(ii) pH = $-\log_{10}(6.3 \times 10^{-8}) + \log_{10}(0.1/0.2)$ = **6.9** **[1]**

or $[H^+] = (6.3 \times 10^{-8}) \times 0.2/0.1 = 1.26 \times 10^{-7}$

∴ pH = $-\log_{10}(1.26 \times 10^{-7})$ = **6.9** **3**

Total 13

2 **(a)** $O_2 + 4H^+ + 4e^- \longrightarrow 2H_2O$ (or equation ÷ 2) **[1]** **1**

(b) ⊕ **[1]** **1**

(c) 1.23 (V) (ignore sign) **[1]** **1**

(d) a better/larger salt bridge *or* a diaphragm *or* larger (area of) electrodes

or increase concentrations/pressure **[1]** **1**

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(e) time = 400 x 24 x 60 x 60 = 34 560 000 seconds **[1]**

charge = current x time = 0.01 x 34 560 000 = 345 600 C ecf **[1]**

moles of H = 345 600/96 500 = 3.6 mol ∴ mass of H = **3.6 g** ecf **[1]** **3**

(f) advantages: less pollution/CO_2/NO_x etc. *or* cleaner by-products

less dependence on fossil fuels/finite resources *any one* **[1]**

disadvantages: more expensive (to develop *or* to run)

takes up more space

poor power-to-volume ratio

hydrogen is difficult to store or to transport *any one* **[1]**

NOT hydrogen is explosive/flammable **2**

Total 9

3 solubilities **decrease** down the group **[1]**

hydration energy of the cation decreases **[1]**

lattice energy stays the same, or decreases less than H.E. **[1]**

making $\Delta H_{solution}$ more endothermic *or* H.E. no longer able to overcome -L.E. **[1]** **4**

Total 4

4 **(a)** an element forming one or more ions with a partially filled/incomplete d-shell **[1]** **1**

(b) **(i)** almost no change (allow *slight* increase or *slight* decrease) **[1]**

(ii) density should increase **[1]**

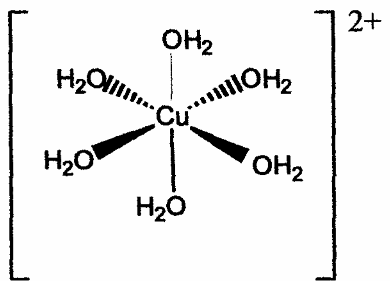
because A_r is increasing but size/volume/radius stays the same **[1]**

(allow partial ecf from **b (i)**) **3**

(c) ………..$3d^9$ **[1]** **1**

(d) **(i)** an ion formed when a ligand (datively) bonds to a (central metal) cation **[1]**

(ii)



[1] **2**

(e) (i) dark/deep/navy/royal/Oxford blue *or* purple [NOT Prussian blue or lilac or mauve] **[1]**

(ii) $4NH_3 + [Cu(H_2O)_6]^{2+} \longrightarrow [Cu(NH_3)_4(H_2O)_2]^{2+} + 4H_2O$

or $\longrightarrow [Cu(NH_3)_4]^{2+} + 6H_2O$ **[1]** **2**

(f) $CuCl_4^{2-}$ is produced **[1]**

the equilibrium is **reversible** *or* ⇌ in equation **[1]**

Cl^- ligands **replace/exchange** with H_2O ligands (in words) **[1]**

(the following equation is worth the first two marks)

$[Cu(H_2O)6]^{2+} + 4Cl^- \rightleftharpoons [CuCl_4]^{2-} + 6H_2O$ **3**

Total 12

5 (a) (i) $AlCl_3$/$FeCl_3$/Al/Fe/I_2 (+ heat) [aq negates] (N.B. NOT $AlBr_3$ etc.) **[1]**

(or names)

(ii) (sun)light/hf/UV (aq negates) **[1]** **2**

(b) $SOCl_2$/PCl_3/PCl_5 [aq negates] **[1]**

(or names) **1**

(c) (i) C > B > A (i.e. a mark in the penultimate box) **[1]**

(ii) (*acyl chloride fastest*) highly δ+ carbon atom joined to 2 electronegative atoms

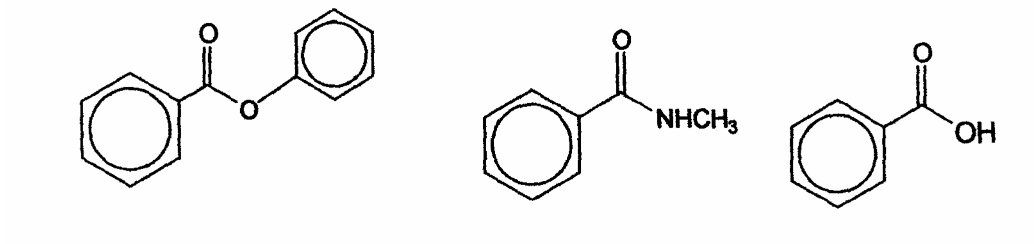
or addition-elimination mechanism is possible **[1]**

(*aryl chloride slowest*) delocalisation of lone pair over ring ⇒ stronger C-Cl bond

or impossibility of 'backside' attack on the C-Cl bond **[1]** **3**

(d) C_6H_5-$CO_2C_6H_5$ **[1]** C_6H_5-$CONHCH_3$ **[1]** C_6H_5-CO_2H **[1]**

OR



3

Total 9

6 (a) (i) E **[1]**

(ii) $CH_3CH_2CH_2CO_2^-(Na^+)$ [NOT C_3H_7COO-Na *or* C_3H_7COOH] **[1]**

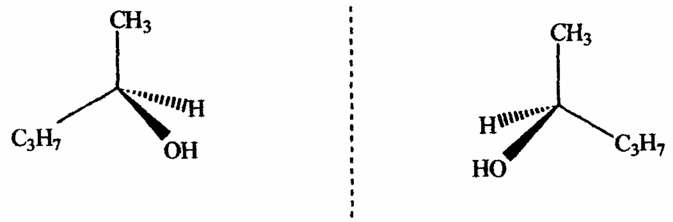
[but allow $CH_3CH_2CH_2CO_2Na$]

CHI_3 *or* name **[1]** **3**

(b) the alcohol from **E** has four different groups around a carbon atom **[1]**

∴ it is chiral/asymmetric *or* it is produced as a 50:50 mixture of mirror images **[1]**

or its mirror images are non-superimposable



formulae: **[1]**

the alcohol from **D** has 2 identical groups on its central carbon atom **[1]**

4 max 3

Total 6

7 (a) orange colour disappears/bromine is decolourised (NOT discoloured, *or* goes clear) **[1]**

(white) precipitate/solid/crystals is formed **[1]** **2**

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(b) e.g. add neutral $FeCl_3$ (aq) – violet colour with phenol

or add universal indicator – red/orange colour with phenol

or add Na metal – fizzing/H_2 evolved with phenol

or add NaOH(aq) to the pure compound – phenol would dissolve

or add H^+ (aq) to the pure compound – phenylamine would dissolve

or add HNO_2 at room temperature – phenylamine would produce gaseous N_2.

or add HNO_2 at 5 °C, followed by an alkaline solution of phenol – phenylamine would produce a coloured (orange) dye **[1]** **1**

(c) IV $KMnO_4$ + heat **[1]**

V $HNO_3 + H_2SO_4$ **[1]** (both) conc^d and at 50 °C < T < 60 °C **[1]**

VI Sn + HC*l* (NOT LiA*l*H_4) **[1]** **4**

Total 7

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