

F work

- ① Write the electronic configuration of Pd?
- ② Why ~~not~~ from  $\text{PCl}_5$ ? but  $\text{NCl}_5$  not form.
- ③ Fluorine is a stronger oxidising agent than chlorine why?
- ④ Arrange ↑ order acidic strength of Seven hydrogen halide?  
 $\text{HF}, \text{HCl}, \text{HBr}, \text{HI}$
- ⑤ why  $\text{SnCl}_4$  is more covalent in nature compare than  $\text{SnCl}_2$ ?
- ⑥ why fluorine exhibit -1 oxidation state?
- ⑦ following arrange increasing oxidation state  
 $\text{HClO}, \text{HClO}_2, \text{HClO}_3, \text{HClO}_4$
- ⑧ why is  $\text{ICl}$  more reactive than  $\text{I}_2$ ?
- ⑨ draw the structure:  
(i)  $\text{XeOF}_4$  (ii)  $\text{ClF}_3$  (iii)  $\text{IF}_7$  (iv)  $\text{XeO}_3$
- ⑩ draw the structure:  
(i)  $\text{H}_2\text{S}_2\text{O}_7$  (ii)  $\text{H}_2\text{S}_2\text{O}_7$
- ⑪ arrange increasing acidic character  
 $\text{H}_2\text{S}, \text{H}_2\text{Te}, \text{H}_2\text{O}$
- ⑫ How is  $\text{SO}_2$  an air pollutant