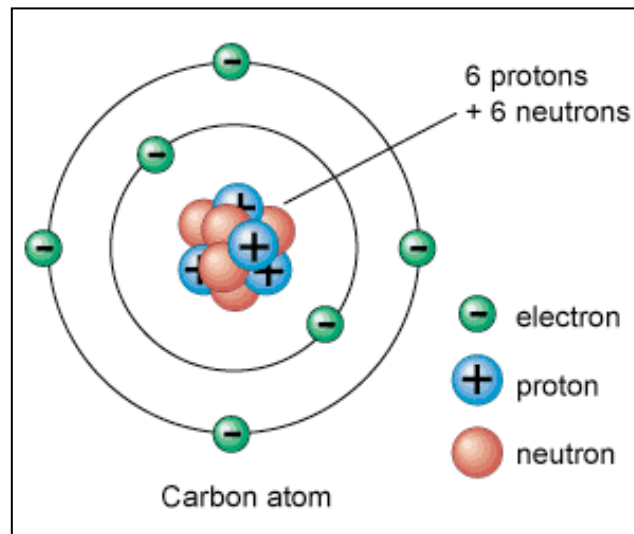


# Chapter 40

## All About Atoms



Lecture 35, 36

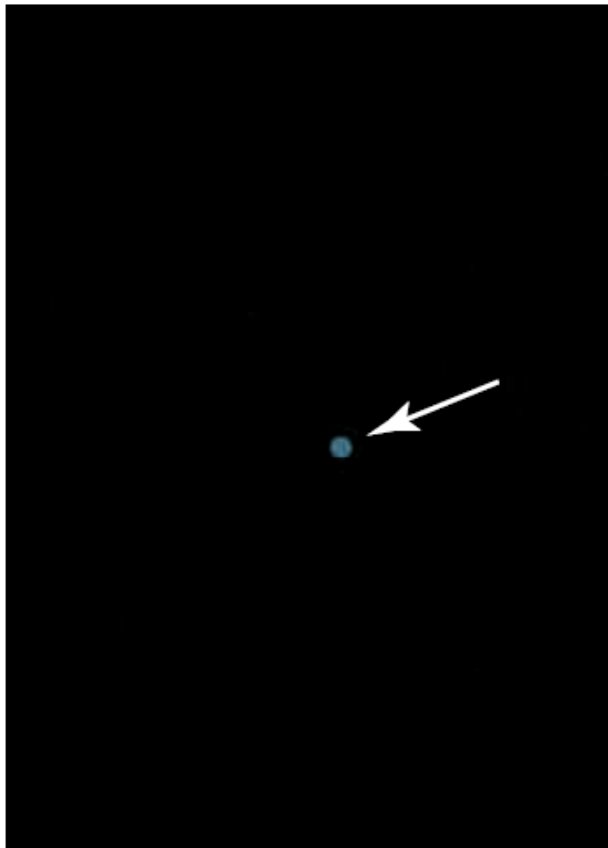
Seon-Hee Seo

2016.12.05-12.07

## 40.2: Some Properties of Atoms

**Atoms are stable.** Essentially all the atoms that form our tangible world have existed without change for billions of years.

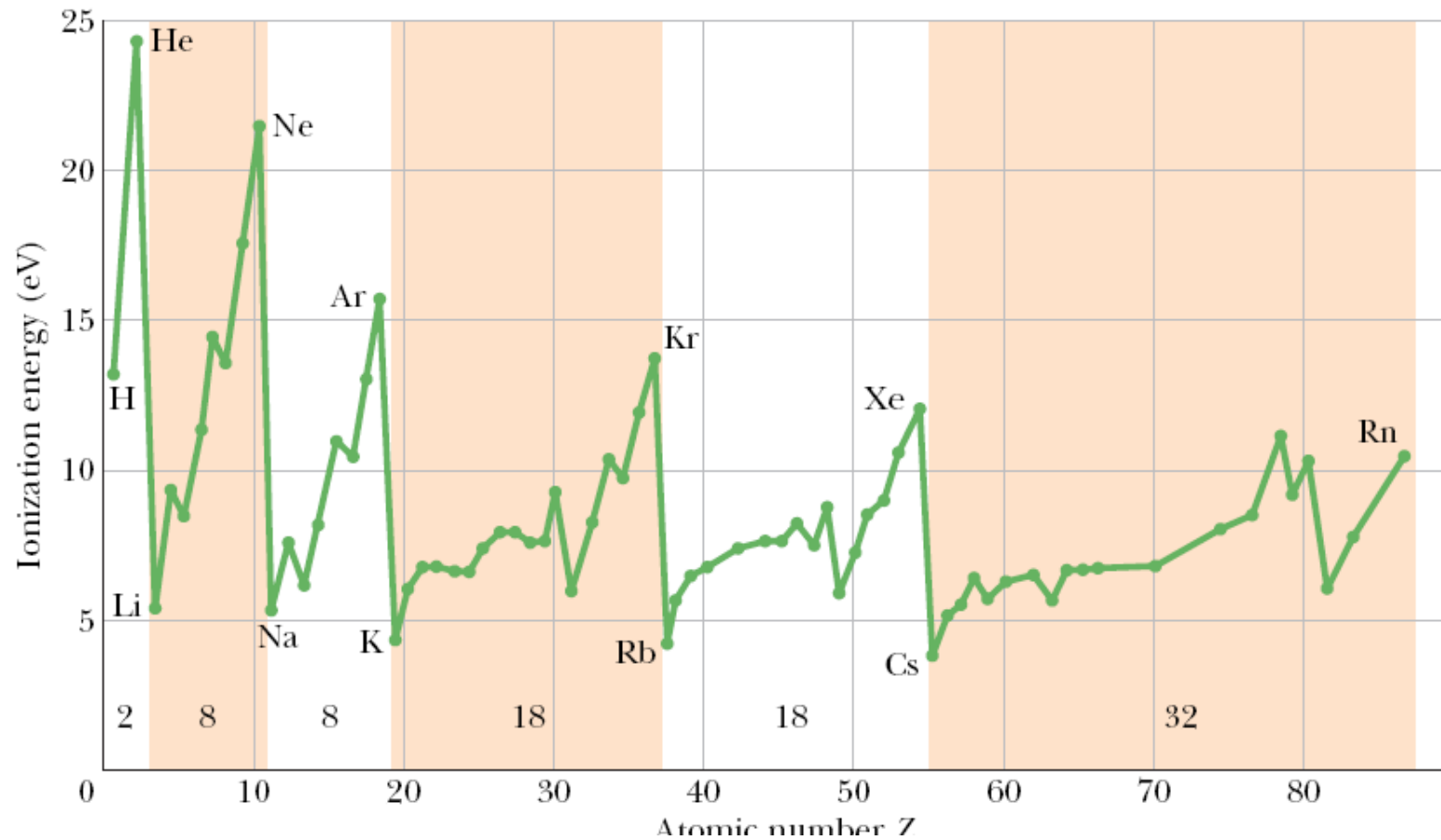
**Atoms combine with each other.** They stick together to form stable molecules and stack up to form rigid solids.



**Fig. 40-1** The blue dot is a photograph of the light emitted from a single barium ion held for a long time in a trap at the University of Washington. Special techniques caused the ion to emit light over and over again as it underwent transitions between the same pair of energy levels. The dot represents the cumulative emission of many photons.  
(Courtesy Warren Nagourney)

## 40.2: Some Properties of Atoms

### Atoms Are Put Together Systematically



**Fig. 40-2** A plot of the ionization energies of the elements as a function of atomic number, showing the periodic repetition of properties through the six complete horizontal periods of the periodic table. The number of elements in each of these periods is indicated.

## 40.2: Some Properties of Atoms

# Atoms Are Put Together Systematically

The elements are arranged in the **periodic table** in **six complete horizontal periods** (and a seventh incomplete period): except for the first, each period starts at the left with a highly reactive **alkali metal** (lithium, sodium, potassium, and so on) and ends at the right with a chemically inert **noble gas** (neon, argon, krypton, and so on).

Quantum physics accounts for the chemical properties of these elements.

The numbers of elements in the six periods are 2, 8, 8, 18, 18, and 32.

Quantum physics predicts these numbers.

## 40.2: Some Properties of Atoms

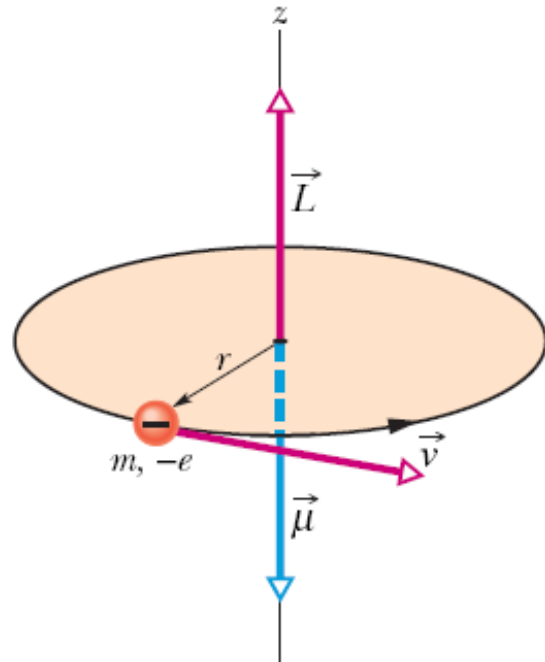
### Atoms Emit and Absorb Light:

An atom can make a transition from one state to another by emitting light (to jump to a lower energy level  $E_{low}$ ) or by absorbing light (to jump to a higher energy level  $E_{high}$ ).

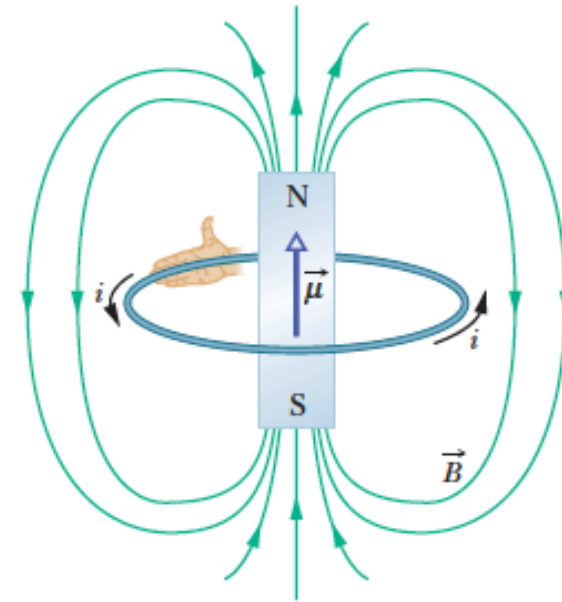
The light is emitted or absorbed as a photon with energy

$$hf = E_{high} - E_{low}.$$

## 40.2: Some Properties of Atoms: Atoms have Angular Momentum and Magnetism



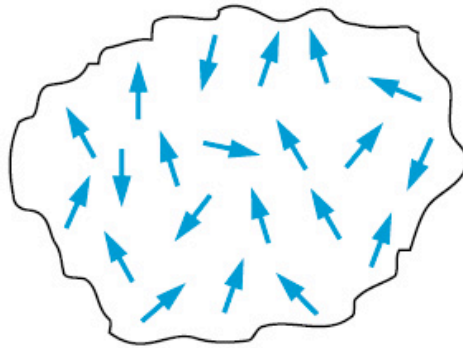
**Fig. 40-3** A classical model showing a particle of mass  $m$  and charge  $-e$  moving with speed  $v$  in a circle of radius  $r$ . The moving particle has an angular momentum  $\vec{L}$  given by  $\vec{r} \times \vec{p}$ , where  $\vec{p}$  is its linear momentum  $m\vec{v}$ . The particle's motion is equivalent to a current loop that has an associated magnetic moment  $\vec{\mu}$  that is directed opposite  $\vec{L}$ .



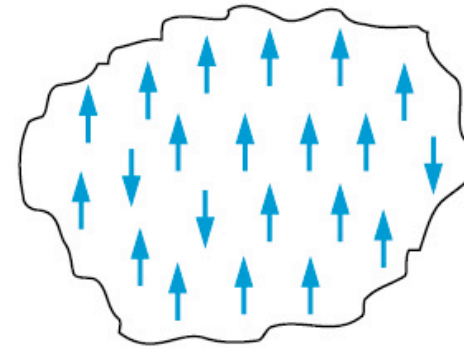
$$\vec{B}(z) = \frac{\mu_0}{2\pi} \frac{\vec{\mu}}{z^3} \quad (\text{current-carrying coil}). \quad (29-27)$$

# 32-8 Ferromagnetism

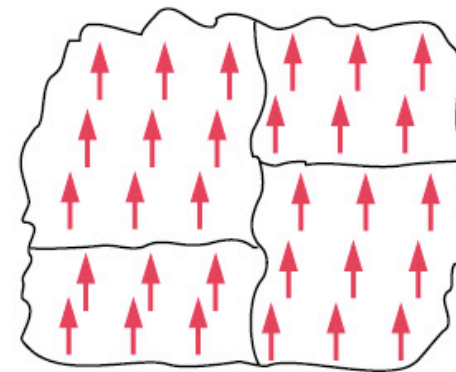
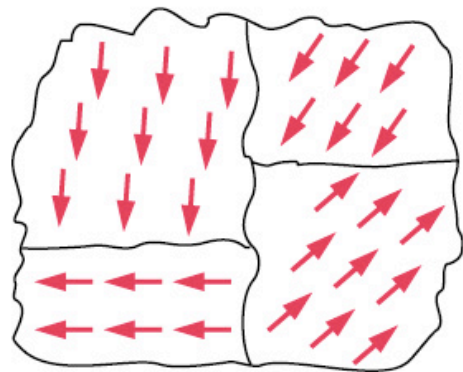
Magnetic field absent



In presence of magnetic field



Paramagnetism



Ferromagnetism

# 32-8 Dia/Para/Ferro magnetism

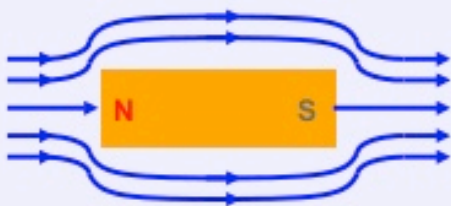
## Comparison of Dia, Para and Ferro Magnetic materials:

### DIA

1. Diamagnetic substances are those substances which are feebly repelled by a magnet.

Eg. Antimony, Bismuth, Copper, Gold, Silver, Quartz, Mercury, Alcohol, water, Hydrogen, Air, Argon, etc.

2. When placed in magnetic field, the lines of force tend to avoid the substance.

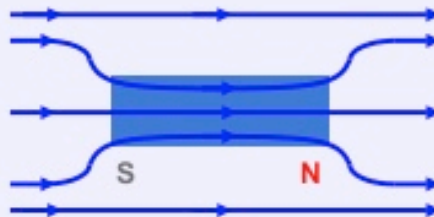


### PARA

Paramagnetic substances are those substances which are feebly attracted by a magnet.

Eg. Aluminium, Chromium, Alkali and Alkaline earth metals, Platinum, Oxygen, etc.

The lines of force prefer to pass through the substance rather than air.

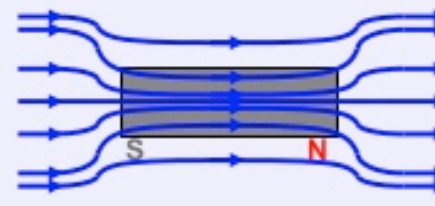


### FERRO

Ferromagnetic substances are those substances which are strongly attracted by a magnet.

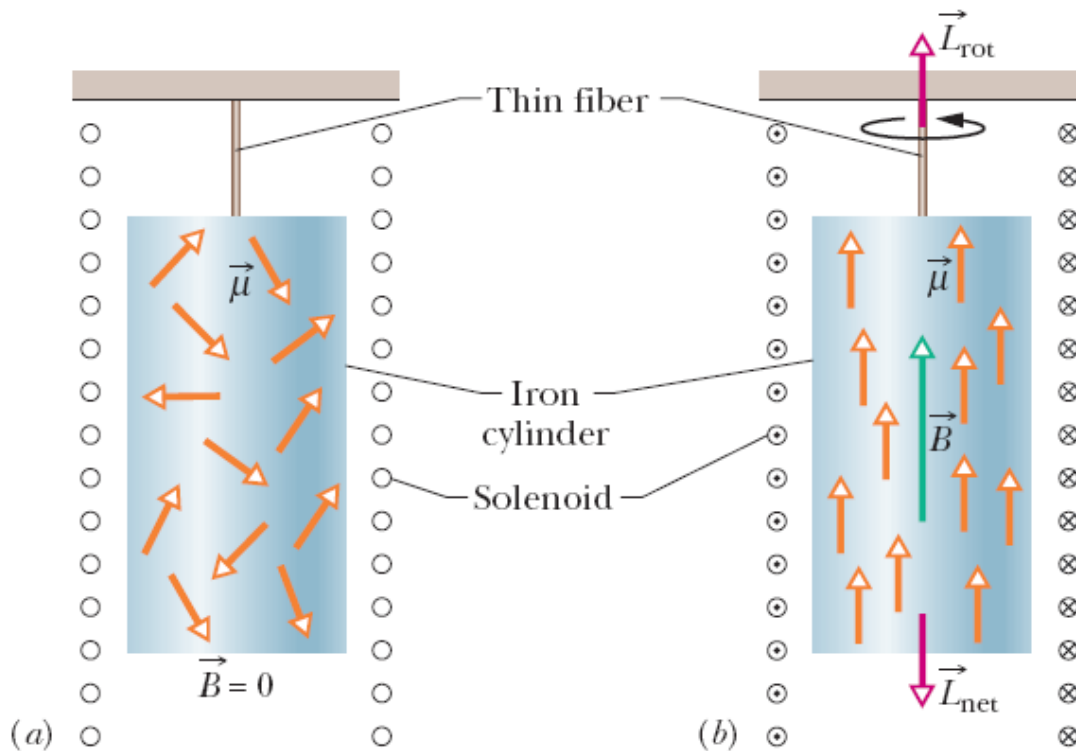
Eg. Iron, Cobalt, Nickel, Gadolinium, Dysprosium, etc.

The lines of force tend to crowd into the specimen.





## 40.2: Some Properties of Atoms: The Einstein-de Haas Experiment



In 1915

Observation of cylinder's rotation  $\rightarrow$  angular momentum and magnetic moment of individual atom are coupled in opposite direction.

Aligning the magnetic moment vectors rotates the cylinder.

**Fig. 40-4** The Einstein–de Haas experimental setup. (a) Initially, the magnetic field in the iron cylinder is zero and the magnetic dipole moment vectors of its atoms are randomly oriented. (b) When a magnetic field is set up along the cylinder's axis, the magnetic dipole moment vectors  $\vec{\mu}$  line up parallel to and the cylinder begins to rotate.

## 40.3: Electron Spin

The electron has an **intrinsic spin angular momentum  $S$** , often called simply **spin**.

The magnitude of  $S$  is quantized and depends on a **spin quantum number  $s$** , which is always  $\frac{1}{2}$  for electrons (and for protons and neutrons).

The component of  $S$  is measured along any axis, is quantized, and depends on a **spin magnetic quantum number  $m_s$** , which can have only the value  $+\frac{1}{2}$  or  $-\frac{1}{2}$

Table 40-1

### Electron States for an Atom

Quantum Number	Symbol	Allowed Values	Related to
Principal	$n$	$1, 2, 3, \dots$	Distance from the nucleus
Orbital	$\ell$	$0, 1, 2, \dots, (n - 1)$	Orbital angular momentum
Orbital magnetic	$m_\ell$	$0, \pm 1, \pm 2, \dots, \pm \ell$	Orbital angular momentum ( $z$ component)
Spin	$s$	$\frac{1}{2}$	Spin angular momentum
Spin magnetic	$m_s$	$\pm \frac{1}{2}$	Spin angular momentum ( $z$ component)

## 40.4: Angular Momenta & Magnetic Dipole Moments

The magnitude  $L$  of the **orbital angular momentum**  $\vec{L}$  of an electron in an atom is quantized; that is, it can have only certain values

$$L = \sqrt{\ell(\ell + 1)}\hbar,$$

The magnetic dipole has an **orbital magnetic dipole moment**  $\vec{\mu}_{orb}$ , is related to the angular momentum by

$$\vec{\mu}_{orb} = -\frac{e}{2m} \vec{L} \longrightarrow \text{magnitude} = \frac{e}{2m} \sqrt{\ell(\ell + 1)}\hbar.$$

If the atom is located in a magnetic field  $\mathbf{B}$ , with a  $z$  axis extending in the direction of the field lines at the atom's location, we can measure the  $z$  components of  $\vec{\mu}_{orb}$  and  $\vec{L}$  along that axis.

$$\mu_{orb,z} = -m_\ell \mu_B.$$

$$L_z = m_\ell \hbar.$$

Here  $\mu_B$  is the **Bohr Magnetron**

$$\mu_B = \frac{eh}{4\pi m} = \frac{e\hbar}{2m} = 9.274 \times 10^{-24} \text{ J/T}$$

## 40.4: Angular Momenta & Magnetic Dipole Moments

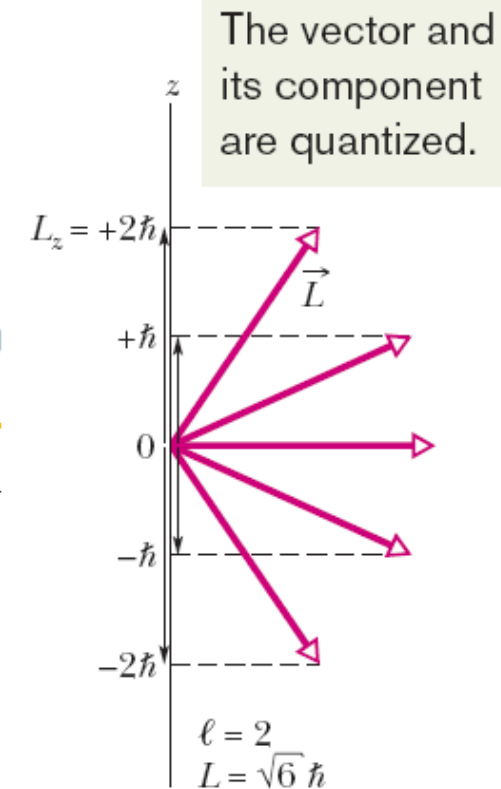
The components  $L_z$  of the angular momentum are also quantized, and they are given by

$$L_z = m_\ell \hbar.$$

Table 40-1

Electron States for an Atom

Quantum Number	Symbol	Allowed Values	Related to
Principal	$n$	$1, 2, 3, \dots$	Distance from the nucleus
Orbital	$\ell$	$0, 1, 2, \dots, (n - 1)$	Orbital angular momentum
Orbital magnetic	$m_\ell$	$0, \pm 1, \pm 2, \dots, \pm \ell$	Orbital angular momentum ( $z$ component)
Spin	$s$	$\frac{1}{2}$	Spin angular momentum
Spin magnetic	$m_s$	$\pm \frac{1}{2}$	Spin angular momentum ( $z$ component)



**Fig. 40-5** The allowed values of  $L_z$  for an electron in a quantum state with  $\ell = 2$ . For every orbital angular momentum vector  $\vec{L}$  in the figure, there is a vector pointing in the opposite direction, representing the magnitude and direction of the orbital magnetic dipole moment  $\vec{\mu}_{\text{orb}}$ .

## 40.4: Angular Momenta & Magnetic Dipole Moments

$$S = \sqrt{s(s+1)}\hbar$$

$$= \sqrt{\left(\frac{1}{2}\right)\left(\frac{1}{2}+1\right)}\hbar = 0.866\hbar,$$

$$\vec{\mu}_s = -\frac{e}{m} \vec{S}. \quad (\text{Spin magnetic dipole moment})$$

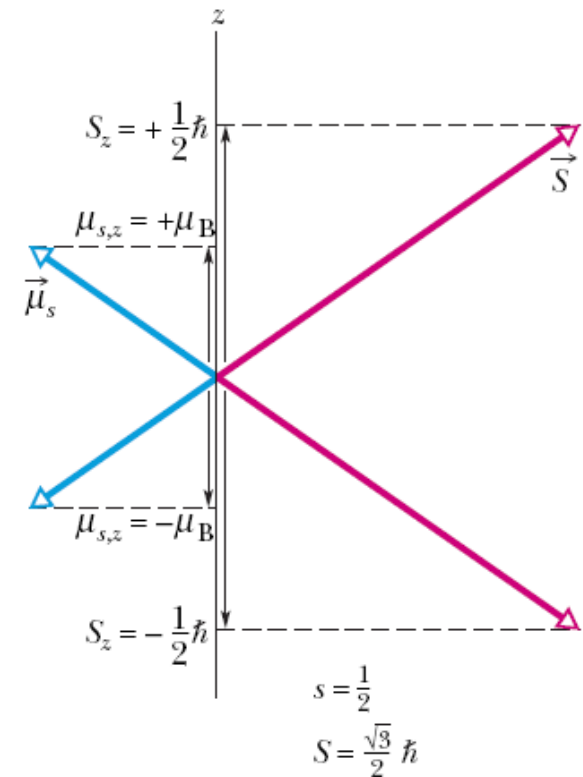
$$= \frac{e}{m} \sqrt{s(s+1)}\hbar.$$

$S_z = m_s \hbar$ , (components of  $\vec{S}$  can be measured along a preferred axis).

$m_s = +\frac{1}{2}$  (Spin angular quantum number can only have two values)

$m_s = -\frac{1}{2}$

$\mu_{s,z} = -2m_s \mu_B$ . (The components of  $\mu_{s,z}$  are also quantized)



**Fig. 40-6** The allowed values of  $S_z$  and  $\mu_z$  for an electron.

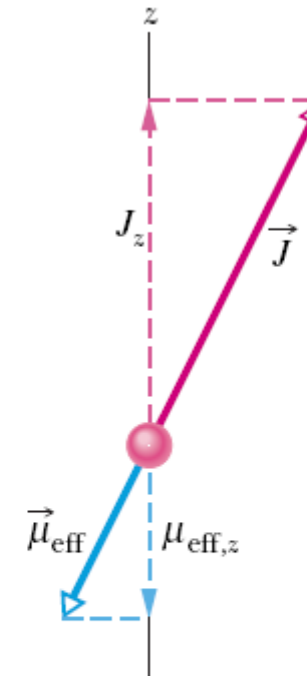
## 40.4: Angular Momenta & Magnetic Dipole Moments

For an atom containing more than one electron, we define a **total angular momentum**,  $\vec{J}$ , which is the vector sum of orbital and their spin angular momenta of the individual electrons.

This number of protons is defined as being the **atomic number**  $Z$  of the element.

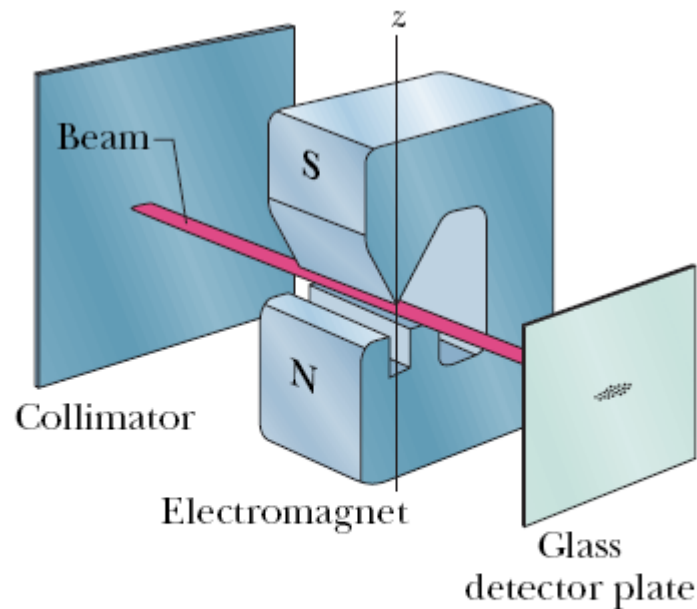
$$\vec{J} = (\vec{L}_1 + \vec{L}_2 + \vec{L}_3 + \cdots + \vec{L}_Z) + (\vec{S}_1 + \vec{S}_2 + \vec{S}_3 + \cdots + \vec{S}_Z).$$

The **effective magnetic dipole moment**,  $\vec{\mu}_{\text{eff}}$ , for the atom is the component of the vector sum of the individual **magnetic dipole moments** in the direction of  $-\vec{J}$ .



**Fig. 40-7** A classical model showing the total angular momentum vector  $\vec{J}$  and the effective magnetic moment vector  $\vec{\mu}_{\text{eff}}$ .

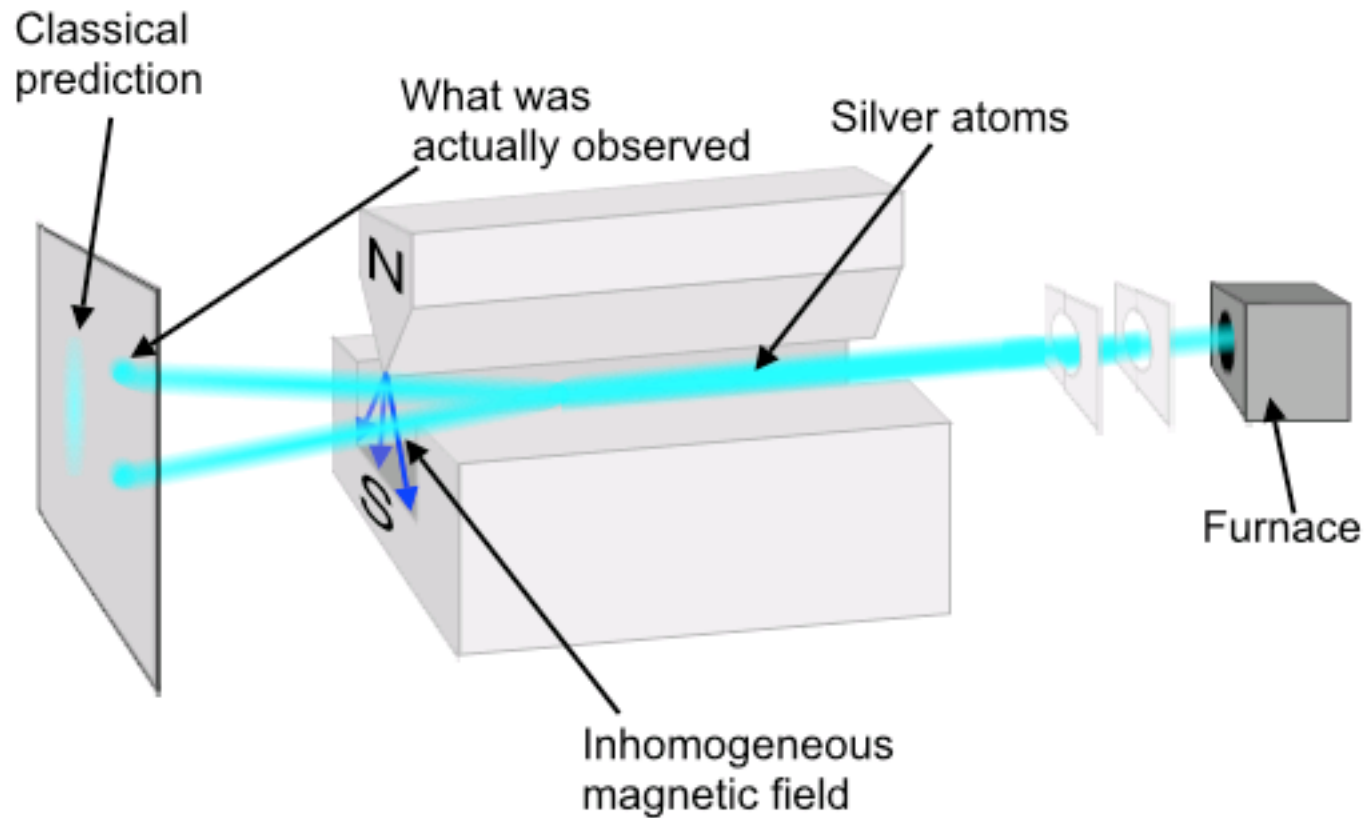
# 40.5: The Stern-Gerlach Experiment



**Fig. 40-8** Apparatus used by Stern and Gerlach.

In 1922, Otto Stern and Walther Gerlach showed experimentally that the magnetic moment of silver atoms is quantized. In the experiment silver is vaporized in an oven, and some of the atoms in that vapor escape through a narrow slit in the oven wall and pass into an evacuated tube. The beam passes between the poles of an electromagnet and then lands on a glass detector plate where it forms a silver deposit. By analyzing the silver deposit on the plate, one can determine what deflections the atoms underwent in the magnetic field.

# 40.5: The Stern-Gerlach Experiment





# 40.5: The Stern-Gerlach Experiment

The type of magnetic force the silver atom experiences in the Stern-Gerlach experiment is due to an interaction between the magnetic field of the electromagnet and the magnetic dipole of the individual silver atom.

The energy  $U$  of the dipole in the magnetic field is  $U = -\vec{\mu} \cdot \vec{B}$ , where  $\mu$  is the magnetic dipole moment of the silver atom.

Therefore  $U = -\mu_z B$ .

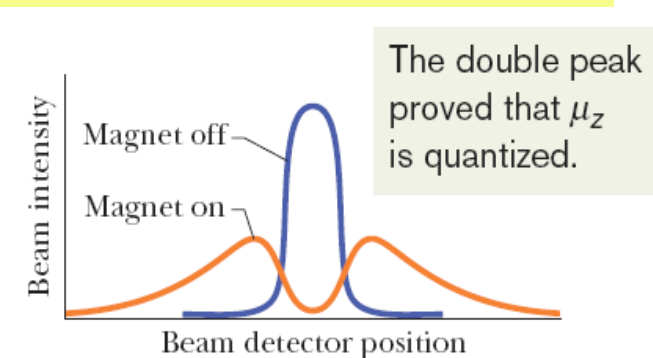
And

$$F_z = -\frac{dU}{dz} = \mu_z \frac{dB}{dz}.$$

$$\mu_{s,z} = -2m_s\mu_B. \quad m_s = \pm \frac{1}{2}$$

$$\mu_{s,z} = -2(+\frac{1}{2})\mu_B = -\mu_B \quad \text{and} \quad \mu_{s,z} = -2(-\frac{1}{2})\mu_B = +\mu_B.$$

$$F_z = -\mu_B \left( \frac{dB}{dz} \right) \quad \text{and} \quad F_z = +\mu_B \left( \frac{dB}{dz} \right),$$



**Fig. 40-9** Results of a modern repetition of the Stern–Gerlach experiment. With the electromagnet turned off, there is only a single beam; with the electromagnet turned on, the original beam splits into two subbeams. The two subbeams correspond to parallel and antiparallel alignment of the magnetic moments of cesium atoms with the external magnetic field.

## Example, Beam Separation in a Stern-Gerlach experiment:

In the Stern–Gerlach experiment of Fig. 40-8, a beam of silver atoms passes through a magnetic field gradient  $dB/dz$  of magnitude 1.4 T/mm that is set up along the  $z$  axis. This region has a length  $w$  of 3.5 cm in the direction of the original beam. The speed of the atoms is 750 m/s. By what distance  $d$  have the atoms been deflected when they leave the region of the field gradient? The mass  $M$  of a silver atom is  $1.8 \times 10^{-25}$  kg.

### KEY IDEAS

(1) The deflection of a silver atom in the beam is due to an interaction between the magnetic dipole of the atom and the magnetic field, because of the gradient  $dB/dz$ . The deflecting force is directed along the field gradient (along the  $z$  axis) and is given by Eqs. 40-19. Let us consider only deflection in the positive direction of  $z$ ; thus, we shall use  $F_z = \mu_B(dB/dz)$  from Eqs. 40-19.

(2) We assume the field gradient  $dB/dz$  has the same value throughout the region through which the silver atoms travel. Thus, force component  $F_z$  is constant in that region, and from Newton's second law, the acceleration  $a_z$  of an atom along the  $z$  axis due to  $F_z$  is also constant.

**Calculations:** Putting these ideas together, we write the acceleration as

$$a_z = \frac{F_z}{M} = \frac{\mu_B(dB/dz)}{M}.$$

Because this acceleration is constant, we can use Eq. 2-15 (from Table 2-1) to write the deflection  $d$  parallel to the  $z$  axis as

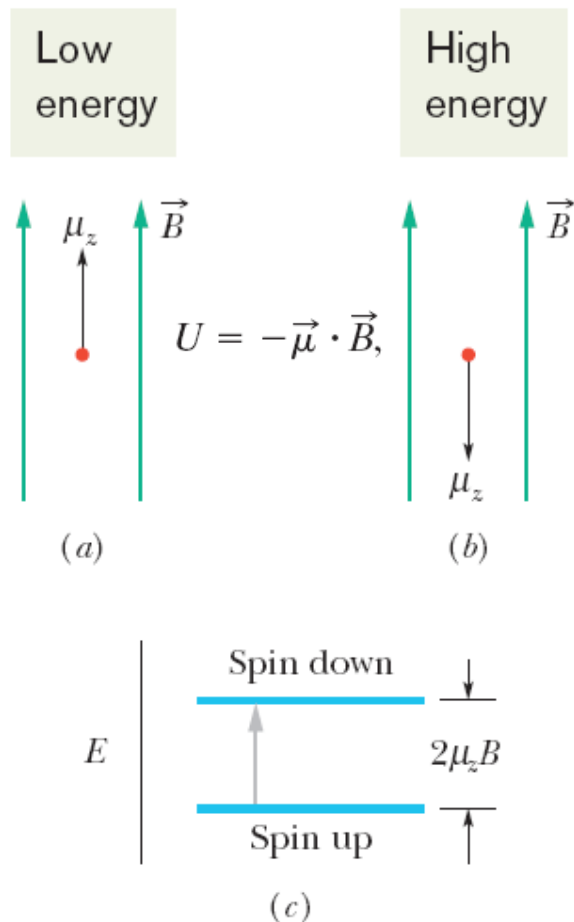
$$d = v_{0z}t + \frac{1}{2}a_z t^2 = 0t + \frac{1}{2}\left(\frac{\mu_B(dB/dz)}{M}\right)t^2. \quad (40-20)$$

Because the deflecting force on the atom acts perpendicular to the atom's original direction of travel, the component  $v$  of the atom's velocity along the original direction of travel is not changed by the force. Thus, the atom requires time  $t = w/v$  to travel through length  $w$  in that direction. Substituting  $w/v$  for  $t$  into Eq. 40-20, we find

$$\begin{aligned} d &= \frac{1}{2}\left(\frac{\mu_B(dB/dz)}{M}\right)\left(\frac{w}{v}\right)^2 = \frac{\mu_B(dB/dz)w^2}{2Mv^2} \\ &= (9.27 \times 10^{-24} \text{ J/T})(1.4 \times 10^3 \text{ T/m}) \\ &\quad \times \frac{(3.5 \times 10^{-2} \text{ m})^2}{(2)(1.8 \times 10^{-25} \text{ kg})(750 \text{ m/s})^2} \\ &= 7.85 \times 10^{-5} \text{ m} \approx 0.08 \text{ mm}. \end{aligned} \quad (\text{Answer})$$

The separation between the two subbeams is twice this, or 0.16 mm. This separation is not large but is easily measured.

## 40.6: Magnetic Resonance

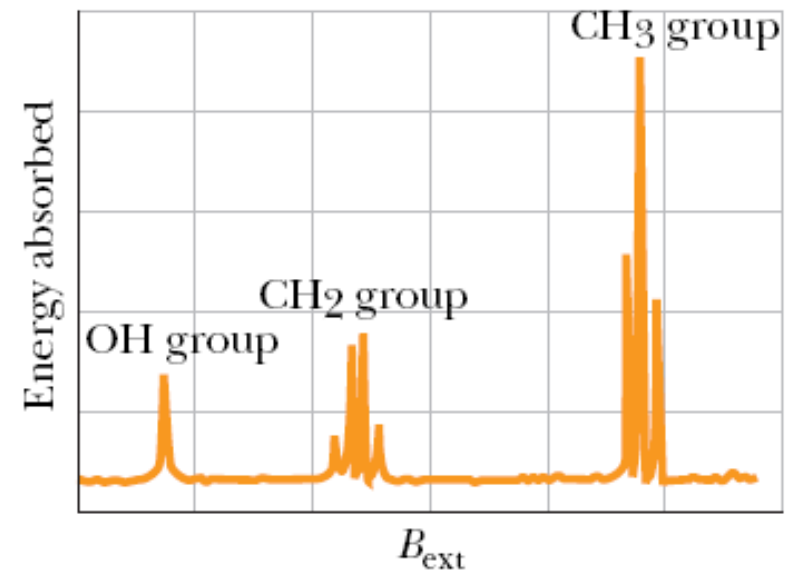


**Fig. 40-10** The  $z$  component of  $\vec{\mu}$  for a proton in the (a) lower-energy (spin-up) and (b) higher-energy (spin-down) state. (c) An energy-level diagram for the states, showing the upward quantum jump the proton makes when its spin flips from up to down.

$$\Delta E = \mu_z B - (-\mu_z B) = 2\mu_z B.$$

$$hf = 2\mu_z B. \quad (\text{Energy of absorbing photon})$$

Such absorption is called **magnetic resonance** or, **nuclear magnetic resonance (NMR)**, and the consequent reversal of  $S_z$  is called **spin-flipping**.



**Fig. 40-11** A nuclear magnetic resonance spectrum for ethanol,  $\text{CH}_3\text{CH}_2\text{OH}$ .

# 40.7: The Pauli Exclusion Principle



No two electrons confined to the same trap can have the same set of values for their quantum numbers.

This principle applies not only to electrons but also to protons and neutrons, all of which have  $s=\frac{1}{2}$ . The principle is known as the **Pauli exclusion principle** after Wolfgang Pauli, who formulated it in 1925.

## 40.8: Multiple Electrons in Rectangular Traps

- ❑ The Pauli exclusion principle disallows any more electrons from occupying that lowest energy level, and the next electron must occupy the next higher level.
- ❑ When an energy level cannot be occupied by more electrons because of the Pauli exclusion principle, we say that level is *full or fully occupied*.
- ❑ In contrast, a level that is not occupied by any electrons is *empty or unoccupied*.
- ❑ For intermediate situations, the level is *partially occupied*.
- ❑ The *electron configuration of a system of trapped* electrons is a listing or drawing either of the energy levels the electrons occupy or of the set of the quantum numbers of the electrons.

## 40.8: Multiple Electrons in Rectangular Traps

Example:

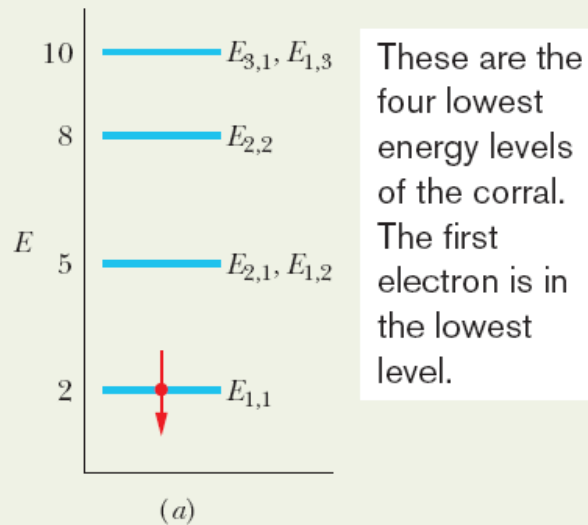
Principal Quantum Number,  $n = 3$

$n$	$\ell$	$m_\ell$	$m_s$	$n$	$\ell$	$m_\ell$	$m_s$
3	0	0	$+\frac{1}{2}$	3	2	-2	$+\frac{1}{2}$
3	0	0	$-\frac{1}{2}$	3	2	-2	$-\frac{1}{2}$
3	1	-1	$+\frac{1}{2}$	3	2	-1	$+\frac{1}{2}$
3	1	-1	$-\frac{1}{2}$	3	2	-1	$-\frac{1}{2}$
3	1	0	$+\frac{1}{2}$	3	2	0	$+\frac{1}{2}$
3	1	0	$-\frac{1}{2}$	3	2	0	$-\frac{1}{2}$
3	1	+1	$+\frac{1}{2}$	3	2	1	$+\frac{1}{2}$
3	1	+1	$-\frac{1}{2}$	3	2	1	$-\frac{1}{2}$
				3	2	2	$+\frac{1}{2}$
				3	2	2	$-\frac{1}{2}$

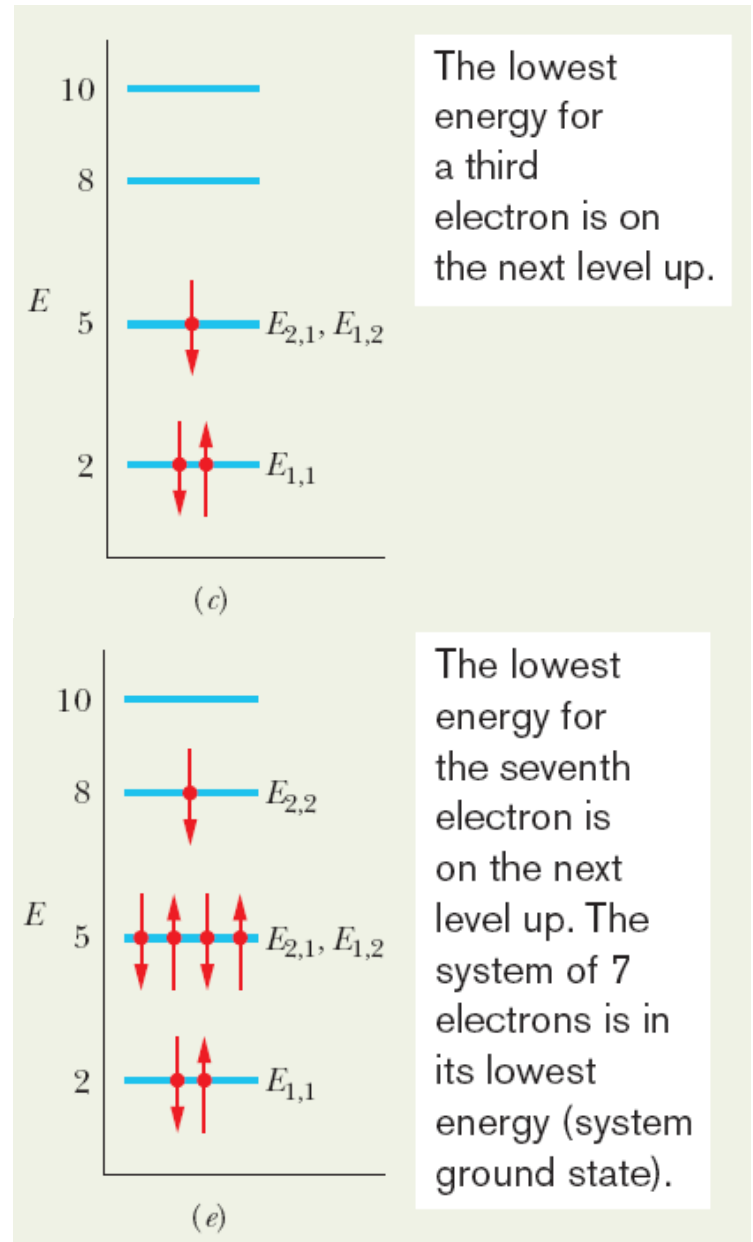
## Example, Energy levels in a 2-D infinite potential well:

Seven electrons are confined to a square corral (two-dimensional infinite potential well) with widths  $L_x = L_y = L$  (Fig. 39-13). Assume that the electrons do not electrically interact with one another.

(a) What is the electron configuration for the ground state of the system of seven electrons?



**Electrons, one by one:** The level for energy  $E_{1,1}$  is fully occupied, and thus the third electron cannot have that energy. Therefore, the third electron goes into the next higher level, which is for the equal energies  $E_{2,1}$  and  $E_{1,2}$  (the level is degenerate). This third electron can have quantum numbers  $n_x$  and  $n_y$  of either 1 and 2 or 2 and 1, respectively. It can also have a quantum number  $m_s$  of either  $+\frac{1}{2}$  or  $-\frac{1}{2}$ . Let us arbitrarily assign it the quantum numbers  $n_x = 2$ ,  $n_y = 1$ , and  $m_s = -\frac{1}{2}$ . We then represent it with a down arrow on the level for  $E_{1,2}$  and  $E_{2,1}$  in Fig. 40-12c.



## Example, Energy levels in a 2-D infinite potential well, cont.:

(b) What is the total energy of the seven-electron system in its ground state, as a multiple of  $h^2/8mL^2$ ?

### KEY IDEA

The total energy  $E_{\text{gr}}$  is the sum of the energies of the individual electrons in the system's ground-state configuration.

**Ground-state energy:** The energy of each electron can be read from Table 39-1, which is partially reproduced in Table 40-2, or from Fig. 40-12e. Because there are two electrons in the first (lowest) level, four in the second level, and one in the third level, we have

$$\begin{aligned} E_{\text{gr}} &= 2\left(2 \frac{h^2}{8mL^2}\right) + 4\left(5 \frac{h^2}{8mL^2}\right) + 1\left(8 \frac{h^2}{8mL^2}\right) \\ &= 32 \frac{h^2}{8mL^2}. \end{aligned} \quad (\text{Answer})$$

(c) How much energy must be transferred to the system for it to jump to its first excited state, and what is the energy of that state?

**First-excited-state energy:** Let us consider the three jumps shown in Fig. 40-12f; all are allowed by the Pauli exclusion principle because they are jumps to either empty or partially occupied states. In one of those possible jumps, an electron jumps from the  $E_{1,1}$  level to the partially occupied  $E_{2,2}$  level. The change in the energy is

Table 40-2

Ground-State Configuration and Energies

$n_x$	$n_y$	$m_s$	Energy <sup>a</sup>
2	2	$-\frac{1}{2}$	8
2	1	$+\frac{1}{2}$	5
2	1	$-\frac{1}{2}$	5
1	2	$+\frac{1}{2}$	5
1	2	$-\frac{1}{2}$	5
1	1	$+\frac{1}{2}$	2
1	1	$-\frac{1}{2}$	2
			Total 32

<sup>a</sup>In multiples of  $h^2/8mL^2$ .

$$\Delta E = E_{2,2} - E_{1,1} = 8 \frac{h^2}{8mL^2} - 2 \frac{h^2}{8mL^2} = 6 \frac{h^2}{8mL^2}.$$

(We shall assume that the spin orientation of the electron making the jump can change as needed.)



## 40.9: Building The Periodic Table

For the purpose of labeling subshells in atoms of elements, the values of  $l$  are represented by letters:

$\ell =$	0	1	2	3	4	5	...
	<i>s</i>	<i>p</i>	<i>d</i>	<i>f</i>	<i>g</i>	<i>h</i>	....

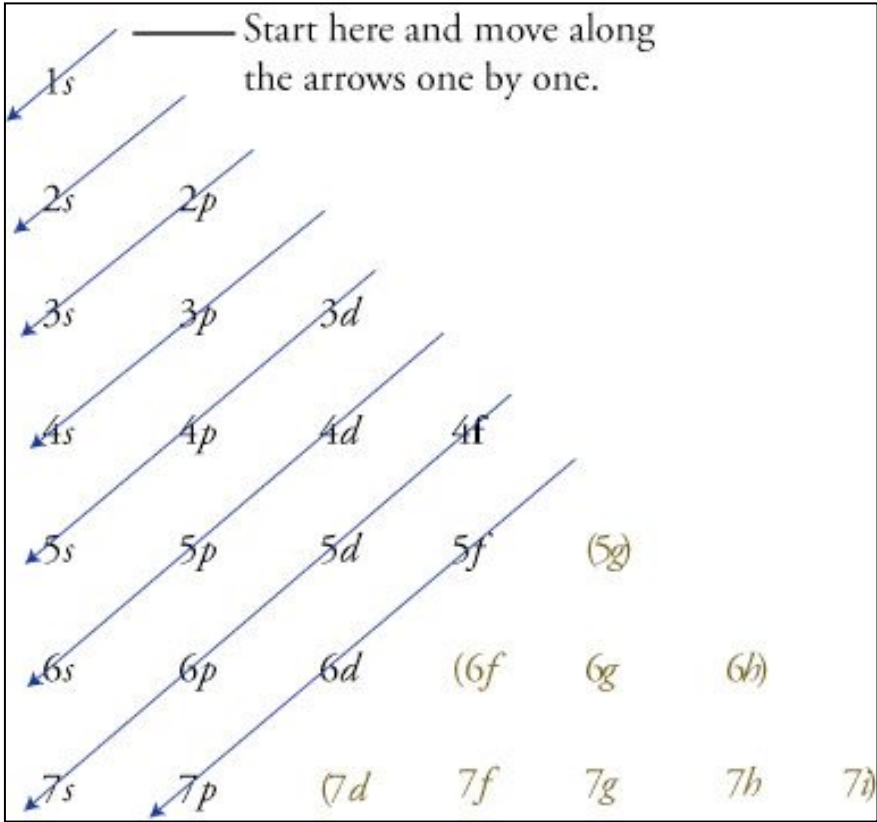
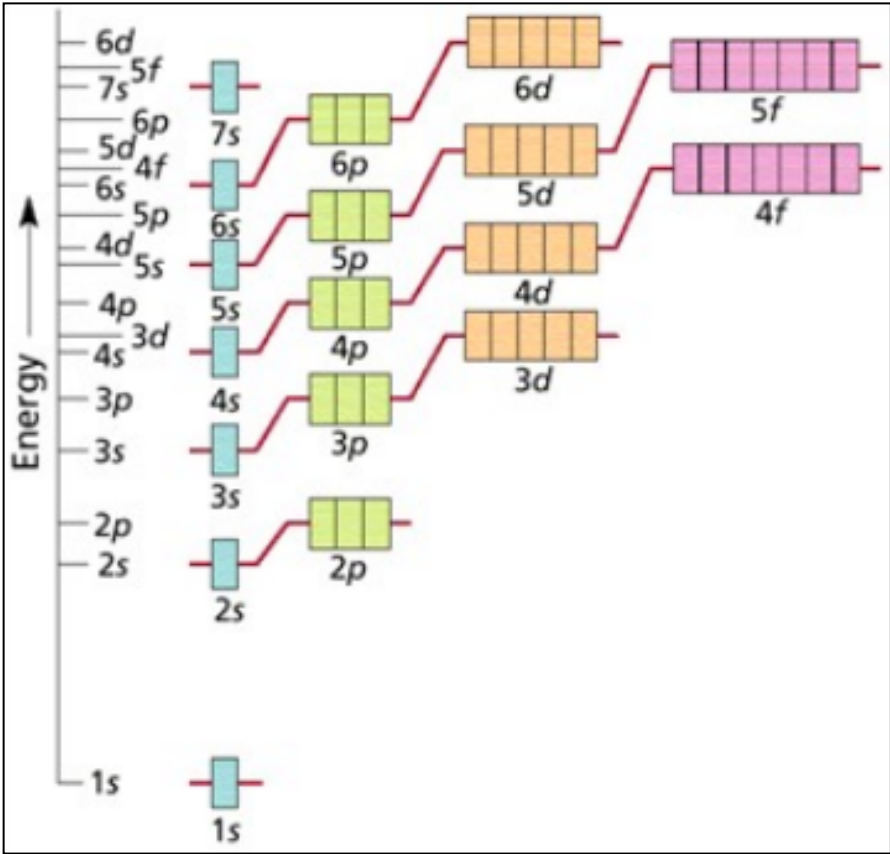
**Neon:** The atom has 10 electrons. It has three closed subshells ( $1s$ ,  $2s$ , and  $2p$ ) and, like the other noble gases that form the right-hand column of the periodic table, is almost chemically inert.

**Sodium:** The atom has 11 electrons. Ten of them form a closed neon-like core, and has zero angular momentum. The remaining electron is largely outside this inert core, in the  $3s$  subshell. This is the valence electron of the atom, and the atom's angular momentum and magnetic dipole moment must be due entirely to the spin of this single electron.

**Chlorine:** This, with 17 electrons, has the outermost 7 electrons in  $3p$  subshell, leaving a “hole” in this state. It is receptive to interacting with other atoms that have a valence electron that might fill this hole. Chlorine, like the other halogens that form column VIIA of the periodic table, is chemically active.

**Iron:**  $\underline{1s^2 \ 2s^2 2p^6 \ 3s^2 3p^6} \ 3d^6 \ 4s^2$ . This atom of 26 electrons, has the first 18 electrons form the five filled subshells that are marked off by the bracket. 6 of the remaining 8 electrons go into the  $3d$  subshell, and the remaining two go into the  $4s$  subshell. The configuration  $3d^6 4s^2$  is of lower energy than  $3d^8$ .

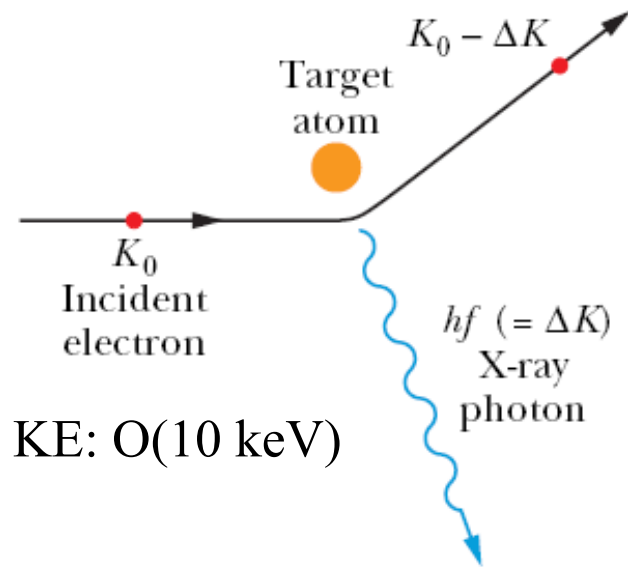
## 40.9: Building The Periodic Table



# 40.9: Building The Periodic Table

Periodic Table of the Elements																		18 VIIIA 8A	
<div>Atomic Number</div> <div>Symbol</div> <div>Name</div> <div>Atomic Mass</div>																			
1 1A 1A	2 IIA 2A											13 IIIA 3A	14 IVA 4A	15 VA 5A	16 VIA 6A	17 VIIA 7A	18 VIIIA 8A		
1 H Hydrogen 1.008																	2 He Helium 4.003		
3 Li Lithium 6.941	4 Be Beryllium 9.012											5 B Boron 10.811	6 C Carbon 12.011	7 N Nitrogen 14.007	8 O Oxygen 15.999	9 F Fluorine 18.998	10 Ne Neon 20.180		
11 Na Sodium 22.990	12 Mg Magnesium 24.305	3 IIIB 3B	4 IVB 4B	5 VB 5B	6 VIB 6B	7 VIIB 7B	8 VIII 8	9 VIII 8	10 VIII 8	11 IB 1B	12 IIB 2B	13 Al Aluminum 26.982	14 Si Silicon 28.086	15 P Phosphorus 30.974	16 S Sulfur 32.066	17 Cl Chlorine 35.453	18 Ar Argon 39.948		
19 K Potassium 39.098	20 Ca Calcium 40.078	21 Sc Scandium 44.956	22 Ti Titanium 47.867	23 V Vanadium 50.942	24 Cr Chromium 51.996	25 Mn Manganese 54.938	26 Fe Iron 55.845	27 Co Cobalt 58.933	28 Ni Nickel 58.693	29 Cu Copper 63.546	30 Zn Zinc 65.38	31 Ga Gallium 69.723	32 Ge Germanium 72.631	33 As Arsenic 74.922	34 Se Selenium 78.971	35 Br Bromine 79.904	36 Kr Krypton 84.798		
37 Rb Rubidium 84.468	38 Sr Strontium 87.62	39 Y Yttrium 88.906	40 Zr Zirconium 91.224	41 Nb Niobium 92.906	42 Mo Molybdenum 95.95	43 Tc Technetium 98.907	44 Ru Ruthenium 101.07	45 Rh Rhodium 102.906	46 Pd Palladium 106.42	47 Ag Silver 107.868	48 Cd Cadmium 112.411	49 In Indium 114.818	50 Sn Tin 118.711	51 Sb Antimony 121.760	52 Te Tellurium 127.6	53 I Iodine 126.904	54 Xe Xenon 131.294		
55 Cs Cesium 132.905	56 Ba Barium 137.328	57-71	72 Hf Hafnium 178.49	73 Ta Tantalum 180.948	74 W Tungsten 183.84	75 Re Rhenium 186.207	76 Os Osmium 190.23	77 Ir Iridium 192.217	78 Pt Platinum 195.085	79 Au Gold 196.967	80 Hg Mercury 200.592	81 Tl Thallium 204.383	82 Pb Lead 207.2	83 Bi Bismuth 208.980	84 Po Polonium [208.982]	85 At Astatine 209.987	86 Rn Radon 222.018		
87 Fr Francium 223.020	88 Ra Radium 226.025	89-103	104 Rf Rutherfordium [261]	105 Db Dubnium [262]	106 Sg Seaborgium [266]	107 Bh Bohrium [264]	108 Hs Hassium [269]	109 Mt Meitnerium [268]	110 Ds Darmstadtium [269]	111 Rg Roentgenium [272]	112 Cn Copernicium [277]	113 Uut Ununtrium unknown	114 Fl Flerovium [289]	115 Uup Ununpentium unknown	116 Lv Livermorium [298]	117 Uus Ununseptium unknown	118 Uuo Ununoctium unknown		
Lanthanide Series		57 La Lanthanum 138.905	58 Ce Cerium 140.116	59 Pr Praseodymium 140.908	60 Nd Neodymium 144.243	61 Pm Promethium 144.913	62 Sm Samarium 150.36	63 Eu Europium 151.964	64 Gd Gadolinium 157.25	65 Tb Terbium 158.925	66 Dy Dysprosium 162.500	67 Ho Holmium 164.930	68 Er Erbium 167.259	69 Tm Thulium 168.934	70 Yb Ytterbium 173.055	71 Lu Lutetium 174.967			
Actinide Series		89 Ac Actinium 227.028	90 Th Thorium 232.038	91 Pa Protactinium 231.036	92 U Uranium 238.029	93 Np Neptunium 237.048	94 Pu Plutonium 244.064	95 Am Americium 243.061	96 Cm Curium 247.070	97 Bk Berkelium 247.070	98 Cf Californium 251.080	99 Es Einsteinium [254]	100 Fm Fermium 257.095	101 Md Mendelevium 258.1	102 No Nobelium 259.101	103 Lr Lawrencium [262]			
Alkali Metal		Alkaline Earth		Transition Metal		Basic Metal		Semimetal		Nonmetal		Halogen		Noble Gas		Lanthanide		Actinide	

# 40.10: X-Rays and the Ordering of the Elements



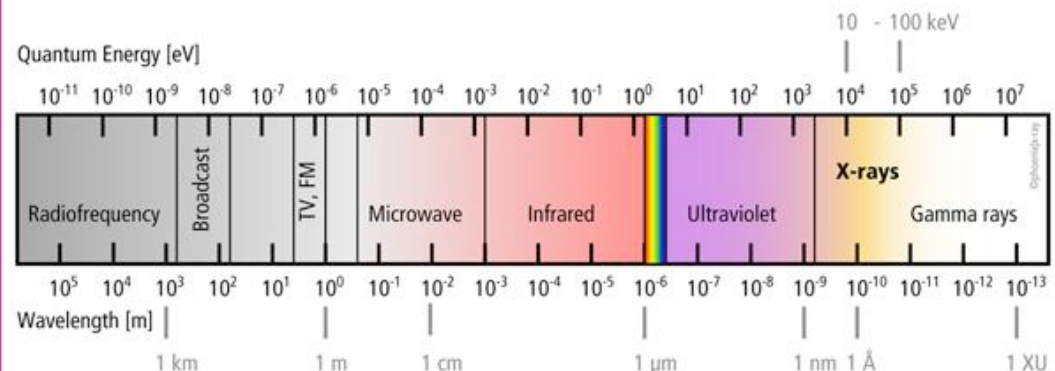
**Fig. 40-14** An electron of kinetic energy  $K_0$  passing near an atom in the target may generate an x-ray photon, the electron losing part of its energy in the process. The continuous x-ray spectrum arises in this way.

The minimum possible x-ray wavelength:

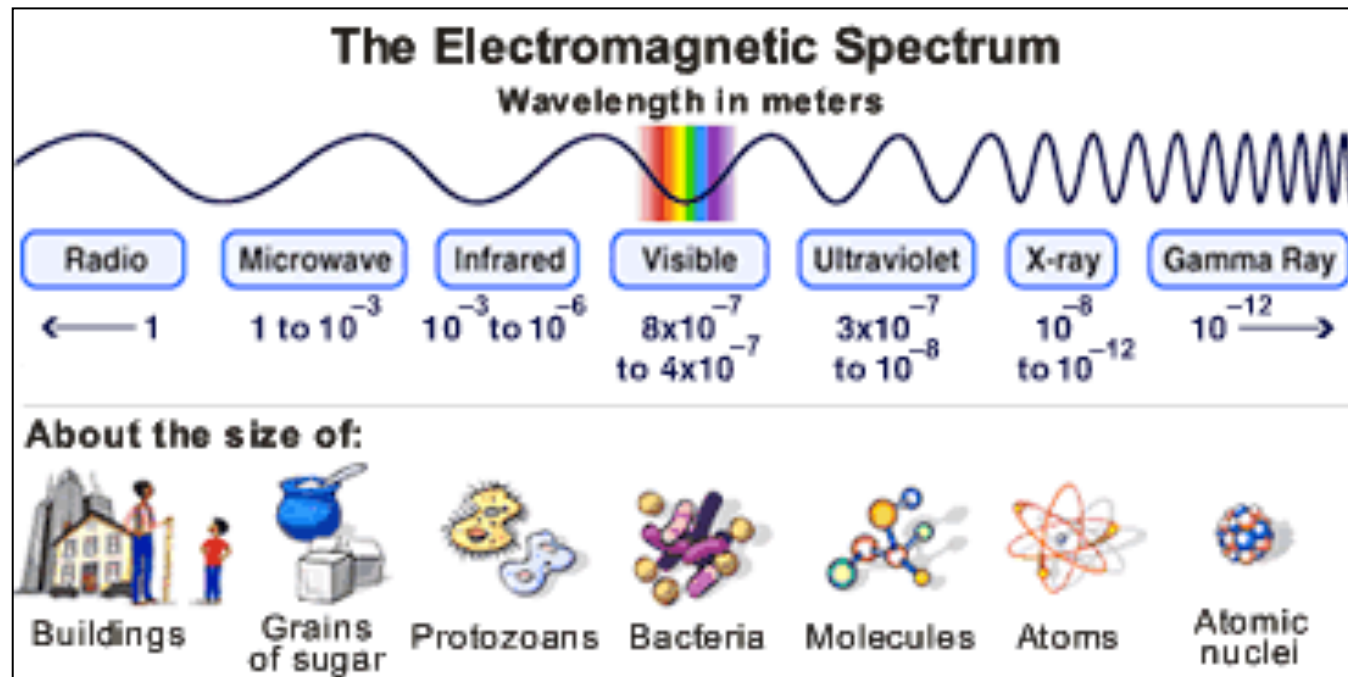
$$K_0 = hf = \frac{hc}{\lambda_{\min}},$$

$$\lambda_{\min} = \frac{hc}{K_0} \quad (\text{cutoff wavelength}).$$

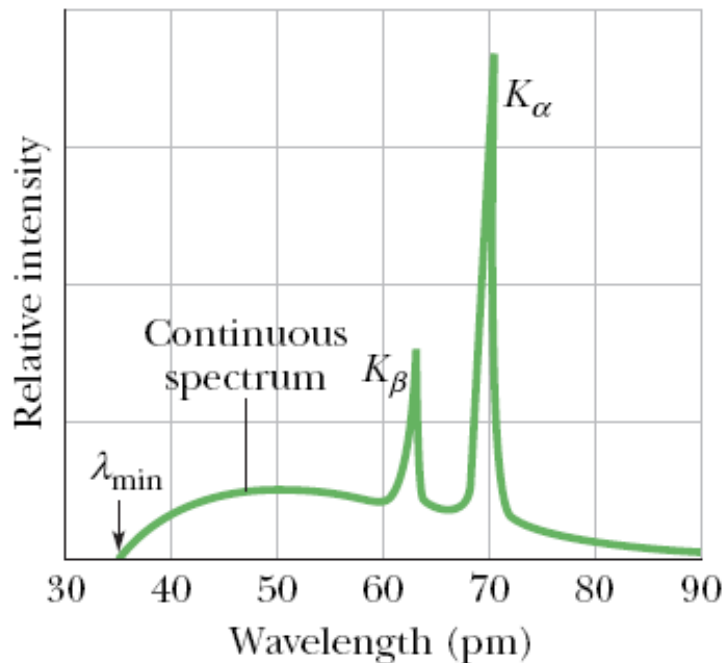
## The Electromagnetic Spectrum



# 40.10: X-Rays and the Ordering of the Elements



## 40.10: X Rays and the Ordering of the Elements: The Characteristic X-ray Spectrum

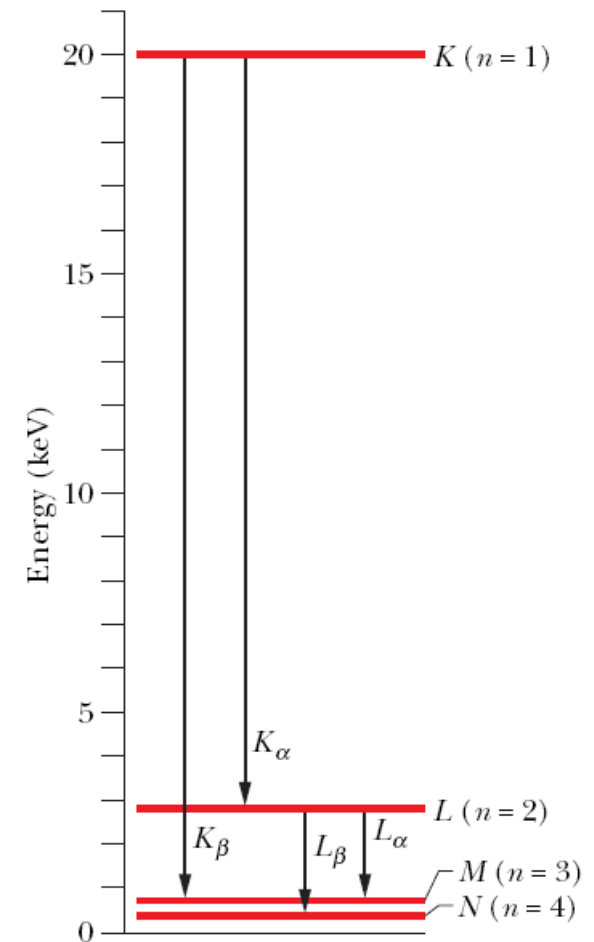


**Fig. 40-13** The distribution by wavelength of the x rays produced when 35 keV electrons strike a molybdenum target

### Two-part process:

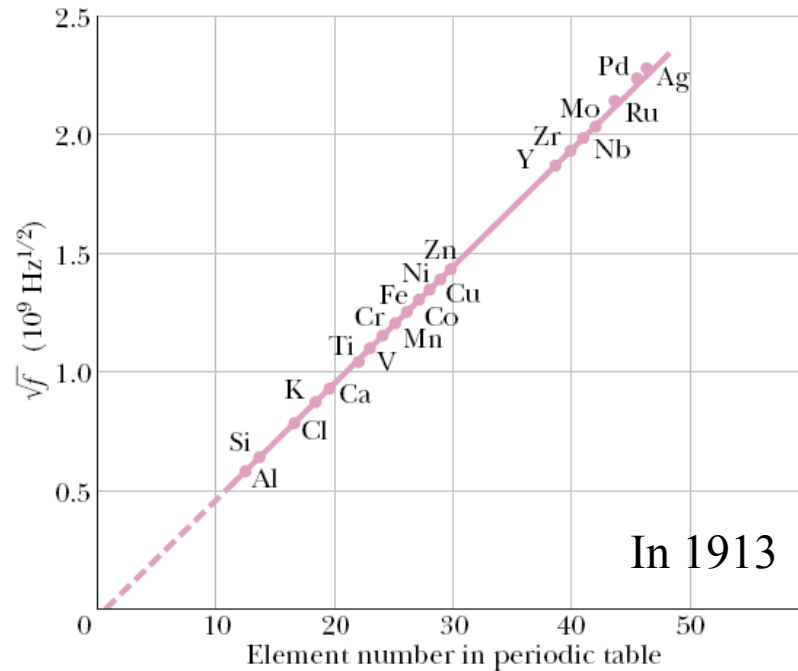
(1) An energetic electron strikes an atom in the target and, while it is being scattered, the incident electron knocks out one of the atom's deep-lying (low  $n$  value) electrons. If the deep-lying electron is in the shell defined by  $n = 1$  (called, for historical reasons, the K shell), there remains a vacancy, or hole, in this shell.

(2) An electron in one of the shells with a higher energy jumps to the K shell, filling the hole in this shell.



**Fig. 40-15** A simplified energy-level diagram for a molybdenum atom, showing the transitions (of holes rather than electrons) that give rise to some of the characteristic x rays of that element. Each horizontal line represents the energy of the atom with a hole (a missing electron) in the shell indicated.

# 40.10: Accounting for Moseley's Plot



**Fig. 40-16** A Moseley plot of the  $K_\alpha$  line of the characteristic x-ray spectra of 21 elements. The frequency is calculated from the measured wavelength.

For hydrogen atom:

$$E_n = -\frac{me^4}{8\epsilon_0^2 h^2} \frac{1}{n^2} = -\frac{13.60 \text{ eV}}{n^2}, \quad \text{for } n = 1, 2, 3, \dots$$

For one of the two innermost electrons in the  $K$  shell of a multi-electron atom, because of the presence of the other  $K$ -shell electron, it “sees” an effective nuclear charge of approximately  $(Z - 1)e$ , where  $e$  is the electronic charge. Therefore the effective energy of the atom is:

$$E_n = -\frac{(13.60 \text{ eV})(Z - 1)^2}{n^2}.$$

Therefore,

$$\begin{aligned} \Delta E &= E_2 - E_1 \\ &= \frac{-(13.60 \text{ eV})(Z - 1)^2}{2^2} - \frac{-(13.60 \text{ eV})(Z - 1)^2}{1^2} \\ &= (10.2 \text{ eV})(Z - 1)^2. \end{aligned}$$

And,

$$\begin{aligned} f &= \frac{\Delta E}{h} = \frac{(10.2 \text{ eV})(Z - 1)^2}{(4.14 \times 10^{-15} \text{ eV} \cdot \text{s})} \\ &= (2.46 \times 10^{15} \text{ Hz})(Z - 1)^2. \end{aligned}$$



$$\sqrt{f} = CZ - C,$$

$C$  is a constant  $(= 4.96 \times 10^7 \text{ Hz}^{1/2})$ .



## Example, Characteristic spectrum in x-ray production:

A cobalt target is bombarded with electrons, and the wavelengths of its characteristic x-ray spectrum are measured. There is also a second, fainter characteristic spectrum, which is due to an impurity in the cobalt. The wavelengths of the  $K_\alpha$  lines are 178.9 pm (cobalt) and 143.5 pm (impurity), and the proton number for cobalt is  $Z_{\text{Co}} = 27$ . Determine the impurity using only these data.

### KEY IDEA

The wavelengths of the  $K_\alpha$  lines for both the cobalt (Co) and the impurity (X) fall on a  $K_\alpha$  Moseley plot, and Eq. 40-27 is the equation for that plot.

**Calculations:** Substituting  $c/\lambda$  for  $f$  in Eq. 40-27, we obtain

$$\sqrt{\frac{c}{\lambda_{\text{Co}}}} = CZ_{\text{Co}} - C \quad \text{and} \quad \sqrt{\frac{c}{\lambda_{\text{X}}}} = CZ_{\text{X}} - C.$$

Dividing the second equation by the first neatly eliminates  $C$ , yielding

$$\sqrt{\frac{\lambda_{\text{Co}}}{\lambda_{\text{X}}}} = \frac{Z_{\text{X}} - 1}{Z_{\text{Co}} - 1}.$$

Substituting the given data yields

$$\sqrt{\frac{178.9 \text{ pm}}{143.5 \text{ pm}}} = \frac{Z_{\text{X}} - 1}{27 - 1}.$$

Solving for the unknown, we find that

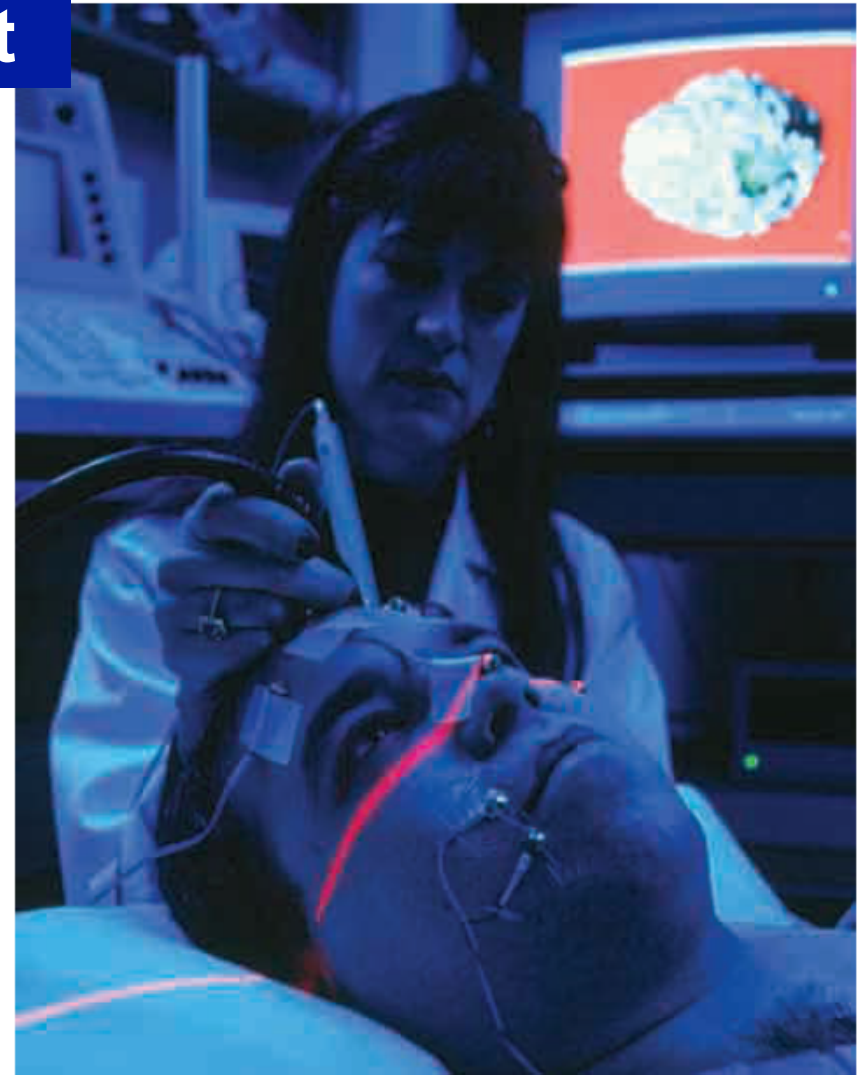
$$Z_{\text{X}} = 30.0. \quad (\text{Answer})$$

Thus, the number of protons in the impurity nucleus is 30, and a glance at the periodic table identifies the impurity as zinc. Note that with a larger value of  $Z$  than cobalt, zinc has a smaller value of the  $K_\alpha$  line. This means that the energy associated with that jump must be greater in zinc than cobalt.



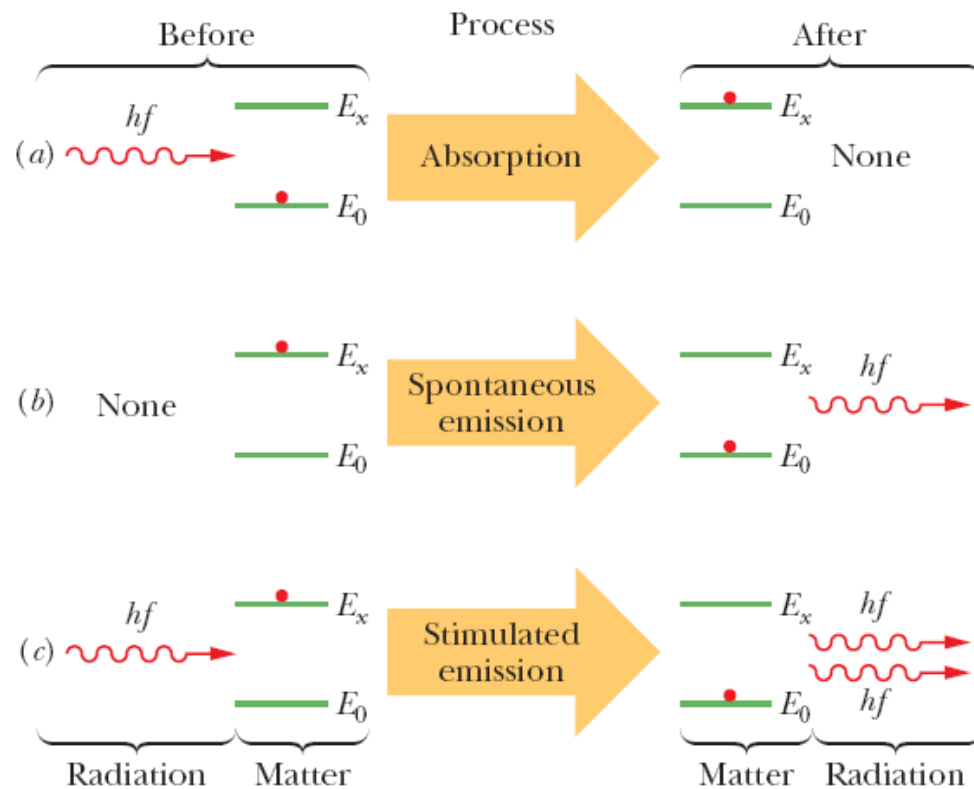
## 40.11: Lasers and Laser Light

1. *Laser light is highly monochromatic.*
2. *Laser light is highly coherent.*
3. *Laser light is highly directional.*
4. *Laser light can be sharply focused.*



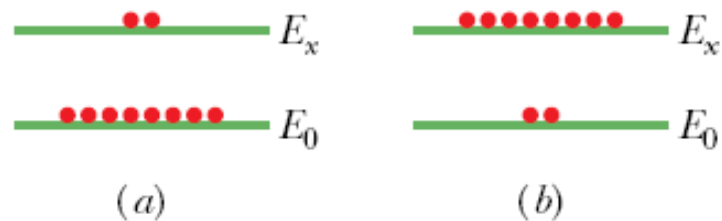
**Fig. 40-17** A patient's head is scanned and mapped by (red) laser light in preparation for brain surgery. During the surgery, the laser-derived image of the head will be superimposed on the model of the brain shown on the monitor, to guide the surgical team into the region shown in green on the model. (Sam Ogden/Photo Researchers)

# 40.12: How Lasers Work



**Fig. 40-18** The interaction of radiation and matter in the processes of (a) absorption, (b) spontaneous emission, and (c) stimulated emission. An atom (matter) is represented by the red dot; the atom is in either a lower quantum state with energy  $E_0$  or a higher quantum state with energy  $E_x$ . In (a) the atom absorbs a photon of energy  $hf$  from a passing light wave. In (b) it emits a light wave by emitting a photon of energy  $hf$ . In (c) a passing light wave with photon energy  $hf$  causes the atom to emit a photon of the same energy, increasing the energy of the light wave.

## 40.12: How Lasers Work

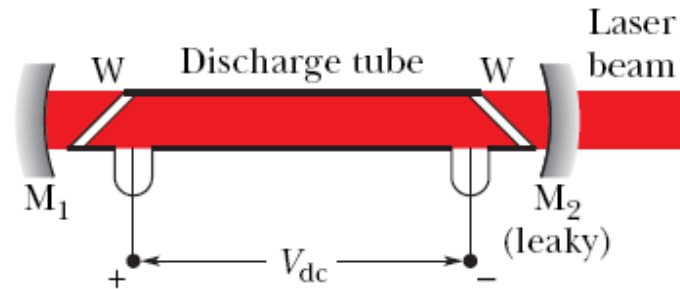


**Fig. 40-19** (a) The equilibrium distribution of atoms between the ground state  $E_0$  and excited state  $E_x$  accounted for by thermal agitation. (b) An inverted population, obtained by special methods. Such a population inversion is essential for laser action.

If the atoms of Fig. 40-19a are flooded with photons of energy  $(E_x - E_0)$ , photons will disappear via absorption by ground-state atoms and photons will be generated largely via stimulated emission of excited-state atoms. Thus, because there are more atoms in the ground state, the *net effect* will be the absorption of photons.

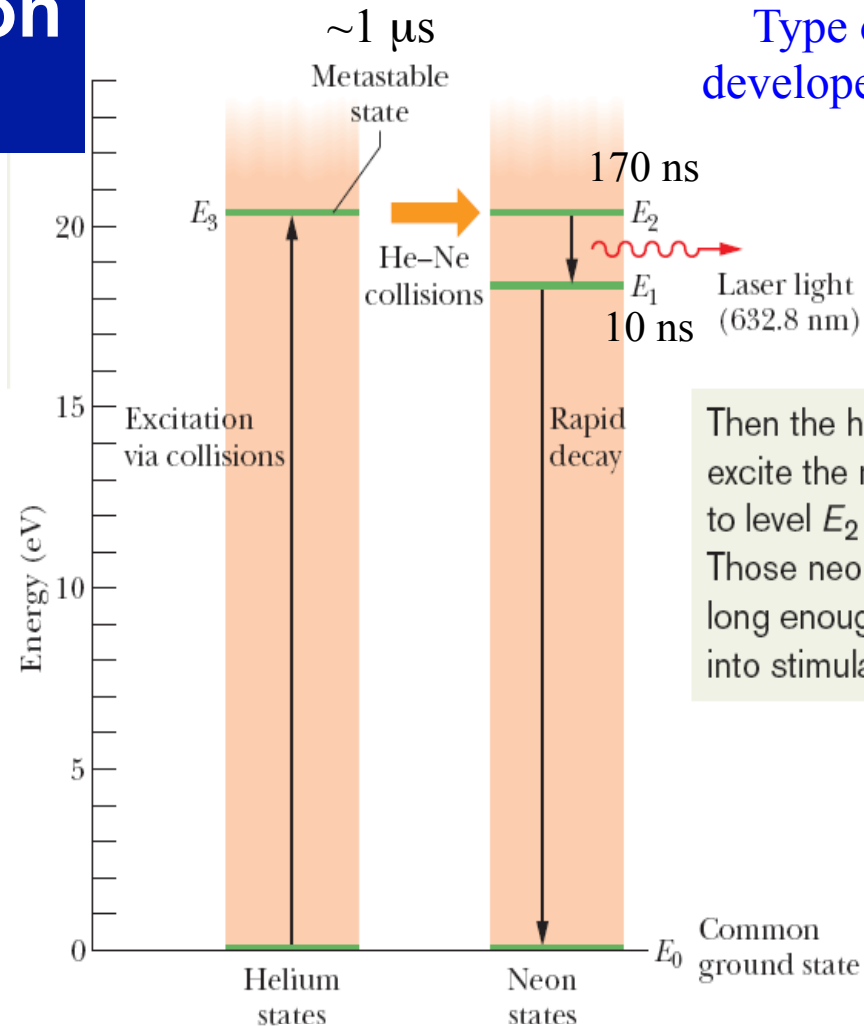
To produce laser light, one must have **more photons emitted than absorbed**; that is, one must have a situation in which stimulated emission dominates. Thus, one needs more atoms in the excited state than in the ground state, as in Fig. 40-19b.

## 40.12: The Helium-Neon Gas Laser



**Fig. 40-20** The elements of a helium–neon gas laser. An applied potential  $V_{dc}$  sends electrons through a discharge tube containing a mixture of helium gas and neon gas. Electrons collide with helium atoms, which then collide with neon atoms, which emit light along the length of the tube. The light passes through transparent windows  $W$  and reflects back and forth through the tube from mirrors  $M_1$  and  $M_2$  to cause more neon atom emissions. Some of the light leaks through mirror  $M_2$  to form the laser beam.

He : Ne = 20 : 80



Type of laser  
developed in 1961

Then the helium atoms excite the neon atoms to level  $E_2$  by collisions. Those neon atoms stay long enough to be forced into stimulated emission.

**Fig. 40-21** Five essential energy levels for helium and neon atoms in a helium–neon gas laser. Laser action occurs between levels  $E_2$  and  $E_1$  of neon when more atoms are at the  $E_2$  level than at the  $E_1$  level.

## Example, Population inversion in a laser:

In the helium–neon laser of Fig. 40-20, laser action occurs between two excited states of the neon atom. However, in many lasers, laser action (*lasing*) occurs between the ground state and an excited state, as suggested in Fig. 40-19*b*.

(a) Consider such a laser that emits at wavelength  $\lambda = 550$  nm. If a population inversion is not generated, what is the ratio of the population of atoms in state  $E_x$  to the population in the ground state  $E_0$ , with the atoms at room temperature?

### KEY IDEAS

(1) The naturally occurring population ratio  $N_x/N_0$  of the two states is due to thermal agitation of the gas atoms (Eq. 40-29):

$$N_x/N_0 = e^{-(E_x - E_0)/kT}. \quad (40-30)$$

To find  $N_x/N_0$  with Eq. 40-30, we need to find the energy separation  $E_x - E_0$  between the two states. (2) We can obtain  $E_x - E_0$  from the given wavelength of 550 nm for the lasing between those two states.

**Calculation:** The lasing wavelength gives us

$$\begin{aligned} E_x - E_0 &= hf = \frac{hc}{\lambda} \\ &= \frac{(6.63 \times 10^{-34} \text{ J}\cdot\text{s})(3.00 \times 10^8 \text{ m/s})}{(550 \times 10^{-9} \text{ m})(1.60 \times 10^{-19} \text{ J/eV})} \\ &= 2.26 \text{ eV}. \end{aligned}$$

To solve Eq. 40-30, we also need the mean energy of thermal agitation  $kT$  for an atom at room temperature (assumed to be 300 K), which is

$$kT = (8.62 \times 10^{-5} \text{ eV/K})(300 \text{ K}) = 0.0259 \text{ eV},$$

in which  $k$  is Boltzmann's constant.

Substituting the last two results into Eq. 40-30 gives us the population ratio at room temperature:

$$\begin{aligned} N_x/N_0 &= e^{-(2.26 \text{ eV})/(0.0259 \text{ eV})} \\ &\approx 1.3 \times 10^{-38}. \end{aligned} \quad (\text{Answer})$$

This is an extremely small number. It is not unreasonable, however. Atoms with a mean thermal agitation energy of only 0.0259 eV will not often impart an energy of 2.26 eV to another atom in a collision.

(b) For the conditions of (a), at what temperature would the ratio  $N_x/N_0$  be 1/2?

**Calculation:** Now we want the temperature  $T$  such that thermal agitation has bumped enough neon atoms up to the higher-energy state to give  $N_x/N_0 = 1/2$ . Substituting that ratio into Eq. 40-30, taking the natural logarithm of both sides, and solving for  $T$  yield

$$\begin{aligned} T &= \frac{E_x - E_0}{k(\ln 2)} = \frac{2.26 \text{ eV}}{(8.62 \times 10^{-5} \text{ eV/K})(\ln 2)} \\ &= 38\,000 \text{ K}. \end{aligned} \quad (\text{Answer})$$