

Thermodynamic Processes Solutions

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1 Cycle Through the Cycles

This problem will develop a useful reference: a list of all quantities associated with thermodynamic processes of ideal gases. Suppose that there are N molecules of an ideal gas with d degrees of freedom (use $\gamma = \frac{d+2}{d}$ where it is more convenient). Suppose the gas starts at (P_0, V_0) . Then $T_0 = P_0 V_0 / (Nk)$. Complete the following table and *draw each process on a P - V diagram*.

Table 1: This table is also available in the workbook on pg. 153.

Quantity	Isochoric	Isovolumetric	Isothermal	Adiabatic
P_f	P_0	P_f (given)	$P_0 \frac{V_0}{V_f}$	$P_0 \left(\frac{T_f}{T_0} \right)^{\frac{\gamma}{\gamma-1}}$
V_f	V_f (given)	V_0	V_f (given)	$V_0 \left(\frac{T_0}{T_f} \right)^{\frac{1}{\gamma-1}}$
T_f	$T_0 \frac{V_f}{V_0}$	$T_0 \frac{P_f}{P_0}$	T_0	T_f (given)
ΔE	$\frac{d}{2} N k_B T_0 \left(\frac{V_f}{V_0} - 1 \right)$	$\frac{d}{2} N k_B T_0 \left(\frac{P_f}{P_0} - 1 \right)$	0 J	$\frac{d}{2} N k_B (T_f - T_0)$
Q	$N k_B T_0 \left(\frac{V_f}{V_0} - 1 \right) \left(\frac{d}{2} - 1 \right)$	$\frac{d}{2} N k_B T_0 \left(\frac{P_f}{P_0} - 1 \right)$	$N k_B T_0 \ln \left(\frac{V_f}{V_0} \right)$	0
W	$P_0 (V_f - V_0)$	0 J	$N k_B T_0 \ln \left(\frac{V_f}{V_0} \right)$	$-\frac{d}{2} N k_B (T_f - T_0)$
ΔS			$N k_B \ln \left(\frac{V_f}{V_0} \right)$	0

2 Problems

2.1 Heat from the Ocean

- (a) Maximum efficiency is through Carnot's theorem combined with the efficiency of a Carnot cycle to get

$$e = 1 - \frac{T_L}{T_H} \quad (1)$$

So in this case evaluating yields

$$\begin{aligned} e &= 1 - (273.15 + 4K)/(273.15 + 22K) \\ &\approx 0.061 \end{aligned}$$

Thus, the maximum theoretical efficiency would be about 6.1%.

- (b) The efficiency can also be written as

$$e = \frac{W}{Q_H} \quad (2)$$

Producing GW of power means 1 GJ of work is being done every second since $[W] = \frac{[J]}{[s]}$. Inverting the above equation, $Q_H = \frac{W}{e}$, so the heat per second is $\frac{dQ_H}{dt} = \frac{P}{e}$. The heat input is related to the volume by

$$\begin{aligned} Q &= mc\Delta T \\ &= \rho cV\Delta T \\ V &= \frac{Q}{\rho c\Delta T} \\ \frac{dV}{dt} &= \frac{P}{e\rho c\Delta T} \end{aligned}$$

Plugging in numbers, the volume processed in one second is

$$\begin{aligned} \frac{VolumeProcessed}{1s} &= \frac{1 \times 10^9 J/s}{(0.061)(1000kg/m^3)(480J/kg * K)(22 - 4K)} \\ &\approx 218m^3 \\ &= 2.18 \times 10^5 L \end{aligned}$$

2.2 Challenge: Adiabatic Atmosphere

From the Ideal Gas Law, $PV = nRT = NkT$. Notice that $\rho = \frac{Nm}{V}$, so

$$P = \rho \frac{kT}{m} \quad (3)$$

Differentiating the given equation of $P\rho^{-\gamma} = \text{constant}$, we have

$$\begin{aligned}\frac{dP}{dh}\rho^{-\gamma} - \gamma\rho^{-\gamma-1}\frac{d\rho}{dh}P &= 0 \\ \frac{dP}{dh} - \gamma\rho^{-1}\frac{d\rho}{dh}P &= 0 \\ \frac{d\rho}{dh} &= \frac{\rho}{P}\frac{1}{\gamma}\frac{dP}{dh}\end{aligned}$$

From Eq. ??, $\frac{\rho}{P} = \frac{m}{kT}$, so

$$\frac{d\rho}{dh} = \frac{m}{kT}\frac{1}{\gamma}\frac{dP}{dh} \quad (4)$$

Now, differentiating Eq. ??,

$$\frac{dP}{dh} = \frac{kT}{m}\frac{d\rho}{dh} + \frac{\rho k}{m}\frac{dT}{dh} \quad (5)$$

Plugging into Eq. ??, we have

$$\begin{aligned}\frac{dP}{dh} &= \frac{kT}{m}\left(\frac{m}{kT}\frac{1}{\gamma}\frac{dP}{dh}\right) + \frac{\rho k}{m}\frac{dT}{dh} \\ \frac{\rho k}{m}\frac{dT}{dh} &= \frac{dP}{dh}\left(1 - \frac{1}{\gamma}\right)\end{aligned}$$

Now, for any substance, the hydrodynamic condition posits that $\frac{dP}{dh} = -\rho g$ (e.g., think of pressure in a lake). Plugging this into the above equation,

$$\begin{aligned}\frac{dT}{dh} &= \frac{m}{\rho k}(-\rho g)\left(\frac{\gamma-1}{\gamma}\right) \\ &= -\frac{mg}{k}\frac{\gamma-1}{\gamma}\end{aligned}$$

If we assume the atmosphere is monoatomic, $\gamma = \frac{5}{3}$, so $\frac{\gamma-1}{\gamma} = \frac{2}{5}$ and we have

$$\frac{dT}{dh} = -2mg/5k \quad (6)$$