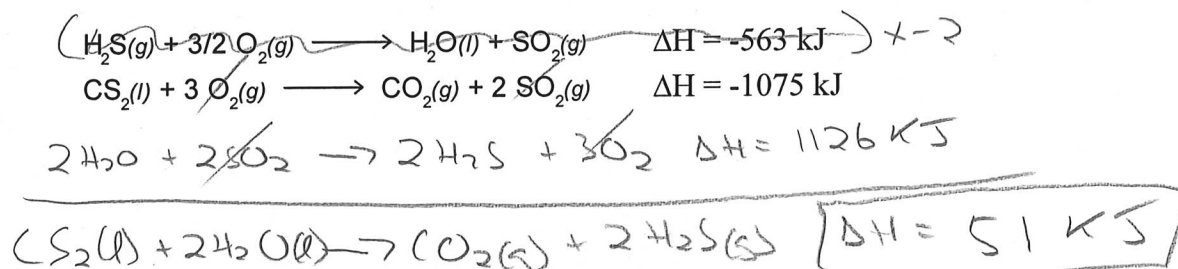
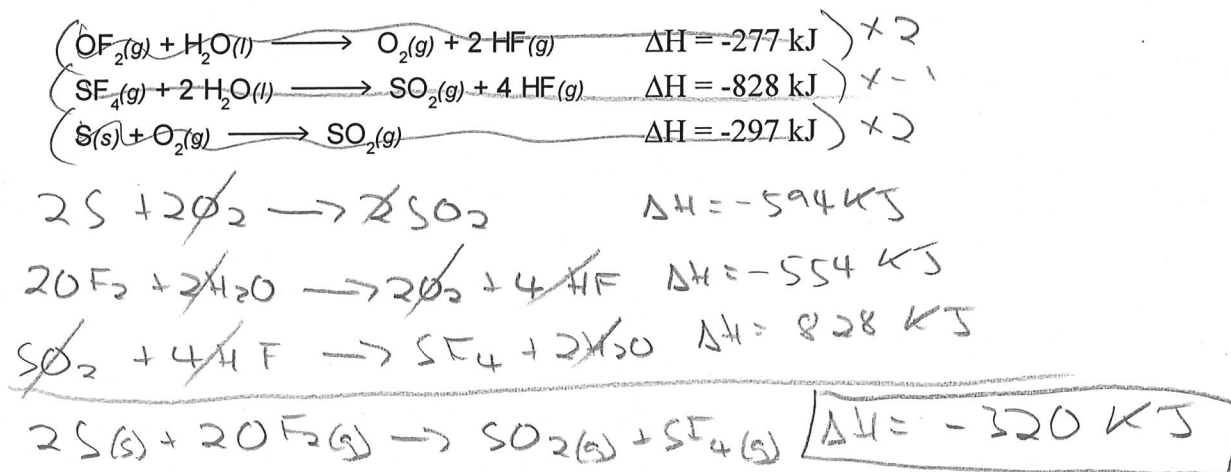


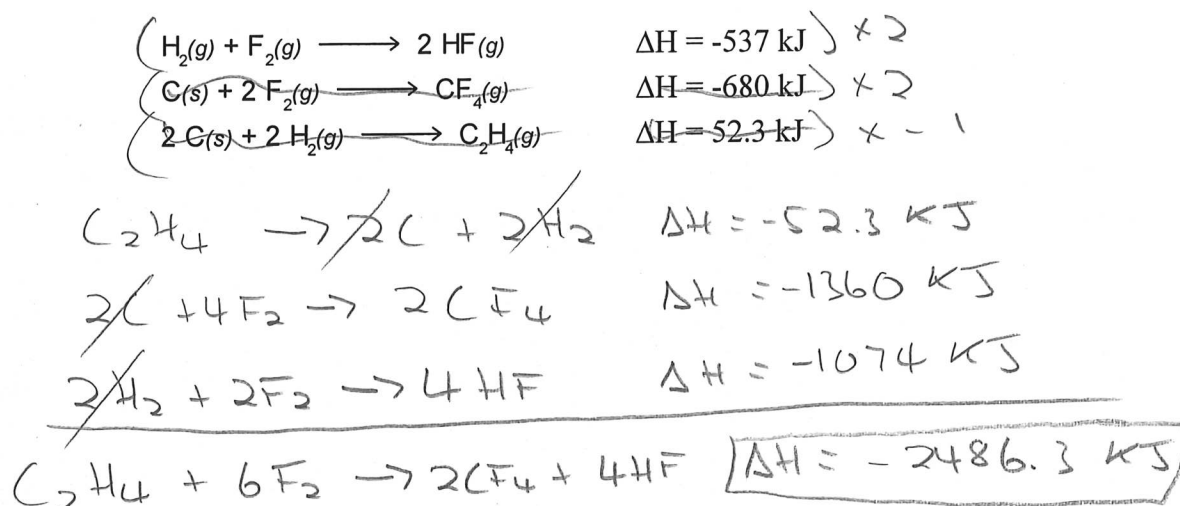
1. Find ΔH for the following reaction : $\text{CS}_2(l) + 2 \text{H}_2\text{O}(l) \longrightarrow \text{CO}_2(g) + 2 \text{H}_2\text{S}(g)$,
using the equations below.



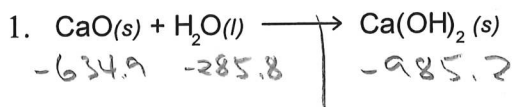
2. Find ΔH for the following equation: $2 \text{S}(s) + 2 \text{OF}_2(g) \longrightarrow \text{SO}_2(g) + \text{SF}_4(g)$, using
the equations below.



3. Calculate ΔH for the following reaction: $\text{C}_2\text{H}_4(g) + 6 \text{F}_2(g) \longrightarrow 2 \text{CF}_4(g) + 4 \text{HF}(g)$,
using the equations below.



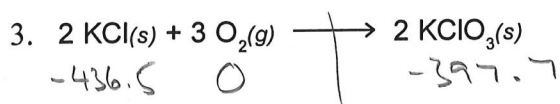
For each of the following reactions, find ΔH for the reaction by using the ΔH_f° values. Use the Heat of Formation table on Schoology.



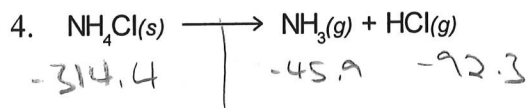
$$\Delta H = [-985.2] - [-634.9 + -285.8] = \boxed{-64.5 \text{ kJ}}$$



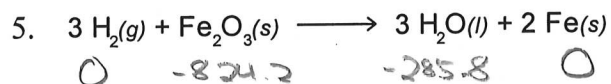
$$\Delta H = [4(-393.5) + 6(-241.8)] - [2(-84.7) + 7(0)] = \boxed{-2855.4 \text{ kJ}}$$



$$\Delta H = [2(-397.7)] - [2(-436.5) + 3(0)] = \boxed{77.6 \text{ kJ}}$$



$$\Delta H = [-45.9 + -92.3] - [-314.4] = \boxed{176.2 \text{ kJ}}$$



$$\Delta H = [3(-285.8)] - [-824.2] = \boxed{-33.2 \text{ kJ}}$$