

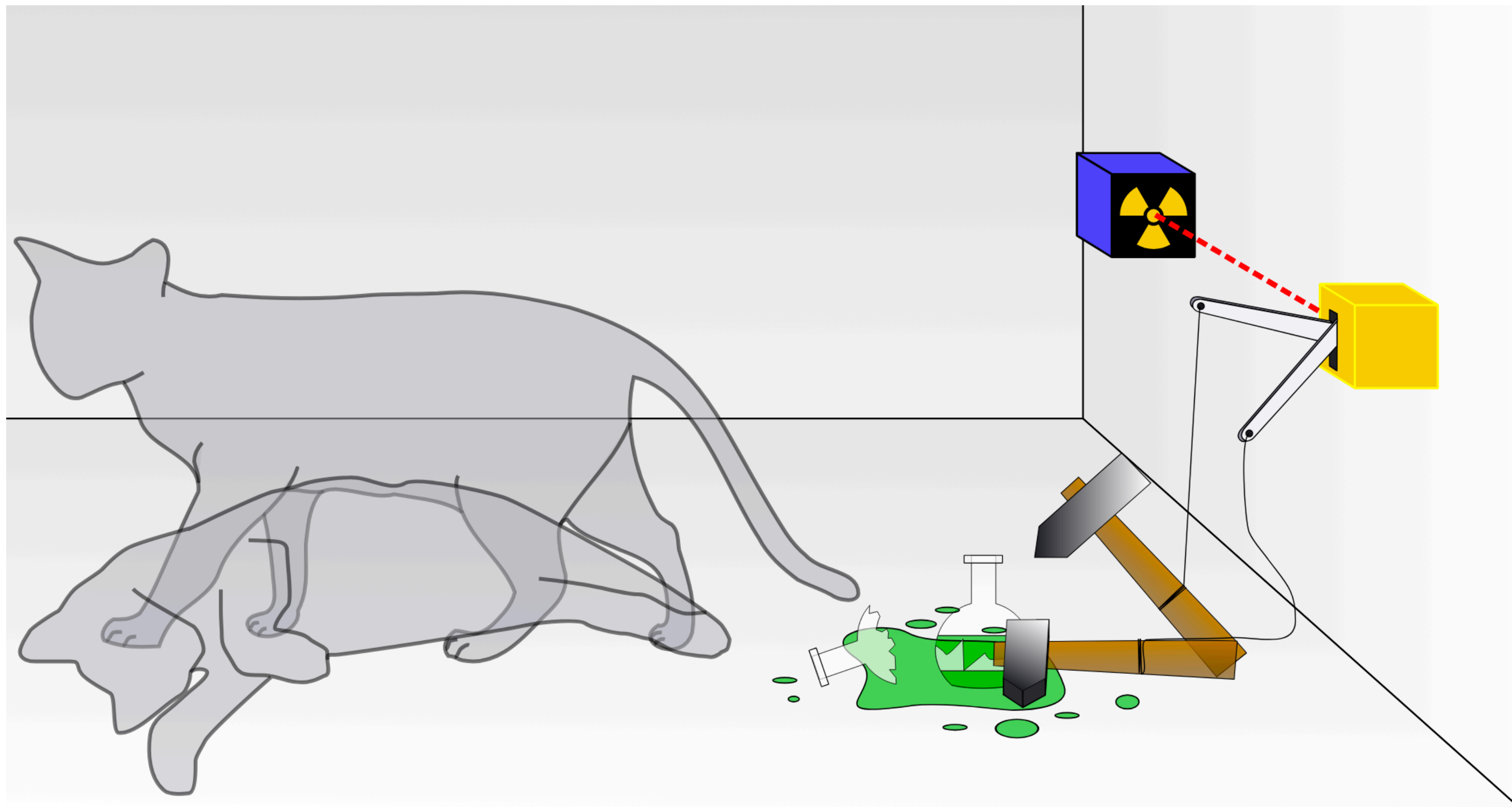
# Simulation 1: Schrödinger Eqn.

**Kevin Nelson**, Jianming Qian, Alexander Takla  
Michigan Math and Science Scholars  
27 July 2023



# Schrödinger's cat

- You may be familiar with Schrödinger's cat, the thought experiment in which a vial of poison is released if a radioactive isotope decays.
- It's said that the cat is both alive and dead until the box is opened. Before opening the cat is in a “quantum superposition” of **both dead and alive**.
- But why?



# Schrödinger's wave equation

$$i\hbar \frac{\partial \Psi}{\partial t} = \left[ \frac{-\hbar^2}{2m} \frac{\partial^2}{\partial x^2} + V \right] \Psi$$



# Schrödinger's wave equation

$$i\hbar \frac{\partial \Psi}{\partial t} = \left[ \frac{-\hbar^2}{2m} \frac{\partial^2}{\partial x^2} + V \right] \Psi$$

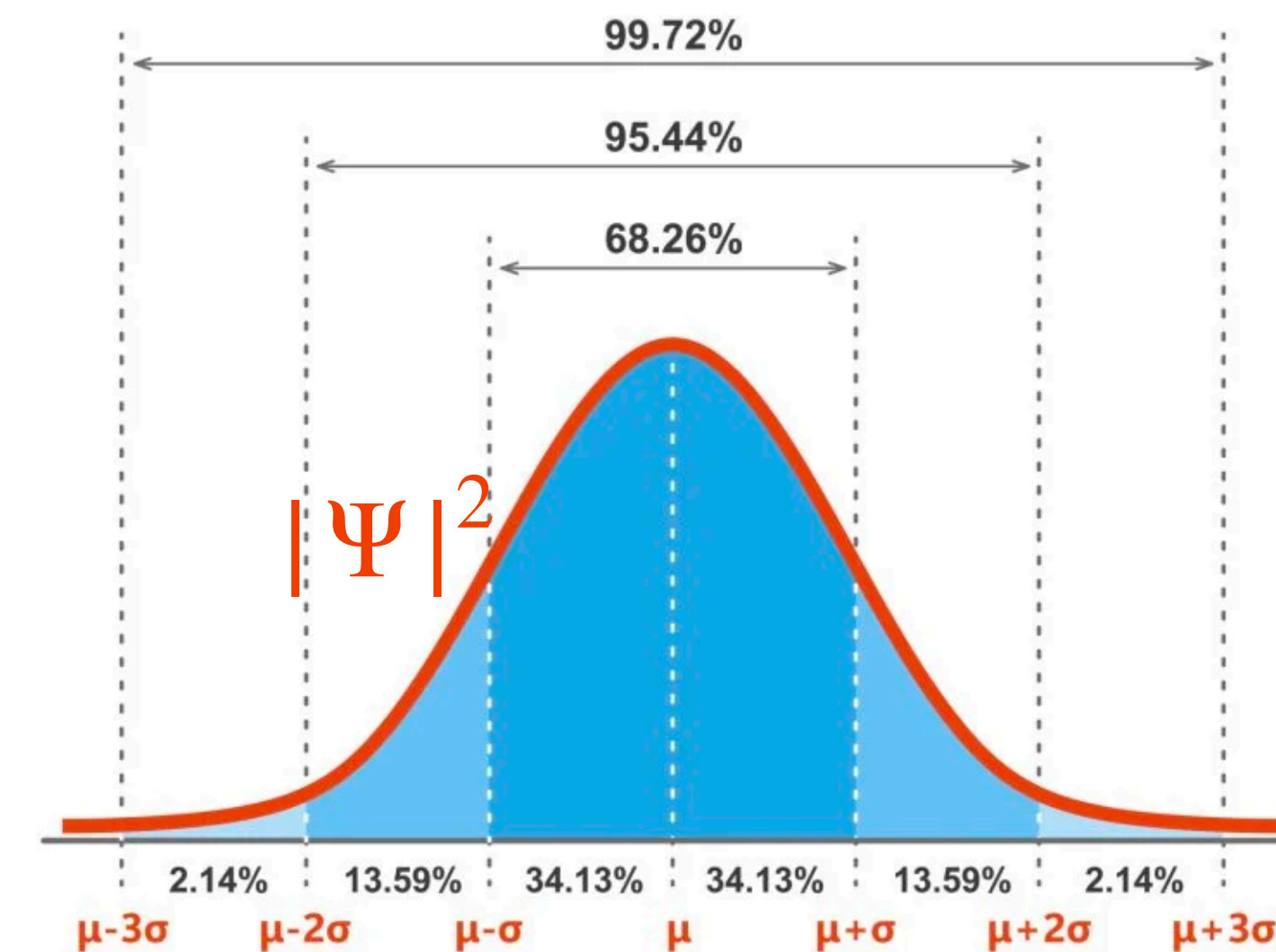
- The Schrödinger Equation describes the quantum wave function of particles.



# Schrödinger's wave equation

$$i\hbar \frac{\partial \Psi}{\partial t} = \left[ \frac{-\hbar^2}{2m} \frac{\partial^2}{\partial x^2} + V \right] \Psi$$

- The Schrödinger Equation describes the quantum wave function of particles.
- Particles do not have a position until it is measured. The square of the wave function tells you the probability of finding the particle in some region

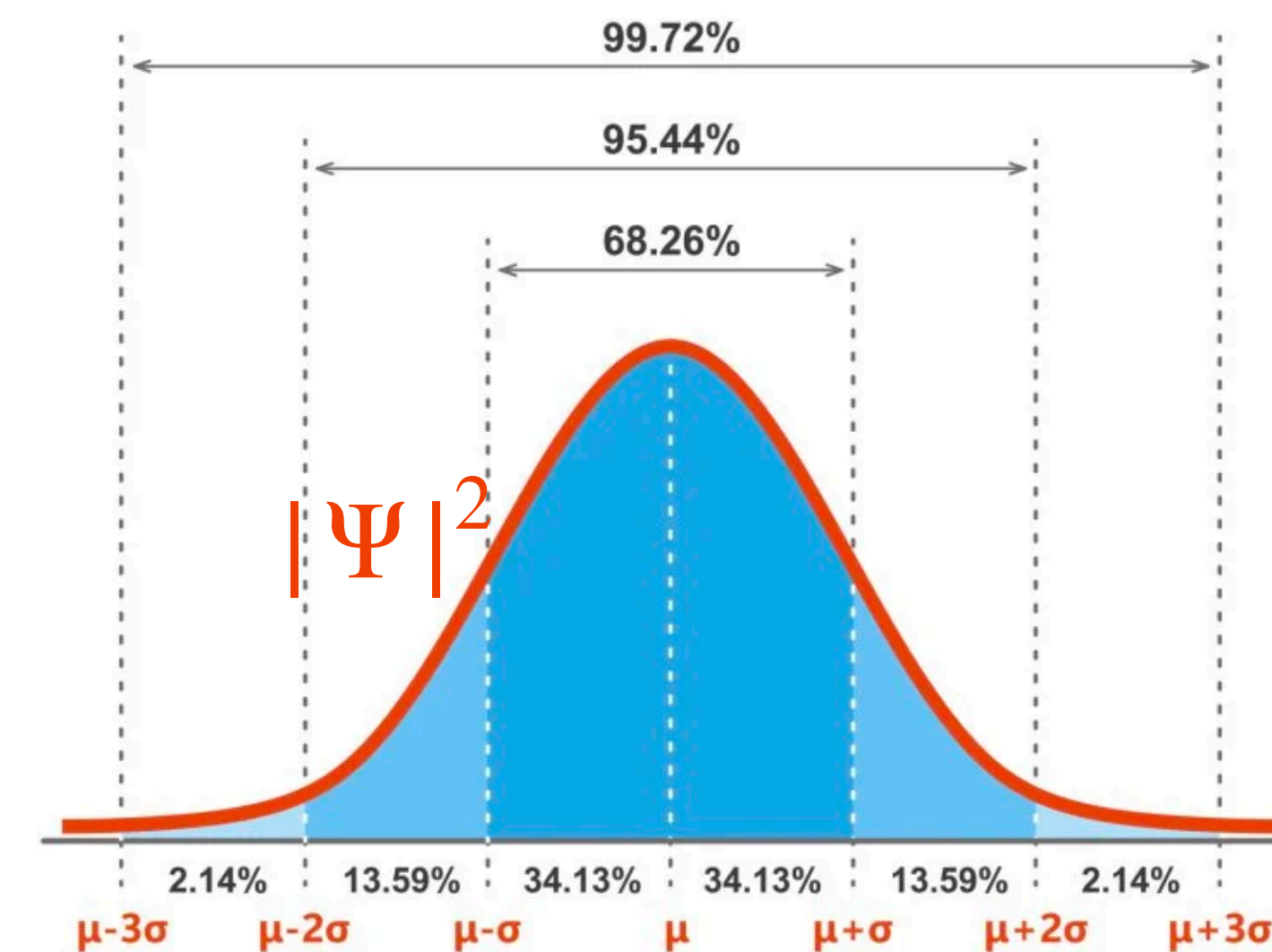




# Schrödinger's wave equation

$$i\hbar \frac{\partial \Psi}{\partial t} = \left[ \frac{-\hbar^2}{2m} \frac{\partial^2}{\partial x^2} + V \right] \Psi$$

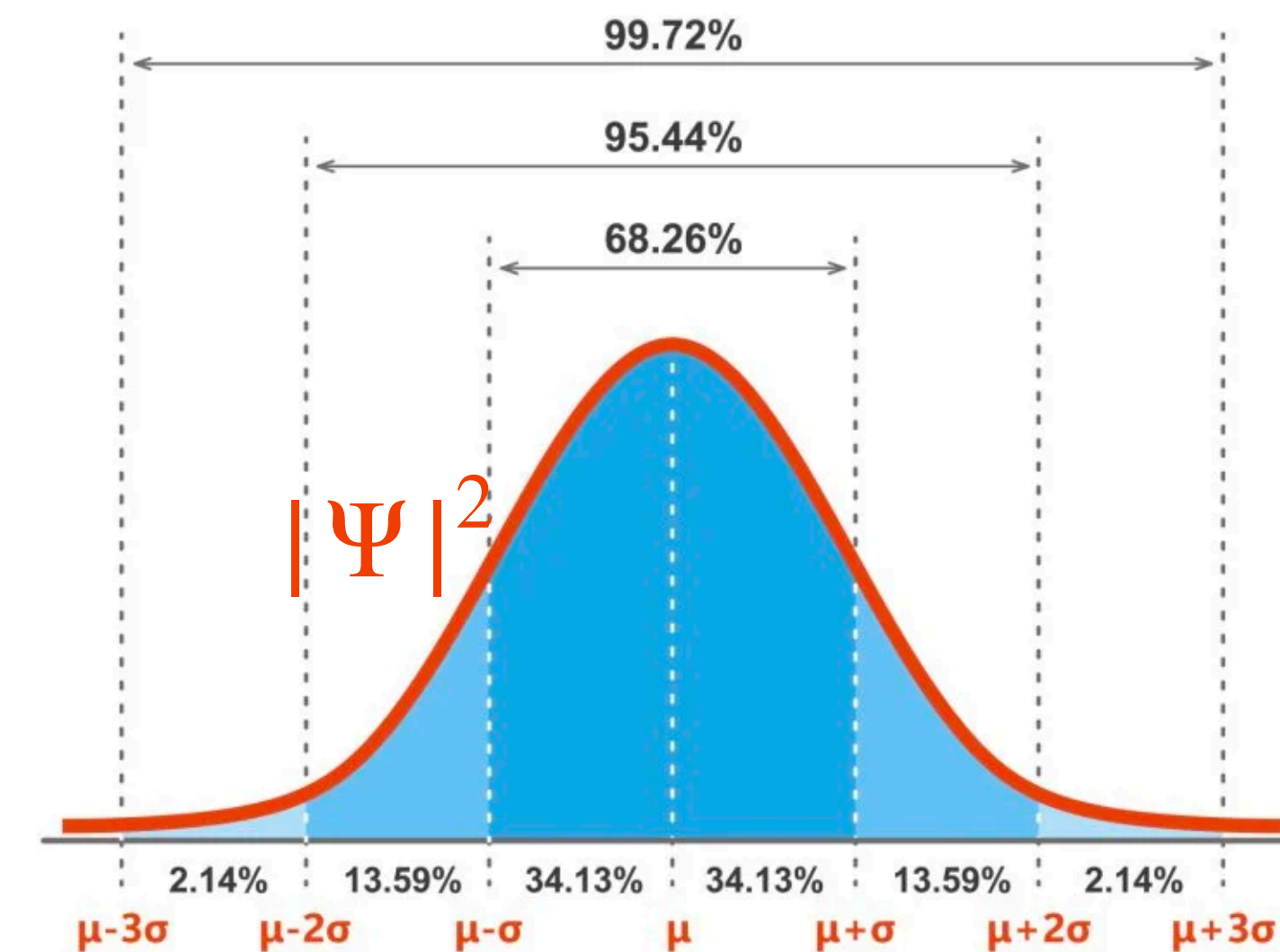
- The Schrödinger Equation describes the quantum wave function of particles.
- Particles do not have a position until it is measured. The square of the wave function tells you the probability of finding the particle in some region
- The constant  $\hbar$  is the smallest possible value of energy. Energy is quantized into tiny packets. You can have  $0, \hbar, 2\hbar, \dots$  but not  $0.13\hbar$  units of energy



# Schrödinger's wave equation

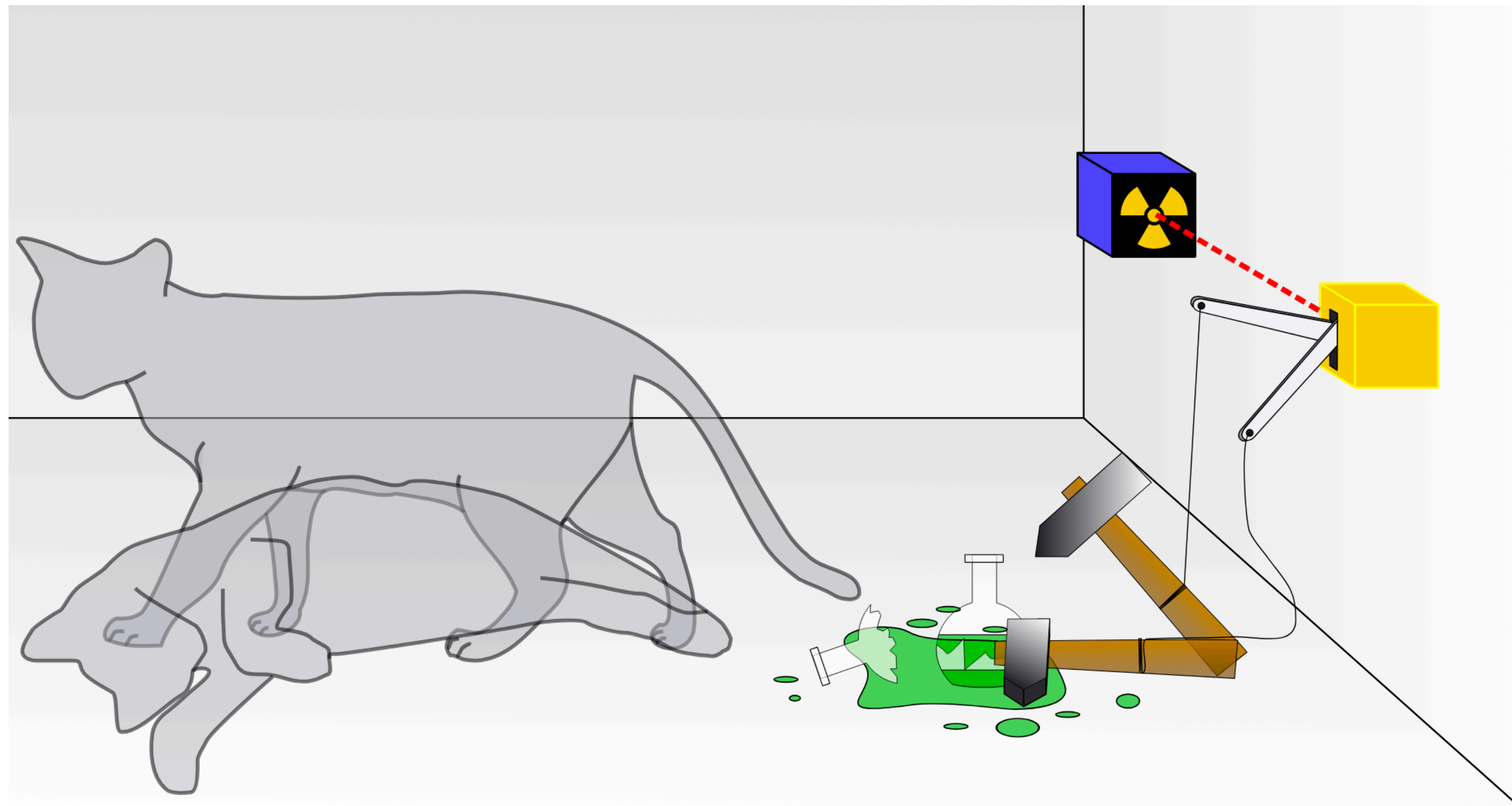
$$i\hbar \frac{\partial \Psi}{\partial t} = \left[ \frac{-\hbar^2}{2m} \frac{\partial^2}{\partial x^2} + V \right] \Psi$$

- The Schrödinger Equation describes the quantum wave function of particles.
- Particles do not have a position until it is measured. The square of the wave function tells you the probability of finding the particle in some region
- The constant  $\hbar$  is the smallest possible value of energy. Energy is quantized into tiny packets. You can have  $0, \hbar, 2\hbar, \dots$  but not  $0.13\hbar$  units of energy
- Note the use of imaginary numbers!  $|\Psi|^2$  is the amplitude of a complex number:  $|a + ib|^2 = \sqrt{a^2 + b^2}$ . Imaginary numbers “are not real” but they are critical to our most accurate theories of reality!!



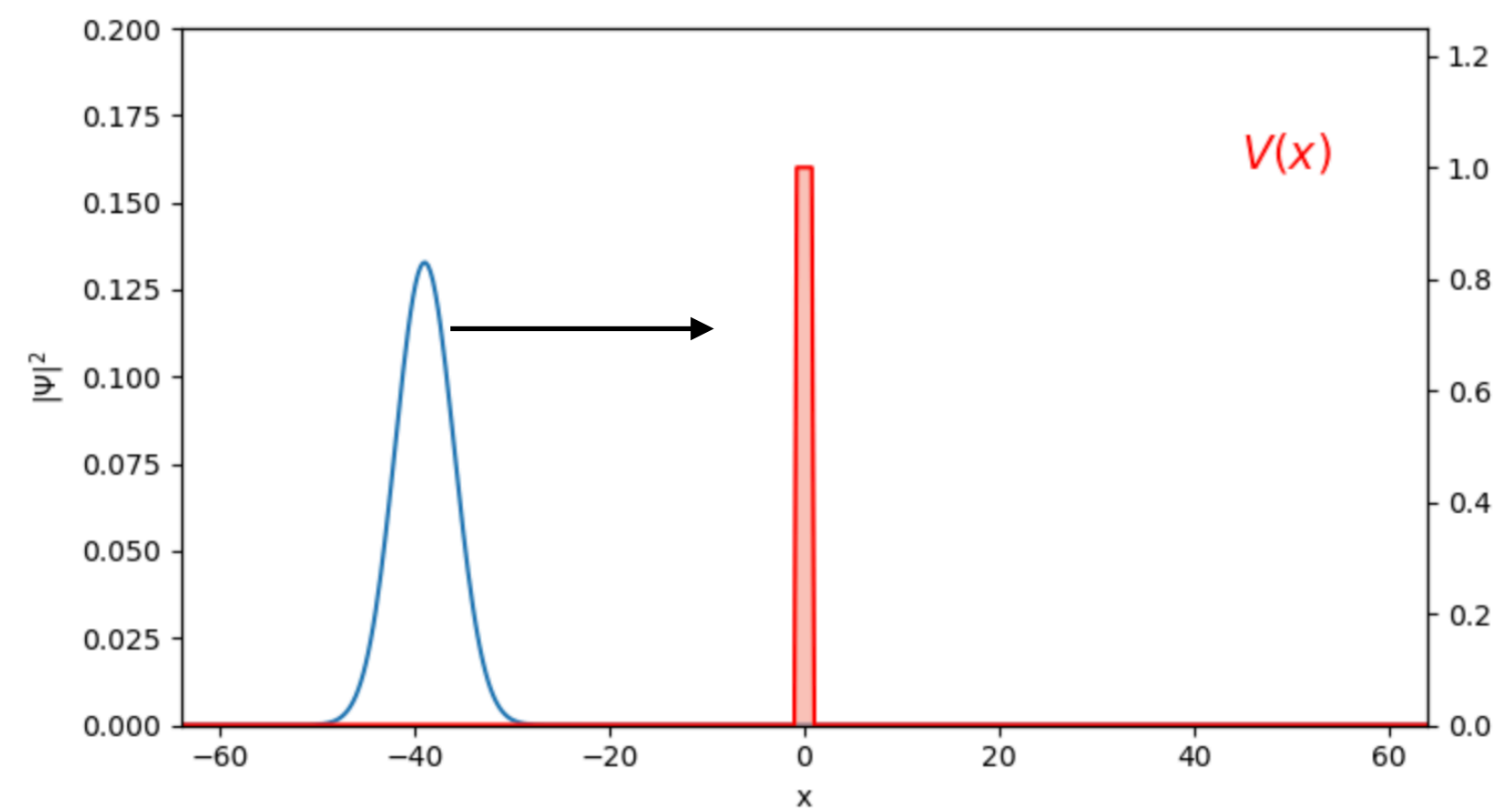
# Schrödinger's cat

- So, because the radioactive isotope has a *probability* of decaying, we can't know for sure whether it does until we open the box.
- Schrödinger originally proposed the thought experiment to highlight the absurdity of quantum superposition, and it remains controversial in interpretation to this day.





# Quantum Tunneling

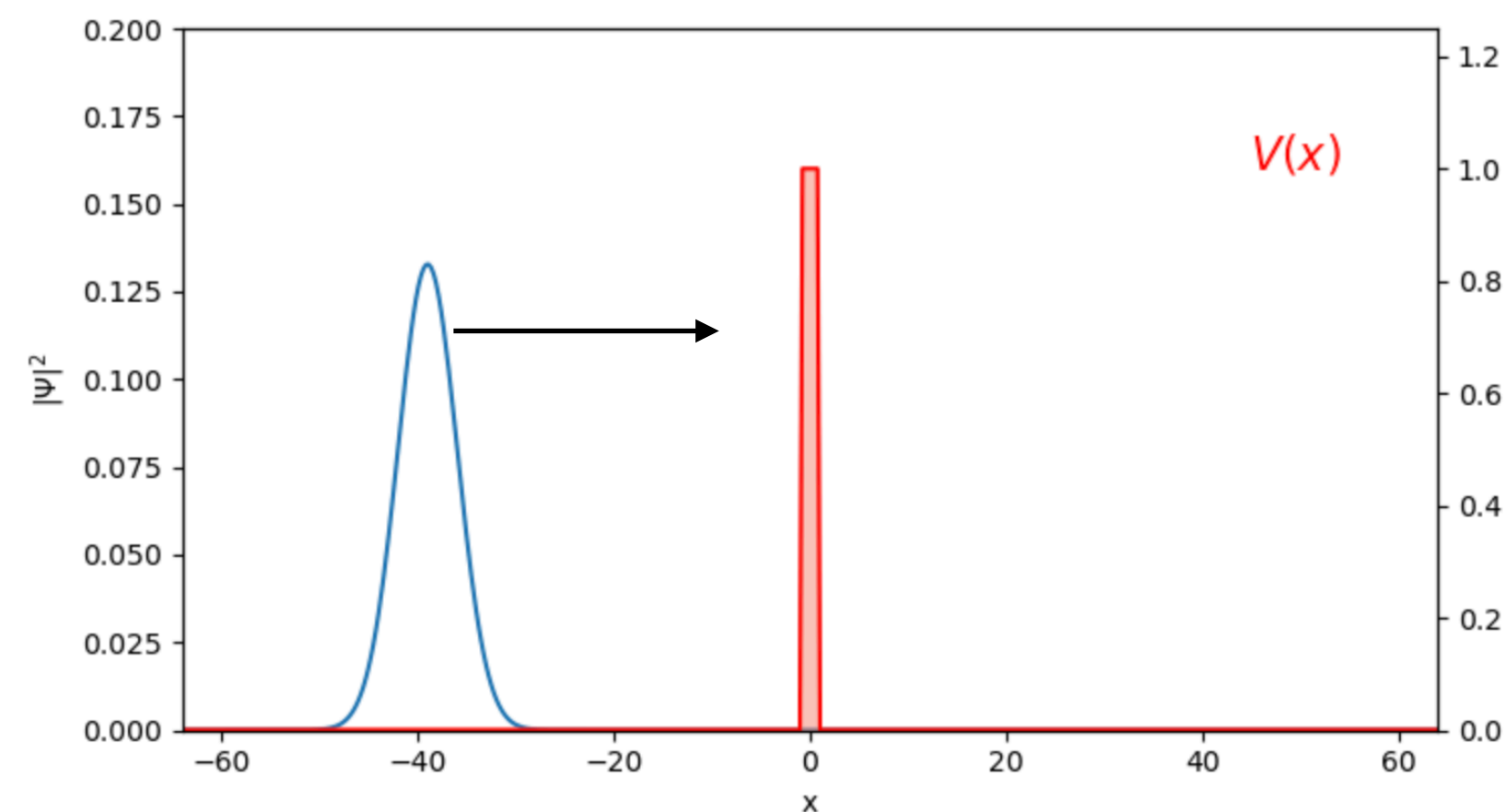


$t = 0$



# Quantum Tunneling

- Imagine we have a quantum wave packet moving to the right towards an energy barrier.
  - The wave packet has half as much energy as the energy barrier.
  - What happens?

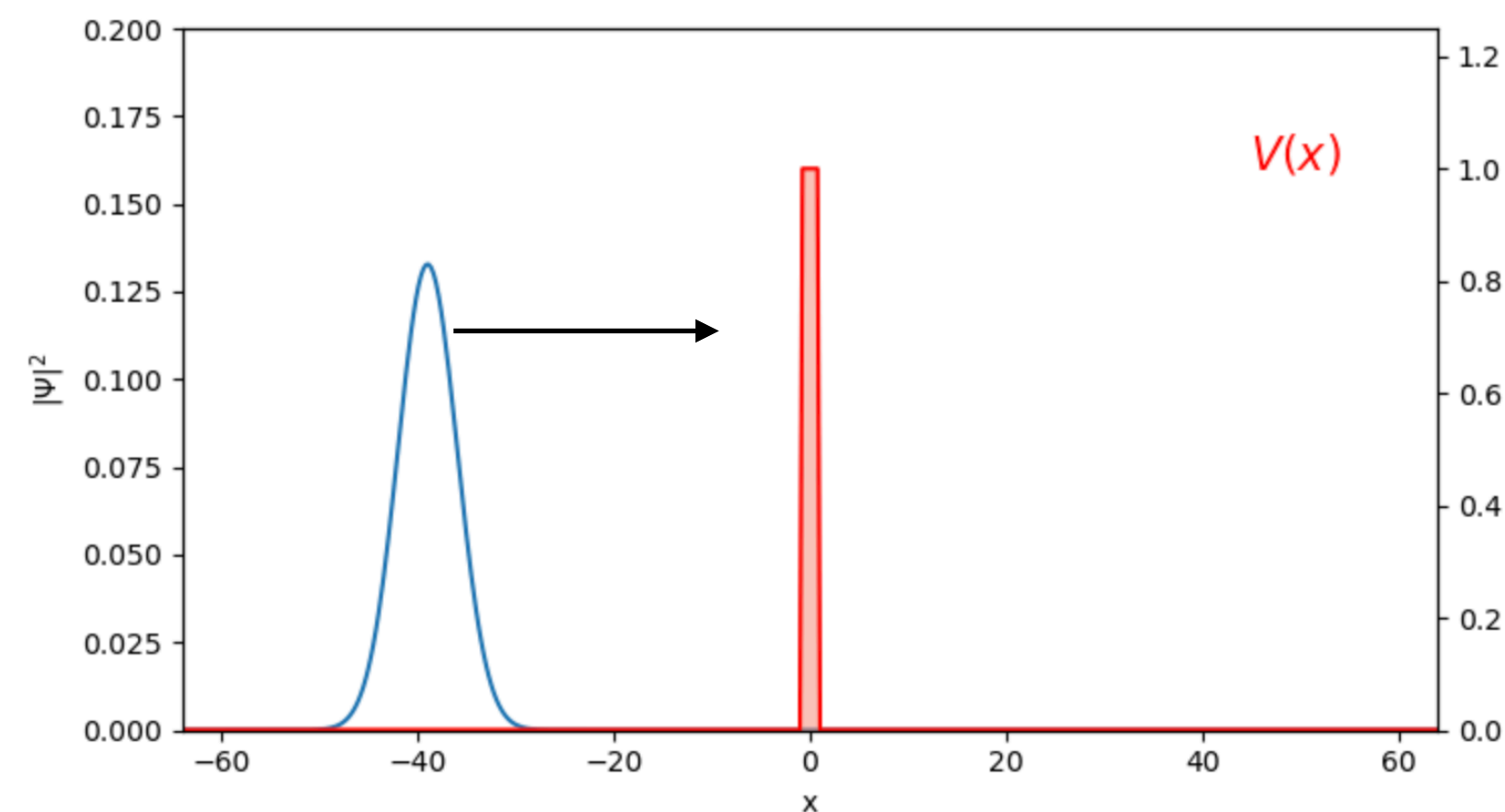


$t = 0$



# Quantum Tunneling

- **Imagine we have a quantum wave packet moving to the right towards an energy barrier.**
  - The wave packet has half as much energy as the energy barrier.
  - What happens?
- **In classical physics, the packet bounces off the barrier and starts moving to the left.**

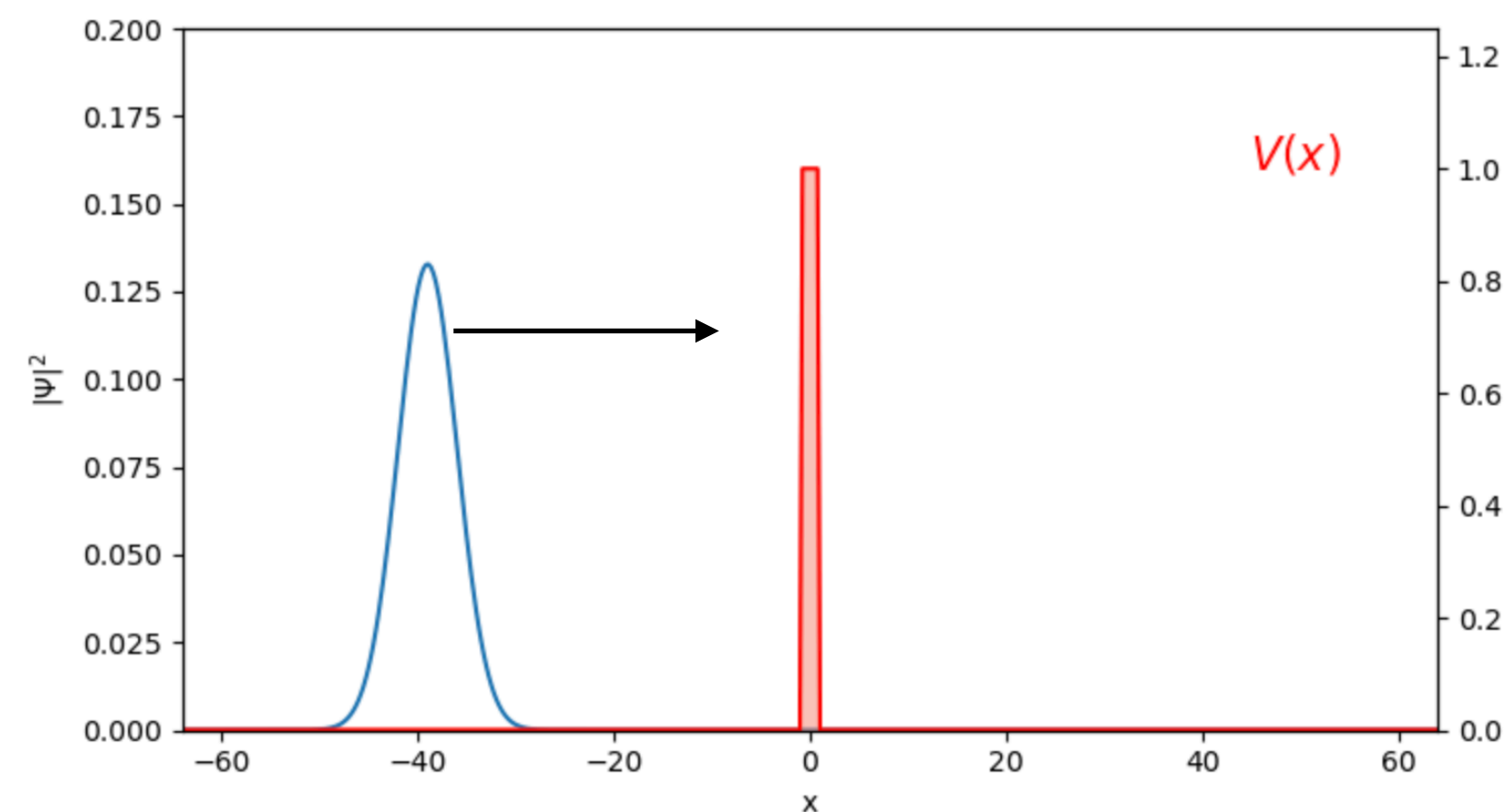


$t = 0$

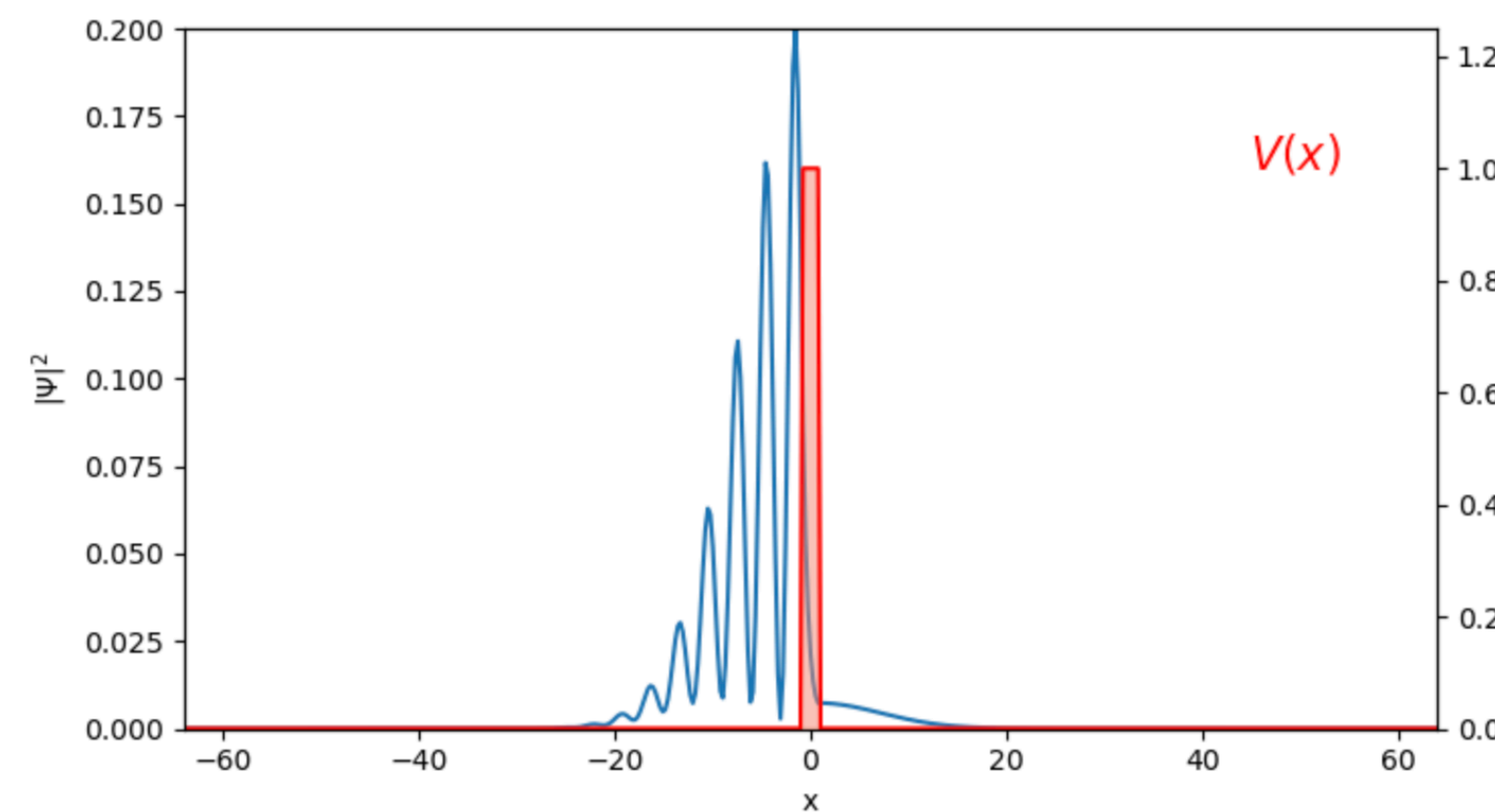


# Quantum Tunneling

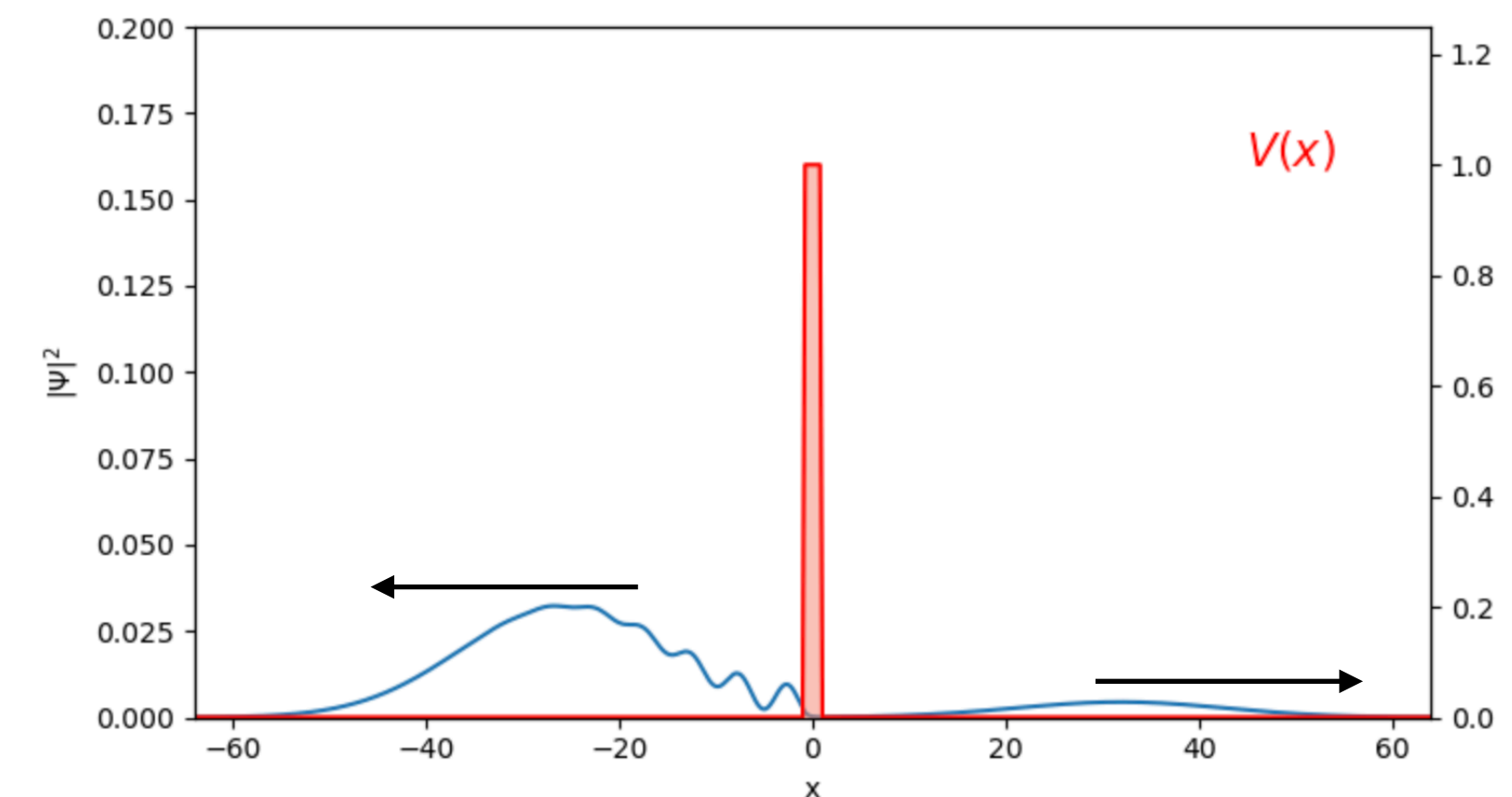
- Imagine we have a quantum wave packet moving to the right towards an energy barrier.
  - The wave packet has half as much energy as the energy barrier.
  - What happens?
- In classical physics, the packet bounces off the barrier and starts moving to the left.
- In quantum physics, this happens *and* some of the wave function leaks through to the right.



$t = 0$



$t = 1$



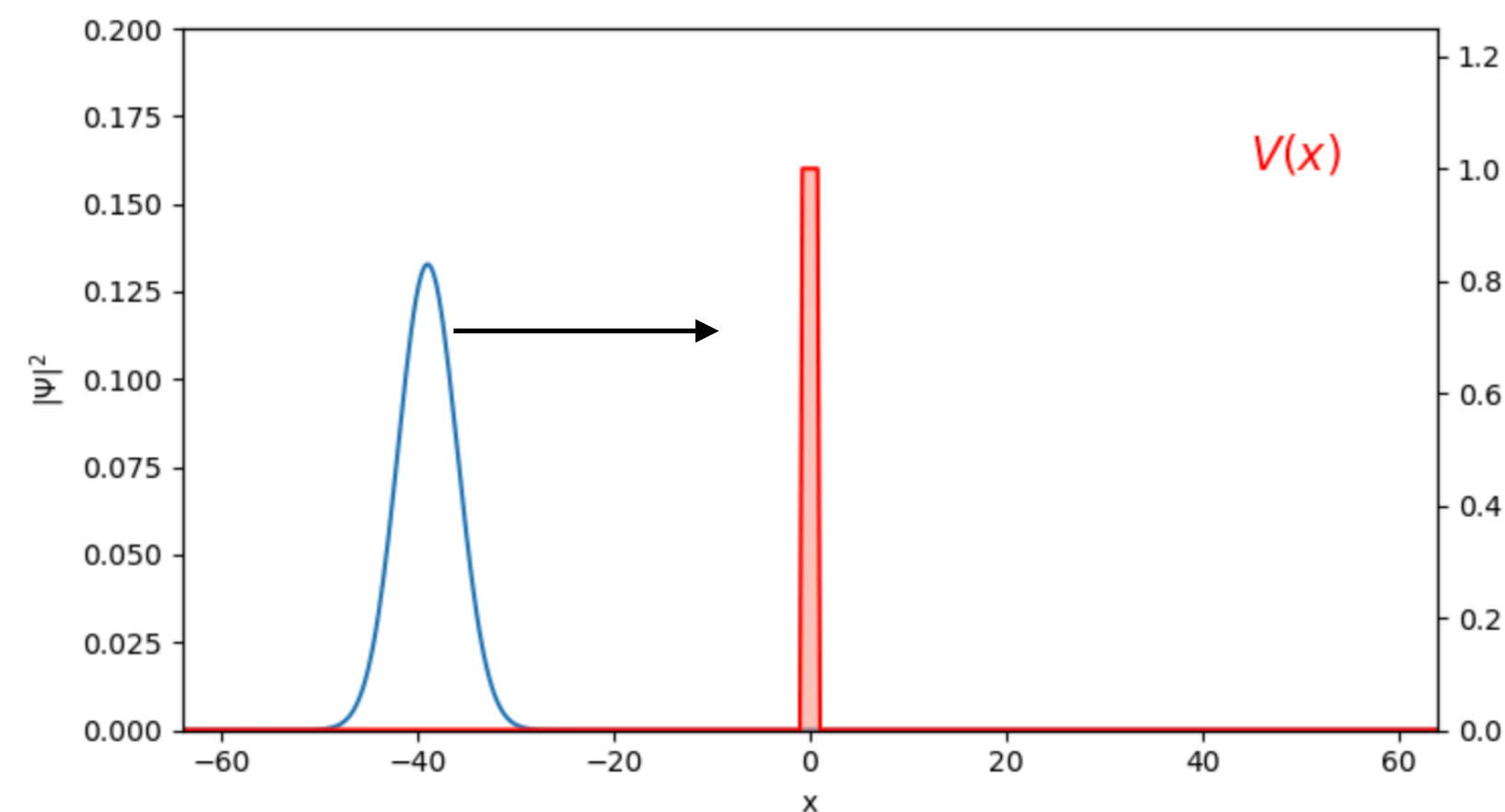
$t = 2$



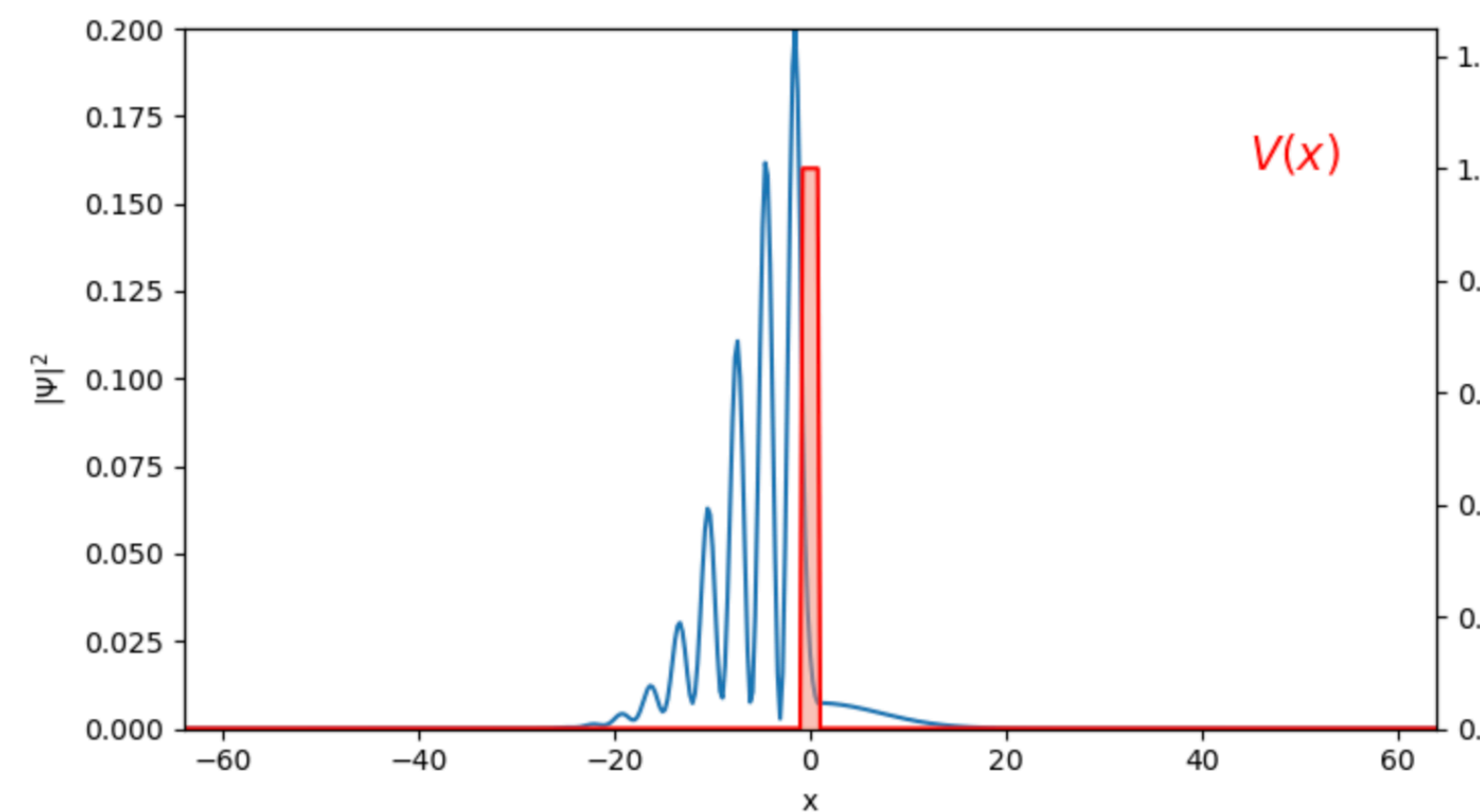


# Quantum Tunneling

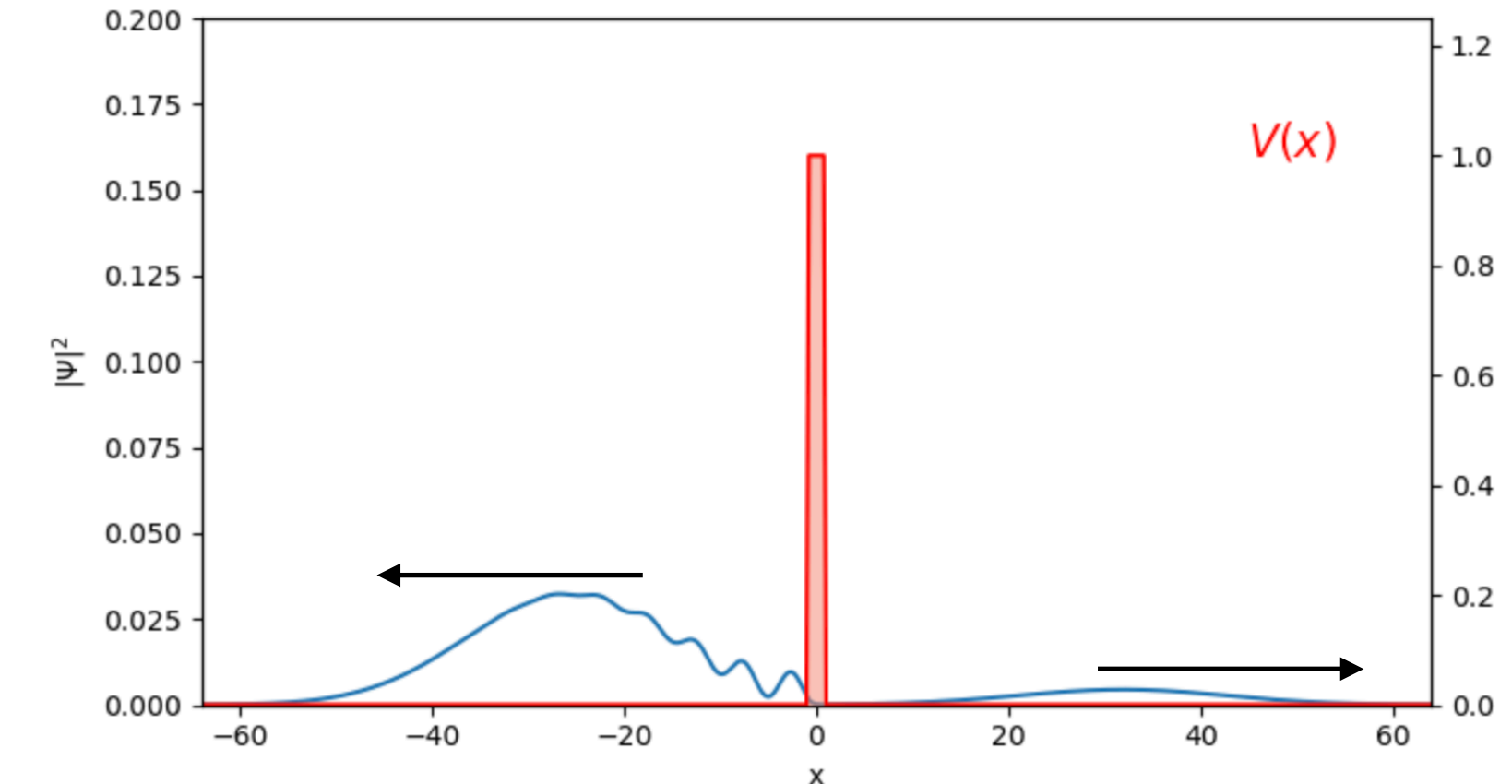
- Imagine we have a quantum wave packet moving to the right towards an energy barrier.
  - The wave packet has half as much energy as the energy barrier.
  - What happens?
- In classical physics, the packet bounces off the barrier and starts moving to the left.
- In quantum physics, this happens *and* some of the wave function leaks through to the right.
- We don't know if the particle bounced off the barrier or tunneled through it until we “open the box”



$t = 0$



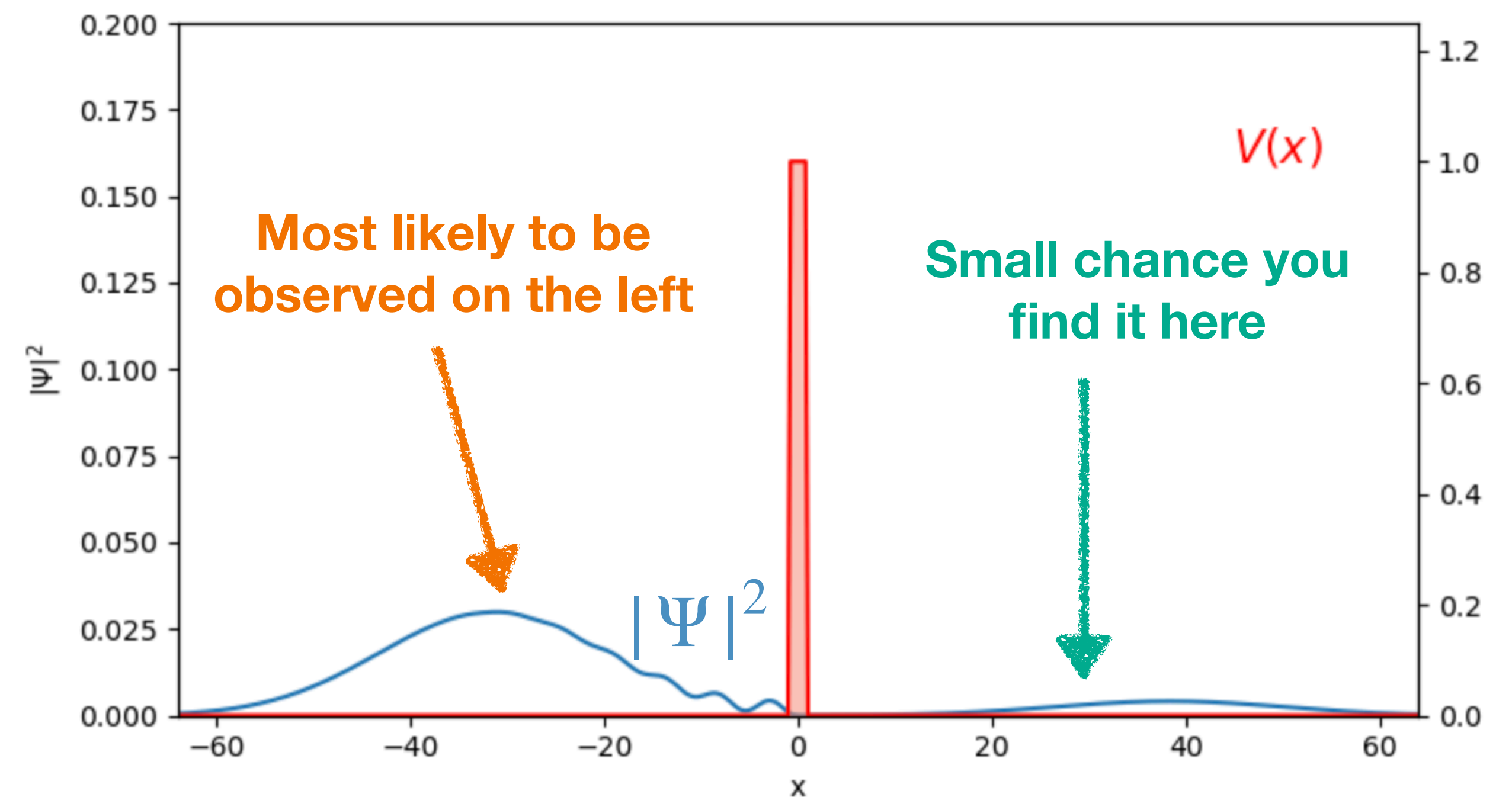
$t = 1$



$t = 2$

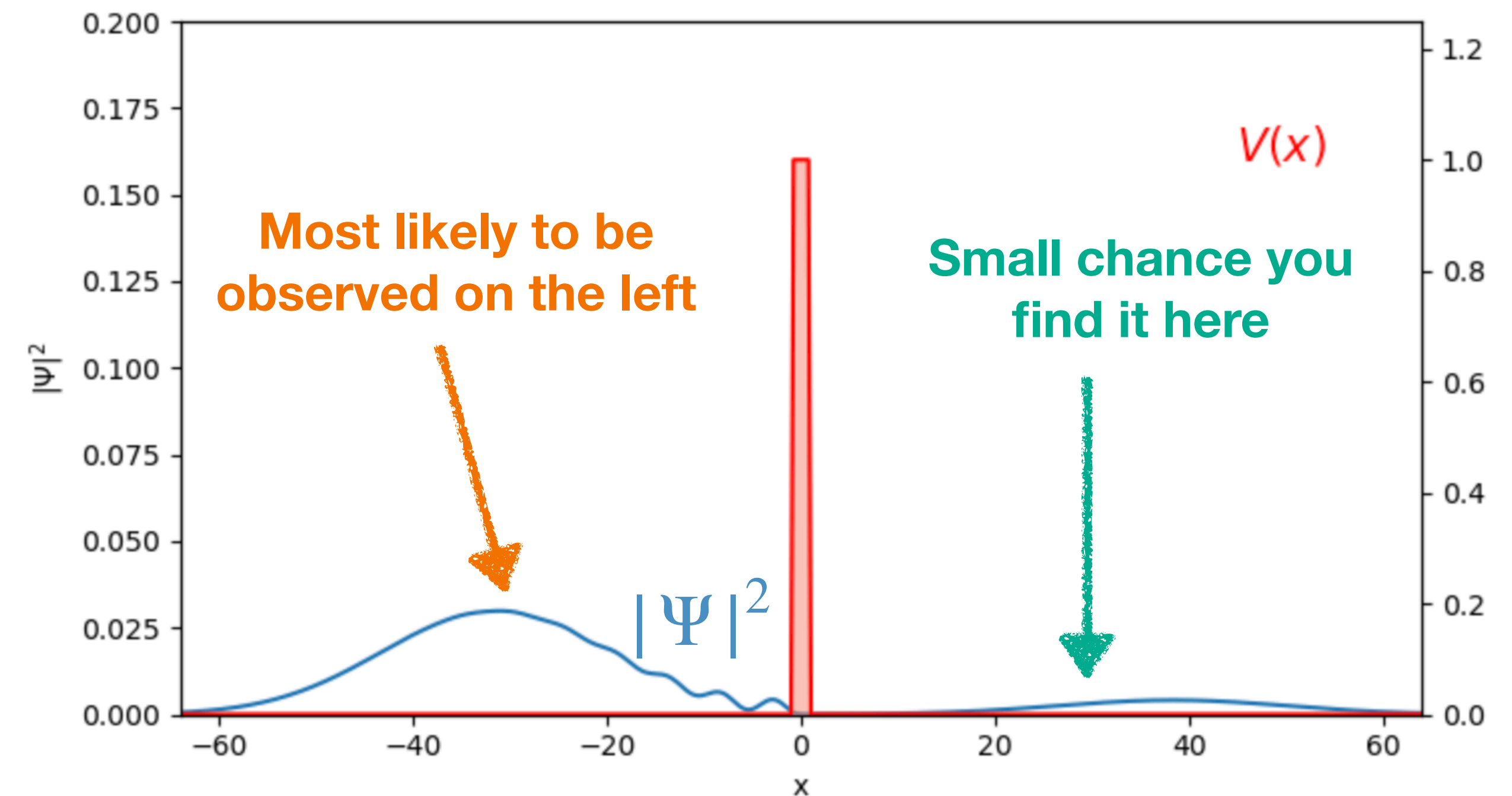


# What if we open the box?



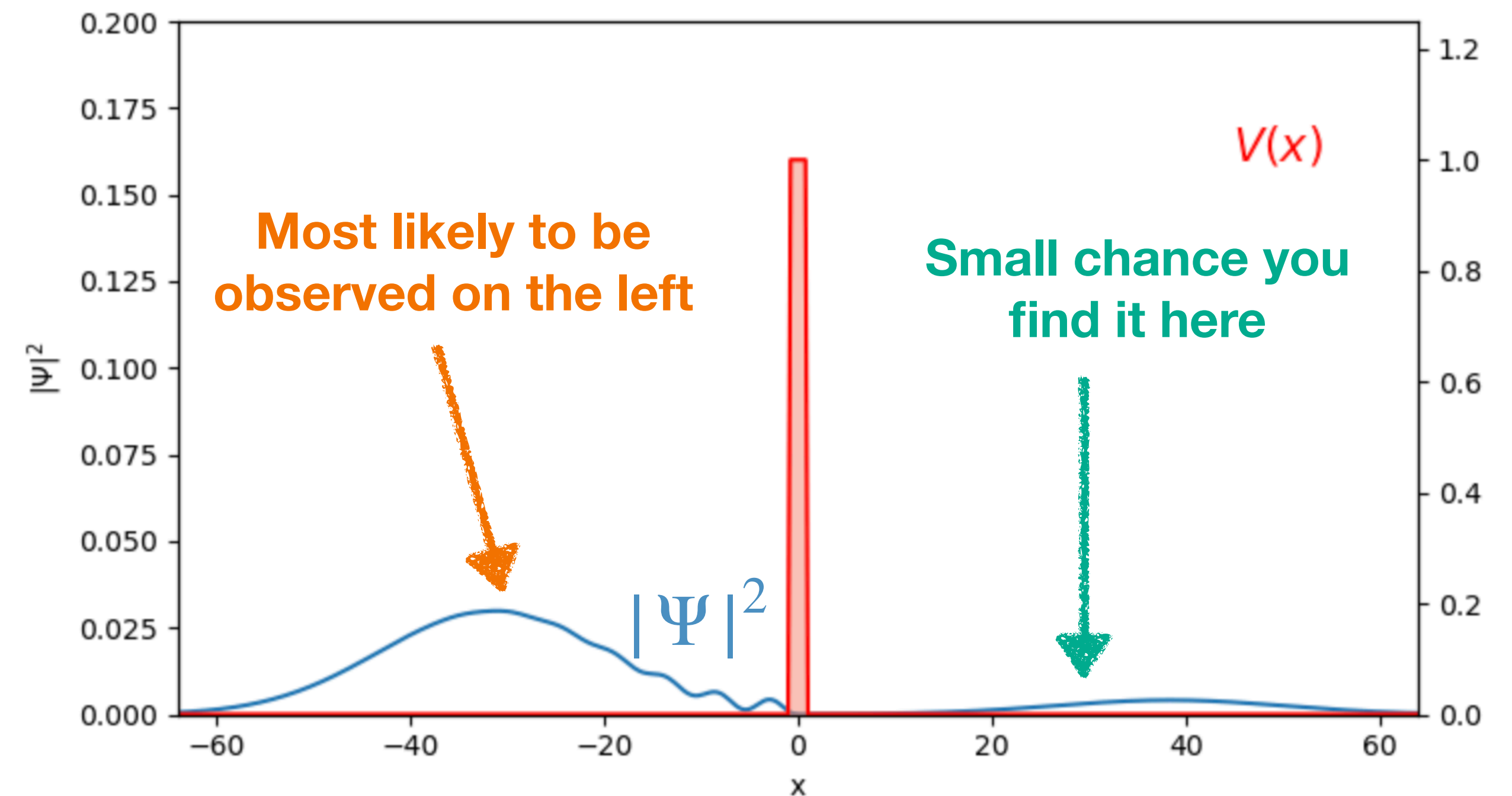
# What if we open the box?

- We are not simulating wave function collapse, but simply the evolution of the quantum wave function.



# What if we open the box?

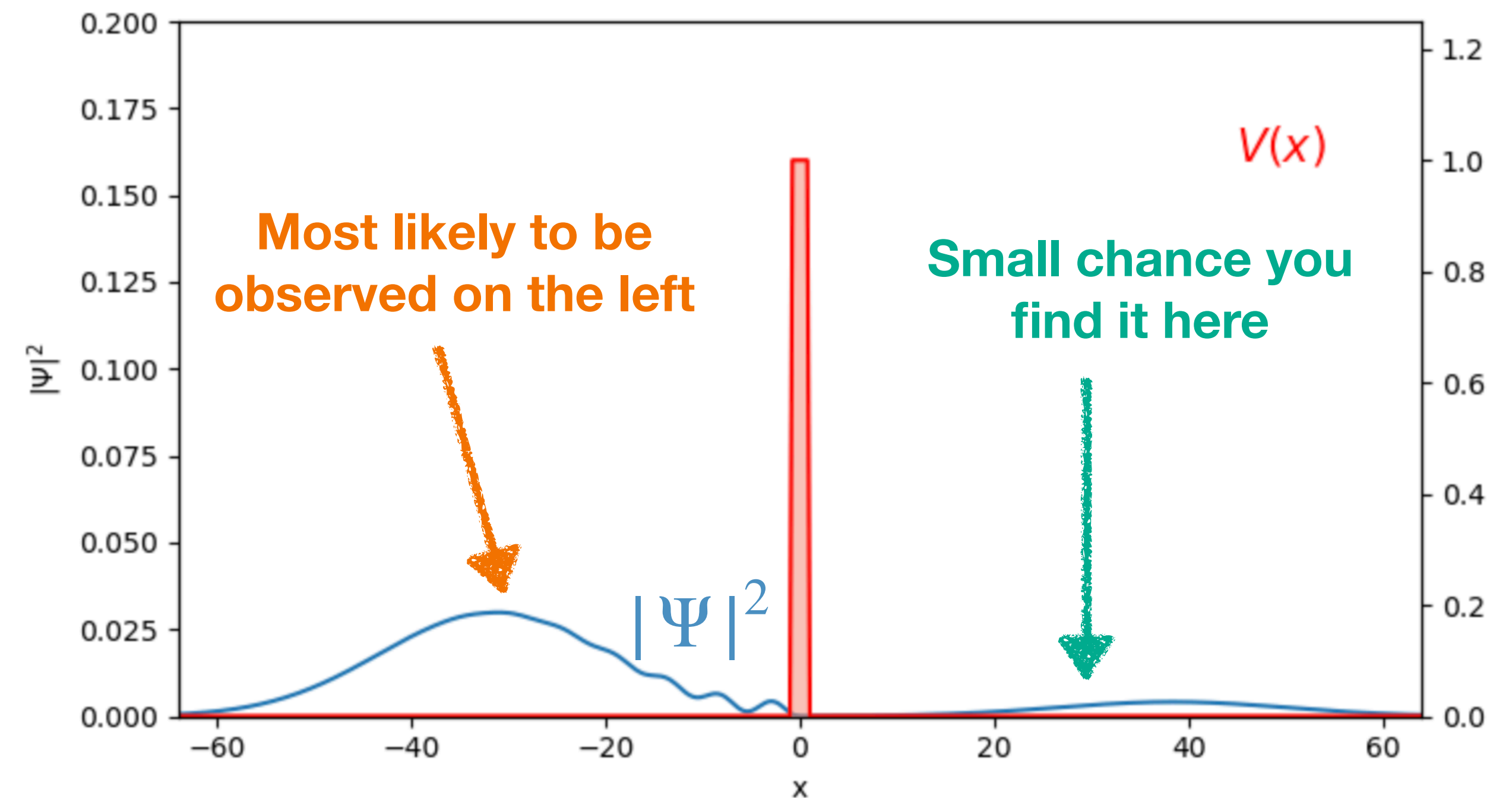
- We are not simulating wave function collapse, but simply the evolution of the quantum wave function.
- If you pause your simulation at a given time, the function currently displayed shows the probability of the particle to be in any particular location (we square the wave function before plotting it)





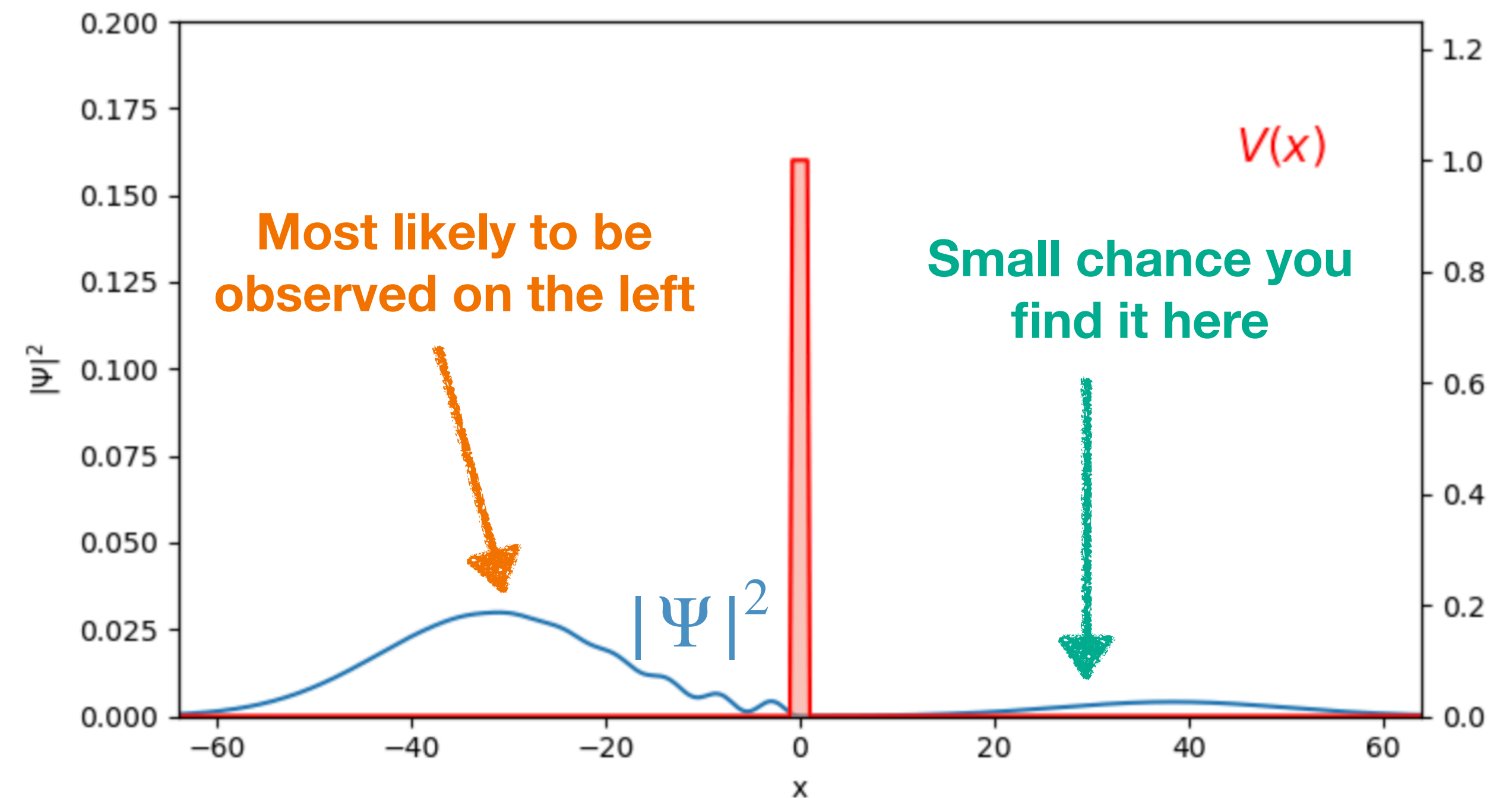
# What if we open the box?

- We are not simulating wave function collapse, but simply the evolution of the quantum wave function.
- If you pause your simulation at a given time, the function currently displayed shows the probability of the particle to be in any particular location (we square the wave function before plotting it)
- When you “open the box” you just get the position of the particle. In this case the most likely answer is  $x=-30$ , but you could sometimes get  $x=-20$  or  $-50$ . If you repeat the experiment rarely, but sometimes, you’ll get a positive number.



# What if we open the box?

- We are not simulating wave function collapse, but simply the evolution of the quantum wave function.
- If you pause your simulation at a given time, the function currently displayed shows the probability of the particle to be in any particular location (we square the wave function before plotting it)
- When you “open the box” you just get the position of the particle. In this case the most likely answer is  $x=-30$ , but you could sometimes get  $x=-20$  or  $-50$ . If you repeat the experiment rarely, but sometimes, you’ll get a positive number.
- When you open the box in Schrödinger’s thought experiment, you find out if the cat is alive or dead. It will be one or the other, not both.



# Schrödinger's wave equation

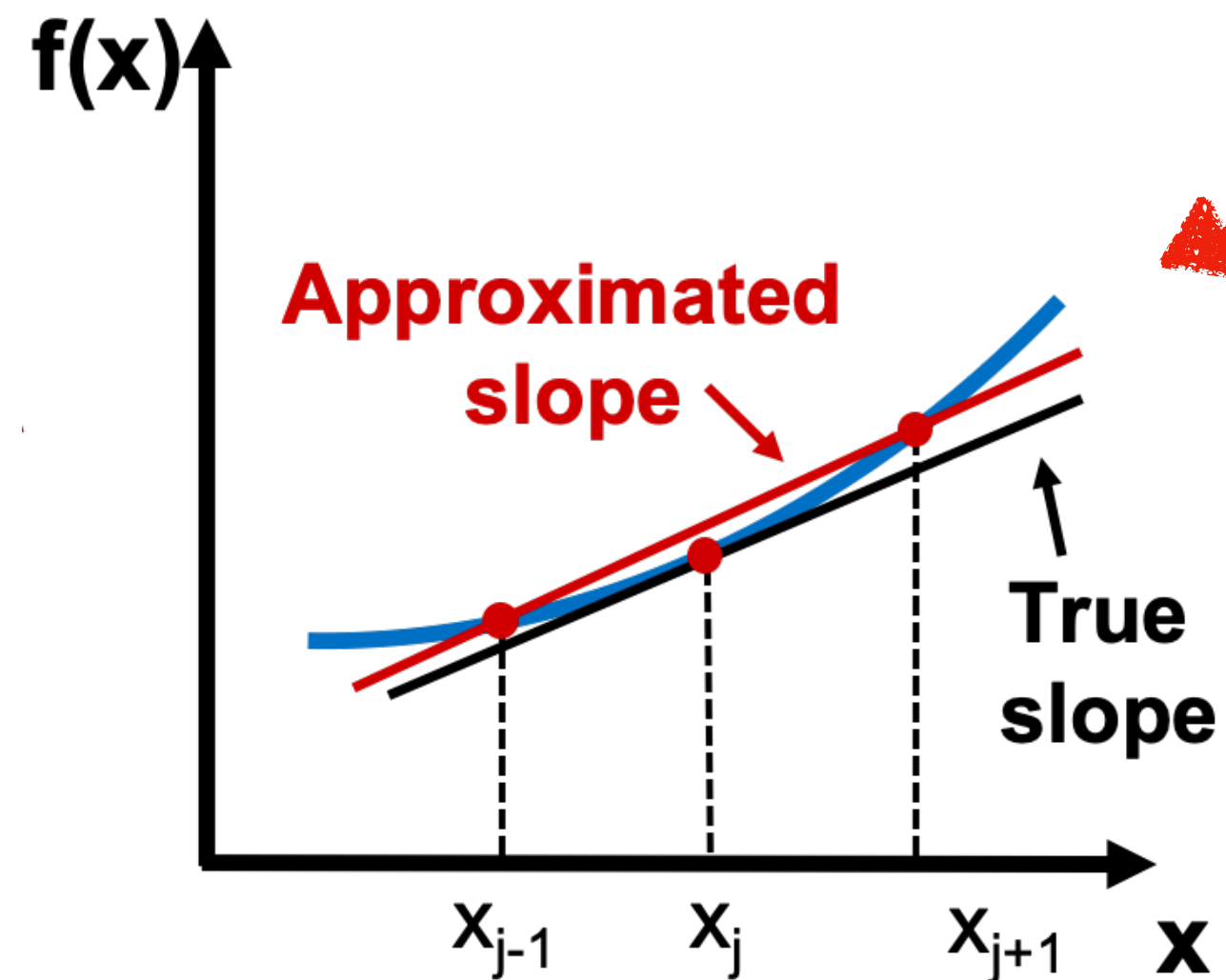
$$i\hbar \frac{\partial \Psi}{\partial t} = \left[ \frac{-\hbar^2}{2m} \frac{\partial^2}{\partial x^2} + V \right] \Psi$$

- The Schrödinger Equation is not only a *wave equation* but a *differential equation*. It is an equation with derivatives of functions inside it.
- Like wave equations, differential equations are *everywhere* in physics:
  - Newtonian physics:  $F = ma = m \frac{d^2x}{dt^2}$
  - Fluid dynamics
  - Schrödinger equation and other quantum wave equations
- Once we have an initial state, we can use this equation in a computer simulation to tell us what the function will look like after some small time step  $\Delta t$ .



# Today's Simulation

- Today you will run a numerical simulation of the Schrödinger Equation to learn about its strange properties
  - Quantization of energy
  - Quantum tunneling
  - Self-interference and wave-particle duality
- The simulation uses approximations of derivatives by discretizing space and time



Derivatives are hard

$$i\hbar \frac{\partial \Psi}{\partial t} = \frac{-\hbar^2}{2m} \frac{\partial^2 \Psi}{\partial x^2}$$

Approximate them by  
Evaluating the function  
At two nearby points

$$\frac{\partial f}{\partial x} \approx \frac{f(x+h) - f(x)}{h}$$

- The numerical simulation will be in the form of a *Jupyter notebook*. Running the notebook will require you to write small amounts of python code. The simulation has already been written.

