



$$\begin{aligned} \text{C} &= 6 \times 12 = 72 \\ \text{H} &= 12 \times 1 = 12 \\ \text{O} &= 6 \times 16 = 96 \\ &\underline{180 \text{ g}} \end{aligned}$$

$$(856 \text{ g } \text{C}_6\text{H}_{12}\text{O}_6) \times \frac{1 \text{ mol } \text{C}_6\text{H}_{12}\text{O}_6}{180 \text{ g } \text{C}_6\text{H}_{12}\text{O}_6}$$

$$4,75 \text{ mol } \text{C}_6\text{H}_{12}\text{O}_6$$

$$\frac{6 \text{ mol } \text{CO}_2}{1 \text{ mol } \text{C}_6\text{H}_{12}\text{O}_6} \times 4,75 \text{ mol } \text{C}_6\text{H}_{12}\text{O}_6 = 28,5 \text{ mol } \text{CO}_2$$

$$\begin{aligned} \text{C} &= 1 \times 12 = 12 \\ \text{O} &= 2 \times 16 = 32 \\ &\underline{44 \text{ g}} \end{aligned}$$

$$28,5 \text{ mol } \text{CO}_2 \times \frac{44 \text{ g } \text{CO}_2}{1 \text{ mol } \text{CO}_2} = \underline{1254 \text{ g } \text{CO}_2}$$