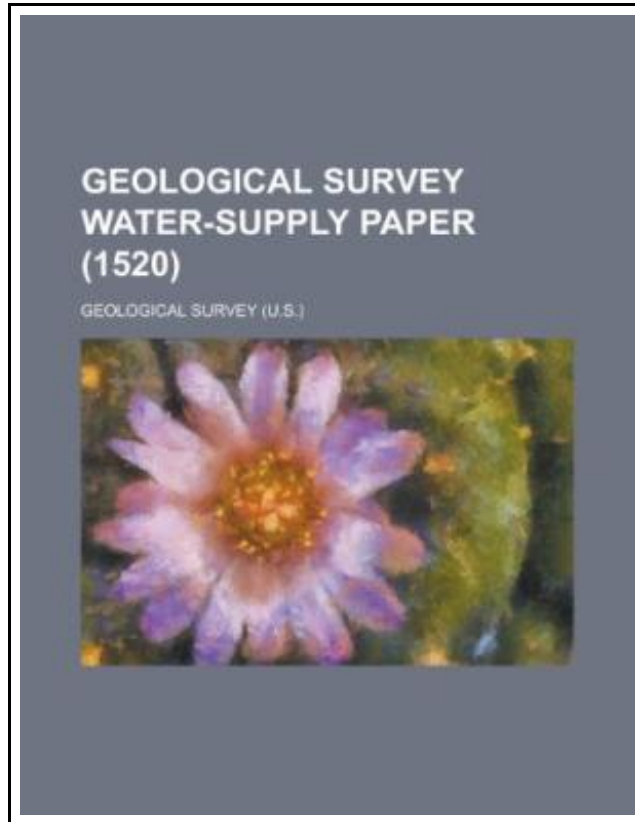


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RareBooksClub. Paperback. Book Condition: New. This item is printed on demand. Paperback. 222 pages. Original publisher: Ada, OK: U. S. Environmental Protection Agency, National Risk Management Research Laboratory, 1999. Subject: In situ remediation. Excerpt: . . . TCE Reductive-Dechlorination The other principal contaminant at the USCG site is TCE. The reductive-dechlorination of chlorinated aliphatics, such as 0 0 TCE, by Fe is thermodynamically favored (Vogel et al. , 1987). The reductive-dechlorination of TCE by Fe to non-toxic hydrocarbon end products occurs by both reductive -elimination and sequential hydrogenolysis (Roberts et al. , 1996; Arnold and Roberts, 1997). The sequential hydrogenolysis pathway results in the production of toxic chlorinated intermediates, such as cis-dichloroethylene (cDCE) and vinyl chloride (VC) (Figure 4). However, less than 10-20 of the TCE mass degrades via this pathway producing cDCE and VC byproducts (Orth and Gillham, 1996), and these byproducts 0 are themselves reductively-dechlorinated by Fe. The major end products of reductive-chlorination of TCE are ethene and ethane. Ethene, ethane and lower concentrations of methane, propene, propane, 1-butene and butane end products have been observed in previous laboratory experiments with granular iron (Orth and Gillham, 1996). The reductive-dechlorination of TCE, cDCE and VC fits a pseudo first-order reaction mechanism. Johnson et al. (1996) 0 describe a pseudo first-order kinetic model for the dehalogenation of various chlorinated hydrocarbons by Fe:
$$\frac{dP}{dt} = -k_a \frac{P}{a} \frac{m}{m_0} \quad (4)$$
 where k is the specific reaction rate constant normalized to the surface area of Fe ($L h m$), a is the surface area of Fe $s a s 2-1 0-1 (m g)$ and is the mass concentration of Fe ($g L$ of solution)....



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