

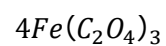
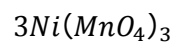
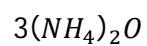
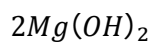
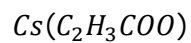
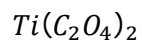
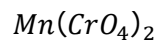
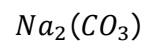
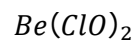
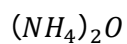
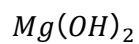
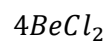
C11 - 1.1 - # of Particles HMW

Fill in the following table.

Element	Atomic #	# of Protons	Atomic Mass	# of Neutrons	# of electrons in <u>Atom</u>	Ion Charge	# of electrons in <u>Ion</u>	Ion Valence electrons
oxygen								
calcium								
silicon								
Iron (II)								
Manganese (IV)								
Nickel (III)								
lithium								
Cobalt (II)								
iodine								
Iridium (IV)								
Lead (II)								
Carbon								
chlorine								
Uranium (VI)								
magnesium								
Chromium (II)								
aluminum								
Platinum (IV)								
helium								
hydrogen								
Plutonium (V)								
Gold (I)								
radium								
neon								
potassium								
tin								
radon								
sodium								
Sulfur								
Mercury								

C11 - 1.1 - # of Atoms WS

How many of each atom are there in the following molecules?



C11 - 1.1 - Bohr Diagram HMW

Draw a Bohr diagram for the following atoms.

O

H

Li

Al

N

He

Ne

Fe

K

Na

Be

Draw a Bohr diagram for the following ions

O⁻²

H⁺

Li⁺

Al⁺³

N⁻³

He

Ne

Fe⁺²

K⁺

Na⁺¹

Be⁺²

Fe⁺³

C11 - 1.1 - Lewis Diagram HMW

Draw a Lewis diagram for the following atoms.

O

H

Li

Al

N

He

Ne

Fe

K

Na

Be

Draw a Lewis diagram for the following ions

O⁻²

H⁺

Li⁺

Al⁺³

N⁻³

He

Ne

Fe⁺²

K⁺

Na⁺¹

Be⁺²

Fe⁺³

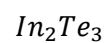
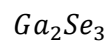
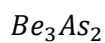
C11 - 1.5 - Naming Ionic WS

Name the ionic compounds:



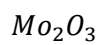
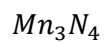
C11 - 1.5 - Naming Ionic WS

Name the ionic compounds:



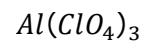
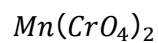
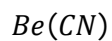
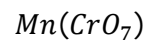
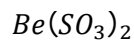
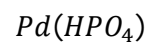
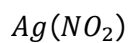
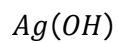
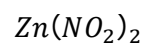
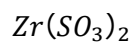
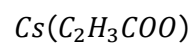
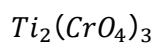
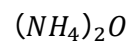
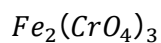
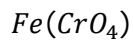
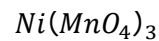
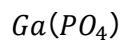
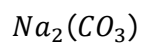
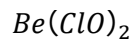
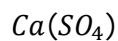
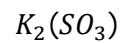
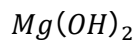
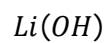
C11 - 1.5 - Naming Multivalent WS

Name the multivalent ionic compounds:



C11 - 1.5 - Naming Polyatomic WS

Name the Polyatomic compounds:



C11 - 1.5 - Naming Covalent WS

Name the covalent compounds:



C11 - 1.5 - Writing Ionic Formulas WS

Write the formula for each ionic compounds:

Magnesium fluoride

Barium selenide

Cesium bromide

Potassium iodide

Silver oxide

Potassium bromide

Radium chloride

Rubidium chloride

Hydrogen chloride

Calcium oxide

Strontium Asatide

Sodium oxide

Sodium fluoride

Lithium fluoride

Beryllium oxide

Calcium arsenide

Zirconium sulfide

Magnesium sulfide

Scandium phosphide

Lithium oxide

Aluminum iodide

Calcium phosphide

Magnesium oxide

Potassium phosphide

Cesium sulfide

Gallium Chloride

Tin oxide

Cerium oxide

Sodium chloride

Francium nitride

Rubidium phosphide

Thorium oxide

Yttrium sulfide

Zinc oxide

Cadmium oxide

C11 - 1.5 - Writing Multivalent Formulas WS

Write the formula for each multivalent ionic compounds:

Titanium (III) oxide

Iron (II) oxide

Cobalt (II) bromide

Cobalt (III) bromide

Cobalt (II)selenide

Nickel (*III*) fluoride

Nickel (II) oxide

Manganese (IV) nitride

Nickel (III) chloride

Gold Nitride

Lead (IV) iodide

Iron (II) iodide

Gold (III) oxide

Copper (II) sulfide

Palladium (II) phosphide

Palladium (IV) sulfide

Platinum (II) Oxide

Vandium (V) sulfide

Copper (II) sulfide

Palladium chloride

Chromium (*III*) iodide

Platinum (IV) sulfide

Bismuth (V) Sulfide

Copper fluoride

Polonium (IV) asatide

Lead (IV) oxide

Plutonium (VI) oxide

Uranium (IV) oxide

Manganese (IV) phosphide

Antimony (III) chloride

C11 - 1.5 - Writing Polyatomic Formulas WS

Write the formula for each polyatomic compounds:

Copper (II) sulfate

Copper sulfate

Calcium oxalate

Iron (III) nitrite

Hydrogen sulfate

Lithium chromate

Sodium phosphate

Iron (II) nitrate

Titanium (IV) bisulfite

Zinc acetate

Zinc silicate

Calcium phosphate

Cobalt (II) permanganate

Titanium (IV) sulfate

Hydrogen cyanide

Nickel (*II*) carbonate

Sodium dichromate

Calcium hydroxide

Barium chlorate

lithium bisulfate

Iron perchlorate

Titanium iodate

Chromium carbonate

Ammonium oxide

Ammonium fluoride

Chromium bicarbonate

Ammonium phosphide

Beryllium bicarbonate

Silver oxalate

Scandium silicate

Manganese chlorite

C11 - 1.5 - Writing Covalent Formulas WS

Write the formula for each covalent compound:

Nitrogen dioxide

Carbon monoxide

Chlorine trifluoride

Carbon dioxide

Dihydrogen oxide

Tetrasulfur dinitride

Dinitrogen Trioxide

Diphosphorus dextoxide

Dinitrogen Tetroxide

Carbon disulfide

Sulfur Tetrafluoride

Bromine Fluoride

Dinitrogen Trioxide

Carbon Tetraoxide

Fluorine Pentoxide

Sulfur Tetroxide

Pentaboron Dioxide

Arsenic hexoxide

Silicon nonochloride

Bromine Decafluoride

Phosphorus triselenide