

Homework 8

PHYSICS 461
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1. Starting from the radial Schrodinger equation

$$-\frac{\hbar^2}{2m_e} \left(R'' + \frac{2}{r} R' \right) - \frac{\ell(\ell+1)}{r^2} R - \frac{\phi^2}{r} R(r) = ER$$

Letting the constants go to 1,

$$\frac{1}{2} R'' + \frac{1}{2} R' + \left[\frac{1}{r} - \frac{\ell(\ell+1)}{2r^2} \right] R = -ER$$

For $R(r) = e^{-r}$ and $\ell = 0$ and $E = -1/2$, the differential equation becomes

$$\frac{1}{2} R(r) - \frac{1}{r} R(r) + \frac{1}{r} R(r) = \frac{R(r)}{2}$$

...which is true.

2. (a) The Rydberg constant is $207\times$ larger, so the wavelength is

$$\lambda = \frac{1}{207 \times R_\infty} = 440.2 \text{ pm}$$

This is in the x-ray to gamma ray region.

- (b) Using the reduced mass,

$$R_\mu = \left(\frac{m_\mu m_p}{m_\mu + m_p} \right) \frac{e^4}{8\epsilon_0^2 h^3 c}$$
$$\lambda = R^{-1} = 481.9 \text{ pm}$$

It's pretty close to the infinite nuclear mass approximation.

3. (a) Using R_μ from Problem 2, we can equate the energy to the Coulomb potential as

$$E = R_\mu h c \frac{Z^2}{n^2} = \frac{Z e^2}{4\pi\epsilon_0 r}$$
$$r = \frac{e^2}{4\pi\epsilon_0 R_\mu Z h c} = 6.89 \text{ fm}$$

- (b) For muonic hydrogen, $Z = 1$ and then from the equation used in (a), the radius is 564 fm.

4. (a) The energy levels are given by

$$E_n = -\frac{\mu e^4}{8h^2 \epsilon_0^2} \frac{1}{n^2} = -\frac{6.8 \text{ eV}}{n^2}$$

For $n = 1, 2, 3$,

$$E_1 = -6.8 \text{ eV}$$

$$E_2 = -1.7 \text{ eV}$$

$$E_3 = -0.75 \text{ eV}$$

(b) For the α and β lines,

$$E_\alpha = E_2 - E_1 = 5.1 \text{ eV}$$

$$\lambda_\alpha = \frac{hc}{E_2 - E_1} = 243 \text{ nm}$$

$$\lambda_\beta = \frac{hc}{E_3 - E_1} = 204 \text{ nm}$$

5. (a) The core electrons shield (p 218) the nuclear charge completely.
 (b) We can equate the energy to the Coulomb potential,

$$\begin{aligned} \frac{Ry}{2000^2} &= \frac{e^2}{4\pi\epsilon_0 r} \\ r &= \frac{e^2 2000^2}{4\pi\epsilon_0 Ry} \\ &= 427 \mu\text{m} \end{aligned}$$

The H atom is $5 \times 10^{-11} \text{ m}$, so this is much larger.

(c) From Wikipedia,

$$\begin{aligned} mvr &= n\hbar \\ v &= 1158 \text{ m} \cdot \text{s}^{-1} \end{aligned}$$

(d) For $n = 1$, $v = 2 \times 10^6 \text{ m/s}$, which makes sense as the electron is much closer in orbit.

6. (a) For the inner $2s$ electron, the energy is $E_{2s} = Ry \frac{Z^2}{n^2} = Ry$. For the outer $4p$ electron, the effective charge is 1 and the energy is $E_{4p} = Ry/16$. Adding these, the energy is

$$E_{2s4p} = 1 + 13.61 \text{ eV} \left(1 + \frac{1}{16}\right) = 14.5 \text{ eV}$$

The ground state of He (from Wikipedia) is -79 eV , so the change in energy is $\Delta E = 64.5 \text{ eV}$, corresponding to $\lambda = 19 \text{ nm}$.

(b) Doing this classically, we can take the energy from (a) and set it to $mv^2/2$,

$$\begin{aligned} \Delta E &= \frac{mv^2}{2} \\ v &= \sqrt{\frac{2\Delta E}{m}} = \sqrt{\frac{2 \times 64.5 \text{ eV}}{0.510 \text{ MeV}}} \\ &= 4.8 \times 10^6 \text{ m} \cdot \text{s}^{-1} \end{aligned}$$

7. From the L shell, the electron drops of $L \rightarrow K$ corresponding to the K_α line 3.69 keV . Similarly for the M and N shells, the energies are 0.341 keV and 0.024 keV respectively.
8. The emitted electron is roughly 300 eV , which is roughly the L_α line, which makes sense.
9. I can't get the generalized way to work out, so I'm just going to give an example with the $2p3s$ state. Under LS coupling, we get four terms: 1P_1 , 3P_0 , 3P_1 , 3P_2 . Under jj coupling, there's four terms as well: for $j_1 = j_2 = 1/2$, we get two states $j = 0, 1$. For $j_1 = 3/2, j_2 = 1/2$, we get two other states, $j = 1, 2$.

10. (a) The combined $L = 1$ and $S = 1$, so J goes from 0, 1, 2, and the terms are 1P_1 , 3P_0 , 3P_1 , 3P_2 .
- (b) For $j_1 = 1 + \frac{1}{2} = \frac{3}{2}$, $j_2 = \frac{1}{2}$, we can couple these with $J = j_1 + j_2$ to $|j_1 - j_2|$. For $j_1 = j_2 = 1/2$, we get two states $j = 0, 1$. For $j_1 = 3/2, j_2 = 1/2$, we get two other states, $j = 1, 2$.
- (c) I think this can be explained by the Paschen-Back effect, except instead of an external field, it's related to the moment from the nucleus?