**Equilibrium Test Review**

30 MCQs 3 FRQs

***MC Hints:*** Total pressure at equilibrium (sum of pressures of reactants + products from ICE chart to find “x”)

Finding K and evaluating >, <, or = 1 meaning?

Effect of changes on equilibrium. (stress 🡪 shift to relieve) Reestablish equilibrium (what has changed?)

Count species in particle diagrams to determine “concentration” for use in K calculation

How do changes in volume/pressure affect systems? Effect of adding material NOT part of the equilibrium?

Given Kp= Kc(RT)Δn understand the relationship between Kp and Kc

Compare Q vs K situations. Evaluate.

How does kinetics connect to equilibrium? Speed? Effect of catalyst?

What can change the Keq?

What does a very large or very small value of Keq indicate?

**FRQs**

1) PV=nRT equilibrium reaction, partial pressures, ICE table, Kp ,Q situation

2) Read, understand graph of data for reaction Evaluate a claim. Kp/Q again

3) Concentration equilibrium reaction, Beer’s Law, ICE chart - quadratic NOT needed