



GEORG-AUGUST-UNIVERSITÄT
GÖTTINGEN

Einführung ins wissenschaftliche Arbeiten

**Teilchensimulationen von
Polymermischungen in begrenzten
Geometrien mit zeitabhängigen
Randbedingungen**

**Particle simulations of polymer mixtures in
confined geometries with time dependent
boundary conditions**

prepared by

Justus Multhaup

from Boffzen

at the Institut für Theoretische Physik

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First Rreferee: Prof. Dr. Marcus Müller

Second referee: Prof. Dr. Stefan Klumpp

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Nomenclature

Lateinische Buchstaben

Variable	Bedeutung	Einheit
A	Querschnittsfläche	m^2
c	Geschwindigkeit	m/s

Griechische Buchstaben

Variable	Bedeutung	Einheit
α	Winkel	$^\circ$; $-$
ϱ	Dichte	kg/m^3

Indizes

Index	Bedeutung
m	Meridian
r	Radial

Abkürzungen

Abkürzung	Bedeutung
2D	zweidimensional
3D	dreidimensional

Nomenclature

Abk"urzung	Bedeutung
max	maximal

1. Introduction

2. Theory

2.1. Polymeric mixtures

Polymer mixtures consist of two or more chemically different polymer types. The mechanical and thermodynamic properties can vary greatly with several factors such as composition, molecular weight and interactions between the polymers. This makes them desirable for manufacturing materials with tailored properties.

If the composition is uniform everywhere, then the mixture is called homogeneous. In this case the properties do not change throughout the mixture. In a heterogeneous mixture, in contrast, the composition is non-uniform, leading to visible boundaries which may have very different properties. This phenomenon is also called macro-phaseseparation. From an entropic viewpoint, mixing is always favored. However, energetic interactions between polymers can either favor or suppress mixing. Whether a mixture is homogeneous or heterogeneous therefore depends on the balance between entropy and energy [5, S. 137].

2.1.1. Flory Huggins Theory

Whether mixing or phase-separation will be favored can be predicted by determining the free energy change associated with mixing the components. This free energy change can be computed within the lattice model developed by Flory and Huggins [4]. Within the Flory-Huggins framework, no volume change is assumed upon mixing. With this assumption, it is convenient to represent the system on a lattice. The lattice site volume v_0 corresponds to the smallest molecular unit and every macromolecule takes up one or multiple lattice sites. Consider a binary mixture with n_A polymers of species A and chain length N_A and n_B polymers of species B and chain length N_B . The free energy of mixing per lattice site ΔF_{mix} is then given by the Flory-Huggins equation of polymer solutions [5, S. 143]:

2. Theory

$$\frac{\Delta F_{mix}}{k_B T} = \frac{\phi}{N_A} \ln \phi + \frac{1-\phi}{N_B} \ln(1-\phi) + \chi \phi(1-\phi). \quad (2.1)$$

Here, $\phi = \frac{n_A N_A}{n_A N_A + n_B N_B}$ is the monomer fraction of species A, k_B is the Boltzmann constant, T is the system temperature and χ is the Flory interaction parameter which characterizes the interaction between different polymer species and can be obtained from experiments. A positive value of χ opposes mixing while a negative value promotes it, knowing the value of χ therefore allows a qualitative prediction of the phase separation behaviour. Note that so far, no space dependence of ϕ has been assumed. To fully capture the complexity of the system, the Flory-Huggins model has to be extended to include spatial variations of ϕ , which gives rise to a free energy functional:

$$\frac{F[\phi]}{k_B T} = \rho_0 \int d^3 \mathbf{r} \left\{ \frac{\phi}{N_A} \ln \phi + \frac{1-\phi}{N_B} \ln(1-\phi) + \chi \phi(1-\phi) \right\}. \quad (2.2)$$

Here, ρ_0 is the average monomer density in the system and. In the more sophisticated de Gennes-Flory-Huggins-model, an additional term proportional to $[\nabla \phi]^2$ is added to the free energy density to ensure that unphysical, sharp changes in the local densities are penalized [2].

The chemical potential is given by the functional derivative $\mu = \frac{\delta F}{\delta \phi}$, which with (2.2) gives:

$$\frac{\mu}{\rho_0 k_B T} = \frac{\ln \phi}{N} - \frac{\ln(1-\phi)}{N} + \chi(1-2\phi). \quad (2.3)$$

2.2. Collective diffusion

Consider a binary mixture of polymers with $N_A = N_B = N$. Since the number of monomers in the system is constant, the continuity equation holds:

$$\frac{\partial \phi}{\partial t} + \nabla \mathbf{J} = 0. \quad (2.4)$$

Here, \mathbf{J} is the local current of species A. Near equilibrium, one postulates a linear

relation between \mathbf{J} and the local chemical potential difference μ [2]:

$$\mathbf{J}(\mathbf{r}) = - \int_V \frac{\Lambda(\mathbf{r}, \mathbf{r}')}{k_B T} \nabla' \mu(\mathbf{r}') d\mathbf{r}' . \quad (2.5)$$

The Onsager coefficient $\Lambda(\mathbf{r}, \mathbf{r}')$ relates the force acting on a monomer at position \mathbf{r}' due to the gradient of chemical potential to the density at position \mathbf{r} . It therefore determines the dynamics of the system. For an incompressible symmetric mixture with $N_A = N_B = N$, in the simplest case, the current is coupled only locally to the chemical potential gradient. The Onsager coefficient then takes the simple form [1, 2]

$$\Lambda = DN\phi(\mathbf{r}')(1 - \phi(\mathbf{r}'))\delta(\mathbf{r} - \mathbf{r}') . \quad (2.6)$$

Here, D is the single chain diffusion coefficient.

3. Simulation technique

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4. Collective diffusion of symmetric homopolymers

4.1. Reference system

In this section, the collective diffusion properties of noninteracting homopolymers with $N_A = N_B = N$ and $\chi = 0$ are investigated. As a reference system, a simulation box with 10000 polymers and dimensions $L_x \times L_y \times L_z = 9.25 \times 3 \times 3 R_e^3$ is considered with spatial discretization of $\Delta L = 0.125 R_e$, so the invariant degree of polymerisation is $\sqrt{N} \approx 120$. Periodic boundary conditions are applied in the lateral y and z directions, whereas impassable walls are applied in the x direction. Initially, the polymers are distributed homogeneously in the system. To stimulate diffusion, conversion zones are introduced close to the walls at $x < 0.25 R_e$ and $x > 9 R_e$. If the center of mass coordinate \mathbf{r}_{cm} of a polymer of type A lies in the conversion zone at $x < 0.25 R_e$, it is converted to type B with probability $p(A \rightarrow B) = r\phi(\mathbf{r}_{cm})$ [3]. Analogously, conversion from B to A takes place in the conversion zone at $x > 9 R_e$ at the same rate r . The total currents \mathbf{J}_A and \mathbf{J}_B are measured by tracking the number of polymer-conversions.

4.2. Collective diffusion coefficient

From (2.3), (2.5) and (2.6), the local monomer current of type A becomes:

$$\mathbf{J}_A = -D\rho_0\nabla\phi. \quad (4.1)$$

Together with (2.4), this gives the well-known diffusion equation:

$$\frac{\partial\phi(\mathbf{r}, t)}{\partial t} - D\rho_0\Delta\phi = 0. \quad (4.2)$$

4. Collective diffusion of symmetric homopolymers

In the steady state, this simply yields $\Delta\phi = 0$. Due to the periodic boundary conditions in the lateral directions, the current only has an x component. Therefore, a one dimensional, linear density profile is obtained in the steady state. The slope is obtained from (4.1) and the intersection is obtained from the condition that $\phi(L_x/2) = 0.5$ since the density is normalized.

5. Discussion

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6. Summary

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A. erster Anhang

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Danksagung

Dank...

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(Justus Multhaup)