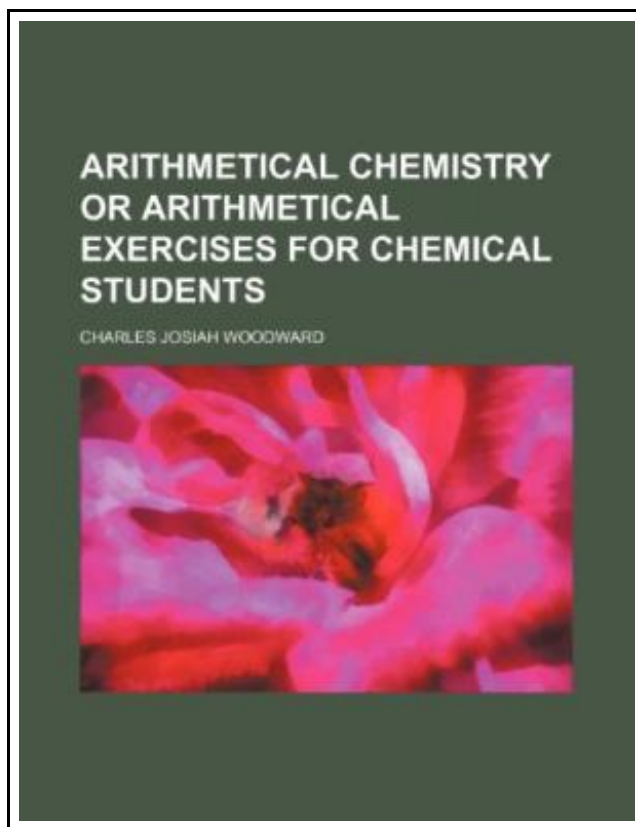


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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book \*\*\*\*\* Print on Demand \*\*\*\*\*.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1883 Excerpt: . Aq = HClAq= 17320 HClAq+ KHOAq= KClAq = 13740 53060 Now heat from the action of Cl on KI solution is simply excess of heat of the one set of changes above that in the other set, i.e.,  $53060 - (32864 + x) = 26209$   $20196 - x = 26209$   $x = -6013$  That is when hydrogen and iodine combine, for every molecule in grammes formed, 6013 units of heat are absorbed. EXERCISE XI. 1. In an experiment to determine the calorific power of charcoal, 5 grammes of charcoal burnt in oxygen to COa raised 3 kgs. of water from 10C to 23-46 C. What is the calorific power of charcoal as determined by this experiment? 2. What is the calorific power of a solid fuel composed of 20 per cent, of hydrogen and 80 per cent, of carbon; a if burnt in oxygen, b if burnt in air? 3. How many pounds of the fuel of question 2 are required when burned in air to completely evaporate 200 gallons of water at 100C. t 4. A coal has the following percentage composition: Carbon 90 34, hydrogen 521, Oxygen 4-45. Calculate its calorific power. From the combustible elements deduct the hydrogen necessary to combine with the ox.vKen to form water, and makefile calculation with the remaindai. 5. 12-8236 parts by weight of a sample of gas from a blast furnace contained, Nitrogen Carbonic anhydride Carbonic oxide Hydrogen Marsh gas Determine calorific power of this mixture. 6. When one gramme of...



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