

# Chapter 1

## Introduction

### 1.1 Background

Radicals are chemical species which tend to be highly reactive due to the presence of one or more unpaired electrons. Living systems depend on radical processes as part of normal metabolism<sup>1</sup> but biomaterials, such as proteins, are susceptible to radical-induced damage. Radical-induced oxidation of biomaterials has been implicated in a number of degenerative disease states, including cancer, Alzheimer's disease, Parkinson's disease, and multiple sclerosis.<sup>2-5</sup>

In biological systems, radicals are derived from many sources. Exogenous sources include solar radiation and air pollutants, while endogenous sources include *in vivo* transition metal-ion redox processes, such as the electron transport chain involved in cellular respiration.<sup>6</sup> Some redox centres in the electron transport chain may uncontrollably transfer an electron to molecular oxygen, forming the superoxide anion ( $\text{O}_2^{\bullet-}$ ). Superoxide is not a strong oxidant on its own, however, it may disproportionate spontaneously, or catalytically through metalloenzymes such as superoxide dismutase, leading to the formation of highly reactive oxygen-centred radicals. In fact, most oxygen-centred radicals derive from similar reactions of  $\text{O}_2$  with redox-active

metals.<sup>1</sup>

Oxygen-centred radicals, known as reactive oxygen species (ROSs) in biology, are particularly important and common due to the nature of aerobic respiration. The ROSs which are of primary concern are the highly reactive hydroxyl radicals ( $\text{HO}^\bullet$ ), alkoxyl radicals ( $\text{RO}^\bullet$ ), superoxide ( $\text{HOO}^\bullet/\text{O}_2^{\bullet-}$ ), and peroxy radicals ( $\text{ROO}^\bullet$ ).<sup>1</sup> The oxidation of proteins by ROSs occurs through a radical chain mechanism which has been studied experimentally in detail.<sup>7,8</sup> This chain reaction occurs when an ROS initiates a radical chain reaction through hydrogen atom transfer (HAT), electron transfer, or addition reactions, leading to rapid propagation. HAT is the most relevant reaction, and thus is the focus of my work. In particular, this work is primarily concerned with developing an understanding of the fundamental chemistry involved in protein oxidation.

Proteins are the most abundant biomaterial in most biological systems,<sup>9</sup> thus understanding their degradation is essential to understanding degenerative disease. Because proteins are composed of as many as 20 common amino acid side-chains, as well as the common peptide backbone, there are a large number of possible reactions. Some of the reactions involved in protein oxidation are shown in Scheme 1.1.

Initial abstraction (Reaction **a**) often occurs at the  $\alpha$ -carbon position ( $\alpha\text{-CH}$ ), forming a carbon-centred radical ( $\alpha\text{-C}^\bullet$ ) which is partially delocalised in the  $\pi$ -system of the neighbouring amide and carbonyl groups. Many studies have indicated that the stability of  $\alpha\text{-C}^\bullet$  is determined by stereo-electronic considerations related to the planarity of the amide group. As such, steric bulk in the side-chains, as well as local protein structure



(helix, sheet, etc.) can constrain radical geometries. For example, the most stable  $\alpha$ -carbon radicals occur at glycine residues in antiparallel  $\beta$ -sheets, whereas other bulkier residues and secondary structures lead to loss of captodative stabilisation.<sup>10</sup> Amino acid side-chains are also susceptible to oxidation. Those side-chains containing sulphur,<sup>11</sup> as well as tyrosine (which has a fairly weak phenolic O-H bond of about 89 kcal mol<sup>-1</sup>),<sup>12</sup> are particularly susceptible to oxidation.

Propagation of the radical chain reaction occurs through various processes. In the presence of molecular oxygen, rapid addition occurs at the newly formed  $\alpha$ -C $\cdot$  (Reaction **b**), generating a peroxy radical, which can carry forward through further HAT reactions (Reaction **c**).<sup>13</sup> The mechanism involved in the radical mediated oxidation of proteins has been studied experimentally using techniques involving ionising radiation.<sup>14,15</sup> The course of this process is complexly dependent on the availability of either singlet oxygen ( $^1\text{O}_2$ ), or superoxide ( $\text{O}_2^{\cdot-}$ ) or the protonated form, peroxy radical ( $\cdot\text{OOH}$ ). A detailed analysis of this process is outside the scope of this thesis, but ultimately, these reactions lead to the generation of a hydroxyl-amide (Reaction **d**).

The radical chain reaction can be terminated through several mechanisms, including protein-protein cross-linking (Reaction **e**), protein fragmentation (Reaction **f**), or reactions with antioxidants ( $\text{A-H}$ , Reaction **g**). The sum total of all these processes contribute to the accumulation of oxidised proteins which are associated with many degenerative diseases.<sup>16</sup> HAT reactions which are important steps in the initiation, propagation, and termination reactions of protein oxidation are investigated in this thesis.

## 1.2 Studying HAT reactions

Developing an understanding of protein and other biomolecular oxidation requires an understanding of the deceptively simple HAT reactions involved. Formal HAT reactions are a fundamental radical chemical transformation which have been studied for more than a century.<sup>17,18</sup> From an experimental perspective, HAT reactions which involve oxygen-centred radicals and non-radical organic substrates, are reasonably well characterised: the effects of different solvents are well understood.<sup>19</sup> However, the main challenge faced by many experiments is elucidating the mechanistic details of a reaction. This is a problem that can be examined by quantum chemistry, which is the approach that I shall take. Background on the theory used in this thesis is given in Chapter 2.

In order to investigate HAT reactions, we need to consider the mechanism in detail. For a simple HAT reaction, there exists several possible mechanisms by which this transformation can occur. The two most common concerted mechanisms are direct HAT<sup>i</sup> and proton-coupled electron transfer (PCET). At the basic level, direct HAT involves the transfer of an electron and proton through the same set of acceptor/donor orbitals, while PCET involves the transfer of an electron and proton through different sets of orbitals. In practise, this distinction is poorly described, and this topic is still in active discussion in the literature.<sup>20–30</sup>

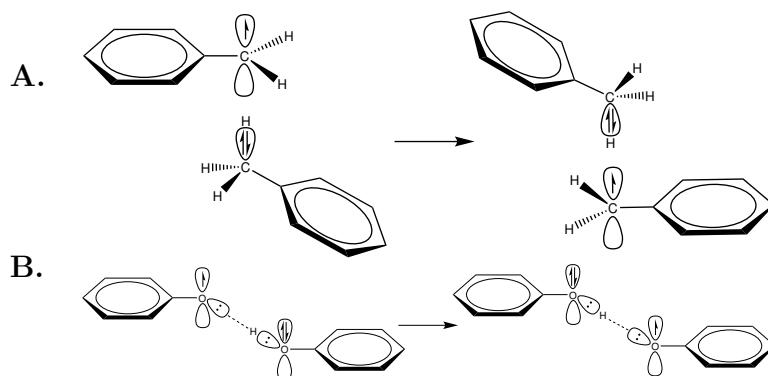
The quintessential example when comparing direct HAT to PCET comes

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<sup>i</sup>Not to be confused with the net reaction of formal hydrogen atom transfer. The abbreviation HAT will be used interchangeably, although the distinction should be clear from context.

## 1.2. Studying HAT reactions

from the work of Mayer et al.,<sup>21</sup> which describes the self-exchange reactions of benzyl-toluene and phenoxyl-phenol, shown in Scheme 1.2. In this work, the transition state (TS) structures were obtained through theoretical studies. These complexes are oriented so that the aromatic rings are trans relative to one another. In this geometry, the benzyl-toluene pair undergoes direct HAT, with the  $2p - \pi$  orbital of the benzylic carbon radical oriented at the benzylic hydrogen on toluene, with little delocalisation of the radical into the  $\pi$ -system. This is described as direct HAT, as the orbital containing the radical overlaps with the X-H  $\sigma^*$  anti-bonding orbital, and thus the electron and proton are transferred through the same set of orbitals.



Scheme 1.2: Self-exchange reactions of the **A.** benzyl-toluene couple through direct HAT **B.** phenoxyl-phenol couple through PCET.

For the phenoxyl-phenol pair, a fairly strongly hydrogen bonded pre-reaction complex is first formed with a binding enthalpy  $\Delta H$  of  $-8.1 \text{ kcal mol}^{-1}$ . As a result of this strong interaction, the TS structure is such that the phenoxyl radical occupies a  $2p$  orbital, and therefore, cannot overlap with

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the H-O  $\sigma^*$  anti-bonding orbital in order to undergo direct HAT. However, this geometry does allow the orbital containing the radical to overlap with the  $2p$  lone pair of and the phenol moiety, and the conjugated aromatic  $\pi$  systems. As a result, the TS complex singly occupied molecular orbital (SOMO) is of  $\pi$ -symmetry and highly delocalised, therefore, HAT occurs through a PCET mechanism. The reaction has an enthalpic barrier height  $\Delta H^\ddagger$  of 5.0 kcal mol<sup>-1</sup> relative to the hydrogen bonded complex, so that the barrier is 3.1 kcal mol<sup>-1</sup> below the separated reactants.

The work by Mayer et al.<sup>21</sup> suggests that hydrogen bonding is a necessary, but not sufficient, condition for PCET to occur. This then implies that PCET is not possible between molecules which do not possess hydrogen bonding moieties, such as carbon atoms. Work by other authors has shown this to be untrue.<sup>24,31</sup> In particular, DiLabio and Johnson<sup>24</sup> demonstrated that this neglected the important contributions of  $\pi - \pi$  interactions and lone pair- $\pi$  interactions. Additional calculations revealed the existence of a TS structure for the benzyl-toluene couple which is 3.7 kcal mol<sup>-1</sup> lower in energy than previously reported. This structure orients the aromatic rings 34° relative to one another, allowing for optimal  $\pi - \pi$  overlap.

Analysis of the TS highest-occupied molecular orbital (HOMO) reveals bonding character between the two  $\pi$ -system, while the SOMO shows anti-bonding character between the  $\pi$ -systems, as well as both C-H bonds. Thus, there exists a net partial bonding interaction between the two  $\pi$ -systems, opening up an additional electronic channel for PCET to occur. DiLabio and Johnson also suggest that the phenol-phenoxyl couple likely prefers a  $\pi$ -stacked TS structure, and compare this to a structural analogue, a nat-

urally occurring tyrosyl-tyrosine couple, which they demonstrated proceeds through a PCET mechanism. Additional work by Muñoz-Rugeles et al.<sup>30</sup> confirmed the existence of a  $\pi$ -stacked TS structure for the phenol-phenoxyl couple. They used an approach which utilises natural population analysis along the intrinsic reaction coordinate, and demonstrated that both the benzyl-toluene couple and phenoxyl-phenol couple favour a  $\pi$ -stacked TS structure and undergo HAT through a PCET mechanism. Interestingly, they also showed that reaction barrier heights for the PCET mechanism are systematically lower than those for related direct HAT mechanism.

Bearing in mind there is not an obvious way to explore the related differences in mechanism experimentally, computational examination of formal HAT reactions enables analysis of the mechanism of these reactions. Through careful investigation, a general distinction between a direct HAT mechanism and PCET mechanism can be achieved. In doing so, important insight is gained from understanding the electronic behaviour of these reactions. In this vein, the investigation of the physico-chemical nature of HAT reactions shall be the central theme of this thesis.

Consider for a moment the potential energy surface (PES) for an arbitrary chemical reaction, which is a complex hypersurface that depends on many variables. Theoretical methods can be used to generate a full PES, however, this quickly becomes computationally infeasible as the number of atoms in a system increases. Typically this problem can be simplified by examining only the relevant degrees of freedom. Often, the two most important coordinates can be isolated, giving a three-dimensional PES. Furthermore, in chemistry we often simplify this problem to two-dimensions using the so-



called intrinsic reaction coordinate, which is the lowest energy cross section of a higher dimension PES. This yields a reaction coordinate diagram, as is illustrated below in Figure 1.1.

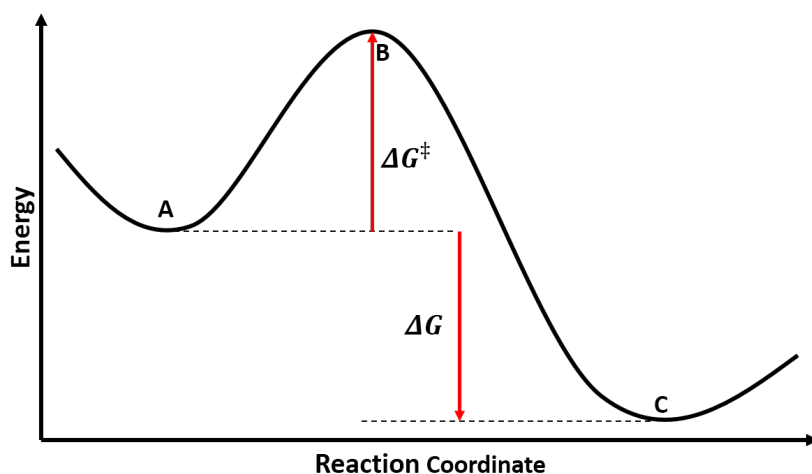


Figure 1.1: A typical reaction coordinate diagram. Stationary states A, B, and C, correspond to the pre-reaction, transition state, and post-reaction complexes, respectively.

In a typical reaction coordinate diagram, the reactants begin to interact and form a pre-reaction complex (A). Given sufficient energy, the reaction will proceed over the top of the energy barrier through a transition state (TS) complex (B). After the chemical transformation is completed, a post-reaction complex (C) is formed until the products are able to separate. This is a somewhat naïve description, as it only broadly describes a chemical transformation. In particular, the roles of substrate-radical and substrate-radical-medium interactions along the reaction coordinate are neglected. This is in fact a key point, as a thorough understanding of these interactions continues to be a hole in the literature.

Consequently, recent work from our group, in collaboration with our experimental colleagues at the University of Rome Tor Vergata, has focused on the importance of substrate-radical interactions in determining the kinetics of HAT reactions. Specifically, it has been shown that the three-dimensional structures of oxygen-centred radicals, as well as the organic substrates, impact the nature of the interactions involved in HAT reaction pathways.<sup>32</sup> In our work, we utilise primarily the benzyloxyl ( $\text{BnO}^\bullet$ ) and cumyloxyl ( $\text{CumO}^\bullet$ ) radicals, which serve as a convenient proxy to biological oxygen-centred radicals. Reactions involving  $\text{BnO}^\bullet$  and  $\text{CumO}^\bullet$  can be monitored using time resolved laser flash photolysis (LFP) techniques. A combination of theoretical and experimental techniques have been used to examine reactions involving  $\text{BnO}^\bullet$  and  $\text{CumO}^\bullet$  with a variety of organic substrates. A detailed discussion of these results shall be reserved for following the chapters, however, a great deal of insight has been gained into the role of structure in both the radicals and substrates, and resulting intermolecular interactions.

## 1.3 Research goals

With respect to the work in this thesis, in Chapter 3 the importance of the left-hand side of Figure 1.1 shall be examined by studying how the pre-reaction complex impacts HAT reactions. There has been no detailed investigation of the importance of pre-reaction complex formation for HAT reaction. This is of particular importance, as oxygen-centred radicals can hydrogen bond with substrates as both acceptors and donors.<sup>33</sup> These hydro-

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gen bonding interactions, in addition to the other non-covalent interactions between the radical and substrate, lead to the formation of a pre-reaction complex. Accordingly, the formation of a pre-reaction complex is a fundamental step in protein oxidation.

The specific aim of Chapter 3 is to investigate the effects of non-covalent binding in the pre-reaction complex, with respect to the well known, but phenomenological Arrhenius equation. As of yet, there is no framework which relates the non-covalently bound pre-reaction complex to kinetic results. I ask the simple question: does there exist a direct correlation between the Arrhenius pre-factor and the non-covalent binding energy? To address this question, we examine the non-covalent binding in the pre-reaction complex in a series of related HAT reactions. Arrhenius parameters for the systems of interest in this work were previously tabulated,<sup>34</sup> and consist of thermoneutral or nearly thermoneutral reactions involving the creation and destruction of oxygen-centred radicals. These reactions are related to the phenol-phenoxyl self-exchange reaction, where a relatively strong pre-reaction complex is expected.

Then in Chapter 4, the right-hand side of Figure 1.1 is considered, where the effects of bond strengths on HAT rate constants are examined. Bond strengths are indicated by bond dissociation enthalpies (BDEs), and are central to the understanding of reactions with respect to thermodynamics. In addition to this, there exists a tremendous amount of literature in which BDEs are linked to chemical reactivity, especially for HAT reactions.<sup>17,23,35–37</sup> There exists a linear free energy relationship (LFER) called the Bell-Evans-Polanyi (BEP) principle,<sup>38,39</sup> which states that the differ-

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ence in activation energy ( $E_a$ ) for two related reactions is proportional to the differences in reaction enthalpy ( $\Delta H$ ):

$$E_a = E_0 + \alpha\Delta H \quad (1.1)$$

where  $E_0$  is the activation energy of a reference reaction, and  $\alpha$ , a constant which characterises the position of the TS along the reaction coordinate. This relationship can be more generally used to compare larger families of reactions. Despite the widespread use of the BEP principle, the applicability of this relationship is not well described.

I probe this applicability, with the aim to determine how generally the BEP principle can be applied. This is achieved by relating accurate, theoretically determined C-H bond strengths of species which undergo abstraction at these positions to the experimentally determined HAT rate constants. HAT reaction rate constants depend on many factors. However, by using measured rate constants from specific conditions (LFP with CumO $\cdot$  at 298K), the difference in reactivity depend mainly on the differences in chemical properties of the substrates of interest. Therefore, we hypothesise that there should exist two BEP relationships for C-H bonds: one in which the incipient radical is delocalised into a  $\pi$ -system (benzylic-allylic), and the other in which the remaining alkyl radicals are largely localised.

Finally, recent experimental results show that non-redox active metal cations, which are found ubiquitously in biological systems, have an inhibitory effect on HAT reactions involving oxygen-centred radicals. This has been demonstrated experimentally for substrates which undergo abstraction

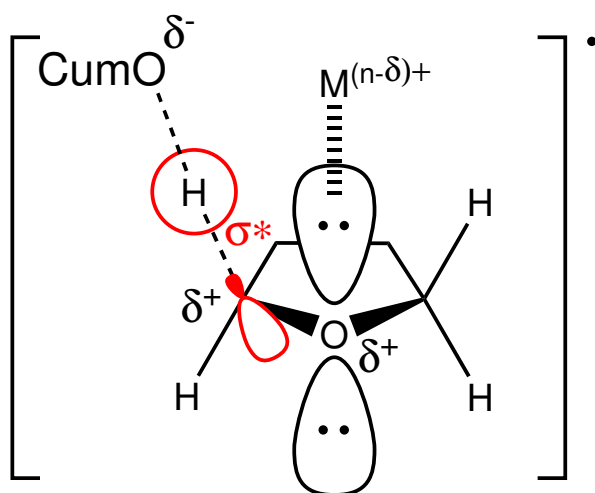
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from sites adjacent to heteroatoms (e.g. amines, amides, and ethers). Under various stoichiometric ratios, these metal cations have effects ranging from full inhibition to partial deactivation of HAT reactivity.<sup>40–42</sup> This effect has been attributed partially to the effects of hyperconjugative overlap. Take for example tetrahydrofuran (THF), shown in Scheme 1.3. Normally, there exists C-H bond weakening hyperconjugative overlap of electron density from one of the oxygen lone-pairs and the adjacent C-H  $\sigma^*$  anti-bonding orbitals. The interaction of a metal cation with the oxygen lone-pairs removes electron density from this interaction, thus increasing the C-H bond strength. As a result, the reactivity of this bond is decreased, as observed from the experimentally measured 3.2-fold decrease in the rate constant for HAT with  $\text{CumO}^\bullet$  in acetonitrile from  $6.65 \times 10^7 \text{ M}^{-1}\text{s}^{-1}$  to  $7.0 \times 10^7 \text{ M}^{-1}\text{s}^{-1}$  in the presence of 1.0 M  $\text{Mg}(\text{ClO}_4)_2$ .<sup>40</sup>

The nature of the interactions between non-redox active metal cations and organic substrates is poorly understood. This problem is explored in Chapter 5, with the aim to understand the fundamental physico-chemical properties which lead to the experimentally observed trends in reactivity. The experimentally observed effects have led us to hypothesise that the presence of non-redox active metal cations has a chemo-protective effect against the radical-induced oxidation of biomaterials such as proteins.

In using theory to study HAT reactions, I hope to contribute to a better understanding of the fundamental properties which govern these reactions, and thus develop insights into the many important biological processes in which HAT takes place.



Scheme 1.3: Hyperconjugative overlap in tetrahydrofuran and the effect of non-redox active metal cations. The metal cation accepts electron density from the heteroatom lone pair, reducing overlap with the C-H  $\sigma^*$  antibonding orbital, shown in red, and increasing the C-H bond strength.

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