Chem Tutorials

Question 6

Moles of EDTA

$0.03212 \cdot 0.0254$=8.15848×10^{-4}= moles of Fe and Cu.

Moles of Cu

$0.01086 \cdot 0.00254$=2.75844×10^{-5}= moles of Cu.

Moles of Fe

$8.15848 \times 10^{-4} - 2.75844 \times 10^{-5}$=$7.882636 \times 10^{-4}$= moles of Fe

Moles of Cu in sample

$4 \cdot 2.75844 \times 10^{-5}$=$1.103376 \times 10^{-4}$= moles of Cu

Mass of Cu in sample

$2.75844 \times 10^{-5} \cdot 63.546 = 0.0070115g$

Mass of Fe in sample

$7.882636 \times 10^{-4} \cdot 55.845 = 0.1760823g$

Mass percentage

Mass percentage of Cu =$\frac{0.1760823}{0.5674}$ =1.236%
Mass percentage of Fe = $\frac{0.0070115}{0.5674}$ =31.03%