C11 - 2.2 - Balancing equations WS

Balance the following equations

$$\underline{\hspace{1cm}} H_2 + \underline{\hspace{1cm}} O_2 \rightarrow \underline{\hspace{1cm}} H_2 O$$

$$\underline{\hspace{1cm}}$$
 $Na + \underline{\hspace{1cm}} Cl_2 \rightarrow \underline{\hspace{1cm}} NaCl$

$$\underline{\hspace{1cm}}$$
 $Ca + \underline{\hspace{1cm}}$ $O_2 \rightarrow \underline{\hspace{1cm}}$ CaO

$$\underline{\hspace{1cm}} H_2O \rightarrow \underline{\hspace{1cm}} H_2 + \underline{\hspace{1cm}} O_2$$

$$\underline{\hspace{1cm}} NH_3 \rightarrow \underline{\hspace{1cm}} N_2 + \underline{\hspace{1cm}} H_2$$

$$C_2H_6O_2 + C_2O_2 + H_2O_2$$

$$\underline{\hspace{1cm}}$$
 $NaCl + \underline{\hspace{1cm}}$ $FeO \rightarrow \underline{\hspace{1cm}}$ $Na_2O + \underline{\hspace{1cm}}$ $FeCl_2$

$$Na + MCl \rightarrow NaCl + M_2$$

$$__Ni(OH)_2 + ___LiF \rightarrow ___NiF_2 + ___Li(OH)$$

$$__Ni(OH)_2 + __LiF \rightarrow __NiF_2 + __Li(OH)$$
 $__Cu_2(SO_4)_3 + __Mg \rightarrow __Cu + __Mg(SO_4)$

C11 - 2.2 - Find Products and Balance WS

Find the products and balance the following chemical reaction.

$$K_3P + Li_{Cl} \longrightarrow$$

$$Ca + Cl_2 \longrightarrow$$

$$Fe(SO_4) + NiCl_3 \longrightarrow$$

$$NaCl \longrightarrow$$

$$Zn+$$
 $O_2 \rightarrow$

$$Fe_2O_3 + CaF_2 \longrightarrow$$

$$HI + NaF \longrightarrow$$

$$N_2 + O_2 \longrightarrow$$

C11 - 2.2 - Create Equations and Balance WS

Write the following word equations into balanced chemical equations

Copper plus sulpher produces copper sulfide	Hydrogen plus oxygen produces water
Iron oxide plus magnesium sulfide produces iron sulfide plus magnesium oxide	Hydrocarbon plus oxygen produces carbon dioxide plus water
Copper sulphate plus iron produces copper loss iron sulphate	Sodium chloride plus potassium hydroxide produces sodium hydroxide plus potassium chloride.
Lithium nitrate plus copper sulphate produces lithium sulphate plus copper nitrate	Nickel + tin chloride produces nickel chloride plus tin

C11 - 2.3 - Reaction Types WS

Name the type of reaction being synthesis, decomposition, single replacement, double replacement, neutralization or combustion.

$$H_2 + O_2 \rightarrow H_2O$$

$$NH_3 \rightarrow N_2 + H_2$$

$$Ca + O_2 \rightarrow CaO$$

$$Na + HCl \rightarrow NaCl + H_2$$

$$HCl + Na(OH) \longrightarrow NaCl + H_2O$$

$$C_2H_6O_2 + O_2 \rightarrow CO_2 + H_2O$$

$$Na + Cl_2 \rightarrow NaCl$$

$$H_2O \rightarrow H_2 + O_2$$

$$Ni(OH)_2 + LiF \rightarrow NiF_2 + Li(OH)$$

$$C_3H_8O_2 + O_2 \to CO_2 + H_2O$$

$$NaCl + FeO \rightarrow Na_2O + FeCl_2$$

$$Cu_2(SO_4)_3 + Mg \rightarrow Cu + Mg(SO_4)$$

$$H_2(SO_4) + K(OH) \rightarrow K_2(SO_4) + H_2O$$