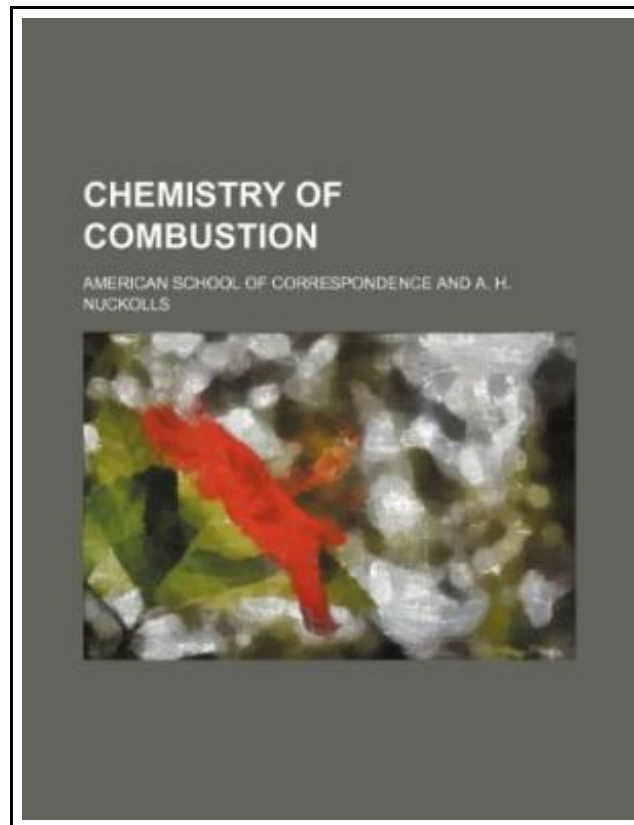


# Chemistry of Combustion



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## ***Reviews***

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***(Mitchell Kuhn III)***

## CHEMISTRY OF COMBUSTION



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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book \*\*\*\*\* Print on Demand \*\*\*\*\*.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1914 Excerpt: .oxidation when exposed to the air; the oxidation heats the phosphorus. The fumes given off are the products of oxidation, and consist of compounds of phosphorus and oxygen. Finally the phosphorus takes fire (if it is desired to stop the experiment cover the phosphorus with an inverted tumbler or cup). Heat carefully a little of the sulphur in the spoon so as to secure a gradual rise in temperature. Try a similar experiment with charcoal. Observe that there is a marked difference in the temperatures at which these substances take fire, that is, enter into active combustion with the oxygen of the air. In the above experiment, the phosphorus took fire very readily, requiring only a slight elevation in temperature to cause its combustion; this elevation in temperature was produced by the heat generated by the oxidation of the phosphorus. The sulphur required a higher temperature to cause combustion, and the charcoal still higher. The first condition, then, necessary to start the combustion of a combustible substance in the presence of air, is that a part of the combustible be at its ignition point, the requisite temperature having been obtained from a flame, spark, hot object, or spontaneous heating. Combustion without Continuous External Heat. When the temperature produced by the combustion of a substance is higher than the ignition point of the substance, such a substance on being ignited will continue to burn, because the heat evolved by the combustion of the portion ignited heats the...



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